



**First Sessional Theory Examination 2020-2021**  
(Odd SEM)

**Subject & Subject Code: Remedial Mathematics (BP106RMB)**

**Day & Date: Tuesday 16/3/2021**

**Class: F.Y/S.Y. B. Pharmacy**

**Semester: I/III**

**Time: 1 hr**

**Max. Marks: 30**

- Instructions:*
1. All questions are compulsory
  2. Draw a well labeled diagram wherever necessary
  3. Right hand side number indicates full marks

**QI Multiple Choice Questions (MCQs)**

10

**1. If A and B are symmetric matrices of the same order, then**

- (a) AB is a symmetric matrix
- (b)  $A - B$  is askew-symmetric matrix
- (c)  $AB + BA$  is a symmetric matrix
- (d)  $AB - BA$  is a symmetric matrix

**2. If  $A = \begin{bmatrix} 2 & 4 \\ X & 0 \end{bmatrix}$  is a symmetric matrix, then x =**

- (a) 4
- (b) 3
- (c) -4
- (d) -3

**3. If A is a square matrix, then  $A - A^T$  is a**

- (a) diagonal matrix
- (b) skew-symmetric matrix
- (c) symmetric matrix
- (d) none of these

**4. If A is any square matrix, then which of the following is skew-symmetric?**

- (a)  $A + A^T$
- (b)  $A - A^T$
- (c)  $AA^T$
- (d)  $A^T A$

**5. Let A is any square Matrix Then  $A^{-1}$  exist if**

- (a)  $|A| \neq 0$
- (b)  $|A| = 0$
- (c) A is singular Matrix
- (d) None

**6. The value of  $\begin{vmatrix} 1 & 2 & 3 \\ -3 & -6 & -9 \\ 2 & 1 & 0 \end{vmatrix}$  is**

- (a) 4
- (b) 3
- (c) -4
- (d) 0



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7. The value of  $\begin{vmatrix} 1 & 1 & 1 \\ 1 & 1 & 1 \\ 1 & 1 & 1 \end{vmatrix}$  is

- (a) 2
- (b) 3
- (c) 0
- (d) 1

8 A square Matrix A is said to Symmetric if

- (a)  $A = A^T$
- (b)  $A = -A^T$
- (c)  $A + A^T$  is identity
- (d)  $A^T A$  exist.

9. If  $\begin{vmatrix} 1 & 2 \\ 3 & x \end{vmatrix} = 0$  then x is

- (a) 2
- (b) 3
- (c) 6
- (d) 1

10 If  $A = \begin{bmatrix} 1 & 2 \\ 3 & 7 \end{bmatrix}$  then  $A^{-1}$  is

- (a)  $\begin{bmatrix} 7 & -2 \\ -3 & 1 \end{bmatrix}$
- (b)  $\begin{bmatrix} 1 & -2 \\ 3 & 7 \end{bmatrix}$
- (c)  $\begin{bmatrix} 1 & 2 \\ -3 & 7 \end{bmatrix}$
- (d)  $\begin{bmatrix} 1 & -2 \\ -3 & 7 \end{bmatrix}$

**QII Long Answers (Answer any 1 out of 2)**

10

1. Solve the following System of equations:  $2x + 3y - z = 1$   
 $4x + y - 3z = 11$   
 $3x - 2y + 5z = 21$

2. Solve the following System of equations:  $-x + 3y - 2z = 5$   
 $4x - y - 3z = -8$   
 $2x + 2y - 5z = 7$

**QIII Short Answers (Answer any 2 out of 3)**

10

1. If  $A = \begin{bmatrix} 1 & 2 & 3 \\ -1 & 0 & 2 \\ -2 & 4 & 1 \end{bmatrix}$  Find  $A^{-1}$  using adjoint method.

2. Find the value of x if  $\begin{vmatrix} 1 & 2 & 3 \\ 1 & x & 4 \\ 2 & -3 & -1 \end{vmatrix} = 0$ .

3. Find minors and cofactors of each elements of  $A = \begin{bmatrix} 0 & 1 & 2 \\ 3 & -1 & 6 \\ 2 & -4 & -8 \end{bmatrix}$





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**First Sessional Theory Examination 2018-2019**

**Subject: Comm. Skills**

**Day & Date: Saturday, 6<sup>th</sup> October 2018**

**Class: FY/ DSA**

**Semester: First**

**Time: 10:30 am to 12:00 pm**

**Max. Marks: 30**

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- Instructions:*
1. All questions are compulsory
  2. Draw a well labeled diagram wherever necessary
  3. Right hand side number indicates full marks

**QI Long Answers (Answer any 1 out of 2)**

**10**

- (1) Define Communication and illustrate its importance or objectives.
- (2) What are the elements of Communication? Explain them in detail with Comm. Cycle.

**QII Short Answers (Answer any 4 out of 6)**

**20**

1. Explain Physiological barrier with its all subtypes in detail with examples.
2. What is linguistic barrier? Explain it in detail with all its sub-types.
3. What is Comm. Styles Matrix? Explain Direct and Spirited Comm. Styles with example.
4. State the importance of Body language with its constituents in detail.
5. Explain Psychological barrier with its all subtypes in detail with examples.
6. What are Systematic and Considerate Comm. Styles? Write a note on factors affecting our perspectives.



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### First Sessional Theory Examination 2020-2021

**Subject: Comm. Skills**

**Day & Date: Monday, 15 March. 2021**

**Class: FY/ DSA**

**Semester: First**

**Time: 10.30 to 11.30**

**Max. Marks: 30**

*Instructions: 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks*

#### **QI Long Answers (Answer any ONE)**

**1\*10**

- (1) *Non-verbal communication enhances the efficacy of professional communication. Explain the statement in relation with all non-verbal codes of communication. (Unit 02)*
- (2) *Explain the process of Communication with its elements in detail. (Unit 01)*

#### **QII Short Answers (Answer any FOUR)**

**4\*5**

1. *What is Communication? Define it in detail. (Unit 01)*
2. *Explain Psychological barrier with all its sub-types and suggest ways to overcome it. (Unit 01)*
3. *What are the types and methods of Communication? (Unit 01)*
4. *State the importance of Body language with its constituents in detail. (Unit 02)*
5. *Explain in detail the factors that affect perspective in day-to-day communication. (Unit 02)*
6. *What are the different communication styles? What do they focus on? (Unit 02)*



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**Subject: Comm. Skills**

**Day & Date: Monday, 15 March. 2021**

**Class: FY/ DSA**

**Semester: First**

**Time: 10.30 to 12.00**

**Max. Marks: 30**

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*Instructions: 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks*

**QI Long Answers (Answer any ONE)**

**1\*10**

- (1) *Non-verbal communication enhances the efficacy of professional communication.* Explain the statement in relation with all non-verbal codes of communication.
- (2) Explain the process of Communication with its elements in detail.

**QII Short Answers (Answer any FOUR)**

**4\*20**

1. What is Communication? Define it in detail.
2. Explain Psychological barrier with all its sub-types and suggest ways to overcome it.
3. What are the types and methods of Communication?
4. State the importance of Body language with its constituents in detail.
5. Explain in detail the factors that affect perspective in day-to-day communication.
6. What are the different communication styles? What do they focus on?



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### **First Sessional Theory Examination 2020-2021**

**Subject: Comm. Skills**

**Day & Date: Monday, 15 March. 2021**

**Class: FY/ DSA**

**Semester: First**

**Time: 10.30 to 12.00**

**Max. Marks: 30**

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*Instructions: 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks*

**QI Long Answers (Answer any ONE)**

**1\*10**

- (1) *Non-verbal communication enhances the efficacy of professional communication.* Explain the statement in relation with all non-verbal codes of communication.
- (2) Explain the process of Communication with its elements in detail.

**QII Short Answers (Answer any FOUR)**

**4\*20**

1. What is Communication? Define it in detail.
2. Explain Psychological barrier with all its sub-types and suggest ways to overcome it.
3. What are the types and methods of Communication?
4. State the importance of Body language with its constituents in detail.
5. Explain in detail the factors that affect perspective in day-to-day communication.
6. What are the different communication styles? What do they focus on?



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### **First Sessional Theory Examination 2021-2022**

**Subject: Communication Skills (BP105T)**

**Day & Date: Friday, 11/2/2022**

**Class: FY/ DSA**

**Semester: I & III**

**Time: 10.30 am to 11.30 am**

**Max. Marks: 30**

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- Instructions: 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks*

**QI) Long Answer Question (Any One) (CO1)**

**1\*10**

1. Explain the process of Communication with its elements in detail.
2. What is communication? Explain communication perspective.

**QII) Short Answer Questions.**

**Section A: (Any One) (CO1)**

**1\*5**

1. Explain Psychological barrier with all its sub-types and suggest ways to overcome it.
2. What are the types and methods of Communication? Illustrate them in detail.

**Section B: (Any Three) (CO2)**

**3\*5**

1. Is Hearing be Listening? Explain how to be effective listener.
2. What are non-verbal codes of communication? Explain them in detail.
3. What is Communication Style Matrix? Explain it in detail.
4. What is listening? What are its types?

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### **First Sessional Theory Examination 2021-2022**

**Subject: Communication Skills (BP105T)**

**Day & Date: Friday, 11/2/2022**

**Class: FY/ DSA**

**Semester: I & III**

**Time: 10.30 am to 11.30 am**

**Max. Marks: 30**

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- Instructions: 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks*

**QI) Long Answer Question (Any One) (CO1)**

**1\*10**

1. Explain the process of Communication with its elements in detail.
2. What is communication? Explain communication perspective.

**QII) Short Answer Questions.**

**Section A: (Any One) (CO1)**

**1\*5**

1. Explain Psychological barrier with all its sub-types and suggest ways to overcome it.
2. What are the types and methods of Communication? Illustrate them in detail.

**Section B: (Any Three) (CO2)**

**3\*5**

1. Is Hearing be Listening? Explain how to be effective listener.
2. What are non-verbal codes of communication? Explain them in detail.
3. What is Communication Style Matrix? Explain it in detail.
4. What is listening? What are its types?

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**First Sessional Theory Examination-(ODD SEM) 2022-2023**

<b>Subject</b> :	Communication skills (BP105T)	<b>Day &amp; Date</b> :	WEDNESDAY, 08/03/2023	
<b>Class</b> :	First Year B. Pharmacy	<b>Semester</b> :	<b>I</b>	<i>Write Your Seat No. Here</i>
<b>Time</b> :	10:30 am – 11:30 am	<b>Max. Marks</b> :	30	<input type="text"/>

<b>Instructions:</b>	1. All questions are compulsory	2. Draw a well labeled diagram wherever necessary
	3. Right hand side number indicates full marks	4. Do not write/tick on the question paper

<b>QI</b>	<b>Long Answers Question (Answer any 1 out of 2)</b>	<b>10 M</b>
1.	Illustrate the process of Communication with its elements in detail. (LL3, CO1)	
2.	Explain the concept of communication along with communication perspectives. (LL3, CO1)	
<b>QII</b>	<b>Short Answers Question</b>	
	<b>Section – A (Answer any ONE)</b>	<b>05 M</b>
1.	Explain the psychological barrier in detail along with your ways to overcome it. (LL2, CO1)	
2.	Interpret the linguistic barrier in detail.	
	<b>Section – A (Answer any Three)</b>	<b>15 M</b>
1.	Explain Verbal and Non-verbal communication in detail. (LL2, CO2)	
2.	Compare any two communication styles. (LL2, CO2)	
3.	Interpret the process of Listening in detail. (LL2, CO2)	
4.	Illustrate the significance of Body Language at workplace. (LL2, CO2)	



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**First Sessional Theory Examination-(ODD SEM) 2023-2024**

<b>Subject</b>	: Communication skills (BP105T)	<b>Day &amp; Date</b>	: FRIDAY,03/11/2023	
<b>Class</b>	: First Year B. Pharmacy	<b>Semester</b>	: I	<b>Write Your Seat No. Here</b>
<b>Time</b>	: 10:30 am – 11:30 am	<b>Max. Marks</b>	: 30	<input type="text"/>

<b>Instructions:</b>	1. All questions are compulsory	2. Draw a well labeled diagram wherever necessary
	3. Right hand side number indicates full marks	4. Do not write/tick on the question paper

<b>QI</b>	<b>Long Answers Question (Answer any 1 out of 2)</b>	<b>10 M</b>
1.	Illustrate the different perspectives in Communication in detail along with examples. (LL3, CO1)	
2.	Explain Verbal and Non-verbal communication in details. (LL3, CO1)	
<b>QII</b>	<b>Short Answers Question</b>	
	<b>Section – A (Answer any ONE)</b>	<b>05 M</b>
1.	Explain the process of communication along with its various elements. Give an example to illustrate. (LL2, CO1)	
2.	Suggest remedies to the barriers to communication	
<b>QIII</b>	<b>Section – A (Answer any Three)</b>	<b>15 M</b>
1.	Explain the barriers of communication. (LL2, CO2)	
2.	Compare any two communication styles. (LL2, CO2)	
3.	Explain Gestalt's theory of Visual perception (LL2, CO2)	
4.	Illustrate the concept of Face- to- Face communication in details. (LL2, CO2)	





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**First Sessional Theory Examination-(ODD SEM) 2023-2024**

<b>Subject</b>	: Communication skills (BP105T)	<b>Day &amp; Date</b>	: FRIDAY,03/11/2023	
<b>Class</b>	: First Year B. Pharmacy	<b>Semester</b>	: I	<b>Write Your Seat No. Here</b>
<b>Time</b>	: 10:30 am – 11:30 am	<b>Max. Marks</b>	: 30	<input type="text"/>

<b>Instructions:</b>	1. All questions are compulsory	2. Draw a well labeled diagram wherever necessary
	3. Right hand side number indicates full marks	4. Do not write/tick on the question paper

<b>QI</b>	<b>Long Answers Question (Answer any 1 out of 2)</b>	<b>10 M</b>
1.	Illustrate the different perspectives in Communication in detail along with examples. (LL3, CO1)	
2.	Explain Verbal and Non-verbal communication in details. (LL3, CO1)	
<b>QII</b>	<b>Short Answers Question</b>	
	<b>Section – A (Answer any ONE)</b>	<b>05 M</b>
1.	Explain the process of communication along with its various elements. Give an example to illustrate. (LL2, CO1)	
2.	Suggest remedies to the barriers to communication	
<b>QIII</b>	<b>Section – A (Answer any Three)</b>	<b>15 M</b>
1.	Explain the barriers of communication. (LL2, CO2)	
2.	Compare any two communication styles. (LL2, CO2)	
3.	Explain Gestalt's theory of Visual perception (LL2, CO2)	
4.	Illustrate the concept of Face- to- Face communication in details. (LL2, CO2)	



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**First Sessional Theory Examination 2019-2020**

**Subject: Comm. Skills (BP105T)**

**Day & Date: Monday, 30 Sept. 2019**

**Class: FY/ DSA**

**Semester: First**

**Time: 10.30 am to 12.00 pm**

**Max. Marks: 30**

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- Instructions: 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks*

**QI Long Answers (Answer any 1 out of 2)**

**10**

- (1) State the importance and objectives of Communication.
- (2) Explain the process of Communication with its elements in detail.

**QII Short Answers (Answer any 4 out of 6)**

**20**

1. What is Communication? Define it in detail.
2. Explain any two barriers that affect the process of Communication.
3. What are the types and methods of Communication?
4. State the importance of Body language with its constituents in detail.
5. What is paralanguage? State its importance in the efficacy of Communication.
6. Write a note on Graphical Language.

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**First Sessional Theory Examination 2018-2019**

**Subject: Human Anatomy and Physiology-I**  
**Class: F. Y. B. Pharma**  
**Time: 10:30-12:00**

**Day & Date: Monday, 1.10.2018**  
**Semester: I**  
**Max. Marks: 30**

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**QI Multiple Choice Questions (MCQs)**

**10**

1. Examples of slightly moveable joints is joints  
A. between the vertebrae B. between the skull bones  
C. between ear ossicles D. between intervertebral discs
2. Bands of connective tissues that are made of collagen are called  
A. osteocytes B. chondrocytes C. ligaments and tendons D. cartilages
3. Joints that move like a hinge on door or back and forth are called  
A. slightly moveable joints B. hinge joints C. ball-and-socket joints D. fixed joint
4. Example of ball-and-socket joint is  
A. vertebral disc joint B. shoulder and hip joints  
C. knee and elbow joints D. spinal cord joint
5. The power house of cell is called  
A) Cell wall B) Mitochondria C) Ribosomes D) Nucleus
6. The functional unit of life is called  
A) Cell B) Egg C) Nucleus D) None of these
7. Tissue is a  
A) Group of organs B) Group of cells C) Group of tissues D) Group of organisms
- 8 The control unit of cell is  
A) Nucleus B) Cell wall C) Cytoplasm D) All of these
9. The part of the bone where blood cells are made.  
A. Marrow B. Joint C. Ligament D. none of the above
10. The ribs, sternum and spine protect these.  
A. Kidneys, bladder, urethra B. Heart, lungs, blood vessel  
C. Small intestine, large intestine D. all the above

**QII Long Answers (Answer any 1 out of 2)**

**10**

1. Draw a neat diagram of skin. Explain in detail anatomy and functions of skin
2. Define transport. Explain in details various mechanism of transport

**QIII Short Answers (Answer any 2 out of 3)**

**10**

1. Write a note of cell with a neat labelled diagram
2. Define homeostasis explain in detail mechanism of homeostasis with suitable illustration.
3. Define joints. Explain in detail synovial joint.



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**First Sessional Theory Examination 2019-2020**

**Subject:** Human Anatomy and Physiology – I (BP101T)

**Day & Date:** Monday, 26.9.19

**Class:** F.Y. B. Pharmacy

**Semester:** I

**Time:** 10.30 am - 12.00 pm

**Max. Marks:** 30

- Instructions:*
1. All questions are compulsory
  2. Draw a well labeled diagram wherever necessary
  3. Right hand side number indicates full marks

**QI Multiple Choice Questions (MCQs)**

10

1. The spherical structured organelle that contains the genetic material is  
A Cell walls      B Ribosomes      C Nucleus      D Mitochondria
2. Which one of the following organelles digests the old organelles that are no longer useful to the cells?  
A Ribosomes      B Mitochondria      C Lysosomes      D Chromatin
3. The Tissue that lines and covers the body is \_\_\_\_\_.  
A. Epithelial      B. Connective      C. Nervous      D. Muscle
4. Blood is classified as a(n) \_\_\_\_\_ tissue.  
A. Nervous      B. Epithelial      C. Muscle      D. Connective
5. Contractility is a property of \_\_\_\_\_ tissue.  
A. Muscle      B. Nervous      C. Epithelial      D. Connective
6. The tissue that is thin enough to allow diffusion is \_\_\_\_\_.  
A. Simple cuboidal epithelium      B. Simple squamous epithelium  
C. Areolar connective tissue      D. Hyaline cartilage
7. The intercellular material surrounding connective tissue cells is \_\_\_\_\_.  
A. Ground substance      B. Matrix      C. Adipose      D. Bone
8. The cell responsible to the formation of cartilage is the \_\_\_\_\_.  
A. Fibroblast      B. Chondroblast      C. Osteoblast      D. Megablast
9. Which cells produce the pigment that contributes to hair color?  
A) Keratinocytes      B) Melanocytes      C) Langerhans cells      D) Merkel cells
10. Example of flat bone  
A) Sternum      B) Clavicle      C) A and B      D) None of the above

**QII Long Answers (Answer any 1 out of 2)**

10

1. Explain in details anatomy and functions of skin with a neat labelled diagram.
2. Define homeostasis. Write in details components and mechanism of homeostasis with suitable examples

**QIII Short Answers (Answer any 2 out of 3)**

10

1. Draw a neat labelled diagram of cell.
2. Define transport, explain any two mechanism of transport across cell membrane.
3. Define skeleton. Write the various bones of axial and appendicular skeleton

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### First Sessional Theory Examination 2020-2021 (Odd SEM)

Subject & Subject Code: Human Anatomy & Physiology I (BP101T)

Day & Date: Monday, 07.02.2022 Class: First Year B. Pharm

Semester: I Time: 10.30 AM to 12.30PM Max. Marks: 30

Instructions: 1. All questions are compulsory.

2. Draw a well labelled diagram wherever necessary.

3. Right hand side number indicates full marks.

#### Q.I Multiple-choice questions (MCQs)

10

- Examples of positive feedback mechanism EXCEPT
  - Blood clotting
  - Childbirth
  - Nerve impulse generation
  - CO<sub>2</sub> regulation
- A sperm cell tail is an example of ---
  - Cilia
  - Flagellac.
  - Microvili
  - Filaments
- Cells communicate over relatively short distances is known as
  - Autocrine signalling
  - Paracrine signalling
  - Synaptic signalling
  - Endocrine signalling
- Spot weld like junctions of epidermis is an example of
  - Tight junctions
  - Adherence junction
  - Desmosomes
  - Hemidesmosomes
- Following is an example of branched tubular glandular epithelium tissue
  - Large intestine
  - Sebaceous gland
  - Sweat gland
  - Gastric gland
- Which of the following epidermal layer is not found in thin skin
  - Stratum corneum
  - Stratum lucidum
  - Stratum spinosum
  - Stratum granulosum
- Bone tissue stores about --- % of the body's ----.
  - 85 % & phosphorous
  - 85% & calcium
  - 99% & calcium
  - 99% & phosphorous
- With increasing age, much of the bone marrow changes from ---
  - Red to yellow
  - Yellow to red
  - Red to blue
  - Yellow to blue
- The human body contains --- pectoral or shoulder girdles
  - One
  - Two
  - Three
  - Four
- Following is an example of non-nucleated cell
  - RBCs
  - WBCs
  - Monocytes
  - Lymphocytes

#### QII Long Answers (Answer any 1 out of 2)

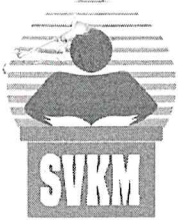
10

- Draw a neat and well labelled diagram of skin. Explain the structure and functions of skin.
- Explain the functions of skeletal system. Explain the different types of bones.

#### QIII Short Answers (Answer any 2 out of 3)

10

- Define and classify tissue with suitable examples. Describe the location, structure and functions of squamous epithelial tissue.
- Draw a neat well labelled diagram of cell. Explain the structure and function of mitochondria.
- Explain the levels of structural organization of human body.



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Phone: 02562- 297802, 297805 Fax: 02562- 287802, Email: examiop@svkm.ac.in

### First Sessional Theory Examination 2020-2021

(Odd SEM)

Subject & Subject Code: Human Anatomy & Physiology I (BP101T)

Day & Date: Wednesday, 10.03.2021

Class: First Year B. Pharm

Semester: I

Time: 10.30 AM to 12.30PM

Max. Marks: 30

Instructions: 1. All questions are compulsory.

2. Draw a well labelled diagram wherever necessary.

3. Right hand side number indicates full marks.

#### Q.I Multiple-choice questions (MCQs)

10

1. Integumentary system contains skin and associated structures EXCEPT

- a. Hair      b. Sweat glands      c. Bones      d. Oil glands

2. Following is an example of negative feedback mechanism

- a. Regulation of carbon dioxide conc.      b. Blood Clotting  
c. Childbirth      d. Generation of Nerve Signals

3. Epithelial tissues that line the stomach, intestines, and urinary bladder contains

- a. Adherence Junctions      b. Tight Junctions  
c. Gap Junctions      d. Desmosomes

4. Somatic cells contain two sets of chromosomes, they are symbolized as

- a. n      b. 2n      c. 4n      d. 3n

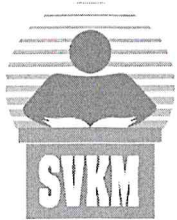
5. Following statements are correct for paracrine signaling

- a. Communicate through release of chemical signals over short distances  
b. Communicate through release of ligand that binds to its own receptors  
c. Communicate through nerve cells over short distances  
d. Transmits signals over long distances through release of hormones

6. Entire outside surface of the cell often has a loose carbohydrate coat called as

- a. Phospholipids      b. Glycocalyx      c. Glycolipids      d. Cholesterol

*Sherikar*  
Mr. Sherikar A.K.  
(Assistant Teacher)



Shri Vile Parle Kelavani Mandal's

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---

7. Axial skeleton have ---- number of bones  
a. 126                      b. 80                      c. 100                      d. 110
8. Only bone cells to undergo cell division  
a. Osteoclasts    b. Osteocytes    c. Osteoblasts    d. Osteogenic Cells
9. Is the longest and largest bone of the upper limb  
a. Femur    b. Radius & Ulna    c. Tibia & Fibula    d. Humerus
10. Is the largest organ of the body in both surface area and weight  
a. Skin                      b. Liver                      c. Lungs                      d. Heart

### Q.II Long Answers (Answer any 1 out of 2)

10

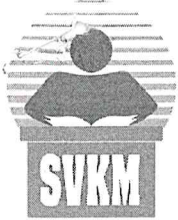
1. Define and classify tissue. Explain the structure, location and functions of any three types of epithelial tissue.
2. Explain the functions of skeletal system. Write in detail about bones of axial skeleton.

### Q.III Short Answers (Answers any 2 out of 3)

10

1. Draw a neat well labelled diagram of skin. Explain the functions of skin.
2. Draw a neat well labelled diagram of cell. Explain the structure and functions of mitochondria.
3. Classify and explain the different types of bones.

*Dhule*  
Mr. Sherikar A-K.  
(Subject In-charge)



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(Odd SEM)

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- a. Hair      b. Sweat glands      c. Bones      d. Oil glands

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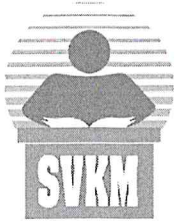
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*Sherikar*  
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8. Only bone cells to undergo cell division  
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9. Is the longest and largest bone of the upper limb  
a. Femur    b. Radius & Ulna    c. Tibia & Fibula    d. Humerus
10. Is the largest organ of the body in both surface area and weight  
a. Skin                      b. Liver                      c. Lungs                      d. Heart

### Q.II Long Answers (Answer any 1 out of 2)

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1. Define and classify tissue. Explain the structure, location and functions of any three types of epithelial tissue.
2. Explain the functions of skeletal system. Write in detail about bones of axial skeleton.

### Q.III Short Answers (Answers any 2 out of 3)

10

1. Draw a neat well labelled diagram of skin. Explain the functions of skin.
2. Draw a neat well labelled diagram of cell. Explain the structure and functions of mitochondria.
3. Classify and explain the different types of bones.

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**First Sessional Theory Examination - (Odd SEM) 2023-2024**

<b>Subject</b> : Human Anatomy & Physiology - I (BP101T)	<b>Day &amp; Date</b> : Monday, 30/10/2023
<b>Class</b> : First Year B. Pharmacy	<b>Semester</b> : I
<b>Time</b> : 10.30 am to 12.00 pm	<b>Max. Marks</b> : 30
	<b>Write Your Seat No. Here</b> <input type="text"/>

**Instructions:** 1. All questions are compulsory  
2. Draw a well-labeled diagram wherever necessary  
3. Right-hand side number indicates full marks  
4. Do not write/tick on the question paper

<b>Q. I: Objective Type Questions (5 X 2)</b>	<b>10 M</b>
1. Enlist the functions of muscular tissue (CO1, LL2)	
2. Write the functions of Golgi Complex (CO1, LL2)	
3. Define the term a. Passive diffusion (CO1, LL2) b. Anemia (CO2, LL2)	
4. Enlist the functions of blood (CO2, LL2)	
5. Draw a neat and well-labelled diagram of skin (CO2, LL2)	

<b>Q. II: Long Answers Questions (Answer any 1 out of 2)</b>	<b>10 M</b>
1. Draw a neat and well-labelled diagram of plasma membrane. Explain in detail about structure and functions of plasma membrane (CO1, LL2)	
2. Define and classify tissue. Explain the structure and functions of epithelial tissue (CO1, LL2)	

<b>Q. III: Short Answers Questions (Answer any 2 out of 3)</b>	<b>10 M</b>
1. Enlist the bones of appendicular system. Write in detail about structure and functions of pelvic girdle. (CO2, LL2)	
2. Define and classify joints. Explain the structure and functions of synovial joints. (CO2, LL2)	
3. Explain the ABO system of blood group. (CO2, LL2)	

**Prepared by**  
Mr. Abdulla Sherikar

**Verified by**  
Mr. Purvaj Barote

**Approved by**  
Dr. Kartik Nakhate



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**First Sessional Theory Examination - (Odd SEM) 2023-2024**

<b>Subject</b> : Human Anatomy & Physiology - I (BP101T)	<b>Day &amp; Date</b> : Monday, 30/10/2023
<b>Class</b> : First Year B. Pharmacy	<b>Semester</b> : I
<b>Time</b> : 10.30 am to 12.00 pm	<b>Max. Marks</b> : 30
<b>Write Your Seat No. Here</b> <input type="text"/>	

**Instructions:** 1. All questions are compulsory  
2. Draw a well-labeled diagram wherever necessary  
3. Right-hand side number indicates full marks  
4. Do not write/tick on the question paper

<b>Q. I: Objective Type Questions (5 X 2)</b>	<b>10 M</b>
1. Enlist the functions of muscular tissue (CO1, LL2)	
2. Write the functions of Golgi Complex (CO1, LL2)	
3. Define the term	
a. Passive diffusion (CO1, LL2)	
b. Anemia (CO2, LL2)	
4. Enlist the functions of blood (CO2, LL2)	
5. Draw a neat and well-labelled diagram of skin (CO2, LL2)	

<b>Q. II: Long Answers Questions (Answer any 1 out of 2)</b>	<b>10 M</b>
1. Draw a neat and well-labelled diagram of plasma membrane. Explain in detail about structure and functions of plasma membrane (CO1, LL2)	
2. Define and classify tissue. Explain the structure and functions of epithelial tissue (CO1, LL2)	

<b>Q. III: Short Answers Questions (Answer any 2 out of 3)</b>	<b>10 M</b>
1. Enlist the bones of appendicular system. Write in detail about structure and functions of pelvic girdle. (CO2, LL2)	
2. Define and classify joints. Explain the structure and functions of synovial joints. (CO2, LL2)	
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Shri. Vile Parle Kelavani Mandal's  
**Institute of Pharmacy, Dhule**

Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

**First Sessional Theory Examination 2019-2020**

Subject	: Pharmaceutical Analysis	Day & Date	: 01/10/2019
Class	: FY/DSY B. Pharmacy	Semester	: I/III*
Time	: 10:30 - 12:00 noon	Max. Marks	: 30

Instructions: 1. All questions are compulsory  
2. Draw a well labelled diagram wherever necessary  
3. Right hand side number indicates full marks

**Q.1 Multiple Choice Questions [10]**

- According to Lewis, bases are \_\_\_\_\_
  - Electron pair donor
  - Electron pair acceptor
  - Proton donor
  - Proton acceptor
- Amphiprotic Solvents are both .....and .....Characters
  - Aprotic, Protophillic
  - Protophillic, Protogenic
  - Protogenic, Aprotic
  - None of the above
- .....is most suitable method to analysed Weak acid and Weak base
  - Aqueous titration
  - Non-aqueous titration
  - Precipitation titration
  - Complexometric titration
- Molarity [M ]is
  - 1 mole / lit of solution
  - 1 g / lit. of solution
  - M.W. of solute in g / lit. of solution
  - 1 g / lit. of solution
- Argentometric titrations mainly utilizes \_\_\_\_\_ as a titrant.
  - 0.02M Disodium EDTA
  - 0.1 M sodium hydroxide
  - 0.1 N Silver chloride
  - 0.1 N perchloric acid
- Precision is expressed by
  - Correctness of measurement
  - Standard value
  - Reproducibility of measurement
  - True value
- \_\_\_\_\_ is a compound of sufficient purity from which a standard solution can be prepared by direct weighing of a quantity of it, followed by dilution to give a defined volume of solution.
  - External standard
  - Primary standard
  - Secondary standard
  - a & c both
- \_\_\_\_\_ are the methods of minimisation of error
  - Amplification methods
  - Standard addition
  - Internal standards
  - All of the above
- Dioxane is a \_\_\_\_\_ solvent
  - Protophillic
  - Protogenic
  - Amphiprotic
  - Aprotic
- Crystal violet is \_\_\_\_\_ indicator
  - Acidic
  - Basic
  - Neutral
  - Acidic & Basic both

**Q.2. Answer the following short questions (any two) [10]**

- What are primary and secondary standards? Explain their types with suitable example
- Write short note on Mohr's method
- Define non aqueous titrations? Explain the types of solvent used.

**Q.3. Answer in detail of following (any one) [10]**

- Explain the various titration curves with suitable example
- Explain types of error with suitable example. Describe the various methods used for minimisation of errors

\*\*\*\*\*



**Shri. Vile Parle Kelvani Mandal's  
Institute of Pharmacy, Dhule**

Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

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**First Sessional Theory Examination 2020-2021**

**Subject:** Pharmaceutical Analysis I (BP102T)

**Day & Date:** 11-03-2021

**Class:** First Year B. Pharm.

**Semester:** I

**Time:** 10:30 am – 12:00 pm

**Max. Marks:** 30

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- Instructions:*
1. All questions are compulsory
  2. Draw a well labeled diagram wherever necessary
  3. Right hand side number indicates full marks

**Q. I. Solve the following Multiple Choice Questions (MCQs)**

**10**

1. Number of gram equivalents of solute present in one litre of a solution is called as .....
  - a) Molarity
  - b) Normality
  - c) Molality
  - d) Percentage weight by volume
2. The features of Primary standard substance includes:
  - a) High purity
  - b) High stability
  - c) Low hygroscopicity
  - d) All of the above
3. Indian pharmacopoeia is published by .....
  - a. World health organization
  - b. Ministry of health and family welfare
  - c. Ministry of science and technology
  - d. Food and drug administration
4. Limit test of iron is based upon the reactions between iron and ....
  - a. Dithiazone
  - b. Thioglycolic acid
  - c. Hydrogen sulphide
  - d. Barium chloride
5. The color of phenolphthalein in alkaline medium is .....
  - a) Pink
  - b) Yellow
  - c) Brown
  - d) Colorless

6. Benzene is an example of .....solvent.
- Aprotic
  - Protogenic
  - Protophilic
  - Amphiprotic
7. Solubility of precipitate is ..... on addition of common ion.
- Decreased
  - Increased
  - Remains Constant
  - None of These
8. In Mohr's method .....is used as an indicator.
- Potassium chromate
  - Ferric ammonium sulphate
  - Ammonium thiocyanate
  - Silver nitrate
9. In Volhard's method .....is used as a titrant
- Potassium chromate
  - Ferric ammonium sulphate
  - Ammonium thiocyanate
  - Silver nitrate
10. The end-point detection in Fajan's method involves color change due to.....
- Formation of colored precipitate
  - Formation of soluble colored complex
  - Adsorption of indicator on the surface of precipitate
  - None of the above

**Q. II. Solve the following (any 1)**

**10**

1. Explain the neutralization curves in acid-base titrations for i) strong acid vs strong base ii) weak acid vs strong base iii) strong acid vs weak base iv) weak acid vs weak base.
2. Explain in detail Mohr's method and Fajan's method.

**Q. III. Solve the following (any 2)**

**10**

1. Explain principle involved in limit test for Chloride and sulphate.
2. Explain in detail the theories of acid-base indicators.
3. Explain the types of solvents used in non-aqueous titrations with suitable examples.



**First Sessional Theory Examination 2021-2022 (Odd SEM)**

**Subject:** Pharmaceutical Analysis-I (BP102-T)

**Day & Date:** Tuesday 08/02/2022

**Class:** F.Y. B. Pharm & D.S.Y.B.Pharm

**Semester:** I & III

**Time:** 10.30 am- 12.00 pm

**Max. Marks:** 30

- Instructions:* 1. All questions are compulsory  
2. Draw a well labelled diagram wherever necessary  
3. Right hand side number indicates full marks

**QI Multiple Choice Questions (MCQs)**

**10**

- Potassium chromate is used as indicator in -----  
a) Fajan's method      b) Gay Lussacs method      c) Mohrs method      d) Volhards method
- error cannot be minimized or identified  
a) Random error      b) Determinate error  
c) Indeterminate error      d) All of the above
- ..... reagent is used for limit test for chloride  
a) BaCl<sub>2</sub>      b) AgNO<sub>3</sub>      c) Methyl Orange      d) EDTA
- Measurement which is close to true value is known as.....  
a) Accuracy      b) Precision      c) Error      d) Average
- A measurement which on repetition gives same or nearly same results are known as.....  
a) Accuracy      b) Precision      c) Error      d) Average
- The acid which dissociate completely in aqueous media is known as .....  
a) Weak acid      b) Strong acid      c) Strong base      d) Weak base
- The number of gm Equivalents to solute in 1 litre of solvent is known as ....  
a) Normality      b) Molarity      c) Molality      d) Formality
- ..... is an example of protophilic solvents?  
a) HCl      b) Acetic acid      c) KOH      d) Methanol
- Crystal Violet is used as indicator in ..... titration  
a) Aqueous      b) non-aqueous      c) Both      d) None of the above
- A substance which can donate proton in aqueous medium is known as .....  
a) Acid      b) base      c) Neutral      d) All of the above

**QII Long Answers Question (Answer any 1 out of 2)**

**10**

- What do you mean by Errors? Explain the types and sources of error in details?
- Define precipitation titrations? Explain Mohr's and Volhard's method?

**QIII Short Answers Question (Answer any 2 out of 3)**

**10**

- Discuss various types of solvents used in non-aqueous titration?
- Write down the preparation and standardization of 0.1 N NaOH as per IP?
- Write down the neutralization curve for SA v/s SB?
  - Enlist the sources of impurities in medicinal agents?





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Mohd Usman

Azim Ansari

Girija Bhavar

**Prepared By**

**Verified By**

**Approved By**

**Mapping of Course Outcome with Second Sessional Theory Examination (2021-2022)**

**Subject: Pharmaceutical Analysis-I**

Course Outcome	Question	Marks
CO 1021:	Q1 – 2,3,4,5,7	22.5
	Q2 – 1	
	Q3- 2, 3-ii	
CO 1022:	Q1 – 1,6,8,9,10	22.5
	Q2 -2	
	Q3 – 1, 3-i	

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**First Sessional Theory Examination 2021-2022 (Odd SEM)**

**Subject:** Pharmaceutical Analysis-I (BP102-T)

**Day & Date:** Tuesday 08/02/2022

**Class:** F.Y. B. Pharm & D.S.Y.B.Pharm

**Semester:** I & III

**Time:** 10.30 am- 12.00 pm

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- error cannot be minimized or identified  
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Mohd Usman

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**Approved By**

**Mapping of Course Outcome with Second Sessional Theory Examination (2021-2022)**

**Subject: Pharmaceutical Analysis-I**

Course Outcome	Question	Marks
CO 1021:	Q1 – 2,3,4,5,7	22.5
	Q2 – 1	
	Q3- 2, 3-ii	
CO 1022:	Q1 – 1,6,8,9,10	22.5
	Q2 -2	
	Q3 – 1, 3-i	

Mohd Usman

Azim Ansari

Girija Bhavar

**Prepared By**

**Verified By**

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**FIRST SESSIONAL THEORY EXAMINATION-(ODD SEM) 2022-2023**

<b>Subject</b> :	Pharmaceutical Analysis (BP102T)	<b>Day &amp; Date</b> :	Sat, 11/03/2023
<b>Class</b> :	First Year B. Pharmacy	<b>Semester</b> :	I
<b>Time</b> :	10:30 am – 12:00 pm	<b>Max. Marks</b> :	30
			<b>Write Your Seat No. Here</b> <input type="text"/>

**Instructions:**

1. All questions are compulsory
2. Draw a well-labeled diagram wherever necessary
3. Right-hand side number indicates full marks
4. Do not write/tick on the question paper

**QI Objective Type Questions**

**10 M**

1. Define: Accuracy and Precision
2. How to prepare 0.1 N Sodium hydroxide solution
3. Write a reaction for weak acid and strong base titration
4. Enlist type of solvents in Non-aqueous titration with one example each.
5. Define acidimetry and limit test

**QII Long Answers Question (Answer any 1 out of 2)**

**10 M**

1. Define Pharmaceutical analysis and describe its scope. Enlist techniques of analysis and describe any three of them.
2. What are the sources of errors and describe methods of minimizing errors.

**QIII Short Answers Question (Answer any 2 out of 3)**

**10 M**

1. Enlist the theories of acid-base indicators and explain any one.
2. Describe Volhard's methods for precipitation titration.
3. Explain masking & de-masking reagents. Give examples of metal ion indicators



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**FIRST SESSIONAL THEORY EXAMINATION-(ODD SEM) 2022-2023**

<b>Subject</b> :	Pharmaceutical Analysis (BP102T)	<b>Day &amp; Date</b> :	Sat, 11/03/2023	
<b>Class</b> :	First Year B. Pharmacy	<b>Semester</b> :	I	<b>Write Your Seat No. Here</b> <input type="text"/>
<b>Time</b> :	10:30 am – 12:00 pm	<b>Max. Marks</b> :	30	

**Instructions:**

1. All questions are compulsory
2. Draw a well-labeled diagram wherever necessary
3. Right-hand side number indicates full marks
4. Do not write/tick on the question paper

**QI Objective Type Questions**

**10 M**

1. Define: Accuracy and Precision
2. How to prepare 0.1 N Sodium hydroxide solution
3. Write a reaction for weak acid and strong base titration
4. Enlist type of solvents in Non-aqueous titration with one example each.
5. Define acidimetry and limit test

**QII Long Answers Question (Answer any 1 out of 2)**

**10 M**

1. Define Pharmaceutical analysis and describe its scope. Enlist techniques of analysis and describe any three of them.
2. What are the sources of errors and describe methods of minimizing errors.

**QIII Short Answers Question (Answer any 2 out of 3)**

**10 M**

1. Enlist the theories of acid-base indicators and explain any one.
2. Describe Volhard's methods for precipitation titration.
3. Explain masking & de-masking reagents. Give examples of metal ion indicators



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**First Sessional Theory Examination-(ODD SEM) 2023-2024**

<b>Subject</b> :	Pharmaceutical Analysis (BP102T)	<b>Day &amp; Date</b> :	Tue, 31/10/2023	
<b>Class</b> :	First Year B. Pharmacy	<b>Semester</b> :	I	<b>Write Your Seat No. Here</b> <input type="text"/>
<b>Time</b> :	10:30 am – 12:00 pm	<b>Max. Marks</b> :	30	

**Instructions:** 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks  
4. Do not write/tick on the question paper

<b>Q. 1</b>	<b>Objective Type Questions</b>	<b>10 M</b>
1.	Define: i) Accuracy; ii) Primary Standard (CO1, LL1)	
2.	What is Impurity? Enlist their types (CO1, LL3)	
3.	Give any two difference between Acidimetric and Alkalimetry type of titration (CO2, LL1)	
4.	Enlist any four indicators used in non-aqueous titration. (CO2, LL1)	
5.	Write the principle and reaction for limit test for chloride (CO1, LL1)	

<b>Q 2</b>	<b>Long Answers Question (Any 1 x 10 Marks)</b>	<b>10 M</b>
1.	Elaborate on Acids and Bases with respect to different theories. Discuss different neutralization curves with suitable examples. (CO2, LL4)	
2.	What are <i>errors</i> ? Discuss their <i>sources</i> and <i>methods of minimizing errors</i> with suitable examples. Comment on relationship between <i>accuracy</i> and <i>precision</i> . (CO1, LL3)	

<b>Q 3</b>	<b>Short Answer Question (Any 2 x 05= 10 M)</b>	<b>10 M</b>
1.	Write a note on Acid-Base indicator theories. (CO2, LL1)	
2.	Describe Volhard's methods in precipitation titration. (CO2, LL2)	
3.	Discuss different Pharmacopoeia and its types (CO1, LL2)	

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<b>Subject</b> :	Pharmaceutical Analysis (BP102T)	<b>Day &amp; Date</b> :	Tue, 31/10/2023
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<b>Time</b> :	10:30 am – 12:00 pm	<b>Max. Marks</b> :	30
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<b>Q 2</b>	<b>Long Answers Question (Any 1 x 10 Marks)</b>	<b>10 M</b>
1.	Elaborate on Acids and Bases with respect to different theories. Discuss different neutralization curves with suitable examples. (CO2, LL4)	
2.	What are <i>errors</i> ? Discuss their <i>sources</i> and <i>methods of minimizing errors</i> with suitable examples. Comment on relationship between <i>accuracy</i> and <i>precision</i> . (CO1, LL3)	

<b>Q 3</b>	<b>Short Answer Question (Any 2 x 05= 10 M)</b>	<b>10 M</b>
1.	Write a note on Acid-Base indicator theories. (CO2, LL1)	
2.	Describe Volhard's methods in precipitation titration. (CO2, LL2)	
3.	Discuss different Pharmacopoeia and its types (CO1, LL2)	

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**First Sessional Theory Examination 2020-2021**

**Subject: Pharmaceutics I BP103T      Day & Date: Friday, 12<sup>th</sup> March 2021**

**Class: F Y B. Pharm                      Semester: I**

**Time: 90 min                                Max. Marks: 30**

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*Instructions: 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks*

**Q.No.1 Multiple choice questions (10 X 1=10)**

1. .... is known as 'Father of Pharmacy education in India'.  
a) M. L. Shroff, b) R. N. Chopra, c) B. N. Ghosh, d) None of the above
2. The fourth edition of IP was published in ....  
a) 1985 b) 1996 c) 1960 d) 19963.
3. Pharmacy Council of India (PCI) was established in ....  
a) 1947 b) 1948 c) 1949 d) 1950.
4. Which of the following excipient act as preservative.....  
a) Sorbitan monolactrate b) Tyloxopol c) Benzethonium chloride d) Thiomersal
5. When two or more drugs are used in combination to increase the pharmacological action, the phenomenon is known as .....  
a) Synergism b) Tolerance c) Potentiation c) Idiosyncrasy
6. Nasal drops should have pH .....  
a) 5-6 b) 6-7 c) 7-8 d) 8-9
7. Which of the following is natural colorant .....  
a) Amaranth b) Erythrosine c) Carotenoids d) Eosin
8. The dosage form intended to be rubbed on the skin for rubeficient and counterirritant effect are called  
a) Lotions b) Liniments c) Gels d) Dentifrices
9. 1 $\bar{3}$  (apothecaries) =  
a) 437.5 grains b) 480 grains c) 7000 grains d) 5760 grains
10. Daily adult dose of drug is 120 mg. How much of it should be given to a child of 11 years?  
a) 57.39, b) 60, c) 50.39, d) 37



**Q. No. 2 Long Answers (Answer 1 out of 2)**

**(1 x 10 = 10)**

1. Define dosage form. Explain need of dosage form. Describe various solid dosage forms.
2. Define posology. Write a note on factors affecting posology

**Q. No. 3 Short Answers (Answer 2 out of 3)**

**(2 x 5 = 10)**

1. Write down parts of prescription
2. Discuss various excipient used in the formulation of liquid dosage forms.
3. Write short note on (any one)
  - a) Syrups
  - b) Topical oral preparations



## Shri. Vile Parle Kelavani Mandal's Institute of Pharmacy, Dhule

Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

### First Sessional Theory Examination 2019-2020

**Subject: Pharmaceutics I BP103T**

**Day & Date: Friday, 27<sup>th</sup> Sept 2019**

**Class: F Y B. Pharm**

**Semester: I**

**Time: 10.30 pm to 12.00 pm**

**Max. Marks: 30**

- Instructions:*
1. All questions are compulsory
  2. Draw a well labeled diagram wherever necessary
  3. Right hand side number indicates full marks

**Q.No.1 Multiple choice questions**

**(10 X 1=10)**

1. The fifth edition of Indian Pharmacopoeia was published in  
 a) 2007                      b) 1996                      c) 1985                      d) 1966
2. Which part of prescription contains instructions for the patients?  
 a) Superscription,                      b) Subscription,                      c) Inscription,                       d) Signa
3. Which of the following is antifoaming agent?  
 a) Simethicone,                      b) Polysorbate,                      c) Methylcellulose,                      d) saccharin
4. .... is commonly used viscous base in throat paint.  
 a) Glycerin,                      b) Water,                      c) Alcohol,                      d) Hydro-alcoholic mixture.
5. A drug that neutralizes excess gastric acid are .....  
 a) Antacid,                      b) Analgesic,                      c) Antiamebic,                      d) Antiepileptic
6. Daily adult dose of drug is 120mg. How much of it should be given to a child of 11 years?  
 a) 57.39,                      b) 60,                      c) 50.39,                      d) 37
7. 60 gr = .....  
 a) 1 ℥,                      b) 2 ℥,                       c) 3 ℥,                      d) 4 ℥
8. The semisolid unit dosage form intended for vaginal administration are called  
 a) Suppositories                       b) Pessaries                      c) Bougies                      d) Cones
9. How will you mix potent drug with large amount of diluent?  
 a) Spatulation                       b) Geometric dilution                      c) Sifting                      d) Tumbling.
10. Which of the following dosage form is not recommended for ocular administration  
 a) Lotions,                       b) Liniments,                      c) Suspensions,                      d) Solutions

**Q. No. 2 Long Answers (Answer 1 out of 2)**

**(1 x 10 = 10)**

1. Define dosage form. Explain need of dosage form. Describe various liquid dosage forms.
2. Define prescription. Write down parts of prescription and methods of handling of prescription.

**Q. No. 3 Short Answers (Answer 2 out of 3)**

**(2 x 5 = 10)**

1. Write a note on factors affecting posology.
2. Discuss methods of dispensing of powders involving special problems.
3. Write short note on (any one)
  - a) Write down silent features of Indian Pharmacopoeia.
  - b) Differentiate between: 1) Lotions and liniments. 2) Syrups and elixirs.



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**First Sessional Theory Examination-(ODD SEM) 2022-2023**

<b>Subject</b> : Pharmaceutics-I (BP103T)	<b>Day &amp; Date</b> : Monday, 13/03/2023
<b>Class</b> : First Year B. Pharmacy	<b>Semester</b> : I
<b>Time</b> : 10.30 am to 12.00 noon	<b>Max. Marks</b> : 30
	<b>Write Your Seat No. Here</b> <input type="text"/>

**Instructions:**

1. All questions are compulsory
2. Draw a well-labeled diagram wherever necessary
3. Right-hand side number indicates full marks
4. Do not write/tick on the question paper

**Q. I: Objective Type Questions (5 X 2)**

**10 M**

- 1 Define following:
  - a) Avoirdupois system
  - b) Apothecaries system
- 2 Define and Classify the dosage form.
- 3 Define the term Posology. Enlist the various factors affecting the dose of the drug.
- 4 Define solid dosage form. Gives advantages and disadvantages of solid dosage forms.
- 5 What is the dose for a 9-month-old infant if the average adult dose of a drug is 500 mg? (By using Fried's formula)

**Q. II: Long Answers Questions (Answer any 1 out of 2)**

**10 M**

- 1 Define prescription. Explain the parts of the prescription and methods of handling the prescription.
- 2 Discuss the historical development of the pharmacy profession in India considering education and industry.

**Q. III: Short Answers Questions (Answer any 2 out of 3)**

**10 M**

- 1 Define and classify powders. Discuss various methods for mixing powders.
- 2 Write a note on various excipients used in the formulation of liquid dosage forms.
- 3 Explain in brief about solubility enhancement techniques.



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## First Sessional Theory Examination-(ODD SEM) 2023-2024

<b>Subject</b> :	Pharmaceutics-I (BP103T)	<b>Day &amp; Date</b> :	Wednesday, 01/11/2023	
<b>Class</b> :	First Year B.Pharm	<b>Semester</b> :	I	<b>Write Your Seat No. Here</b>
<b>Time</b> :	10.30 am to 12.00 pm	<b>Max. Marks</b> :	30	<input type="text"/>

<b>Instructions:</b>	1. All questions are compulsory	2. Draw a well labeled diagram wherever necessary
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<b>QI</b>	<b>Objective Type Questions (5 x 2) = 5 x 2 = 10 (Answer all the questions)</b>	<b>10 M</b>
1.	Give salient features of first edition of Indian pharmacopoeia (LL1, CO1)	
2.	Give formula for Pediatric dose calculations based on age (LL1, CO1)	
3.	Define a. Pharmacopoeia      b. Mouthwashes (LL1, CO1)	
4.	Convert the following imperial values to metric system (LL1, CO2) a. 1 fluid ounce = _____ ml      b. 1 Pound (Avoir) = _____ gm.	
5.	What are effervescent granules (LL1, CO2)	
<b>QII</b>	<b>Long Answers Question (Answer any 1 out of 2)</b>	<b>10 M</b>
1.	Classify dosage forms and describe all with examples (LL1, CO2)	
2.	What is Prescription? Explain in detail Parts of Prescription and Errors in Prescription. (LL1, CO2)	
<b>Q.III</b>	<b>Short Answers Question (Answer any 2 out of 3) -</b>	<b>10M</b>
1.	Definition and classify Powders. Write a short note on Dusting powder. (LL2, CO1)	
2.	In what proportion should alcohol 95% and 65% strengths be mixed to make 200 ml of 80% alcohol? (LL2, CO1)	
3.	What are liquid dosage form? Give classification, advantages and disadvantages of liquid dosage form. (LL2, CO1)	



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3.	What are liquid dosage form? Give classification, advantages and disadvantages of liquid dosage form. (LL2, CO1)	

**First Sessional Theory Examination 2018-2019**

**Subject: Pharmaceutics I BP103T**

**Day & Date: Thu, 4<sup>th</sup> Oct 2018**

**Class: F Y B. Pharm**

**Semester: I**

**Time: 90 min**

**Max. Marks: 30**

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- Instructions:*
1. All questions are compulsory
  2. Draw a well labeled diagram wherever necessary
  3. Right hand side number indicates full marks

**Q.No.1 Multiple choice questions**

**(10 X 1=10)**

1. The third edition of Indian Pharmacopoeia was published in  
a) 2007 b) 1996 c) 1985 d) 1966
2. The semisolid unit dosage form intended for vaginal administration are called  
a) Suppositories b) Pessaries c) Bougies d) Cones
3. Drugs that suppresses nausea and vomiting are called  
a) Analgesic b) Anthelmintic c) Antibiotic, d) Antiemetic
4. 1 $\bar{3}$  (apothecaries) =  
a) 437.5 grains b) 480 grains c) 7000 grains d) 5760 grains
5. The powders on mixing turn to liquid due to depression of melting point are called  
a) Hygroscopic powders b) Efflorescent powders c) Eutectic mixtures d) Compound powders
6. The adult dose of drug is 100mg, what will be the dose for a child having body surface area 0.57m<sup>2</sup>?  
a) 33 mg b) 66 mg c) 50 mg d) 100mg.
7. Which of the following co-solvents are used to increase the solubility of a drug?  
a) Ehanol, b) Sorbitol, c) Glycerine, d) All the above
8. Which of the following dosage form is not recommended for ocular administration  
a) Lotions, b) Liniments, c) Suspensions, d) Solutions
9. Which part of prescription contains name and quantities of the prescribed ingredients?  
a) Superscription, b) Subscription, c) Inscription, d) Signa
10. .... is known as 'Father of Medicine'.  
a) Hippocrates, b) Theophrastus, c) Paracelsus, d) Benjamin Franklin

**Q. No. 2 Long Answers (Answer 1 out of 2)**

**(1 x 10 = 10)**

1. Define dosage form. Classify dosage form. Describe various solid dosage forms.
2. Define the term Posology. Explain in detail various factors affecting dose of drug.

**Q. No. 3 Short Answers (Answer 2 out of 3)**

**(2 x 5 = 10)**

1. Define prescription. Write a note on parts of prescription.
2. a) Explain various methods of powders mixing.  
b) Differentiate between lotions and liniments.
3. Write short note on (any one)  
a) History of Indian Pharmacopoeia.  
b) Excipient used in liquid dosage form.



**First Sessional Theory Examination 2018-2019**

**Subject:** Pharmaceutical Inorganic Chemistry (BP104T)

**Day & Date:** Fri, 05/10/2018

**Class:** First Year B. Pharm.

**Semester:** I

**Time:** 10.30 am – 12.00 pm

**Max. Marks:** 30

- Instructions:* 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks

**Q.I. Multiple Choice Questions (MCQs)**

**10**

1. Lewis base is.....  
a) electron pair donor    b) electron pair acceptor    c) proton donor    d) proton acceptor
2. The principal function of chloride is.....  
a) maintenance of proper hydration    b) maintenance of osmotic pressure  
c) normal electrolyte balance    d) all of the above
3. Citric acid is used in limit test for iron to prevent.....  
a) precipitation of iron by ammonia    b) reduction of iron  
c) precipitation of iron by thioglycollic acid    d) complexation with thioglycollic acid
4. Indian pharmacopoeia is published by.....  
a) Ministry of health and family welfare    b) food and drug administration  
c) world health organization    d) ministry of science and technology
5. Lead acetate cotton plug is used in limit test for arsenic to.....  
a) trap hydrogen sulphide    b) trap thiosulphate    c) trap hydrogen sulphate    d) trap arsine gas
6. High serum potassium level is called as.....  
a) hypocalcemia    b) hypokalemia    c) hypercalcemia    d) hyperkalemia
7. Choose the correct composition of ORS as per current WHO formula  
i) NaCl ii) KCl iii) Sodium citrate iv) sodium bicarbonate v) anhydrous glucose  
a) i, ii, iii, v    b) i, ii, iv, v    c) I, ii, iii, iv    d) ii, iii, iv, v
8. calcium gluconate is prepared from.....  
a) Lactic acid and CaCO<sub>3</sub>    b) Oxalic acid and CaCO<sub>3</sub>    c) Gluconic acid and CaCO<sub>3</sub>  
d) Gluconic acid and Ca(OH)<sub>2</sub>
9. Limit test for heavy metals is based on reaction between.....  
a) metallic impurities and hydrogen sulphide    b) metallic impurities and hydrogen sulphate  
c) metallic impurities and acetic acid    d) metallic impurities and ammonia
10. The first Indian pharmacopoeia was published by Indian Pharmacopoeia committee in .....  
a) 1948    b) 1955    c) 1960    d) 1966

**Q. II. Solve any One**

**10**

1. Explain different theories of acid and base. Explain the mechanism of buffer action.  
2. Discuss the method of preparation, uses and assay of sodium chloride and calcium gluconate.

**Q. III. Solve any Two (5 Marks each)**

**10**

1. Explain the principle behind limit test of Iron and Sulphate.  
2. Explain the principle of arsenic limit test. Draw a neat labelled diagram of Gutzeit apparatus.  
3. Explain the sources of impurities.



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**First Sessional Theory Examination 2019-2020**

**Subject:** Pharmaceutical Inorganic Chemistry (BP104T) **Day & Date:** 28-09-2019

**Class:** First Year B. Pharm.

**Semester:** I

**Time:** 10:30 am – 12:00 pm

**Max. Marks:** 30

- Instructions:**
1. All questions are compulsory
  2. Draw a well labeled diagram wherever necessary
  3. Right hand side number indicates full marks

**QI. Solve the following Multiple Choice Questions (MCQs) 10**

1. The sources of impurities include .....  
a) raw materials      b) manufacturing process      c) instability of product      d) all of these
2. Lewis base is .....  
a. electron pair acceptor      b. electron pair donor      c. proton acceptor      d. proton donor
3. Indian pharmacopoeia is published by .....  
a) world health organization      b) ministry of health and family welfare  
c) ministry of science and technology      d) food and drug administration
4. Citric acid is used in the limit test for iron to prevent .....  
a) precipitation of iron by ammonia      b) precipitation of iron by thioglycolic acid  
c) oxidation of iron      d) reduction of iron
5. Limit test for heavy metals is based on reaction between heavy metal and .....  
a) hydrogen sulphide      b) hydrogen sulphate      c) ammonia      d) dithiazone
6. Lead acetate cotton plug is used in arsenic limit test to trap .....  
a) arsine gas      b) hydrogen sulphide gas      c) thiosulphate      d) hydrogen sulphate
7. The major functions of electrolytes in body are .....  
a) to maintain osmotic pressure      b) to maintain acid-base balance  
c) to maintain fluid volume      d) all of the above
8. The concentration of sodium chloride solution that is isotonic with body fluids is .....  
a) 0.9 % w/v      b) 1.6 % w/v      c) 0.7 % w/v      d) 2% w/v
9. Good dental cleaning agent generally have.....  
a) coarse to fine particle size      b) good abrasive property      c) both a and b      d) none of these
10. Assay of calcium gluconate is performed by ..... titration  
a) Redox      b) complexometric      c) acid-base      d) precipitation

**QII. Solve any one of the following 10**

1. Explain the different theories of acids and bases. Define Buffer. Explain the mechanism of buffer action and buffer capacity.
2. Explain the role of sodium and potassium in the body. Write a note on regulation of acid-base balance in body fluids.

**QIII Solve the following (any 2) 10**

1. What are dental caries? Explain the role of fluoride in the treatment of dental caries.
2. Write a note on ORS. Explain the preparation, properties, assay and uses of Sodium chloride
3. Explain the principle involved in limit test of Arsenic with sequence of reactions involved. Draw a neat labelled diagram of Gutzeit apparatus for limit test of Arsenic.



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**First Sessional Theory Examination 2020-2021**

**Subject:** Pharmaceutical Inorganic Chemistry (BP104T) **Day & Date:** 13-03-2021

**Class:** First Year B. Pharm.

**Semester:** I

**Time:** 10:30 am – 12:00 pm

**Max. Marks:** 30

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*Instructions:* 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks

**Q. I. Solve the following Multiple Choice Questions (MCQs) 10**

1. The first Indian pharmacopoeia was published by Indian Pharmacopoeia committee in .....
  - a. 1948
  - b. 1955
  - c. 1960
  - d. 1966
2. Limit test for sulphate involves reaction between soluble sulphate and .....
  - a. Silver nitrate
  - b. Acetic acid
  - c. Barium chloride
  - d. Hydrochloric acid
3. Which of the following is a major extracellular cation?
  - a. Sodium
  - b. Chloride
  - c. Potassium
  - d. Phosphate
4. The role of sodium in body is .....
  - a. To maintain osmotic pressure
  - b. To transmission of nerve impulses
  - c. To maintain fluid volume
  - d. All of the above
5. Calcium gluconate is prepared from.....
  - a. Lactic acid and  $\text{CaCO}_3$
  - b. Oxalic acid and  $\text{CaCO}_3$
  - c. Gluconic acid and  $\text{CaCO}_3$
  - d. Gluconic acid and  $\text{Ca(OH)}_2$

6. Fluoride inhibits caries formation via ....
  - a. Decreasing acid solubility of enamel
  - b. Bacterial inhibition
  - c. Both a and b
  - d. Increasing acid solubility of enamel
7. Absence of HCl in gastric secretion is called as .....
  - a. Hyperacidity
  - b. Achlorhydria
  - c. Constipation
  - d. Diarrhoea
8. Antacid should not be .....
  - a. Constipative
  - b. Laxative
  - c. Both a and b
  - d. Protective
9. What is incorrect about sodium bicarbonate?
  - a. It is systemic antacid
  - b. It is called as baking soda
  - c. Produces effervescence with acid
  - d. Does not get absorbed systemically
10. The sources of impurities include .....
  - a. Raw materials
  - b. Manufacturing process
  - c. Instability of product
  - d. All of the above

**Q. II. Solve any one of the following**

**10**

1. Explain various sources of impurities. Explain the principle involved in limit test for Arsenic with reactions. Draw a neat-labelled diagram of Gutzeit apparatus
2. Explain different theories of acids and bases. Add a note of mechanism of buffer action.

**Q. III. Solve the following (any 2)**

**10**

1. What are ideal properties of antacids? Classify antacid and write a note on Aluminium hydroxide gel.
2. Write a note on ORS. Explain the preparation, properties, assay and uses of Sodium chloride
3. Explain the principle and reaction involved in limit test of Iron





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## First Sessional Theory Examination-(ODD SEM) 2022-2023

<b>Subject</b> : Pharmaceutical Inorganic Chemistry (BP104T)	<b>Day &amp; Date</b> : Tues, 14/03/2023	<b>Write Your Roll No. Here</b> <input type="text"/>
<b>Class</b> : First Year B. Pharmacy	<b>Semester</b> : I	
<b>Time</b> : 10:30 am – 12:00 pm	<b>Max. Marks</b> : 30	

### Instructions:

1. All questions are compulsory
2. Draw a well-labeled diagram wherever necessary
3. Right-hand side number indicates full marks
4. Do not write/tick on the question paper

### Q. I: Objective Type Questions

10 M

1. Define the following term. (CO1, CO2, LL1)  
Pharmacopoeia ii) Desensitizing agents
2. Define Achlorhydria and gives its treatment. (CO2, LL1)
3. What is the role of citric acid and ammonia in the limit test of iron? (CO1, LL1)
4. Enlist the composition of ORS (CO2, LL1)
5. Write the principle and reaction behind the limit test of chloride and sulphate. (CO1, LL1)

### Q. II: Long Answers Questions (Answer any 1 out of 2)

10 M

1. Define Buffers. Explain the mechanism of buffer action with an example. Briefly discuss the role of Buffers in Pharmacy. (CO1, LL3)
2. Define Impurities. Enlist and explain the various sources and types of impurities. (CO1, LL3)

### Q. III: Short Answers Questions (Answer any 2 out of 3)

10 M

1. Explain the role of fluoride in the treatment of Tooth Decay. (CO2, LL3)
2. Define and discuss on Antacid used in Combination Therapy. (CO2, LL3)
3. Write note on Major Extracellular Physiological ions. (Any Two) (CO2, LL3)



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**First Sessional Theory Examination-(ODD SEM) 2022-2023**

<b>Subject</b> : Pharmaceutical Inorganic Chemistry (BP104T)	<b>Day &amp; Date</b> : Tues, 14/03/2023	<b>Write Your Roll No. Here</b> <input type="text"/>
<b>Class</b> : First Year B. Pharmacy	<b>Semester</b> : I	
<b>Time</b> : 10:30 am – 12:00 pm	<b>Max. Marks</b> : 30	

**Instructions:**

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**Q. I: Objective Type Questions**

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Pharmacopoeia ii) Desensitizing agents
2. Define Achlorhydria and gives its treatment. (CO2, LL1)
3. What is the role of citric acid and ammonia in the limit test of iron? (CO1, LL1)
4. Enlist the composition of ORS (CO2, LL1)
5. Write the principle and reaction behind the limit test of chloride and sulphate. (CO1, LL1)

**Q. II: Long Answers Questions (Answer any 1 out of 2)**

**10 M**

1. Define Buffers. Explain the mechanism of buffer action with an example. Briefly discuss the role of Buffers in Pharmacy. (CO1, LL3)
2. Define Impurities. Enlist and explain the various sources and types of impurities. (CO1, LL3)

**Q. III: Short Answers Questions (Answer any 2 out of 3)**

**10 M**

1. Explain the role of fluoride in the treatment of Tooth Decay. (CO2, LL3)
2. Define and discuss on Antacid used in Combination Therapy. (CO2, LL3)
3. Write note on Major Extracellular Physiological ions. (Any Two) (CO2, LL3)



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**First Sessional Theory Examination-(ODD SEM) 2023-2024**

<b>Subject</b> : Pharmaceutical Inorganic Chemistry (BP104T)	<b>Day &amp; Date</b> : Thu, 02/11/2023	
<b>Class</b> : First Year B. Pharmacy	<b>Semester</b> : I	<b>Write Your Roll No. Here</b> <input type="text"/>
<b>Time</b> : 10:30 am – 12:00 pm	<b>Max. Marks</b> : 30	

**Instructions:**

1. All questions are compulsory
2. Draw a well-labeled diagram wherever necessary
3. Right-hand side number indicates full marks
4. Do not write/tick on the question paper

<b>Q. I: Objective Type Questions</b>	<b>10 M</b>
1. Define the following term. (CO1, CO2, LL1) i) Buffers ii) Antacid	
2. Write the chemical reaction for limit test of lead. (CO1, LL3)	
3. Define acid and base according to Lowry Bronsted theory with example. (CO1, LL1)	
4. Enlist the composition of Zinc-Eugenol Cement (CO2, LL3)	
5. Define Hyponatremia and Hypernatremia. (CO2, LL1)	

<b>Q. II: Long Answers Questions (Answer any 1 out of 2)</b>	<b>10 M</b>
1. Explain the types of tonicity and methods of adjusting isotonicity. Describe the mechanism of action of buffers (CO1, LL3)	
2. Define limit test and describe the principle, apparatus and procedure for the limit test of arsenic. (CO1, LL3)	

<b>Q. III: Short Answers Questions (Answer any 2 out of 3)</b>	<b>10 M</b>
1. Define dentifrices and describe the method of M.W., M.F., properties, preparation and uses of CaCO <sub>3</sub> . (CO2, LL3)	
2. Define and discuss on Acidifiers with example. (CO2, LL3)	
3. What is electrolyte combination therapy? Explain a note on ORS. (Any Two) (CO2, LL3)	



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**First Sessional Theory Examination-(ODD SEM) 2023-2024**

<b>Subject</b> : Pharmaceutical Inorganic Chemistry (BP104T)	<b>Day &amp; Date</b> : Thu, 02/11/2023	
<b>Class</b> : First Year B. Pharmacy	<b>Semester</b> : I	<b>Write Your Roll No. Here</b> <input type="text"/>
<b>Time</b> : 10:30 am – 12:00 pm	<b>Max. Marks</b> : 30	

**Instructions:**

1. All questions are compulsory
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3. Right-hand side number indicates full marks
4. Do not write/tick on the question paper

<b>Q. I: Objective Type Questions</b>	<b>10 M</b>
1. Define the following term. (CO1, CO2, LL1) i) Buffers ii) Antacid	
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3. Define acid and base according to Lowry Bronsted theory with example. (CO1, LL1)	
4. Enlist the composition of Zinc-Eugenol Cement (CO2, LL3)	
5. Define Hyponatremia and Hypernatremia. (CO2, LL1)	

<b>Q. II: Long Answers Questions (Answer any 1 out of 2)</b>	<b>10 M</b>
1. Explain the types of tonicity and methods of adjusting isotonicity. Describe the mechanism of action of buffers (CO1, LL3)	
2. Define limit test and describe the principle, apparatus and procedure for the limit test of arsenic. (CO1, LL3)	

<b>Q. III: Short Answers Questions (Answer any 2 out of 3)</b>	<b>10 M</b>
1. Define dentifrices and describe the method of M.W., M.F., properties, preparation and uses of CaCO <sub>3</sub> . (CO2, LL3)	
2. Define and discuss on Acidifiers with example. (CO2, LL3)	
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**First Sessional Theory Examination 2020-2021**

**Subject:** Pharmaceutical Analysis I (BP102T)

**Day & Date:** 11-03-2021

**Class:** First Year B. Pharm.

**Semester:** I

**Time:** 10:30 am – 12:00 pm

**Max. Marks:** 30

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*Instructions:*

1. All questions are compulsory
2. Draw a well labeled diagram wherever necessary
3. Right hand side number indicates full marks

**Q. I. Solve the following Multiple Choice Questions (MCQs)**

**10**

1. Number of gram equivalents of solute present in one litre of a solution is called as .....
  - a) Molarity
  - b) Normality
  - c) Molality
  - d) Percentage weight by volume
2. The features of Primary standard substance includes:
  - a) High purity
  - b) High stability
  - c) Low hygroscopicity
  - d) All of the above
3. Indian pharmacopoeia is published by .....
  - a. World health organization
  - b. Ministry of health and family welfare
  - c. Ministry of science and technology
  - d. Food and drug administration
4. Limit test of iron is based upon the reactions between iron and ....
  - a. Dithiazone
  - b. Thioglycolic acid
  - c. Hydrogen sulphide
  - d. Barium chloride
5. The color of phenolphthalein in alkaline medium is .....
  - a) Pink
  - b) Yellow
  - c) Brown
  - d) Colorless

6. Benzene is an example of .....solvent.
- Aprotic
  - Protogenic
  - Protophilic
  - Amphiprotic
7. Solubility of precipitate is ..... on addition of common ion.
- Decreased
  - Increased
  - Remains Constant
  - None of These
8. In Mohr's method .....is used as an indicator.
- Potassium chromate
  - Ferric ammonium sulphate
  - Ammonium thiocyanate
  - Silver nitrate
9. In Volhard's method .....is used as a titrant
- Potassium chromate
  - Ferric ammonium sulphate
  - Ammonium thiocyanate
  - Silver nitrate
10. The end-point detection in Fajan's method involves color change due to.....
- Formation of colored precipitate
  - Formation of soluble colored complex
  - Adsorption of indicator on the surface of precipitate
  - None of the above

**Q. II. Solve the following (any 1)**

**10**

1. Explain the neutralization curves in acid-base titrations for i) strong acid vs strong base ii) weak acid vs strong base iii) strong acid vs weak base iv) weak acid vs weak base.
2. Explain in detail Mohr's method and Fajan's method.

**Q. III. Solve the following (any 2)**

**10**

1. Explain principle involved in limit test for Chloride and sulphate.
2. Explain in detail the theories of acid-base indicators.
3. Explain the types of solvents used in non-aqueous titrations with suitable examples.



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**First Sessional Theory Examination 2020-2021**

**Subject:** Pharmaceutical Analysis I (BP102T)

**Day & Date:** 11-03-2021

**Class:** First Year B. Pharm.

**Semester:** I

**Time:** 10:30 am – 12:00 pm

**Max. Marks:** 30

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*Instructions:*

1. All questions are compulsory
2. Draw a well labeled diagram wherever necessary
3. Right hand side number indicates full marks

**Q. I. Solve the following Multiple Choice Questions (MCQs)**

**10**

1. Number of gram equivalents of solute present in one litre of a solution is called as .....
  - a) Molarity
  - b) Normality
  - c) Molality
  - d) Percentage weight by volume
2. The features of Primary standard substance includes:
  - a) High purity
  - b) High stability
  - c) Low hygroscopicity
  - d) All of the above
3. Indian pharmacopoeia is published by .....
  - a. World health organization
  - b. Ministry of health and family welfare
  - c. Ministry of science and technology
  - d. Food and drug administration
4. Limit test of iron is based upon the reactions between iron and ....
  - a. Dithiazone
  - b. Thioglycolic acid
  - c. Hydrogen sulphide
  - d. Barium chloride
5. The color of phenolphthalein in alkaline medium is .....
  - a) Pink
  - b) Yellow
  - c) Brown
  - d) Colorless

6. Benzene is an example of .....solvent.
- Aprotic
  - Protogenic
  - Protophilic
  - Amphiprotic
7. Solubility of precipitate is ..... on addition of common ion.
- Decreased
  - Increased
  - Remains Constant
  - None of These
8. In Mohr's method .....is used as an indicator.
- Potassium chromate
  - Ferric ammonium sulphate
  - Ammonium thiocyanate
  - Silver nitrate
9. In Volhard's method .....is used as a titrant
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  - Ferric ammonium sulphate
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  - Silver nitrate
10. The end-point detection in Fajan's method involves color change due to.....
- Formation of colored precipitate
  - Formation of soluble colored complex
  - Adsorption of indicator on the surface of precipitate
  - None of the above

**Q. II. Solve the following (any 1)**

**10**

1. Explain the neutralization curves in acid-base titrations for i) strong acid vs strong base ii) weak acid vs strong base iii) strong acid vs weak base iv) weak acid vs weak base.
2. Explain in detail Mohr's method and Fajan's method.

**Q. III. Solve the following (any 2)**

**10**

1. Explain principle involved in limit test for Chloride and sulphate.
2. Explain in detail the theories of acid-base indicators.
3. Explain the types of solvents used in non-aqueous titrations with suitable examples.



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**First Sessional Theory Examination 2020-2021**

**Subject:** Pharmaceutical Inorganic Chemistry (BP104T) **Day & Date:** 13-03-2021

**Class:** First Year B. Pharm.

**Semester:** I

**Time:** 10:30 am – 12:00 pm

**Max. Marks:** 30

---

*Instructions:*

1. All questions are compulsory
2. Draw a well labeled diagram wherever necessary
3. Right hand side number indicates full marks

**Q. I. Solve the following Multiple Choice Questions (MCQs)**

**10**

1. The first Indian pharmacopoeia was published by Indian Pharmacopoeia committee in .....
  - a. 1948
  - b. 1955
  - c. 1960
  - d. 1966
2. Limit test for sulphate involves reaction between soluble sulphate and .....
  - a. Silver nitrate
  - b. Acetic acid
  - c. Barium chloride
  - d. Hydrochloric acid
3. Which of the following is a major extracellular cation?
  - a. Sodium
  - b. Chloride
  - c. Potassium
  - d. Phosphate
4. The role of sodium in body is .....
  - a. To maintain osmotic pressure
  - b. To transmission of nerve impulses
  - c. To maintain fluid volume
  - d. All of the above
5. Calcium gluconate is prepared from.....
  - a. Lactic acid and  $\text{CaCO}_3$
  - b. Oxalic acid and  $\text{CaCO}_3$
  - c. Gluconic acid and  $\text{CaCO}_3$
  - d. Gluconic acid and  $\text{Ca(OH)}_2$

6. Fluoride inhibits caries formation via ....
  - a. Decreasing acid solubility of enamel
  - b. Bacterial inhibition
  - c. Both a and b
  - d. Increasing acid solubility of enamel
7. Absence of HCl in gastric secretion is called as .....

  - a. Hyperacidity
  - b. Achlorhydria
  - c. Constipation
  - d. Diarrhoea

8. Antacid should not be .....

  - a. Constipative
  - b. Laxative
  - c. Both a and b
  - d. Protective

9. What is incorrect about sodium bicarbonate?
  - a. It is systemic antacid
  - b. It is called as baking soda
  - c. Produces effervescence with acid
  - d. Does not get absorbed systemically
10. The sources of impurities include .....

  - a. Raw materials
  - b. Manufacturing process
  - c. Instability of product
  - d. All of the above

**Q. II. Solve any one of the following**

**10**

1. Explain various sources of impurities. Explain the principle involved in limit test for Arsenic with reactions. Draw a neat-labelled diagram of Gutzeit apparatus
2. Explain different theories of acids and bases. Add a note of mechanism of buffer action.

**Q. III. Solve the following (any 2)**

**10**

1. What are ideal properties of antacids? Classify antacid and write a note on Aluminium hydroxide gel.
2. Write a note on ORS. Explain the preparation, properties, assay and uses of Sodium chloride
3. Explain the principle and reaction involved in limit test of Iron





**First Sessional Theory Examination 2021-2022  
(Odd SEM)**

**Subject & Subject Code: Pharmaceutical Inorganic Chemistry (BP104T)**

**Day & Date:** Thursday 10/02/2022

**Class:** F.Y. B. Pharmacy

**Semester:** I

**Time:** 10:30 AM to 12:00 PM

**Max. Marks:** 30

*Instructions:* 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks

**QI Multiple Choice Questions (MCQs)**

**10**

- Ammonium chloride is an example of -----  
A) Stimulant expectorant B) Sedative expectorant  
C) Both a and b D) Antacid
- Citric acid is added in limit test of iron to \_\_\_\_\_  
A) Produce pink colour B) To form complex with iron and prevent the precipitation by ammonia Solution  
C) React with iron D) Produce acidity
- Dental cement consists of \_\_\_\_\_  
A) Zn + Eugenol B) Na + Mannitol C) Al + Cresol D) None of above
- Alcohol is used in barium sulphate reagent to prevent -----  
A) Saturation B) Supersaturation C) Turbidity D) Opalascence
- Impurities in pharmaceutical preparation may be due to following  
A) Raw material B) Manufacturing process  
C) Chemical instability D) All of the above
- In which method tonicity is calculated by adding water to the drugs to make an isotonic solution  
A) Sodium chloride equivalent method B) Cryoscopic method  
C) White Vincent method D) Potentiometric method
- First edition of IP was published in  
A) 1960 B) 1948 C) 1955 D) 1966
- The body fluid within the cells is called as \_\_\_\_\_  
A) plasma B) Intracellular fluid C) Extracellular fluid D) Water
- To prevent dental caries it becomes necessary to use -----  
A) Sodium Fluoride B) Gold particles C) Silver nitrate D) Copper sulphate
- Which of the following is ideal properties of dental product  
A) Easy to use B) Must not produce any stain on teeth  
C) Must not produce nausea or vomiting D) All of the above

**QII Long Answers (Answer any 1 out of 2)**

**10**

- Define the limit test and discuss in detail the limit test for arsenic with suitable diagram of gutzeit apparatus.
- Define the term isotonicity and discuss in detail different methods used to measure the isotonicity.

**QIII Short Answers (Answer any 2 out of 3)**

**10**

- Define anti - carries agents. Discuss how fluoride produces anti-carries activity
- Define and Classify antacids with suitable example and give the ideal properties of Antacid.
- Discuss in detail Metabolic Acidosis and Alkalosis.



**First Sessional Theory Examination 2021-2022  
(Odd SEM)**

**Subject & Subject Code: Pharmaceutical Inorganic Chemistry (BP104T)**

**Day & Date:** Thursday 10/02/2022

**Class:** F.Y. B. Pharmacy

**Semester:** I

**Time:** 10:30 AM to 12:00 PM

**Max. Marks:** 30

*Instructions:* 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks

**QI Multiple Choice Questions (MCQs)**

**10**

- Ammonium chloride is an example of -----  
A) Stimulant expectorant B) Sedative expectorant  
C) Both a and b D) Antacid
- Citric acid is added in limit test of iron to \_\_\_\_\_  
A) Produce pink colour B) To form complex with iron and prevent the precipitation by ammonia Solution  
C) React with iron D) Produce acidity
- Dental cement consists of \_\_\_\_\_  
A) Zn + Eugenol B) Na + Mannitol C) Al + Cresol D) None of above
- Alcohol is used in barium sulphate reagent to prevent -----  
A) Saturation B) Supersaturation C) Turbidity D) Opalascence
- Impurities in pharmaceutical preparation may be due to following  
A) Raw material B) Manufacturing process  
C) Chemical instability D) All of the above
- In which method tonicity is calculated by adding water to the drugs to make an isotonic solution  
A) Sodium chloride equivalent method B) Cryoscopic method  
C) White Vincent method D) Potentiometric method
- First edition of IP was published in  
A) 1960 B) 1948 C) 1955 D) 1966
- The body fluid within the cells is called as \_\_\_\_\_  
A) plasma B) Intracellular fluid C) Extracellular fluid D) Water
- To prevent dental caries it becomes necessary to use -----  
A) Sodium Fluoride B) Gold particles C) Silver nitrate D) Copper sulphate
- Which of the following is ideal properties of dental product  
A) Easy to use B) Must not produce any stain on teeth  
C) Must not produce nausea or vomiting D) All of the above

**QII Long Answers (Answer any 1 out of 2)**

**10**

- Define the limit test and discuss in detail the limit test for arsenic with suitable diagram of gutzeit apparatus.
- Define the term isotonicity and discuss in detail different methods used to measure the isotonicity.

**QIII Short Answers (Answer any 2 out of 3)**

**10**

- Define anti - carries agents. Discuss how fluoride produces anti-carries activity
- Define and Classify antacids with suitable example and give the ideal properties of Antacid.
- Discuss in detail Metabolic Acidosis and Alkalosis.



**First Sessional Theory Examination 2021-2022  
(Odd SEM)**

**Subject & Subject Code: Pharmaceutical Inorganic Chemistry (BP104T)**

**Day & Date:** Thursday 10/02/2022

**Class:** F.Y. B. Pharmacy

**Semester:** I

**Time:** 10:30 AM to 12:00 PM

**Max. Marks:** 30

*Instructions:* 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks

**QI Multiple Choice Questions (MCQs)**

**10**

- Ammonium chloride is an example of -----  
A) Stimulant expectorant B) Sedative expectorant  
C) Both a and b D) Antacid
- Citric acid is added in limit test of iron to \_\_\_\_\_  
A) Produce pink colour B) To form complex with iron and prevent the precipitation by ammonia Solution  
C) React with iron D) Produce acidity
- Dental cement consists of \_\_\_\_\_  
A) Zn + Eugenol B) Na + Mannitol C) Al + Cresol D) None of above
- Alcohol is used in barium sulphate reagent to prevent -----  
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- Impurities in pharmaceutical preparation may be due to following  
A) Raw material B) Manufacturing process  
C) Chemical instability D) All of the above
- In which method tonicity is calculated by adding water to the drugs to make an isotonic solution  
A) Sodium chloride equivalent method B) Cryoscopic method  
C) White Vincent method D) Potentiometric method
- First edition of IP was published in  
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- The body fluid within the cells is called as \_\_\_\_\_  
A) plasma B) Intracellular fluid C) Extracellular fluid D) Water
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A) Sodium Fluoride B) Gold particles C) Silver nitrate D) Copper sulphate
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**QII Long Answers (Answer any 1 out of 2)**

**10**

- Define the limit test and discuss in detail the limit test for arsenic with suitable diagram of gutzeit apparatus.
- Define the term isotonicity and discuss in detail different methods used to measure the isotonicity.

**QIII Short Answers (Answer any 2 out of 3)**

**10**

- Define anti - carries agents. Discuss how fluoride produces anti-carries activity
- Define and Classify antacids with suitable example and give the ideal properties of Antacid.
- Discuss in detail Metabolic Acidosis and Alkalosis.



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**First Sessional Theory Examination-(ODD SEM) 2022-2023**

<b>Subject</b> :	Remedial Biology (BP106RBT)	<b>Day &amp; Date</b> :	THURSDAY, 09/03/2023	
<b>Class</b> :	First Year B. Pharmacy	<b>Semester</b> :	<b>I</b>	<i>Write Your Seat No.</i>
<b>Time</b> :	10:30 am – 11:30 am	<b>Max. Marks</b> :	30	<input type="text"/>

<b>Instructions:</b>	1. All questions are compulsory	2. Draw a well labeled diagram wherever necessary
	3. Right hand side number indicates full marks	4. Do not write/tick on the question paper

<b>QI</b>	<b>Long Answers Question (Answer any 1 out of 2)</b>	<b>10 M</b>
1.	Describe kingdom of life with suitable example (LL3, CO1)	
2.	Elaborate types of Inflorescences with suitable example (LL3, CO1)	
<b>QII</b>	<b>Short Answers Question</b>	
	<b>Section – A (Answer any ONE)</b>	<b>05 M</b>
1.	Write a note on classification and uses of Fruit (LL2, CO1)	
2.	Illuminate the general histological characters of Dicot and Monocot root. (LL2, CO1)	
	<b>Section – B (Answer any Three)</b>	<b>15 M</b>
1.	Write a note on composition and functions of the Lymphatic system (LL2, CO2)	
2.	Comment on Digestive enzymes (LL2, CO2)	
3.	Explain the mechanism of breathing (LL2, CO2)	
4.	Draw well labeled structure of Human Urinary System. (LL2, CO2)	



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**First Sessional Theory Examination-(ODD SEM) 2022-2023**

<b>Subject</b> :	Remedial Biology (BP106RBT)	<b>Day &amp; Date</b> :	THURSDAY,09/03/2023	
<b>Class</b> :	First Year B. Pharmacy	<b>Semester</b> :	<b>I</b>	<i>Write Your Seat No.</i>
<b>Time</b> :	10:30 am – 11:30 am	<b>Max. Marks</b> :	30	<input type="text"/>

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	<b>Section – A (Answer any ONE)</b>	<b>05 M</b>
1.	Write a note on classification and uses of Fruit (LL2, CO1)	
2.	Illuminate the general histological characters of Dicot and Monocot root. (LL2, CO1)	
	<b>Section – B (Answer any Three)</b>	<b>15 M</b>
1.	Write a note on composition and functions of the Lymphatic system (LL2, CO2)	
2.	Comment on Digestive enzymes (LL2, CO2)	
3.	Explain the mechanism of breathing (LL2, CO2)	
4.	Draw well labeled structure of Human Urinary System. (LL2, CO2)	



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**First Sessional Theory Examination-(ODD SEM) 2022-2023**

<b>Subject</b> :	Remedial Biology (BP106RBT)	<b>Day &amp; Date</b> :	THURSDAY,09/03/2023	
<b>Class</b> :	First Year B. Pharmacy	<b>Semester</b> :	<b>I</b>	<i>Write Your Seat No.</i>
<b>Time</b> :	10:30 am – 11:30 am	<b>Max. Marks</b> :	30	<input type="text"/>

<b>Instructions:</b>	1. All questions are compulsory	2. Draw a well labeled diagram wherever necessary
	3. Right hand side number indicates full marks	4. Do not write/tick on the question paper

<b>QI</b>	<b>Long Answers Question (Answer any 1 out of 2)</b>	<b>10 M</b>
1.	Describe kingdom of life with suitable example (LL3, CO1)	
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<b>QII</b>	<b>Short Answers Question</b>	
	<b>Section – A (Answer any ONE)</b>	<b>05 M</b>
1.	Write a note on classification and uses of Fruit (LL2, CO1)	
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1.	Write a note on composition and functions of the Lymphatic system (LL2, CO2)	
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**First Sessional Theory Examination 2020-2021**  
(Odd SEM)

**Subject & Subject Code: Remedial Biology (BP106RBT)**

**Day & Date: 16.03.2021 Tuesday**

**Class: F.Y. B. Pharm**

**Semester: I**

**Time: 10.30 to 11.30**

**Max. Marks: 30 marks**

*Instructions: 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks*

**Q1. Answer any one (10 marks)**

- 1. Explain the morphology and anatomy of flower with well labelled diagram.**
- 2. Give structure of human heart and explain in brief about cardiac cycle**

**Q2. Answer any four (20 marks)**

- 1. Enlist difference between Eukaryotic and Prokaryotic cells**
- 2. Write a note on binominal nomenclature**
- 3. Explain about ECG**
- 4. Give a short note on mechanism of breathing**
- 5. Explain the role of Digestive Enzymes**
- 6. Give composition of blood**

*Bahadur  
(S.S. Rathod)  
Subject Incharge*

# Shri. Vile Parle Kelavani Mandal's

## Institute of Pharmacy, Dhule

Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

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### First Sessional Theory Examination 2018-2019

**Subject: Remedial Biology**

**Day & Date: Monday, 8<sup>th</sup> October 2018**

**Class: F. Y. B. Pharma**

**Semester: I**

**Time: 10:30 am -12:00 pm**

**Max. Marks: 30**

---

*Instructions: 1. All questions are compulsory*

*2. Draw a well labeled diagram wherever necessary*

*3. Right hand side number indicates full marks*

**QI Answer any one**

**10**

1. Short note on human heart and blood vessels. Describe in details Cardiac cycle
2. Explain the kingdoms of life and basis of classification. Short note on Salient features of Plantae and Virus.

**QII Answer any four**

**20**

1. Short note on breathing process and its regulation
2. What is blood and explain its Composition and function.
3. Describe Role of digestive enzymes in digestion process
4. Explain General Anatomy of Root, stem, leaf of monocotyledons & Dicotyledones.
5. Short note on Electrocardiogram
6. Definition and characters of organisms.

*First Sessional Theory Examination 2018-2019*

**Class: F.Y.B.Pharmacy**

**Semester: First**

**Subject : Remedial Mathematics**

**Max. Marks : 30**

**Date: October. 8<sup>th</sup>, 2018**

**Time : 10.30 am - 12.00 pm.**

---

• **Instructions :**

1. Both the questions are compulsory.
  2. Figures to the right indicates full marks.
- 

1. Attempt any two of the following. (5+5=10)

(a) Solve by Cramer's rule

$$x + y + z = 3, \quad x - y + z = 1, \quad x + y - 2z = 0.$$

(b) Using log-tables find the value of

$$\frac{(30.2153)^{\frac{3}{2}}(0.00516)^{\frac{2}{3}}}{\sqrt{0.02157}}$$

(c) Resolve into partial fractions  $\frac{x}{(x+1)^3(x-2)^2}$ .

2. Attempt any five of the following. (4x5=20)

(a) Show that  $7\log\frac{16}{15} + 5\log\frac{25}{24} + 3\log\frac{81}{80} + 21\log 1 = \log 2$ .

(b) Resolve into partial fractions  $\frac{x^2+1}{x(x-1)(x-3)}$ .

(c) Solve the equation, 
$$\begin{vmatrix} x+a & b & c \\ a & x+b & c \\ a & b & x+c \end{vmatrix} = 0.$$

(d) Solve  $\log_{10}(x+1) - \log_{10}(x-1) = 1$ .

(e) Without expanding the determinant show that 
$$\begin{vmatrix} 0 & b & -c \\ -b & 0 & a \\ c & -a & 0 \end{vmatrix} = 0.$$

(f) Is the matrix  $\begin{bmatrix} 1 & 2 & 3 \\ 0 & 1 & 2 \\ 0 & 0 & 1 \end{bmatrix}$  idempotent? Justify.

(g) If  $A = \begin{bmatrix} 3 & 5 & -1 \\ 6 & 4 & 2 \end{bmatrix}$  and  $B = \begin{bmatrix} 2 & 5 & -1 \\ 1 & 3 & -2 \end{bmatrix}$ , verify that  $(A+B)^t = A^t + B^t$ .

\*\*\*\*\*



# Shri. Vile Parle Kelavani Mandal's Institute of Pharmacy, Dhule

Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

## First Sessional Theory Examination 2019-2020

**Subject:** Remedial Mathematics (BP106RMT) **Day & Date:** Tuesday, 01/10/19

**Class:** First Year B. Pharm

**Semester:** I/III\*

**Time:** 01-30 pm to 03.00 pm

**Max. Marks:** 30

- Instructions:*
1. All questions are compulsory
  2. Draw a well labeled diagram wherever necessary
  3. Right hand side number indicates full marks

**QI Attempt any two of the following**

(5+5=10)

(1). Find the value of x if  $\begin{vmatrix} 1 & 2 & 3 \\ 1 & x & 4 \\ 2 & -3 & -1 \end{vmatrix} = 0$ . x=3

(2). Find the value of  $\begin{vmatrix} 2 & -3 & 4 \\ 6 & 1 & -2 \\ 0 & -2 & -6 \end{vmatrix} = -176$

(3). Find inverse of  $A = \begin{bmatrix} 7 & -3 & 2 \\ 4 & 0 & 1 \\ 1 & 2 & 3 \end{bmatrix}$

**QII Attempt any five of the following**

(4×5=20)

(1). Define the following terms with example .

- (i) Diagonal Matrix. (ii) Symmetric Matrix  
(iii) Skew symmetric Matrix. (iv) Transpose of a matrix.

$$D_1 = \begin{vmatrix} 1 & 0 & -1 \\ 2 & 9 & 3 \\ 1 & 2 & 1 \end{vmatrix} = 3 - (4 - 9) = 3 - (-5) = 8$$

(2). Define the following terms with example

- (i) Singular Matrix (ii) Non singular Matrix.  
(iii) Transpose of a matrix. (iv) Identity Matrix.

$$D_2 = \begin{vmatrix} 1 & 1 & 0 \\ 2 & 1 & 9 \\ 1 & -1 & 2 \end{vmatrix} = 11 - (-5) = 16$$

(3). Find minors and cofactors of each elements of  $A = \begin{bmatrix} 0 & 1 & 2 \\ 3 & -1 & 6 \\ 2 & -4 & -8 \end{bmatrix}$

(4). Solve by Cramer's rule  $x + y - z = 0$  D = 8

$2x + y + 3z = 9$  D<sub>x</sub> = 8 x = 1

$x - y + z = 2$  D<sub>y</sub> = 8 y = 1

D<sub>z</sub> = 16 z = 2

(5) If  $A = \begin{bmatrix} 3 & 1 & 1 \\ -1 & 5 & -1 \\ -1 & -1 & 3 \end{bmatrix}$  And  $B = \begin{bmatrix} 1 & 3 & 6 \\ 1 & 4 & 5 \\ 1 & 5 & 4 \end{bmatrix}$  Verify that  $(A+B)^T = A^T + B^T$

(6). Find The value of Determinant Of The Matrix  $\begin{bmatrix} 2 & 5 & 7 \\ 8 & 1 & 0 \\ 12 & 11 & 14 \end{bmatrix} = 0$

(7). Verify that  $A(B+C) = AB+AC$ . Where  $A = \begin{bmatrix} 1 & 2 \\ 3 & -1 \end{bmatrix}$ ,  $B = \begin{bmatrix} -1 & 2 \\ 1 & 0 \end{bmatrix}$ ,  $C = \begin{bmatrix} 2 & -3 \\ 4 & -8 \end{bmatrix}$ .



**First Sessional Theory Examination 2021-2022**  
(Odd SEM)

**Subject & Subject Code: Remedial Mathematics (BP106RMT)**

**Day & Date: Saturday, 12/02/2022**

**Class: F.Y/S.Y. B. Pharmacy**

**Semester: I/III**

**Time: 1 hr**

**Max. Marks: 30**

- Instructions: 1. All questions are compulsory*  
*2. Draw a well labeled diagram wherever necessary*  
*3. Right hand side number indicates full marks*

**QI Multiple Choice Questions (MCQs)**

10

**1. If  $5^3 = 125$  Then its logarithmic form is**

- (a)  $\log_5 125 = 3$   
(b)  $\log_5 25 = 3$   
(c)  $\log_3 125 = 5$   
(d)  $\log_{125} 5 = 3$

**2. If  $A = \begin{bmatrix} 2 & 4 \\ X & 0 \end{bmatrix}$  is a symmetric matrix, then  $x =$**

- (a) 4  
(b) 3  
(c) -4  
(d) -3

**3. If  $A^T = -A$  Then A is**

- (a) diagonal matrix  
(b) skew-symmetric matrix  
(c) symmetric matrix  
(d) none of these

**4. If  $\log_{\left(\frac{1}{2}\right)}\left(\frac{1}{8}\right) = 3$  then its exponential form is**

- (a)  $\left(\frac{1}{3}\right)^3 = \frac{1}{27}$       (b)  $\left(\frac{1}{2}\right)^3 = 8$   
(c)  $\left(\frac{1}{2}\right)^3 = \frac{1}{8}$       (d)  $\left(\frac{1}{2}\right)^4 = \frac{1}{8}$   
(d)

**5. Let A is any square Matrix Then  $A^{-1}$  exist if**

- (a)  $|A| \neq 0$   
(b)  $|A| = 0$   
(c) A is singular Matrix  
(d) None

**6. The value of  $\begin{vmatrix} 1 & 2 & 3 \\ -3 & -6 & -9 \\ 2 & 1 & 0 \end{vmatrix}$  is**

- (a) 4  
(b) 3  
(c) -4  
(d) 0





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7.  $\lim \frac{x^2-5x+6}{x^2-4} =$

- (a)  $\frac{1}{8}$  (b)  $\frac{3}{4}$   
(c)  $\frac{1}{4}$  (d)  $-\frac{1}{4}$   
(d)  $-\frac{1}{4}$

8.  $\lim \frac{x^5-3^5}{x-3} =$

- (a) 410  
(b) 405  
(c) 415  
(d) 420

9. If  $\begin{vmatrix} 1 & 2 \\ 3 & x \end{vmatrix} = 0$  then x is

- (a) 2  
(b) 3  
(c) 6  
(d) 1

10 If  $A = \begin{bmatrix} 1 & 2 \\ 3 & 7 \end{bmatrix}$  then  $A^{-1}$  is

- (a)  $\begin{bmatrix} 7 & -2 \\ -3 & 1 \end{bmatrix}$   
(b)  $\begin{bmatrix} 1 & -2 \\ 3 & 7 \end{bmatrix}$   
(c)  $\begin{bmatrix} 1 & 2 \\ -3 & 7 \end{bmatrix}$   
(d)  $\begin{bmatrix} 1 & -2 \\ -3 & 7 \end{bmatrix}$

**QII Long Answers (Answer any 1 out of 2)**

10

1. Solve the following System of equations:  $x + y + z = 6$   
 $2x + 3y - z = 5$   
 $6x - 2y - 3z = -7$

Using crammers rule.

2. Solve the following System of equations:  $-x + 3y - 2z = 5$   
 $4x - y - 3z = -8$   
 $2x + 2y - 5z = 7$

Using crammers rule.

**QIII Short Answers (Answer any 2 out of 3)**

10

1. Resolve into the partial fraction:  $\frac{x^2-3x+1}{(x-1)^2(x-2)}$   
2. Resolve into the partial fraction:  $\frac{x-6}{(x^2+9)(x+2)}$

3. Prove that:  $2\log\left(\frac{15}{18}\right) - \log\left(\frac{25}{162}\right) + \log\left(\frac{4}{9}\right) = \log 2.$





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**First/Second Sessional Theory Examination-(ODD SEM) 2022-2023**

<b>Subject</b> :	<b>Remedial Mathematics(BP106RMT)</b>	<b>Day &amp; Date</b> :	Thursday, 09/03/2023	
<b>Class</b> :	First Year B. Pharmacy	<b>Semester</b> :	<b>I</b>	<i>Write Your Seat No. Here</i>
<b>Time</b> :	10:30 am – 11:30 am	<b>Max. Marks</b> :	30	<input type="text"/>

<b>Instructions:</b>	1. All questions are compulsory	2. Draw a well labeled diagram wherever necessary
	3. Right hand side number indicates full marks	4. Do not write/tick on the question paper

<b>QII</b>	<b>Long Answers Question (Answer any 1 out of 2)</b>	<b>10 M</b>
1.	Resolve into Partial Fraction $\frac{9x-7}{(x+3)(x^2+1)}$	(LL3, CO1)
2.	Using the adjoint Method, Find inverse of $A = \begin{bmatrix} 1 & -1 & 2 \\ 2 & -1 & 1 \\ 1 & 1 & -1 \end{bmatrix}$	(LL3, CO2)
<b>QIII</b>	<b>Short Answers Question (Answer any 4 out of 6)</b>	<b>20 M</b>
1.	Resolve into the partial fraction: $\frac{x^2-3x+1}{(x-1)^2(x-2)}$	(LL3, CO1)
2.	Find the value of x, if $\log(x+5)+\log(x-5)=4\log 2+2\log 3$	(LL3, CO1)
3.	Prove that : $2\log\left(\frac{15}{18}\right)-\log\left(\frac{25}{162}\right)+\log\left(\frac{4}{9}\right)=\log 2$	(LL3, CO1)
4.	Solve the following System of equations: $\begin{aligned} x + y + z &= 6 \\ 2x + 3y - z &= 5 \\ 6x - 2y - 3z &= -7 \end{aligned}$ Using crammers rule.	(LL3, CO2)
5.	Find Characteristics equation of $A = \begin{bmatrix} 8 & -6 & 2 \\ -6 & 7 & -4 \\ 2 & -4 & 3 \end{bmatrix}$	(LL3, CO2)
6.	Verify Cayley Hamilton Theorem for $A = \begin{bmatrix} 1 & 1 \\ 1 & 0 \end{bmatrix}$	(LL3, CO2)



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**First/Second Sessional Theory Examination-(ODD SEM) 2022-2023**

<b>Subject</b> :	<b>Remedial Mathematics(BP106RMT)</b>	<b>Day &amp; Date</b> :	Thursday, 09/03/2023	
<b>Class</b> :	First Year B. Pharmacy	<b>Semester</b> :	<b>I</b>	<i>Write Your Seat No. Here</i>
<b>Time</b> :	10:30 am – 11:30 am	<b>Max. Marks</b> :	30	<input type="text"/>

<b>Instructions:</b>	1. All questions are compulsory	2. Draw a well labeled diagram wherever necessary
	3. Right hand side number indicates full marks	4. Do not write/tick on the question paper

<b>QII</b>	<b>Long Answers Question (Answer any 1 out of 2)</b>	<b>10 M</b>
1.	Resolve into Partial Fraction $\frac{9x-7}{(x+3)(x^2+1)}$	(LL3, CO1)
2.	Using the adjoint Method, Find inverse of $A = \begin{bmatrix} 1 & -1 & 2 \\ 2 & -1 & 1 \\ 1 & 1 & -1 \end{bmatrix}$	(LL3, CO2)
<b>QIII</b>	<b>Short Answers Question (Answer any 4 out of 6)</b>	<b>20 M</b>
1.	Resolve into the partial fraction: $\frac{x^2-3x+1}{(x-1)^2(x-2)}$	(LL3, CO1)
2.	Find the value of x, if $\log(x+5)+\log(x-5)=4\log 2+2\log 3$	(LL3, CO1)
3.	Prove that : $2\log\left(\frac{15}{18}\right)-\log\left(\frac{25}{162}\right)+\log\left(\frac{4}{9}\right)=\log 2$	(LL3, CO1)
4.	Solve the following System of equations: $\begin{aligned} x + y + z &= 6 \\ 2x + 3y - z &= 5 \\ 6x - 2y - 3z &= -7 \end{aligned}$ Using crammers rule.	(LL3, CO2)
5.	Find Characteristics equation of $A = \begin{bmatrix} 8 & -6 & 2 \\ -6 & 7 & -4 \\ 2 & -4 & 3 \end{bmatrix}$	(LL3, CO2)
6.	Verify Cayley Hamilton Theorem for $A = \begin{bmatrix} 1 & 1 \\ 1 & 0 \end{bmatrix}$	(LL3, CO2)



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## First Sessional Theory Examination-(Odd SEM) 2023-2024

<b>Subject</b> : Remedial Mathematics	<b>Day &amp; Date</b> : Saturday, 4/11/2023
<b>Class</b> : First Year B. Pharmacy	<b>Semester</b> : I/III <span style="float: right;"><b>Write Your Seat No.</b></span>
<b>Time</b> : 10:30 am to 11:30 am	<b>Max. Marks</b> : 30 <span style="float: right;"><b>Here</b> <span style="border: 1px solid black; display: inline-block; width: 50px; height: 15px;"></span></span>

**Instructions:** 1. All questions are compulsory 2. Draw a well-labeled diagram wherever necessary  
3. Right-hand side number indicates full marks 4. Do not write/tick on the question paper

<b>Q. I: Long Answers Question (Answer any 1 out of 2)</b>	<b>10 M</b>
a) <b>Solve the equations using Cramer's Rule of determinants:</b> $x - y + z = 0, 2x - y + z = 1, x + y + z = 4$	
b) i) Find the value of x if: $\begin{vmatrix} 1 & 1 & 1 \\ 3 & x & 3 \\ 1 & x & 2 \end{vmatrix} = 0$	
ii) If $f(x) = \frac{x+5}{3x-4}$ and $t = \frac{5+4x}{3x-2}$ show that $f(t) = x$	
<b>Q. II: Short Answers Question (Answer any 4 out of 6)</b>	<b>20 M</b>
a) <b>Resolve into partial fraction</b> $\frac{13x+19}{(x+3)(x-2)(x+1)}$	
b) <b>Resolve into partial fraction</b> $\frac{1}{x^3+1}$	
c) <b>Express the following as a single logarithm using Properties of logarithms</b> $\log_2 4 + 2 \log_2 3 - 3 \log_2 2$	
d) <b>Find the value of <math>f(1/2)</math> &amp; <math>f(1/4)</math>, if <math>f(x) = 16^x - \log_2 x</math></b>	
e) i) $\lim_{x \rightarrow 0} \frac{3^x - 2^x}{\tan 4x}$ ii) $\lim_{x \rightarrow 0} \frac{3^{\tan x} - 1}{2x}$	
f) <b>Evaluate:</b> $\begin{matrix} 2 & 4 & 6 \\ 1 & 2 & 5 \\ 6 & -3 & 9 \end{matrix}$	



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**Second Sessional Theory Examination 2020-2021**

**Subject: Pharmaceutics I BP103T      Day & Date: Tuesday, 27<sup>th</sup> April 2021**

**Class: F Y B. Pharm                      Semester: I**

**Time: 90 min                                  Max. Marks: 30**

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*Instructions: 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks*

**Q.No.1 Multiple choice questions**

**(10 X 1=10)**

1. In the mixing of thymol and menthol the following type of incompatibility occurs,  
a. Chemical incompatibility                      b. Therapeutic incompatibility  
c. Physical incompatibility                      d. None of these.
2. Upward creaming is observed in?  
a. W/O b. O/W c. Both d. Micro
3. For ideal suspension , the sedimentation volume should be  
a. Zero b. Equal to one c. More than one d. Less than one
4. Which of the following is an example of Hydrophilic bases?  
a. Hydrogenated oils b. Emulsified cocoa butter c. Glycero-gelatin base d. None of these
5. In stokes' relationship, a parameter that greatly changes the velocity of settling is:  
a. Density of the liquid b. Density of particle c. Radius of the particle d. Viscosity of the medium
6. Structured vehicle is included in the formulation of a suspension, in order to:  
a. Decreases the interfacial tension                      b. Prevents the caking of the sediment  
c. Prevents the sedimentation of particles                      d. Reduces the size by chemical means
7. In the preparation of vanishing creams, which types of bases are used generally?  
a. Absorption bases b. Water removable bases c. Hydrocarbon bases d. None
8. The melting range of cocoa butter lies between  
a. 12 – 13°C b. 20 – 30°C c. 5 – 10°C d. 30 – 36° C

9. Tetracycline interacts with calcium to form chelates is an example of .....

- a. Chemical incompatibility
- b. Therapeutic incompatibility
- c. Physical incompatibility
- d. Tolerance incompatibility.

10. What is the difference between vanishing cream and cold cream?

- a. Both are examples of oil-in-water emulsions
- b. Vanishing cream is an oil-in-water emulsion whereas cold cream is a water-in-oil emulsion
- c. Vanishing cream is a water-in-oil emulsion whereas cold cream is an oil-in-water emulsion
- d. Both are examples of water-in-oil emulsions

**Q. No. 2 Long Answers (Answer 1 out of 2)**

**(1 x 10 = 10)**

- 1. Discuss stability problems of emulsion and methods to overcome.
- 2. Write in detail various suppositories bases used in the formulation of suppositories.

**Q. No. 3 Short Answers (Answer 2 out of 3)**

**(2 x 5 = 10)**

- 1. Define pharmaceutical incompatibility. Write in short about physical incompatibility.
- 2. Discuss various emulsifying agents.
- 3. Write short note on (any one)
  - a) Paste
  - b) Gels



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**Second Sessional Theory Examination-(ODD SEM) 2023-2024**

<b>Subject</b> : Pharmaceutical Inorganic Chemistry (BP104T)	<b>Day &amp; Date</b> : Thursday 21/12/2023	
<b>Class</b> : First Year B. Pharmacy	<b>Semester</b> : I	<b>Write Your Roll No. Here</b>
<b>Time</b> : 10:30 am – 12:00 pm	<b>Max. Marks</b> : 30	<input type="text"/>
<b>Instructions:</b> 1. All questions are compulsory 2. Draw a well-labeled diagram wherever necessary 3. Right-hand side number indicates full marks 4. Do not write/tick on the question paper		
<b>Q. I: Objective Type Questions</b>		<b>10 M</b>
1.	Define the following term. (CO3, CO4, LL3) i) Emetics ii) alpha particle decay	
2.	Draw the structure and uses of sodium potassium tartrate. (CO3, LL3)	
3.	Define hematinic with example. (CO3, LL3)	
4.	Write any four therapeutic application of radiopharmaceuticals. (CO4, LL3)	
5.	Define Antidote and enlist the types of antidote with example. (CO4, LL3)	
<b>Q. II: Long Answers Questions (Answer any 1 out of 2)</b>		<b>10 M</b>
1. Define radioactivity and describe in detail methods for measurement of radioactivity. (CO4, LL3)		
2. Write a note on a) Heavy metal poisoning b) Astringent with one example. (CO4, LL3)		
<b>Q. III: Short Answers Questions (Answer any 2 out of 3)</b>		<b>10 M</b>
1. Define and classify antimicrobial agents and describe M.W., M.F., properties, preparation and uses of hydrogen peroxide or boric acid. (CO3, LL3)		
2. Define and classify expectorants with example. (CO3, LL3)		
3. Describe and classify the cathartics based on their mechanism of action. (CO3, LL3)		

Prepared By

Checked By

Verified By





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**Second Resessional Theory Examination 2020-2021**

**Subject: Pharmaceutics I BP103T      Day & Date: Wednesday, 05<sup>th</sup> May 2021**

**Class: F Y B. Pharm                      Semester: I**

**Time: 90 min                                  Max. Marks: 30**

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*Instructions: 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks*

**Q.No.1 Multiple choice questions**

**(10 X 1=10)**

1. Which of the following is not used as an emulsifying agent?  
a. Surfactant b. Hydrophilic colloid c. Electrolytes d. Finely divided solids
2. Vaginal suppositories also called as  
a. Pessaries b. Simple suppositories c. Bougies d. None
3. ....is stable form of Cocoa butter.  
a.  $\alpha$ -form b.  $\beta$ -form c.  $\gamma$ -form d. All
4. The tail of a surfactant molecule is oil-loving or known as \_\_\_\_\_?  
a. Hydrophilic b. Lipophobic c. Lipophilic d. Hydrophilic
5. In case of coalescence \_\_\_\_\_  
a. Dispersed droplet does not fuse b. Globules size decrease  
c. No of globules increase d. Dispersed droplets tend to fuse
6. Upward creaming is observed in?  
a. W/O b. O/W c. Both d. Micro
7. Which of the following is an example of Synthetic base?  
a. Witespol b. Emulsified cocoa butter  
c. Glycero-gelatin base d. Hydrogenated oils
8. In the preparation of a structured vehicle, which one of the following substances is used?

a. bismuth subnitrate b. ethyl alcohol c. glycerin d. methyl cellulose

9. The value of sedimentation volume (F) gives a knowledge about the \_\_\_\_\_ of the suspension

a. physical stability b. Incompatibility c. Solubility d. All of the above

10. In Dry gum method, 4:2:1 consist of

a. 4 parts oil, 2 parts water, and 1 part emulsifier

b. 4 parts water, 2 parts oil, and 1 part emulsifier

c. 4 parts emulsifier, 2 parts water, and 1 part emulsifier

d. 4 parts oil, 2 parts emulsifier and 1 part water

**Q. No. 2 Long Answers (Answer 1 out of 2)**

**(1 x 10 = 10)**

1. Classify and discuss with advantages and disadvantages various semisolid bases used in the preparation of semisolid dosage form.

2. Define suppositories. Write down types and advantages and disadvantages of suppositories. Explain method of preparation.

**Q. No. 3 Short Answers (Answer 2 out of 3)**

**(2 x 5 = 10)**

1. Write down various test used for identification of emulsion.

2. Discuss physical incompatibility.

3. Write short note on emulsifying agents.



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### Second Sessional Theory Examination 2020-2021

**Subject: Comm. Skills**

**Day & Date: Friday, 23 April 2021**

**Class: FY/ DSA**

**Semester: First**

**Time: 10.30 to 12.00**

**Max. Marks: 30**

*Instructions: 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks*

#### **QI Long Answers (Answer any ONE)**

**1\*10**

- (1) Explain in detail how to make presentation effective w.r.t. four P's of effective presentation and how to deal with fear and structure the presentation. (Unit 04)
- (2) How to make Writing and Listening more effective? Enumerate the techniques and methods required to make Writing and Listening more productive. (Unit 03)

#### **QII Short Answers (Answer any FOUR)**

**4\*5**

1. Explain the Selection GD in detail with all its characteristics. (Unit 05)
2. What is Listening? Explain its sub-types with their purpose in detail. (Unit 03)
3. Explain the role of Team member and Leader wrt GD. (Unit 05)
4. What strategies will you adopt to contribute in Group Interaction? (Unit 02)
5. How will you prepare for Selection Interview? Explain the Interview process in detail. (Unit 04)
6. What are the Dos and Don'ts of Group Discussion? (Unit 05)

-----END-----



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**Second Sessional Theory Examination 2018-2019**

**Subject: Comm. Skills**

**Day & Date: Thursday, 22 Nov. 2018**

**Class: FY/ DSA**

**Semester: First**

**Time: 10.30 to 12.00**

**Max. Marks: 30**

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*Instructions: 1. All questions are compulsory*

*2. Draw a well labeled diagram wherever necessary*

*3. Right hand side number indicates full marks*

**QI. Long Answers (Answer any 1 out of 2)**

**10**

- (1) What is Group Discussion? State its objective, purpose & importance in selection procedure.
- (2) Explain effective Listening & Writing skills in detail.

**QII. Short Answers (Answer any 4 out of 6)**

**20**

1. Explain Do's and Don'ts of Interviews.
2. What are the 4Ps of Presentation? Explain them in detail.
3. Explain the role of 'know your audience' in Presentation & Writing effectively.
4. Explain the listening subtypes with techniques to improve Listening.
5. What is 'writing effectively'? Explain the elements involved in it.
6. State the importance of Body Language & Verbal Comm. in Interviews.



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**Second Sessional Theory Examination 2019-2020**

**Subject: Comm. Skills**

**Day & Date: Saturday, 09 Nov. 2019**

**Class: FY/ DSA**

**Semester: First**

**Time: 10.30 am to 12.00 pm**

**Max. Marks: 30**

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- Instructions: 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks*

**QI. Long Answers (Answer any ONE)**

**1\*10**

- (1) What is Group Discussion? State its objective, purpose & importance in selection procedure.
- (2) Explain Dr Eileen Russo's Communication Style Matrix in detail with different Communication Styles.

**QII. Short Answers (Answer any FOUR)**

**4\*5**

1. Explain Do's and Don'ts of Interviews.
2. What are the 4Ps of Presentation? Explain them in detail.
3. Explain the role of 'know your audience' in Presentation & Writing effectively.
4. Explain the listening subtypes with techniques to improve Listening.
5. What are the Perspectives? Explain their importance in Communication.
6. State the importance of Body Language & Verbal Comm. in Interviews.

\*\*\*\*BEST OF LUCK\*\*\*\*





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**Second Sessional Theory Examination 2021-2022**

**Subject: Communication Skills (BP105T)**

**Day & Date: Thursday, 07 April 2022**

**Class: First Year and Direct Scnd Year B.Pharm**

**Semester: I & III**

**Time: 02.00PM to 03.00PM**

**Max. Marks: 30**

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*Instructions: 1. All questions are compulsory*

*2. Draw a well labeled diagram wherever necessary*

*3. Right hand side number indicates full marks*

**QI Long Answers (Answer any ONE) (CO3)**

**1\*10**

- (1) Explain in detail how to make presentation effective w.r.t. four P's of effective presentation.  
(Unit 04)
- (2) Enumerate the factors that assist to prepare and appear for Interview. (Unit 04)

**QII Short Answer Questions**

**Section A: (Any One) (CO3)**

**1\*5**

1. Explain the role on non-verbals in Selection Interview. (Unit 04)
2. How will you deal with performance anxiety during Interviews and Presentation? (Unit 04)

**Section B: (Any One) (CO4)**

**3\*5**

1. Explain the Selection GD in detail with all its characteristics. (Unit 05)
2. Explain the role of Team member and Leader wrt GD. (Unit 05)
3. What strategies will you adopt to contribute in Group Interaction? (Unit 05)
4. What are the Dos and Don'ts of Group Discussion? (Unit 05)

-----END-----





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**Mapping of Course Outcome with First Sessional Theory Examination (AY 2021-22)**

**Subject: Communication Skills**

<b>Course Outcome</b>	<b>Questions</b>	<b>Marks</b>
<b>CO 1053:</b> Identify interview-process and effective presentation. ( <i>Level: 1</i> )	<b>QI: 1, 2, 3, 4</b>	<b>15</b>
<b>CO 1054:</b> Recognize contribution in Group Discussion. ( <i>Level: 1</i> )	<b>QII: 1, 2, 3, 4</b>	<b>15</b>

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**Second Sessional Theory Examination-(ODD SEM) 2022-2023**

<b>Subject</b> :	Communication skills (BP105T)	<b>Day &amp; Date</b> :	Friday, 21/04/2023	
<b>Class</b> :	First Year B. Pharmacy & DSA	<b>Semester</b> :	<b>I</b>	<b>Write Your Seat No. Here</b>
<b>Time</b> :	10:30 am – 11:30 am	<b>Max. Marks</b> :	30	<input type="text"/>

<b>Instructions:</b>	1. All questions are compulsory	2. Draw a well labeled diagram wherever necessary
	3. Right hand side number indicates full marks	4. Do not write/ tick on the question paper

**QI Long Answers Question (Answer any 1 out of 2) 10 M**

1. Illustrate how to deliver and design an effective presentation. (LL3, CO3)
2. Explain interview process with your ways to prepare and appear of an interview. (LL3, CO3)

**QII Short Answers Question**

**Section – A (Answer any ONE) 05 M**

1. Illustrate Do's and Don'ts of an interview. (LL3, CO3)
2. Interpret your ways to cope-up with stage fear and anxiety during presentation. (LL3, CO3)

**Section – A (Answer any Three) 15 M**

1. Interpret the role of Non-verbal communication during Group Discussion. (LL3, CO4)
2. Explain your methods of contribution in Group Discussion. (LL3, CO4)
3. Illustrate Do's and Don'ts of Group Discussion. (LL3, CO4)
4. Summarize the significance of tone, voice and language in Group Discussion. (LL3, CO4)



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**Second Sessional Theory Examination-(ODD SEM) 2023-2024**

<b>Subject</b>	: Communication skills (BP105T)	<b>Day &amp; Date</b>	: FRIDAY, 22/12/2023	
<b>Class</b>	: First Year B. Pharmacy	<b>Semester</b>	: I	<b>Write Your Seat No. Here</b> <input type="text"/>
<b>Time</b>	: 10:30 a.m. – 11:30 a.m.	<b>Max. Marks</b>	: 30	

<b>Instructions:</b>	1. All questions are compulsory	2. Draw a well labeled diagram wherever necessary
	3. Right hand side number indicates full marks	4. Do not write/tick on the question paper

<b>QI</b>	<b>Long Answers Question (Answer any 1 out of 2)</b>	<b>10 M</b>
1.	Define self-awareness. Explain in detail the ways of becoming an active listener. (LL2, CO4)	
2.	State the purpose of an interview and enlist the do's and don'ts of facing the interview. (LL2, CO4)	
<b>QII</b>	<b>Short Answers Question</b>	
	<b>Section – A (Answer any ONE)</b>	<b>05 M</b>
1.	Defend the following sentence: Knowing your audience/ readers is of prime importance in order to make your writing effective. (LL2, CO4)	
2.	Suggest the remedies of dealing with stage fright/ performance anxiety. (LL2, CO4)	
<b>QII</b>	<b>Section – B (Answer any Three)</b>	<b>15 M</b>
1.	Write a short note on listening in difficult situations. (LL2, CO3)	
2.	Explain in brief about the techniques of presentation delivery. (LL2, CO3)	
3.	Enlist the tips for framing a good e-mail subject line. (LL2, CO3)	
4.	Illustrate the concept of group discussion. State do's and don'ts of GD (two each). (LL3, CO3)	



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### Second Sessional Theory Examination 2020-2021

(Odd SEM)

Subject & Subject Code: Human Anatomy & Physiology I (BP101T)

4Day & Date: Monday, 26.04.2021

Class: First Year B. Pharm

Semester: I

Time: 10.30 AM to 12.30PM

Max. Marks: 30

Instructions: 1. All questions are compulsory.

2. Draw a well labelled diagram wherever necessary.

3. Right hand side number indicates full marks.

#### Q.I Multiple-choice questions (MCQs)

10

1. The temperature of blood is ----

- a. 35<sup>0</sup>C                      b. 36<sup>0</sup>C                      c. 37<sup>0</sup>C                      d. 38<sup>0</sup>C

2. Blood is about ---- % formed elements and --- % blood plasma respectively

- a. 75 & 25                      b. 50 & 50                      c. 45 & 55                      d. 60 & 40

3. ----- contributes to higher hematocrits in males

- a. Testosterone                      b. Estrogen                      c. Progesterone                      d. Prolactin

4. About 90 per cent of all erythropoietin is formed in the ----

- a. Liver                      b. Kidney                      c. Lungs                      d. Heart

5. Tissues that lack lymphatic capillaries include avascular tissues such as ---

- a. Liver                      b. Kidney                      c. Red bone marrow                      d. Lungs

6. The superior vena cava enters the \_\_\_\_\_ of the heart.

- a. Right ventricle                      b. Right Atrium                      c. Left atrium                      d. Left ventricle



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7. Which of the following contains blood with a relatively low oxygen content?  
a. Pulmonary artery    b. Pulmonary vein    c. Left ventricle    d. Aorta
8. The time taken for an impulse to travel from the S-A to the A-V node is evidenced in the ---  
a. QRS complex    b. S-T interval    c. P-Q interval    d. QRS-T interval
9. The amount of blood ejected by the left ventricle each minute is calculated from which formula?  
a. Heart rate + stroke volume    b. Heart rate X stroke volume  
c. Heart rate divided by stroke volume    d. Heart rate - stroke volume
10. In an ECG, where does the depolarization of the atria occur?  
a. P wave    b. QRS complex    c. T wave    d. U wave

**Q.II Long Answers (Answer any 1 out of 2) 10**

1. Draw a neat well labelled diagram of heart. Explain the structure and function of heart.
2. Write the functions of blood. Explain about the ABO system of blood group.

**Q.III Short Answers (Answers any 2 out of 3) 10**

1. Differentiate between parasympathetic and sympathetic nervous system
2. Write the location and functions of cranial nerves.
3. Draw a neat well labelled diagram of spinal cord. Explain about mechanism of reflex arc.

\*\*\* END\*\*\*



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**Second Sessional Theory Examination 2018-2019**

**Subject: Human Anatomy and Physiology-I**

**Day & Date:**

**Class: F. Y. B. Pharma**

**Semester: I**

**Time: 11:00-12:30**

**Max. Marks: 30**

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*Instructions: 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks*

**QI Write a short note on (Any five) 10**

1. Classify WBCs
2. Write a note on heart valves
3. Define the following term a) Blood pressure b) Heart rate
4. Write the composition of blood
5. Differentiate between artery and vein
6. Write a note on various types of anemia

**QII Long Answers (Answer any 1 out of 2) 10**

1. Explain in detail ABO and RH blood grouping and its significance. Enlist various coagulation factors.
2. Draw a neat-labelled diagram of internal structure of eye. Add a note on physiology of vision

**QIII Short Answers (Answer any 2 out of 3) 10**

1. Write a note on ECG with neat labelled diagram
2. Explain blood coagulation pathway.
3. Write the anatomy and physiology of RBCs.





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### Second Sessional Theory Examination 2020-2021

(Odd SEM)

Subject & Subject Code: Human Anatomy & Physiology I (BP101T)

4Day & Date: Monday, 26.04.2021

Class: First Year B. Pharm

Semester: I

Time: 10.30 AM to 12.30PM

Max. Marks: 30

Instructions: 1. All questions are compulsory.

2. Draw a well labelled diagram wherever necessary.

3. Right hand side number indicates full marks.

#### Q.I Multiple-choice questions (MCQs)

10

1. The temperature of blood is ----

- a. 35°C                      b. 36°C                      c. 37°C                      d. 38°C

2. Blood is about ---- % formed elements and --- % blood plasma respectively

- a. 75 & 25                      b. 50 & 50                      c. 45 & 55                      d. 60 & 40

3. ----- contributes to higher hematocrits in males

- a. Testosterone                      b. Estrogen                      c. Progesterone                      d. Prolactin

4. About 90 per cent of all erythropoietin is formed in the ----

- a. Liver                      b. Kidney                      c. Lungs                      d. Heart

5. Tissues that lack lymphatic capillaries include avascular tissues such as ---

- a. Liver                      b. Kidney                      c. Red bone marrow                      d. Lungs

6. The superior vena cava enters the \_\_\_\_\_ of the heart.

- a. Right ventricle                      b. Right Atrium                      c. Left atrium                      d. Left ventricle

*Shikha*  
ms. S. Lerikar A.K.  
(Subject In-charge)



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7. Which of the following contains blood with a relatively low oxygen content?  
a. Pulmonary artery    b. Pulmonary vein    c. Left ventricle    d. Aorta
8. The time taken for an impulse to travel from the S-A to the A-V node is evidenced in the ---  
a. QRS complex    b. S-T interval    c. P-Q interval    d. QRS-T interval
9. The amount of blood ejected by the left ventricle each minute is calculated from which formula?  
a. Heart rate + stroke volume    b. Heart rate X stroke volume  
c. Heart rate divided by stroke volume    d. Heart rate - stroke volume
10. In an ECG, where does the depolarization of the atria occur?  
a. P wave    b. QRS complex    c. T wave    d. U wave

### Q.II Long Answers (Answer any 1 out of 2)

10

1. Draw a neat well labelled diagram of heart. Explain the structure and function of heart.
2. Write the functions of blood. Explain about the ABO system of blood group.

### Q.III Short Answers (Answers any 2 out of 3)

10

1. Differentiate between parasympathetic and sympathetic nervous system
2. Write the location and functions of cranial nerves.
3. Draw a neat well labelled diagram of spinal cord. Explain about mechanism of reflex arc.

*Sherika*

Mr. Sherikar A.K.

(Subject In-charge)

\*\*\* END\*\*\*



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### Second Sessional Theory Examination 2020-2021

(Odd SEM)

Subject & Subject Code: Human Anatomy & Physiology I (BP101T)

4Day & Date: Monday, 26.04.2021

Class: First Year B. Pharm

Semester: I

Time: 10.30 AM to 12.30PM

Max. Marks: 30

Instructions: 1. All questions are compulsory.

2. Draw a well labelled diagram wherever necessary.

3. Right hand side number indicates full marks.

#### Q.I Multiple-choice questions (MCQs)

10

1. The temperature of blood is ----

- a. 35°C                      b. 36°C                      c. 37°C                      d. 38°C

2. Blood is about ---- % formed elements and --- % blood plasma respectively

- a. 75 & 25                      b. 50 & 50                      c. 45 & 55                      d. 60 & 40

3. ----- contributes to higher hematocrits in males

- a. Testosterone                      b. Estrogen                      c. Progesterone                      d. Prolactin

4. About 90 per cent of all erythropoietin is formed in the ----

- a. Liver                      b. Kidney                      c. Lungs                      d. Heart

5. Tissues that lack lymphatic capillaries include avascular tissues such as ---

- a. Liver                      b. Kidney                      c. Red bone marrow                      d. Lungs

6. The superior vena cava enters the \_\_\_\_\_ of the heart.

- a. Right ventricle                      b. Right Atrium                      c. Left atrium                      d. Left ventricle

*Shikha*  
ms. S. Lerikar A.K.  
(Subject In-charge)



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7. Which of the following contains blood with a relatively low oxygen content?  
a. Pulmonary artery    b. Pulmonary vein    c. Left ventricle    d. Aorta
8. The time taken for an impulse to travel from the S-A to the A-V node is evidenced in the ---  
a. QRS complex    b. S-T interval    c. P-Q interval    d. QRS-T interval
9. The amount of blood ejected by the left ventricle each minute is calculated from which formula?  
a. Heart rate + stroke volume    b. Heart rate X stroke volume  
c. Heart rate divided by stroke volume    d. Heart rate - stroke volume
10. In an ECG, where does the depolarization of the atria occur?  
a. P wave    b. QRS complex    c. T wave    d. U wave

### Q.II Long Answers (Answer any 1 out of 2)

10

1. Draw a neat well labelled diagram of heart. Explain the structure and function of heart.
2. Write the functions of blood. Explain about the ABO system of blood group.

### Q.III Short Answers (Answers any 2 out of 3)

10

1. Differentiate between parasympathetic and sympathetic nervous system
2. Write the location and functions of cranial nerves.
3. Draw a neat well labelled diagram of spinal cord. Explain about mechanism of reflex arc.

*Sherika*

Mr. Sherikar A.K.

(Subject In-charge)

\*\*\* END\*\*\*





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### Second Sessional Theory Examination 2021-2022 (Odd SEM)

Subject & Subject Code: Human Anatomy & Physiology I (BP101T)

Day & Date: Monday, 4.04.2022

Class: First Year B. Pharm

Semester: I

Time: 2:00 PM to 3.30PM

Max. Marks: 30

Instructions: 1. All questions are compulsory.

2. Draw a well labelled diagram wherever necessary.

3. Right hand side number indicates full marks.

#### Q.I Multiple-choice questions (MCQs)

10

- The components first line defence innate immunity is  
a. Fever      b. NK cells      c. Inflammation      d. skin
- In skin, the lymphatic vessels lie in ----- tissue  
a. Subcutaneous      b. Epidermis      c. Dermis      d. Keratin
- The membrane that surrounds and protects the heart is the –  
a. Myocardium      b. Endocardium      c. Epicardium      d. Pericardium
- is also called the right atrioventricular valve  
a. Bicuspid      b. Tricuspid      c. Mitral      d. Semilunar valve
- The --- surface of heart is deep to the sternum and ribs  
a. Inferior      b. Anterior      c. Right border      d. Left border
- All somatic motor neurons release only ---- as their neurotransmitter  
a. Norepinephrine      b. Epinephrine      c. Acetylcholine      d. Both a and b
- The cell bodies of preganglionic neurons are located in ---  
a. Brain      b. Autonomic ganglia      c. Visceral organs      d. Vertebral column
- is the largest autonomic plexus  
a. Superior mesenteric      b. Inferior mesenteric      c. Renal      d. Celiac
- There are ---- pairs of spinal nerves  
a. 12      b. 21      c. 31      d. 13
- Somatic nervous system consists of ---- axons  
a. Single & myelinated      b. Multiple and myelinated  
c. Single & unmyelinated      d. Multiple & unmyelinated

#### QII Long Answers (Answer any 1 out of 2)

10

- Explain the functions of lymphatic system. Explain the structure and functions of lymph nodes.
- Draw a neat labelled diagram of heart. Explain the anatomical structure of heart.

#### QIII Short Answers (Answer any 2 out of 3)

10

- Differentiate between sympathetic and parasympathetic nervous system.
- Write the number, name, location and function of cranial nerves.
- Draw a neat labelled diagram of eye. Explain the physiology of vision.



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### Second Sessional Theory Examination 2021-2022 (Odd SEM)

Subject & Subject Code: Human Anatomy & Physiology I (BP101T)

Day & Date: Monday, 4.04.2022

Class: First Year B. Pharm

Semester: I

Time: 2:00 PM to 3.30PM

Max. Marks: 30

Instructions: 1. All questions are compulsory.

2. Draw a well labelled diagram wherever necessary.

3. Right hand side number indicates full marks.

#### Q.I Multiple-choice questions (MCQs)

10

- The components first line defence innate immunity is  
a. Fever      b. NK cells      c. Inflammation      d. skin
- In skin, the lymphatic vessels lie in ----- tissue  
a. Subcutaneous      b. Epidermis      c. Dermis      d. Keratin
- The membrane that surrounds and protects the heart is the –  
a. Myocardium      b. Endocardium      c. Epicardium      d. Pericardium
- is also called the right atrioventricular valve  
a. Bicuspid      b. Tricuspid      c. Mitral      d. Semilunar valve
- The --- surface of heart is deep to the sternum and ribs  
a. Inferior      b. Anterior      c. Right border      d. Left border
- All somatic motor neurons release only ---- as their neurotransmitter  
a. Norepinephrine      b. Epinephrine      c. Acetylcholine      d. Both a and b
- The cell bodies of preganglionic neurons are located in ---  
a. Brain      b. Autonomic ganglia      c. Visceral organs      d. Vertebral column
- is the largest autonomic plexus  
a. Superior mesenteric      b. Inferior mesenteric      c. Renal      d. Celiac
- There are ---- pairs of spinal nerves  
a. 12      b. 21      c. 31      d. 13
- Somatic nervous system consists of ---- axons  
a. Single & myelinated      b. Multiple and myelinated  
c. Single & unmyelinated      d. Multiple & unmyelinated

#### QII Long Answers (Answer any 1 out of 2)

10

- Explain the functions of lymphatic system. Explain the structure and functions of lymph nodes.
- Draw a neat labelled diagram of heart. Explain the anatomical structure of heart.

#### QIII Short Answers (Answer any 2 out of 3)

10

- Differentiate between sympathetic and parasympathetic nervous system.
- Write the number, name, location and function of cranial nerves.
- Draw a neat labelled diagram of eye. Explain the physiology of vision.





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**Second Sessional Theory Examination-(ODD SEM) 2022-2023**

<b>Subject</b> : Human Anatomy & Physiology - I (BP101T)	<b>Day &amp; Date</b> : Monday, 17/04/2023
<b>Class</b> : First Year B. Pharmacy	<b>Semester</b> : I
<b>Time</b> : 10.30 am to 12.00 pm	<b>Max. Marks</b> : 30
	<b>Write Your Seat No. Here</b> <input type="text"/>

**Instructions:** 1. All questions are compulsory  
2. Draw a well-labeled diagram wherever necessary  
3. Right-hand side number indicates full marks  
4. Do not write/tick on the question paper

**Q. I: Objective Type Questions (5 X 2)**

**10 M**

1. Write about structure of lymph node. (CO3, LL3)
2. Draw a neat and well labelled diagram of interior of heart. (CO3, LL3)
3. Write location and functions of heart valves. (CO3, LL3)
4. Draw a neat and well labelled diagram of eye. (CO2, LL2)
5. Differentiate between somatic and autonomic motor neurons. (CO2, LL2)

**Q. II: Long Answers Questions (Answer any 1 out of 2)**

**10 M**

1. Write in detail about name, number, nature, location and functions of cranial nerves. (CO2, LL2)
2. Write in detail about comparison of sympathetic and parasympathetic divisions of autonomic nervous system. (CO2, LL2)

**Q. III: Short Answers Questions (Answer any 2 out of 3)**

**10 M**

1. Explain the composition, formation and circulation of lymph. (CO3, LL3)
2. Write about electrocardiogram. (CO3, LL3)
3. Explain about cardiac cycle. (CO3, LL3)



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## Second Sessional Theory Examination - (Odd SEM) 2023-2024

<b>Subject</b> : Human Anatomy & Physiology - I (BP101T)	<b>Day &amp; Date</b> : Monday, 18/12/2023
<b>Class</b> : First Year B. Pharmacy	<b>Semester</b> : I
<b>Time</b> : 10.30 am to 12.00 pm	<b>Max. Marks</b> : 30
	<b>Write Your Seat No. Here</b> <input type="text"/>

**Instructions:** 1. All questions are compulsory  
2. Draw a well-labeled diagram wherever necessary  
3. Right-hand side number indicates full marks  
4. Do not write/tick on the question paper

### Q. I: Objective Type Questions (5 X 2)

10 M

1. Write the functions of lymphatic system (CO4, LL2)
2. Draw diagram of spleen. Enlist the functions of spleen (CO4, LL2)
3. Answer the followings
  - a. Enlist the phases of cardiac muscle action potential (CO4, LL2)
  - b. Enlist the functions of ear (CO3, LL2)
4. Enlist the functions of nose (CO3, LL2)
5. Draw a neat and well-labelled diagram of eye (CO3, LL2)

### Q. II: Long Answers Questions (Answer any 1 out of 2)

10 M

1. Draw a neat and well-labelled diagram of heart. Explain in detail about circulatory system of heart (CO4, LL2)
2. Explain about electrocardiogram and cardiac cycle (CO4, LL2)

### Q. III: Short Answers Questions (Answer any 2 out of 3)

10 M

1. Differentiate between sympathetic and parasympathetic nervous system. (CO3, LL2)
2. Write the name, number, location and functions of cranial nerves. (CO3, LL2)
3. Draw neat and well-labelled diagram of ear. Explain the mechanism of hearing. (CO3, LL2)



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**Institute of Pharmacy, Dhule**

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**Second Sessional Theory Examination 2019-2020**

**Subject : Pharmaceutical Analysis Day & Date : 13/11/2019**  
**(BP102T)**  
**Class : F. Y. B. Pharmacy Semester : I**  
**Time : 10:30 am to 12:00 noon Max. Marks : 30**

Instructions: 1. All questions are compulsory  
2. Draw a well labelled diagram wherever necessary  
3. Right hand side number indicates full marks

Q.1 Multiple Choice Questions [10]

- a In polarographic method DME is used as \_\_\_\_\_ electrode  
a. reference ~~b. indicator~~  
c. partial reference d. None of the above
- b Standard hydrogen electrode (SHE / NHE) is \_\_\_\_\_ electrode.  
a. Indicator ~~b. Reference~~  
c. Second kind d. None of the above
- c Nernst equation can be used to calculate \_\_\_\_\_  
a. Conductance b. Resistance  
~~c. Electrode potential~~ d. Residual current
- d In \_\_\_\_\_ titration salicylic acid is used as a specific indicator  
a. Redox ~~b. complexometric~~  
c. acid base d. precipitation
- e Potentiometry can be utilized in below type of reactions except \_\_\_\_\_  
a. Neutralization b. Redox  
c. Precipitation ~~d. Gravimetric analysis~~
- f \_\_\_\_\_ is not a self-indicator  
a. Potassium permanganate b. Ceric ammonium sulfate  
c. Iodine solution ~~d. Potassium dichromate~~
- g In gravimetric analysis the term 'digestion' carried out to obtain \_\_\_\_\_  
~~a. pure precipitate~~ b. small crystals  
c. impure precipitate d. a & c both
- h \_\_\_\_\_ are involved in Bromatometry type of titrations  
a.  $\text{Br}^-$  ~~b.  $\text{BrO}_4^-$~~   
c.  $\text{Br}_2$  d. All of the above
- i Resistance can be expressed as \_\_\_\_\_  
a. S b.  $\Omega^{-1}$   
~~c.  $\Omega$~~  d. a & b both
- j The difference between the residual current and limiting current is called as \_\_\_\_\_ current.  
~~a. migration~~ b. residual  
c. kinetic d. b & c both

Q.2. Answer the following short questions (any two) [10]

- a Write a short note on masking and demasking agents. Enlist the factors affecting selectivity of EDTA.
- b Describe the construction, working and application of DME.
- c Explain principle involved in various types of *iodine titration*.
- d Explain the principle and application of diazotization titration

Q.3. Answer in detail of following (any one) [10]

- a Explain the term *specific* and *molar conductance*. Describe in detail about various conductometric titration with suitable example.
- b Discuss the steps involved in gravimetric analysis? Comment on the differences between lyophilic and lyophobic sols obtained during precipitation
- c Explain the principle behind potentiometer. Write a construction and working of reference electrode and indicator electrode

\*\*\*\*\*



**Second Sessional Theory Examination 2020-2021**  
**(Odd SEM)**

**Subject:** Pharmaceutical Analysis I (BP102T)

**Day & Date:** 22-04-2021

**Class:** First Year B. Pharm

**Semester:** I

**Time:** 10:30 am – 12:00 pm

**Max. Marks:** 30

*Instructions: 1. All questions are compulsory*

*2. Draw a well labeled diagram wherever necessary*

*3. Right hand side number indicates full marks*

**Q1. Multiple Choice Questions (MCQs)**

**10**

1. Closeness of observed value with true value is .....
  - a) accuracy
  - b) linearity
  - c) precision
  - d) relative error
2. Given the following are the type of systemic error except
  - a) Error of methods
  - b) Instrumental method
  - c) Personal error
  - d) Random error
3. In which step, ions or elements are aggregated in Gravimetric analysis?
  - a) Supersaturation
  - b) Nucleation
  - c) Particle growth
  - d) None of the above
4. When more than two ions in solution and simultaneously second ion is also precipitated it is called as .....
  - a) Co precipitation
  - b) Post precipitation
  - c) A and B
  - d) None of the above
5. The reagent which undergoes reduction is an ..... agent and the reagent which undergoes oxidation is a .....
  - a) Oxidizing, Reducing
  - b) Reducing, Oxidizing
  - c) Complexing, Reducing
  - d) None of these
6. Which method are used to determination of primary amine?



- a) Diazotization Titration
  - b) Karl fischer titration
  - c) Redox titration
  - d) All of the above
7. Indicator used in standardization of sodium thiosulphate is
- a) Starch
  - b) Phenolphthaline
  - c) ceric sulphate
  - d) Methylene red
8. When an analyte that is a reducing agent is titrated directly with a standard iodine solution, the method is called
- a) iodimetry
  - b) iodometry
  - c) bromatometry
  - d) cerimetry
9. Complexing agent is .....
- a) Electron donating ion
  - b) Electron accepting ion
  - c) Both a & b
  - d) None of the above
- 10 The diffusion current in polarography is expressed by the equation
- a) Beer's law
  - b) Nernst equation
  - c) Bragg's equation
  - d) Ilkovic equation

**QII. Solve the following (any 1)**

**10**

1. Define gravimetric analysis. Explain in detail the steps involved in gravimetric analysis
2. State the basic principle involved in polarography. Write down the construction and working of dropping mercury electrode

**QIII Short Answers (Answer any 2 out of 3)**

**10**

1. Explain the titration involving iodine
2. Write down the construction and working of normal hydrogen electrode
3. Write principle and application of Diazotization titrations.





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**Second Sessional Theory Examination 2018-2019**

<b>Subject</b>	: <b>Pharmaceutical Analysis</b>	<b>Day &amp; Date</b>	: <b>17/11/2018</b>
<b>Class</b>	: <b>F. Y. B. Pharmacy</b>	<b>Semester</b>	: <b>I</b>
<b>Time</b>	: <b>10:30 am to 12:00 noon</b>	<b>Max. Marks</b>	: <b>30</b>

- Instructions:*
1. All questions are compulsory
  2. Draw a well labelled diagram wherever necessary
  3. Right hand side number indicates full marks

Q.1 Multiple Choice Questions [10]

- a Iodometric titration utilizes \_\_\_\_\_
- a. Iodine solution  
b. Liberated iodine  
c. Both a & b  
d. None of the above
- b Hydrogen electrode is \_\_\_\_\_ electrode.
- a. Indicator  
b. Reference  
c. Second kind  
d. None of the above
- c Mordant Black II mixture is used in \_\_\_\_\_ titration
- a. Precipitation  
b. Neutralization  
c. Redox  
d. Complexometric
- d EDTA is \_\_\_\_\_ type of ligand
- a. Unidentate  
b. Bidentate  
c. Tetradentate  
d. Hexadentate
- e DME stands for \_\_\_\_\_
- a. Diffusion mercury  
b. Dropping mercury electrode  
c. Diffusion measuring  
d. Direct mercury
- f \_\_\_\_\_ is self-indicator
- a. Potassium permanganate  
b. Ceric ammonium sulfate  
c. Iodine solution  
d. a & b both
- g In gravimetric analysis the term 'Ignition' refers to heating precipitate \_\_\_\_\_
- a. upto 250°  
b. 250-500°  
c. 250-1200°  
d. b & c both
- h \_\_\_\_\_ are involved in Bromatometry type of titrations
- a.  $\text{Br}^-$   
b.  $\text{BrO}_4^-$   
c.  $\text{Br}_2^\dagger$   
d. All of the above
- i Conductance can be expressed as \_\_\_\_\_
- a. S  
b.  $\Omega^{-1}$   
c.  $\Omega$   
d. a & b both
- j Diffusion current ( $i_d$ ) is defined as \_\_\_\_\_.
- a.  $i_i$   
b.  $i_r$   
c.  $i_i - i_r$   
d.  $i_m$

Q.2. Answer the following short questions (any two) [10]

- a What do you mean by permanganate titrations explain with suitable example?
- b Write short note on principle of *polarography*
- c Write down the construction and working of conductivity meter

Q.3. Answer in detail of following (any one) [10]

- a Explain the term *specific* and *molar conductance*. Describe in detail about various conductometric titration with suitable example.
- b Write down the construction and working of dropping mercury electrode involved in polarography.

\*\*\*\*\*



**Second Sessional Theory Examination 2018-2019**

<b>Subject</b>	: <b>Pharmaceutical Analysis-I</b>	<b>Day &amp; Date</b>	: <b>17/11/2018</b>
<b>Class</b>	: <b>F. Y. B. Pharmacy</b>	<b>Semester</b>	: <b>I</b>
<b>Time</b>	: <b>10:30 am to 12:00 noon</b>	<b>Max. Marks</b>	: <b>30</b>

*Instructions:* 1. All questions are compulsory  
2. Draw a well labelled diagram wherever necessary  
3. Right hand side number indicates full marks

Q.1 Multiple Choice Questions [10]

- a Iodometric titration utilizes \_\_\_\_\_
- a. Iodine solution ~~b.~~ Liberated iodine  
c. Both a & b d. None of the above
- b Standard hydrogen electrode (SHE / NHE) is \_\_\_\_\_ electrode.
- a. Indicator ~~b.~~ Reference  
c. Second kind d. None of the above
- c Mordant Black II mixture is used in \_\_\_\_\_ titration
- a. Precipitation b. Neutralization  
c. Redox ~~d.~~ Complexometric
- d EDTA is \_\_\_\_\_ type of ligand
- a. unidentate b. bidentate  
c. tetradentate ~~d.~~ hexadentate
- e Potentiometric titrations can be utilizes in \_\_\_\_\_ type of reactions.
- a. Neutralization b. Redox  
c. Precipitation ~~d.~~ All of the above
- f \_\_\_\_\_ is self-indicator
- a. Potassium permanganate ~~b.~~ Ceric ammonium sulfate  
c. Iodine solution ~~d.~~ a & b both
- g In gravimetric analysis the term 'drying' refers to heating precipitate \_\_\_\_\_
- ~~a.~~ upto 250° b. 250-500°  
c. 250-1200° d. b & c both
- h \_\_\_\_\_ are involved in Bromatometry type of titrations
- a. Br<sup>-</sup> b. BrO<sub>4</sub><sup>-</sup>  
c. Br<sub>2</sub><sup>†</sup> ~~d.~~ All of the above
- i Conductance can be expressed as \_\_\_\_\_
- a. S b. Ω<sup>-1</sup>  
c. Ω ~~d.~~ a & b both
- j Diffusion current (*i<sub>d</sub>*) is defined as \_\_\_\_\_.
- a. *i<sub>i</sub>* b. *i<sub>r</sub>*  
~~c.~~ *i<sub>i</sub>*-*i<sub>r</sub>* d. *i<sub>m</sub>*

Q.2. Answer the following short questions (any two) [10]

- a Explain the principle involved in iodimetric and iodometric titrations? Why starch indicator added towards the end point only in iodometric and not in iodimetric titration.
- b Write short note on principle of *polarography*
- c Explain the various types of *EDTA titrations*.
- d Define *any five* of the following:  
**i. Precipitation, ii. Co-precipitation, iii. Post precipitation, iv. Tyndall effect**  
**v. Peptisation, vi. Coagulation value**

Q.3. Answer in detail of following (any one) [10]

- a Explain the term *specific* and *molar conductance*. Describe in detail about various conductometric titration with suitable example.
- b Discuss the steps involved in gravimetric analysis? Comment on the differences between lyophilic and lyophobic sols obtained during precipitation



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Behind Gurudwara, Mumbai-Agra Highway, Dhule  
**Phone:** 02562- 297802, 297805 **Fax:** 02562- 287802, **Email:** examiop@svkm.ac.in

**Second Sessional Theory Examination 2020-2021**  
**(Odd SEM)**

**Subject:** Pharmaceutical Analysis I (BP102T)

**Day & Date:** 22-04-2021

**Class:** First Year B. Pharm

**Semester:** I

**Time:** 10:30 am – 12:00 pm

**Max. Marks:** 30

*Instructions: 1. All questions are compulsory*

*2. Draw a well labeled diagram wherever necessary*

*3. Right hand side number indicates full marks*

**QI. Multiple Choice Questions (MCQs)**

**10**

1. Closeness of observed value with true value is .....

  - a) accuracy
  - b) linearity
  - c) precision
  - d) relative error

2. Given the following are the type of systemic error except

  - a) Error of methods
  - b) Instrumental method
  - c) Personal error
  - d) Random error

3. In which step, ions or elements are aggregated in Gravimetric analysis?

  - a) Supersaturation
  - b) Nucleation
  - c) Particle growth
  - d) None of the above

4. When more than two ions in solution and simultaneously second ion is also precipitated it is called as .....

  - a) Co precipitation
  - b) Post precipitation
  - c) A and B
  - d) None of the above

5. The reagent which undergoes reduction is an ..... agent and the reagent which undergoes oxidation is a .....

  - a) Oxidizing, Reducing
  - b) Reducing, Oxidizing
  - c) Complexing, Reducing
  - d) None of these

6. Which method are used to determination of primary amine?



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- a) Diazotization Titration  
b) Karl fischer titration  
c) Redox titration  
d) All of the above
7. Indicator used in standardization of sodium thiosulphate is  
a) Starch  
b) Phenolphthaline  
c) ceric sulphate  
d) Methylene red
8. When an analyte that is a reducing agent is titrated directly with a standard iodine solution, the method is called  
a) iodimetry  
b) iodometry  
c) bromatometry  
d) cerimetry
9. Complexing agent is .....  
a) Electron donating ion  
b) Electron accepting ion  
c) Both a & b  
d) None of the above
- 10 The diffusion current in polarography is expressed by the equation  
a) Beer's law  
b) Nernst equation  
c) Bragg's equation  
d) Ilkovic equation

**QII. Solve the following (any 1)**

**10**

1. Define gravimetric analysis. Explain in detail the steps involved in gravimetric analysis
2. State the basic principle involved in polarography. Write down the construction and working of dropping mercury electrode

**QIII Short Answers (Answer any 2 out of 3)**

**10**

1. Explain the titration involving iodine
2. Write down the construction and working of normal hydrogen electrode
3. Write principle and application of Diazotization titrations.



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**Second Sessional Theory Examination 2021-2022 (Odd SEM)**

**Subject:** Pharmaceutical Analysis-I (BP102-T)

**Day & Date:** Friday 01/04/2022

**Class:** F.Y. B. Pharm

**Semester:** I

**Time:** 02.00 pm – 03.30 pm

**Max. Marks:** 30

- Instructions:* 1. All questions are compulsory  
2. Draw a well labelled diagram wherever necessary  
3. Right hand side number indicates full marks

**QI Multiple Choice Questions (MCQs)**

**10**

1. .... is a chelating agent  
a) Salicylic acid      b) Glycerol c) Benzoic acid d) EDTA
2. ----- is used as a titrant in conductometry  
a) Sodium thiosulphate      b) Oxalic acid c) Perchloric acid      d) None of these
3. The complexometric titration where EDTA is used carried out at basic pH. Why?  
a) For the stability of complex formed      b) Reaction rate is optimum in basic pH  
c) There is less number of side reaction      d) All
4. .... is the weight of analyte per unit weight of the precipitate.  
a) Gravimetric factor      b) Precipitation factor      c) Electrogravimetry factor d) None of the above
5. Following are properties of good precipitate except  
a) Easily filtered and washed free of contamination. b) Significant loss of the analyte occurs during filtration and washing. c) Unreactive with constituents of the atmosphere. d) both a and b
6. Which of the following is used as indicator electrode in polarography  
a) Glass      b) Silver      c) Dropping Mercury      d) platinum
7. SI unit of conductance is  
a) Mho      b) Seimens      c) Volt      d) None of the above
8. Example for reference electrode except  
a) Antimony electrode b) Silver-silver electrode c) Calomel electrode d) None of the above
9. Each electrochemical cell is composed of  
a) Two half cells b) Half cells c) Both A and B d) None of the above
- 10) For Qualitative analysis by polarography, characteristic parameter used is  
a) Diffusion current      b) Half wave potential      c) Voltage      d) None of the above

**QII Long Answers Question (Answer any 1 out of 2)**

**10**

1. What is redox titration? Enlist its type? Explain iodometry and iodometric titration?
2. Write a short note on conductometry titration?

**QIII Short Answers Question (Answer any 2 out of 3)**

**10**

1. Discuss diazotization titration?
2. What are masking and demasking agents?
3. Describe the construction and working of Dropping Mercury Electrode?



**Mapping of Course Outcome with Second Sessional Theory Examination (2021-2022)**

**Subject: Pharmaceutical Analysis-I**

<b>Course Outcome</b>	<b>Question</b>	<b>Marks</b>
<b>CO 1023:</b>	Q1 – 1,2,3,4,5	25
	Q2 – 1	
	Q3- 1, 2,	
<b>CO 1024:</b>	Q1 – 6,7,8,9,10	20
	Q2 -2	
	Q3 – 3	

Mohd Usman

**Prepared By**

Azim Ansari

**Verified By**

Girija Bhavar

**Approved By**





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**Second Sessional Theory Examination-(ODD SEM) 2023-2024**

<b>Subject</b>	: Pharmaceutical Analysis (BP102T)	<b>Day &amp; Date</b>	: Tuesday, 19/12/2023	
<b>Class</b>	: First Year B. Pharmacy	<b>Semester</b>	: I	<b>Write Your Seat No. Here</b>
<b>Time</b>	: 10:30 am – 12:00 pm	<b>Max. Marks</b>	: 30	<input type="text"/>

<b>Instructions:</b>	1. All questions are compulsory 3. Right hand side number indicates full marks	2. Draw a well labeled diagram wherever necessary 4. Do not write/tick on the question paper
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<b>QI</b>	<b>Objective Type Questions</b>	<b>10 M</b>
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1.	Define Masking and Demasking with its example (CO3, LL1)
----	--

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2.	Describe co-precipitation in gravimetric analysis (CO3, LL2)
----	--

--	--

3.	Define: Self Indicators and Half wave potential (CO3, LL1; CO4, LL1)
----	--

--	--

4.	Draw a well labelled diagram of Calomel electrode (CO4, LL1)
----	--

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5.	Write Ilkovic equation and explain all terms (CO4, LL3)
----	---

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<b>QII</b>	<b>Long Answers Question (Answer any 1 out of 2) (CO3, LL2)</b>	<b>10 M</b>
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1.	Explain in detail the principle involved in Iodometry and cerimetry titrations with its applications
----	--

2.	Explain the principle and steps involved in gravimetric analysis and describe post-precipitation.
----	---

<b>QIII</b>	<b>Short Answers Question (Answer any 2 out of 3)</b>	<b>10 M</b>
-------------	---	-------------

1.	Explain conductivity cells and write applications of conductometric titration (CO4, LL2)
----	--

2.	Enumerate the indicator electrodes used in potentiometric titration and explain the Glass electrode (CO4, LL1)
----	--

--	--

3.	Describe rotating dropping mercury electrode and its applications (CO4, LL1)
----	--





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**Second Sessional Theory Examination-(ODD SEM) 2022-2023**

<b>Subject</b> : Pharmaceutics-I (BP103T)	<b>Day &amp; Date</b> : Wednesday, 19/04/2023	<b>Write Your Seat No. Here</b> <input type="text"/>
<b>Class</b> : First Year B. Pharmacy	<b>Semester</b> : I	
<b>Time</b> : 10.30 pm to 12.00 noon	<b>Max. Marks</b> : 30	

**Instructions:**

1. All questions are compulsory
2. Draw a well-labeled diagram wherever necessary
3. Right-hand side number indicates full marks
4. Do not write/tick on the question paper

**Q. I: Objective Type Questions (5 X 2)**

**10 M**

- 1 Differentiate between flocculated and deflocculated Suspension (CO3, LL2)
- 2 State the ideal properties of suppositories bases (CO3, LL2)
- 3 Define displacement value and cream.(CO3 and CO4, LL1)
- 4 Enlist the factor affecting on dermal penetration of semisolid dosage forms.(CO4, LL1)
- 5 Define paste and Enlist the method of preparation of ointment(CO4, LL1)

**Q. II: Long Answers Questions (Answer any 1 out of 2)**

**10 M**

- 1 Define emulsion. Explain different methods of preparation and identification tests.(CO3, LL3)
- 2 Define the term pharmaceutical incompatibility and discuss physical, chemical, and therapeutic incompatibility with examples. (CO3, LL3)

**Q. III: Short Answers Questions (Answer any 2 out of 3)**

**10 M**

- 1 Discuss in brief about Jellies and add note on types of Jellies (CO4, LL3)
- 2 Differentiate between ointment and paste. Add note on ointment bases.(CO4, LL3)
- 3 Give the evaluation of semisolid dosage forms.(CO4, LL3)



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**Second Sessional Theory Examination-(ODD SEM) 2022-2023**

<b>Subject</b> : Pharmaceutics-I (BP103T)	<b>Day &amp; Date</b> : Wednesday, 19/04/2023	<b>Write Your Seat No. Here</b> <input type="text"/>
<b>Class</b> : First Year B. Pharmacy	<b>Semester</b> : I	
<b>Time</b> : 10.30 pm to 12.00 noon	<b>Max. Marks</b> : 30	

**Instructions:**

1. All questions are compulsory
2. Draw a well-labeled diagram wherever necessary
3. Right-hand side number indicates full marks
4. Do not write/tick on the question paper

**Q. I: Objective Type Questions (5 X 2)**

**10 M**

- 1 Differentiate between flocculated and deflocculated Suspension (CO3, LL2)
- 2 State the ideal properties of suppositories bases (CO3, LL2)
- 3 Define displacement value and cream.(CO3 and CO4, LL1)
- 4 Enlist the factor affecting on dermal penetration of semisolid dosage forms.(CO4, LL1)
- 5 Define paste and Enlist the method of preparation of ointment(CO4, LL1)

**Q. II: Long Answers Questions (Answer any 1 out of 2)**

**10 M**

- 1 Define emulsion. Explain different methods of preparation and identification tests.(CO3, LL3)
- 2 Define the term pharmaceutical incompatibility and discuss physical, chemical, and therapeutic incompatibility with examples. (CO3, LL3)

**Q. III: Short Answers Questions (Answer any 2 out of 3)**

**10 M**

- 1 Discuss in brief about Jellies and add note on types of Jellies (CO4, LL3)
- 2 Differentiate between ointment and paste. Add note on ointment bases.(CO4, LL3)
- 3 Give the evaluation of semisolid dosage forms.(CO4, LL3)

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Institute of Pharmacy, Dhule

Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

Second Sessional Theory Examination 2018-2019

Subject: Pharmaceutics I BP103T

Day & Date: 19.11.2018

Class: F Y B. Pharm

Semester: I

Time: 90 min

Max. Marks: 30

- Instructions:
1. All questions are compulsory
  2. Draw a well labeled diagram wherever necessary
  3. Right hand side number indicates full marks

Q.No.1 Multiple choice questions

(10 X 1=10)

1. The amount of alcohol in elixirs ranges from  
a) 4-40 % b) 76% c) 70% d) 5-50%
2. .... are liquid dosage forms and externally inserted into body cavity to promote defecation.  
a) Enemas b) Pessaries c) Suppositories d) Bougies
3. Regarding the rate of sedimentation of pharmaceutical suspensions designed for oral administration, which of the following statements are true?  
a. The rate of sedimentation is increased as the diameter of the dispersed drug particles is increased.  
b. The rate of sedimentation is increased as the viscosity of the continuous phase is increased.  
c. The rate of sedimentation is affected by the concentration of buffer salts.  
d. None of the above.
4. Bottle method is used for formulation of emulsion containing ..... oil.  
a) Fixed, b) Mineral, c) Volatile, d) None of the above.
5. The dosage form intended to be rubbed on the skin for rubeficient and counterirritant effect are called.....  
a) Lotions b) Liniments c) Gels d) Dentifrices
6. .... is used as lubricant for gelato-glycerin suppository bases.  
a) Glycerin, alcohol and liquid paraffin/ arachis oil, b) soap solution, c) mineral oil, d) None of the above.
7. .... form is most stable form of cocoa butter.  
a)  $\alpha$  b)  $\beta$  c)  $\gamma$  d)  $\sigma$
8. Calculate the displacement value of zinc oxide in theobroma oil suppositories containing 40% of zinc oxide prepared in a 1g mould (weight of 10 suppositories is 10g and weight of medicated suppositories is 14.72g).  
a) 5, d) 4, c) 6, d) 2
9. .... are used to deodorize oral cavity.  
a) mouth-washes, b) gargles, c) throat-paints, d) None of the above
10. Tetracycline taken with milk, is a type of ..... incompatibility.  
a) Therapeutic, b) Physical, c) Chemical, d) Adjusted.

Q. No. 2 Long Answers (Answer 1 out of 2)

(1 x 10 = 10)

1. Discuss stability problems of emulsion and methods to overcome.
2. Define suspension. Write down ideal properties of suspension. Write a note on formulation of suspension.

Q. No. 3 Short Answers (Answer 2 out of 3)

(2 x 5 = 10)

1. Define emulsion. Write down ideal properties of emulsion. Write down any four test used for identification emulsion.
2. Discuss various bases used in the formulation of suppositories.
3. Write short note on (any one)  
a) Elixir.  
b) Physical incompatibility.



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**IN- SEMESTER THEORY ASSESSMENT 2018 -2019**

**Class:** First Year B. Pharmacy

**Semester:** I

**Division:** A

**Subject:** Pharmaceutical Analysis-I

**Total Marks:** 05

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**Assessment Type:** Quiz

**Date:** 25-10-2018

1. Write the important feature of ethylene diamine tetra acetic acid which make it choice of chelating agent
- 

**Assessment Type:** Class Test

**Date:** 26-10-2018

1. Discuss in details the steps involved in gravimetric analysis. Add note on difference between lyophilic and lyophobic colloids.
- 

**Assessment Type:** Assignment

**Date:** 27-10-2018

1. Explain in detail the neutralization curves of various acid-base titration



**Second Sessional Theory Examination 2018-2019**

**Subject: Pharmaceutics I BP103T**

**Day & Date: 19.11.2018**

**Class: F Y B. Pharm**

**Semester: I**

**Time: 90 min**

**Max. Marks: 30**

- Instructions:**
1. All questions are compulsory
  2. Draw a well labeled diagram wherever necessary
  3. Right hand side number indicates full marks

**Q.No.1 Multiple choice questions**

**(10 X 1=10)**

1. .... are diluted with warm water before use.  
a) Gargles b) Throat paints c) Mouthwashes d) All of above
2. In suspensions , ..... are added for suspending insoluble powder in a vehicle.  
a) Suspending agents, b) emulsifying agents, c) surfactants, d) electrolytes
3. .... Is used as a flocculating agent.  
a) Tweens, b) Butylated hydroxyl anisole, c) glycerine, d) None of the above.
4. .... are semisolid unit dosage forms and externally inserted into body cavity to promote defecation.  
a) Enemas b) Pessaries c) Suppositories d) Bougies
5. Emulsifying agents which reduces interfacial tension are classified as.....  
a) Sufactants, b) Finely divided solids, c) Polysaccharides, d) Sterols
6. Regarding the rate of sedimentation of pharmaceutical suspensions designed for oral administration, which of the following statements are true?
  - i. The rate of sedimentation is increased as the diameter of the dispersed drug particles is increased.
  - ii. The rate of sedimentation is increased as the viscosity of the continuous phase is increased.
  - iii. The rate of sedimentation may be increased by centrifugation.
  - iv. The rate of sedimentation may be increased by decreasing density difference.
- a) i & iii b) i & ii, c) ii & iii, d) ii & iv
7. .... is used as o/w emulsifiers.  
a) Acacia, b) Span, c) Magnesium oxide, d) None of the above.
8. .... is used as lubricant for gelato-glycerin suppository bases.  
a) Glycerin, alcohol and liquid paraffin/ arachis oil, b) soap solution, c) mineral oil, d) None of the above.
9. .... form is most stable form of cocoa butter.  
a)  $\alpha$  b)  $\beta$  c)  $\gamma$  d)  $\sigma$
10. Tetracycline taken with milk, is a type of ..... incompatibility.  
a) Therapeutic, b) Physical, c) Chemical, d) Adjusted.

**Q. No. 2 Long Answers (Answer 1 out of 2)**

**(1 x 10 = 10)**

3. Define emulsion. Write down ideal properties of emulsion. Write a note on emulsifying agents.
4. Define suppositories. Write down types and advantages and disadvantages of suppositories. Explain method of preparation.

**Q. No. 3 Short Answers (Answer 2 out of 3)**

**(2 x 5 = 10)**

1. Define suspension. Write down ideal properties of suspensions.
2. Discuss therapeutic incompatibility.
3. Write short note on Syrup.

**Second Sessional Theory Examination 2018-2019**

**Subject: Pharmaceutics I BP103T**

**Day & Date: 19.11.2018**

**Class: F Y B. Pharm**

**Semester: I**

**Time: 90 min**

**Max. Marks: 30**

- Instructions:**
1. All questions are compulsory
  2. Draw a well labeled diagram wherever necessary
  3. Right hand side number indicates full marks

**Q.No.1 Multiple choice questions**

**(10 X 1=10)**

1. .... are diluted with warm water before use.  
a) Gargles b) Throat paints c) Mouthwashes d) All of above
2. In suspensions , ..... are added for suspending insoluble powder in a vehicle.  
a) Suspending agents, b) emulsifying agents, c) surfactants, d) electrolytes
3. .... Is used as a flocculating agent.  
a) Tweens, b) Butylated hydroxyl anisole, c) glycerine, d) None of the above.
4. .... are semisolid unit dosage forms and externally inserted into body cavity to promote defecation.  
a) Enemas b) Pessaries c) Suppositories d) Bougies
5. Emulsifying agents which reduces interfacial tension are classified as.....  
a) Sufactants, b) Finely divided solids, c) Polysaccharides, d) Sterols
6. Regarding the rate of sedimentation of pharmaceutical suspensions designed for oral administration, which of the following statements are true?  
i. The rate of sedimentation is increased as the diameter of the dispersed drug particles is increased.  
ii. The rate of sedimentation is increased as the viscosity of the continuous phase is increased.  
iii. The rate of sedimentation may be increased by centrifugation.  
iv. The rate of sedimentation may be increased by decreasing density difference.  
a) i & iii b) i & ii, c) ii & iii, d) ii & iv
7. .... is used as o/w emulsifiers.  
a) Acacia, b) Span, c) Magnesium oxide, d) None of the above.
8. .... is used as lubricant for gelato-glycerin suppository bases.  
a) Glycerin, alcohol and liquid paraffin/ arachis oil, b) soap solution, c) mineral oil, d) None of the above.
9. .... form is most stable form of cocoa butter.  
a)  $\alpha$  b)  $\beta$  c)  $\gamma$  d)  $\sigma$
10. Tetracycline taken with milk, is a type of ..... incompatibility.  
a) Therapeutic, b) Physical, c) Chemical, d) Adjusted.

**Q. No. 2 Long Answers (Answer 1 out of 2)**

**(1 x 10 = 10)**

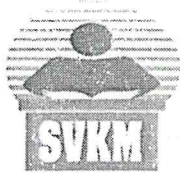
3. Define emulsion. Write down ideal properties of emulsion. Write a note on emulsifying agents.
4. Define suppositories. Write down types and advantages and disadvantages of suppositories. Explain method of preparation.

**Q. No. 3 Short Answers (Answer 2 out of 3)**

**(2 x 5 = 10)**

1. Define suspension. Write down ideal properties of suspensions.
2. Discuss therapeutic incompatibility.
3. Write short note on Syrup.





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**Second Sessional Theory Examination 2019-2020**

**Subject: Pharmaceutics I BP103T**

**Day & Date: Thursday, 7<sup>th</sup> Nov 2019**

**Class: F Y B. Pharm**

**Semester: I**

**Time: 90 min**

**Max. Marks: 30**

- Instructions:*
1. All questions are compulsory
  2. Draw a well labeled diagram wherever necessary
  3. Right hand side number indicates full marks

**Q. No.1 Multiple choice questions**

(10 X 1=10)

1. In dry gum method containing fixed oil ratio of oil: water: acacia to formulate primary emulsion is.....

- a) 4:2:1,                      b) 3:2:1,                      c) 2:2:1,                      d) 2:2:1

2. .... form is most stable form of cocoa butter.

- a)  $\alpha$                                        b)  $\beta$                                       c)  $\gamma$                                       d)  $\sigma$

3. If water soluble dye (amaranth) is mixed with the emulsion and examined under the microscope, the dispersed globules appears red and ground colorless indicates .....

- a) o/w type,                       b) w/o type,                      c) w/o/w,                      d) None of the above.

4. Glycero-gelatin suppository USP contains.....

- a) 70% glycerin, 20 % gelatin and 10 % water,                      b) 70% glycerin, 14 % gelatin and 16 % water,  
c) 40% glycerin, 32.5 % gelatin and 28.5 % water,                      d) 60% glycerin, 20 % gelatin and 20 % water.

5. If two drugs having opposing pharmacological effect are prescribed together results in.....

- a) Physical incompatibility,                      b) Chemical incompatibility,  
 c) Therapeutic incompatibility,                      d) None of the above.

6. Regarding the rate of sedimentation of pharmaceutical suspensions designed for oral administration, which of the following statements are true?

- i. The rate of sedimentation is increased as the diameter of the dispersed drug particles is increased.
- ii. The rate of sedimentation is increased as the viscosity of the continuous phase is increased.
- iii. The rate of sedimentation may be increased by centrifugation.
- iv. The rate of sedimentation may be increased by decreasing density difference.

- a) i & iii                      b) i & ii,                      c) ii & iii,                      d) ii & iv

7. Which of the following is hydrocarbon base

- a) Lanolin,                       b) White soft paraffin,                      c) Macrogols,                      d) Non of the above

8. .... 2% dispersion is used as suspending agent for indiffusible solids.

- a) Bentonite,                      b) Benzoic acid,                      c) Acacia,                      d) Sodium acetate.

9. Emulsifying agents which reduces interfacial tension are classified as.....

- a) Sufactants,                      b) Finely divided solids,                      c) Polysaccharides,                      d) Sterols

10. .... are semisolid unit dosage forms and externally inserted into body cavity to promote defecation.

- a) Enemas                      b) Pessaries                       c) Suppositories                      d) Bougies

**Q. No. 2 Long Answers (Answer 1 out of 2)**

(1 x 10 = 10)

1. Discuss stability problems of emulsion and methods to overcome.
2. Define ointment. Write down ideal properties of ointment. Discuss various ointment bases along with examples.

**Q. No. 3 Short Answers (Answer 2 out of 3)**

(2 x 5 = 10)

1. Write a note on suppository bases.
2. Write down various test used for identification of emulsions.
3. Write a note on (any one)
  - a. Suspending agents
  - b. Therapeutic incompatibility

**Second Sessional Theory Examination 2018-2019**

**Subject:** Pharmaceutical Inorganic Chemistry (BP104T)    **Day & Date:** Wed, 21/11/2018  
**Class:** First Year B. Pharm.    **Semester:** I  
**Time:** 10.30 am – 12.00 pm    **Max. Marks:** 30

*Instructions:* 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks

**Q.I. Multiple Choice Questions (MCQs)**

**10**

1. Fluoride inhibits caries formation via.....
  - a) decreasing acid solubility of enamel
  - b) bacterial inhibition
  - c) both a & b
  - d) increasing acid solubility of enamel
2. Calcium carbonate is used as.....
  - a) cleaning agent in dentifrice
  - b) an antacid
  - c) both a & b
  - d) none of the above
3. Absence of HCl in gastric secretions is called as .....
  - a) Hyperacidity
  - b) Achlorhydria
  - c) Constipation
  - d) Diarrhoea
4. Antacid should not be .....
  - a) constipative
  - b) laxative
  - c) both a & b
  - d) protective
5. assay of sodium bicarbonate is carried out by ..... titration
  - a) acid base
  - b) precipitation
  - c) complexometric
  - d) redox
6. saline cathartics act by.....
  - a) local irritation of the intestinal tract
  - b) swells in intestine
  - c) acting as lubricants
  - d) increasing osmotic load of intestine
7. ammonium chloride can be used as
  - a) expectorant
  - b) systemic acidifier
  - c) diuretic
  - d) all of the above
8. which of the following antidote act by adsorption?
  - a) activated charcoal
  - b) sodium thiosulphate
  - c) sodium nitrite
  - d) none of the above
9. the drug used in the treatment of cyanide poisoning is .....
  - a) sodium nitrite
  - b) sodium thiosulphate
  - c) both a & b
  - d) none of these
10. Aluminium hydroxide gel is.....
  - a) absorbable antacid
  - b) non-absorbable antacid
  - c) soluble in water
  - d) soluble in alcohol

**Q. II. Solve any One**

**10**

1. What are dental caries? Explain the role of fluoride in the treatment of dental caries. Write a note on sodium fluoride.
2. Define and classify antacids. Enlist the ideal properties and write a note on Aluminium hydroxide gel.

**Q. III. Solve any Two (5 Marks each)**

**10**

1. Define and classify antidotes with examples. Write in short on cyanide poisoning and its treatment.
2. What are emetics? Explain copper sulphate with respect to properties and assay.
3. Define expectorants. Explain the properties, assay and uses of ammonium chloride.





**Second Sessional Theory Examination 2019-2020**

**Subject:** Pharmaceutical Inorganic Chemistry (BP104T)

**Day & Date:** Fri, 08/11/2019

**Class:** First Year B. Pharm.

**Semester:** I

**Time:** 10.30 am – 12.00 pm

**Max. Marks:** 30

- Instructions:** 1. All questions are compulsory  
2. Right hand side number indicates full marks

**Q.I. Multiple Choice Questions (MCQs)**

**10**

- .....radiations are deflected in magnetic field  
a)  $\alpha$                       b)  $\beta$                       c)  $\gamma$                       d) both a & b
- Which of the following is an example of astringent?  
a) Zinc sulphate    b) Potash alum                      c) both a & b                      d) none of these
- Tincture iodine solution is .....  
a) aqueous iodine solution    b) strong iodine solution    c) Weak iodine solution    d) all of these
- Antacid should not be .....  
a) constipative                      b) laxative                      c) both a & b                      d) protective
- Unit of radioactivity is .....  
a) Calorie                      b) Newton                      c) Curie                      d) Joules
- Saline cathartics act by.....  
a) local irritation of the intestinal tract                      b) swells in intestine  
c) acting as lubricants                      d) increasing osmotic load of intestine
- Ammonium chloride can be used as  
a) expectorant    b) systemic acidifier    C) diuretic                      d) all of the above
- The drug used in the treatment of cyanide poisoning is .....  
a) sodium nitrite                      b) sodium thiosulphate                      c) both a & b                      d) none of these
- Milk of magnesia IP is an aqueous suspension of hydrated .....  
a)  $MgSO_4$                       b)  $Mg(OH)_2$                       c)  $MgCO_3$                       d)  $Mg$ -trisillicate
- which of the following antidote act by adsorption?  
a) activated charcoal    b) sodium thiosulphate                      c) sodium nitrite                      d) none of the above

**Q. II. Solve any One**

**10**

- Define and classify antacids. Enlist the ideal properties and write a note on Aluminium hydroxide gel.
- Define and classify antidotes with examples. Write in short on cyanide poisoning and its treatment. Explain the properties, assay and uses of sodium thiosulphate

**Q. III. Solve any Two (5 Marks each)**

**10**

- Explain the mechanisms by which antimicrobials act. Write any one method of preparation and assay of hydrogen peroxide.
- What is radioactive decay? Describe properties of  $\alpha$ ,  $\beta$ ,  $\gamma$  radiations.
- Define expectorants. Explain the properties, assay and uses of ammonium chloride.



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**Second Sessional Theory Examination 2020-2021**

**(Odd SEM)**

**Subject:** Pharmaceutical Inorganic Chemistry (BP104T) **Day & Date:** 28-04-2021

**Class:** First Year B. Pharm.

**Semester:** I

**Time:** 10:30 am – 12:00 pm

**Max. Marks:** 30

*Instructions: 1. All questions are compulsory*

*2. Draw a well labeled diagram wherever necessary*

*3. Right hand side number indicates full marks*

**Q.I. Multiple Choice Questions (MCQs)**

**10**

1. Saline cathartics act by .....

- a. Local irritation of intestinal tract
- b. Swelling in intestine
- c. Acting as a lubricant
- d. increasing osmotic load of intestine

2. Hydrogen peroxide act as an antimicrobial agent by ..... mechanism

- a. Protein precipitation
- b. Oxidation
- c. Halogenation
- d. All of the above

3. Which of the following is an example of astringent?

- a. Zinc sulphate
- b. Potash alum
- c. Both a and b
- d. None of these

4. Drug that is used to clear mucus from respiratory tract is called as .....

- a. Expectorant
- b. Antacid
- c. Astringent
- d. Cathartic

5. Which of the following is an example of hematinic?

- a. Ferrous gluconate
- b. Zinc sulphate
- c. Copper sulphate
- d. Magnesium sulphate

6. Which of the following antidote act by adsorption?

- a. Sodium thiosulphate
- b. Activated charcoal



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- c. Sodium nitrite
  - d. None of the above
7. Unit of radioactivity is .....
- a. Calorie
  - b. Newton
  - c. Curie
  - d. Joules
8. ....Radiations are deflected in magnetic field
- a. Alpha
  - b. Beta
  - c. Gamma
  - d. Both alpha and beta
9. The radioisotope used in studies of thyroid gland is .....
- a. Iodine-131
  - b. Sodium-22
  - c. Iron-59
  - d. Gold-198
10. Chemically potash alum is .....
- a. Potassium sulphate
  - b. Potassium aluminium sulphate
  - c. Aluminium chloride
  - d. Potassium permanganate

**Q. II. Solve the following (any 1)**

**10**

1. Define and classify antidotes. Write in short about cyanide poisoning. Write a note on sodium thiosulphate.
2. What is radioactive decay? Describe the properties of radioactive decay particles. Write a note on handling and storage of radiopharmaceuticals.

**Q. III. Solve the following (any 2)**

**10**

1. Explain the mechanism of action of antimicrobial agents with examples.
2. What are expectorants? Explain the method for preparation, properties, Assay and uses of Ammonium chloride.
3. What are hematinics? Explain the method for preparation, properties, Assay and uses of Ferrous sulphate



**Second Sessional Theory Examination 2021-2022  
(Odd SEM)**

**Subject & Subject Code: Pharmaceutical Inorganic Chemistry (BP104T)**

**Day & Date:** Wednesday 06/04/2022

**Class:** F.Y. B. Pharmacy

**Semester:** I

**Time:** 2:00 PM to 3:30 PM

**Max. Marks:** 30

*Instructions:* 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks

**QI Multiple Choice Questions (MCQs)**

**10**

1. When cyanide poisoning treated with sodium nitrite it converts blood hemoglobin to \_\_\_\_\_.  
A) Oxyhemoglobin B) Hemoglobin C) Carboxyhemoglobin D) Methemoglobin
2. Universal Antidote consist of \_\_\_\_\_.  
A) Tannic acid B) Charcoal C) MgO D) All of the above
3. Agent which inhibit growth of microorganism living tissue is called as \_\_\_\_\_.  
A) Disinfectant B) Antiseptic C) Germicide D) Sanitizers
4. Which of the following devices are used for the measurement of radio- activity?  
A) Geiger - Muller counter B) Ionization chamber  
C) Proportional counter D) All of the above
5. Delayed evacuation of the bowels / Faeces is called \_\_\_\_\_.  
A) Acidity B) Constipation C) Peristalsis D) Cathartics
6. Radiopharmaceuticals use in study of thyroid uptake \_\_\_\_\_.  
A) K-42 B) Cr-51 C) I-131 D) S-357
7. The pharmaceutical aid used as astringent and clearing agent is  
A) Agar B) Alum C) Benzyl Alcohol D) All
8. Which of following is called Rochelle salt? \_\_\_\_\_.  
A) Sodium Potassium Tartrate B) Potassium Bitartrate  
C) Potassium citrate D) Potassium oxalate
9. The mechanism of halogenation occurring with antiseptic are \_\_\_\_\_.  
A) Sulfhydryl B) Halogen C) Hypohalite D) Oxidation
10. All of the following drugs used for iron deficiency anaemia EXCEPT:  
A) Ferrous Sulphate B) Ferrous Fumarate  
C) Calcium Gluconate D) Ferrous Gluconate

**QII Long Answers (Answer any 1 out of 2)**

**10**

1. Describe in detail Geiger - Muller counters and Scintillation counters for the measurement of Radio - activity.
2. Define & classify Antidote. Write a note on cyanide poisoning and discuss the physical and chemical properties and assay of sodium thiosulphate.

**QIII Short Answers (Answer any 2 out of 3)**

**10**

1. Define Haematinics, give synonym physical properties and uses of FeSO<sub>4</sub>.
2. Define and classify expectorant, give assay, physical properties and uses of Ammonium chloride.
3. Define Antimicrobial agent. Discuss mechanism of action of inorganic antimicrobial agent.





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**Second Sessional Theory Examination 2020-2021**

**(Odd SEM)**

**Subject:** Pharmaceutical Inorganic Chemistry (BP104T)

**Day & Date:** 28-04-2021

**Class:** First Year B. Pharm.

**Semester:** I

**Time:** 10:30 am – 12:00 pm

**Max. Marks:** 30

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- Instructions:*
1. All questions are compulsory
  2. Draw a well labeled diagram wherever necessary
  3. Right hand side number indicates full marks

**Q.I. Multiple Choice Questions (MCQs)**

**10**

1. Saline cathartics act by .....

  - a. Local irritation of intestinal tract
  - b. Swelling in intestine
  - c. Acting as a lubricant
  - d. increasing osmotic load of intestine

2. Hydrogen peroxide act as an antimicrobial agent by ..... mechanism

  - a. Protein precipitation
  - b. Oxidation
  - c. Halogenation
  - d. All of the above

3. Which of the following is an example of astringent?

  - a. Zinc sulphate
  - b. Potash alum
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- c. Sodium nitrite
  - d. None of the above
7. Unit of radioactivity is .....
- a. Calorie
  - b. Newton
  - c. Curie
  - d. Joules
8. ....Radiations are deflected in magnetic field
- a. Alpha
  - b. Beta
  - c. Gamma
  - d. Both alpha and beta
9. The radioisotope used in studies of thyroid gland is .....
- a. Iodine-131
  - b. Sodium-22
  - c. Iron-59
  - d. Gold-198
10. Chemically potash alum is .....
- a. Potassium sulphate
  - b. Potassium aluminium sulphate
  - c. Aluminium chloride
  - d. Potassium permanganate

**Q. II. Solve the following (any 1)**

**10**

1. Define and classify antidotes. Write in short about cyanide poisoning. Write a note on sodium thiosulphate.
2. What is radioactive decay? Describe the properties of radioactive decay particles. Write a note on handling and storage of radiopharmaceuticals.

**Q. III. Solve the following (any 2)**

**10**

1. Explain the mechanism of action of antimicrobial agents with examples.
2. What are expectorants? Explain the method for preparation, properties, Assay and uses of Ammonium chloride.
3. What are hematinics? Explain the method for preparation, properties, Assay and uses of Ferrous sulphate



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**Second Sessional Theory Examination-(ODD SEM) 2022-2023**

<b>Subject</b>	: Pharmaceutical Analysis (BP102T)	<b>Day &amp; Date</b>	: Tuesday, 18/04/2023	
<b>Class</b>	: First Year B. Pharmacy	<b>Semester</b>	: I	<b>Write Your Seat No. Here</b>
<b>Time</b>	: 10:30 am – 12:00 pm	<b>Max. Marks</b>	: 30	<input type="text"/>

<b>Instructions:</b>	1. All questions are compulsory	2. Draw a well labeled diagram wherever necessary
	3. Right hand side number indicates full marks	4. Do not write/tick on the question paper

**QI Objective Type Questions 10 M**

1. Write the applications of sodium nitrite titration (CO3, LL3)
2. Describe co-precipitation in gravimetric analysis (CO3, LL2)
3. Define: Self Indicators and Half wave potential (CO3, LL1; CO4, LL1)
4. Draw a well labelled diagram of Calomel electrode (CO4, LL1)
5. Write Ilkonic equation and explain all terms (CO4, LL3)

**QII Long Answers Question (Answer any 1 out of 2) (CO3, LL2) 10 M**

1. Explain in detail the principle involve in Idiometry and cerimetry titrations with its applications
2. Explain principle and steps involved in gravimetric analysis and describe post-precipitation.

**QIII Short Answers Question (Answer any 2 out of 3) 10 M**

1. Explain conductivity cells and write applications of conductometric titration (CO4, LL2)
2. Enumerate the indicator electrodes used in potentiometric titration and explain Glass electrode (CO4, LL1)
3. Describe rotating platinum electrode and its applications (CO4, LL1)



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**Second Sessional Theory Examination-(ODD SEM) 2022-2023**

<b>Subject</b>	: Pharmaceutical Analysis (BP102T)	<b>Day &amp; Date</b>	: Tuesday, 18/04/2023	
<b>Class</b>	: First Year B. Pharmacy	<b>Semester</b>	: I	<b>Write Your Seat No. Here</b>
<b>Time</b>	: 10:30 am – 12:00 pm	<b>Max. Marks</b>	: 30	<input type="text"/>

<b>Instructions:</b>	1. All questions are compulsory	2. Draw a well labeled diagram wherever necessary
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**QIII Short Answers Question (Answer any 2 out of 3) 10 M**

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2. Enumerate the indicator electrodes used in potentiometric titration and explain Glass electrode (CO4, LL1)
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**Second Sessional Theory Examination-(ODD SEM) 2022-2023**

<b>Subject</b>	: Pharmaceutical Analysis (BP102T)	<b>Day &amp; Date</b>	: Tuesday, 18/04/2023	
<b>Class</b>	: First Year B. Pharmacy	<b>Semester</b>	: I	<b>Write Your Seat No. Here</b>
<b>Time</b>	: 10:30 am – 12:00 pm	<b>Max. Marks</b>	: 30	<input type="text"/>

<b>Instructions:</b>	1. All questions are compulsory	2. Draw a well labeled diagram wherever necessary
	3. Right hand side number indicates full marks	4. Do not write/tick on the question paper

**QI Objective Type Questions 10 M**

1. Write the applications of sodium nitrite titration (CO3, LL3)
2. Describe co-precipitation in gravimetric analysis (CO3, LL2)
3. Define: Self Indicators and Half wave potential (CO3, LL1; CO4, LL1)
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1. Explain in detail the principle involve in Idiometry and cerimetry titrations with its applications
2. Explain principle and steps involved in gravimetric analysis and describe post-precipitation.

**QIII Short Answers Question (Answer any 2 out of 3) 10 M**

1. Explain conductivity cells and write applications of conductometric titration (CO4, LL2)
2. Enumerate the indicator electrodes used in potentiometric titration and explain Glass electrode (CO4, LL1)
3. Describe rotating platinum electrode and its applications (CO4, LL1)



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**Second Sessional Theory Examination 2018-2019**

**Subject: Remedial Biology**

**Day & Date: Saturday, 24/11/2018**

**Class: F. Y. B.Pharm**

**Semester: I**

**Time: 10:30 am-12:00 pm**

**Max. Marks: 30**

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**Q1. Answer any one**

(10 marks)

1. Elucidate the functions of various hormones secreted by pituitary gland.
2. Explain importance of Rennin angiotensin aldosterone system in regulation of homeostasis

**Q2. Answer any four**

(20 marks)

1. Explain structure and functions of cell.
2. Classification of nervous system and short note on structure of brain and spinal cord
3. Describe macro and micronutrients
4. Explain photosynthesis process
5. Short note on Plant Respiration
6. Short note on Spermatogenesis



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**Second Sessional Theory Examination 2020 2021**  
(Odd SEM)

**Subject & Subject Code: Remedial Biology (BP106RBT)**

**Day & Date: 24.04.2021 Saturday**

**Class: F.Y. B. Pharm Semester: I**

**Time: 10.30 to 11.30 Max.**

**Marks: 30 marks**

- Instructions: 1. All questions are compulsory*  
*2. Draw a well labeled diagram wherever necessary*  
*3. Right hand side number indicates full marks*

Q1. Answer any one (10 marks)

1. Explain in detail about human excretory system. Add a note on Renin-Angiotensin System (RAS)
2. What is respiration? Give its classification and explain in detail about anaerobic respiration.

Q2. Answer any four (20 marks)

1. Write a short note on Peripheral nervous system
2. Enlist various endocrine glands along with their functions
3. Add a note on spermatogenesis
4. Add a note on nitrogen metabolism
5. Add a short note on cell division
6. Give the phases of plant growth

*Sahmod :- (S.S. RaMod)*  
Name and Signature  
Subject-Teacher



**Second Sessional Theory Examination 2020-2021**  
(Odd SEM)

**Subject & Subject Code: Remedial Mathematics (BP106RMB)**

**Day & Date: Tuesday 24/4/2021**

**Semester: I/III**

**Class: F.Y/S.Y. B. Pharmacy**

**Time: 1 Hour 30 Minutes**

**Max. Marks: 30**

- Instructions: 1. All questions are compulsory*  
*2. Draw a well labeled diagram wherever necessary*  
*3. Right hand side number indicates full marks*

**QI Multiple Choice Questions (MCQs)**

**10 Marks**

**1. If  $3x-4y+2=0$  is given equation of line. Then its slope is**

- (a)  $\frac{3}{4}$  (b)  $\frac{1}{4}$   
(c)  $-\frac{3}{4}$  (d)  $\frac{3}{2}$

**2. If  $A = \begin{bmatrix} 1 & 2 & 3 \\ 0 & 1 & -1 \\ 2 & 4 & x \end{bmatrix}$  is a singular matrix, then  $x =$**

- (a) 4 (b) 3 (c) 6 (d) 2

**3. If 2, 3, -1 are eigen values of matrix A Then eigen values of  $adjA$  are**

- (a) 1, 3, 6 (b) -2, -4, 6  
(c) 3, 2, 1 (d) -3, -2, 6

**4. If line L passes through the points (1, 2) and (-2, 3) then slope of line L is**

- (a)  $-\frac{1}{3}$  (b)  $\frac{2}{3}$  (c)  $\frac{1}{3}$  (d)  $-\frac{2}{3}$

**5.  $\theta$  is inclination of line parallel to line  $y=x$  then  $\theta =$**

- (a)  $\frac{\pi}{2}$  (b)  $\frac{\pi}{3}$  (c)  $\frac{\pi}{4}$  (d)  $\frac{\pi}{6}$

**6. If  $y = \sin 2x$  then  $\left(\frac{dy}{dx}\right)$  at  $x = \frac{\pi}{4}$  is**

- (a) 1 (b) 3 (c) -4 (d) 0

**7. If  $\alpha$  and  $\beta$  are eigen values of Matrix  $A = \begin{bmatrix} 1 & 2 \\ 2 & -1 \end{bmatrix}$  then  $\alpha + \beta$  and  $\alpha\beta$  are respectively**

- (a) 0, -5 (c) 0, 5  
(b) 3, 5 (d) 1, 2



8 If  $y = x^3 - x^2 + 2x - 1$ . Then  $\left(\frac{dy}{dx}\right)_{x=2}$  is  
(a) -10 (b) 14 (c) 10 (d) 12

9. If  $y = a^{2x} \sqrt{x}$  then  $\left(\frac{dy}{dx}\right)_{x=1}$  is  
(a)  $\frac{1}{2}a^2(4\log a + 1)$  (b)  $-\frac{1}{2}a^2(4\log a + 1)$   
(c)  $\frac{1}{2}(4\log a + 1)$  (d)  $a^2(4\log a + 1)$

10 If  $y=mx$  and  $y= - x$  are perpendicular to each other then  $m=$   
(a) -1 (b) 1 (c) 2 (d) -2

**QII Long Answers (Answer any 1 out of 2)**

**10 Marks**

1. Solve By Crammers Rule.  $x + y - z = 0$   $2x + y + 3z = 9$   $x - y + z = 2$

2. Find  $\frac{dy}{dx}$  if  $y = \frac{a^{2x}}{\log x} + \sin 2x \cdot \cos 4x + e^{-2x} \cdot \sec 2x$

**QIII Short Answers (Answer any 2 out of 3)**

**10 Marks**

1. Find Eigen values and Eigen vectors of  $A = \begin{bmatrix} 5 & 4 \\ 1 & 2 \end{bmatrix}$

2. Find equation of line passing through (3, 2) and perpendicular to the line  $x - 2y - 3 = 0$

3. Find equation of line passing through (1, 2) and Parallel to the line  $3x + 4y + 10 = 0$

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*Second Sessional Theory Examination 2018-2019*

Class: F.Y.B.Pharmacy

Semester: First

Subject : Remedial Mathematics

Max. Marks : 30

Date: November 24<sup>th</sup>, 2018

Time : 1 hr. 30 min.

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• Instructions :

1. Both the questions are compulsory.
  2. Figures to the right indicates full marks.
- 

1. Attempt any two of the following. (5+5=10)

(a) Solve by matrix method

$$x + y + z = 7, \quad x - y + 2z = 6, \quad x + y - z = 1.$$

(b) Show that  $\lim_{x \rightarrow a} \left( \frac{x^n - a^n}{x - a} \right) = na^{n-1}$ . Using the same, evaluate  $\lim_{x \rightarrow 2} \left( \frac{x^5 - 32}{x - 2} \right)$ .

(c) Find  $A^{-1}$ , if  $A = \begin{bmatrix} 3 & 1 & 2 \\ 2 & 4 & 1 \\ -1 & 3 & 3 \end{bmatrix}$ .

2. Attempt any five of the following. (4x5=20)

(a) Without expanding the determinant, show that  $\begin{vmatrix} 1 & 2 & 3 \\ 4 & 5 & 6 \\ 7 & 8 & 9 \end{vmatrix} = 0$ .

(b) If  $f : \mathbf{R} \rightarrow \mathbf{R}$  such that  $f(x) = 2x + 1$  and  $g : \mathbf{R} \rightarrow \mathbf{R}$  such that  $g(x) = x^2 - 2$ .  
Find i)  $g \circ f$     ii)  $f \circ g$ .

(c) Evaluate :  $\lim_{x \rightarrow 3} \left( \frac{x^2 - 5x + 6}{x^2 - 9} \right)$ .

(d) Evaluate :  $\lim_{x \rightarrow 0} \left( \frac{3^x - 5^x}{3^x - 2^x} \right)$ .

(e) If  $A = \begin{bmatrix} 1 & 5 \\ 2 & 1 \end{bmatrix}$  and  $B = \begin{bmatrix} 2 & 4 \\ 1 & 6 \end{bmatrix}$ , verify that  $(AB)^{-1} = B^{-1}A^{-1}$ .

(f) Using Cramer's rule, solve  $x + y = 3$ ,  $x - y = 2$ .

(g) If  $A = \begin{bmatrix} 1 & 2 \\ 3 & 4 \end{bmatrix}$  and  $B = \begin{bmatrix} 1 & 5 \\ 1 & 6 \end{bmatrix}$ , verify that  $(A + B)^t = A^t + B^t$ .

\*\*\*\*\*





Shri. Vile Parle Kelavani Mandal's  
Institute of Pharmacy, Dhule

Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

**Second Sessional Theory Examination 2019-2020**

**Subject:** Remedial Mathematics (BP106RMT) **Day & Date:** Tuesday,  
11/11/19

**Class:** First Year B. Pharm

**Semester:** I/III\*

**Time:** 10-30 am to 12.00 pm

**Max. Marks:** 30

- Instructions: 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks

**QI Attempt any two of the following**

(5+5=10)

1. Solve By Crammers Rule.  $x + y - z = 0$

$$2x + y + 3z = 9$$

$$x - y + z = 2$$

2. If  $A = \begin{bmatrix} 1 & -2 & 0 \\ 0 & 1 & 6 \\ 2 & -3 & 4 \end{bmatrix}$  Find  $A^{-1}$

3. if  $y(x) = e^{\tan 4x} \cos(\log x)$  Find  $\frac{dy}{dx}$

**QII Attempt any five of the following**

(4×5=20)

1. (i) Define Rectangular Matrix. (ii) Write definition of Logarithm.

(iii) Nonsingular Matrix. (iv) Write equation of line in Double intercept form. 2. Find equation Of a line with inclination  $\frac{\pi}{3}$  and passes through the point (2, 3).

3. If  $y(x) = e^{4x} \log x + \sin(e^{4x}) \cdot x^3$  Find  $\frac{dy}{dx}$

4. Find equation Of a line which passes through the point (-2, -3) and (-7, -9).

5. Find  $\begin{vmatrix} 5 & -4 & 9 \\ -6 & -8 & -2 \\ 0 & -2 & 12 \end{vmatrix}$

6. If  $A = \begin{bmatrix} 2 & 1 & 1 \\ -3 & 4 & -3 \\ -1 & -5 & -5 \end{bmatrix}$  And  $B = \begin{bmatrix} 0 & 6 & -6 \\ -2 & -4 & 0 \\ -8 & 5 & -5 \end{bmatrix}$  Verify that  $(A+B)^T = A^T + B^T$

7. If  $y(x) = \frac{\log(\cot x)}{\sin(4x) \cdot x^3}$  Find  $\frac{dy}{dx}$





**Second Sessional Theory Examination 2021-2022**  
(Odd SEM)

**Subject & Subject Code: Remedial Mathematics (BP106RMT)**

**Day & Date: Friday 08/04/2022**

**Class: F.Y/S.Y. B. Pharmacy.**

**Semester: I/III**

**Time: 1 hr**

**Max. Marks: 30**

- Instructions: 1. All questions are compulsory*  
*2. Draw a well labeled diagram wherever necessary*  
*3. Right hand side number indicates full marks*

**QI Multiple Choice Questions (MCQs)**

**10**

**1. If  $y = \sin 2x$  Then  $\frac{dy}{dx} =$**

- (a)  $2\sin 2x$   
(b)  $2\cos 2x$   
(c)  $2\cos x$   
(d)  $2\sin x$

**2. If  $y = \tan x$  Then  $\frac{dy}{dx} =$**

- (a)  $-\sec^2 x$   
(b)  $\operatorname{cosec}^2 x$   
(c)  $\sec^2 x$   
(d)  $-\operatorname{cosec}^2 x$

**3. Laplace transform of  $t^2$  is**

- (a)  $\frac{1}{s}$                       (b)  $\frac{1}{s^2}$   
(c)  $\frac{2}{s^2}$                       (d)  $\frac{4}{s^2}$

**4. If inclination of line L is  $\frac{\pi}{4}$  Then slop of line L is**

- (a)  $-1$                       (b)  $1$   
(c)  $0$                       (d)  $\frac{1}{2}$   
(d)

**5. If Line passes through points (1 2) and (2 4) then slop of line L is**

- (a)  $-1$   
(b)  $0$   
(c)  $1$   
(d)  $2$

**6. If  $y = \sec x$  Then  $\frac{dy}{dx} =$**

- (a)  $\sec 2x \tan 2x$   
(b)  $\sec x \tan 2x$   
(c)  $\sec x \tan x$   
(d)  $\sec 2x \tan x$



7. If  $y = x^4$  Then  $\frac{dy}{dx} =$

- (a)  $4x^3$                       (b)  $x^3$   
(c)  $4x^2$                       (d) 0

8. Slope of line  $y=4$  is

- (a) -2  
(b) 3  
(c) -3  
(d) 0

9. Laplace Transform of  $\sin 2t$  is

- (a)  $\frac{1}{s^2-4}$                       (b)  $\frac{2}{s^2-4}$   
(c)  $\frac{s}{s^2+4}$                       (d)  $\frac{2}{s^2+4}$

10. Laplace Transform of  $\cos 2t$  is

- (a)  $\frac{1}{s^2+4}$                       (b)  $\frac{s}{s^2-4}$   
(c)  $\frac{1}{s^2-4}$                       (d)  $\frac{s}{s^2+4}$

**QII Long Answers (Answer any 1 out of 2)**

**10**

1. Using adjoint method, Find inverse of the matrix  $A = \begin{bmatrix} 1 & -1 & 2 \\ 2 & -1 & 1 \\ 1 & 1 & -1 \end{bmatrix}$

2. Find equation of line passing through the point  $(-2, 0)$  and perpendicular to the line to the line  $4x-3y=2$

**QIII Short Answers (Answer any 2 out of 3)**

**10**

1. If  $y = \cos x \cdot \log x$ . find  $\frac{dy}{dx}$

2. Verify Cayley Hamilton theorem for  $A = \begin{bmatrix} 1 & 1 \\ 1 & 0 \end{bmatrix}$  hence find its inverse.

3. Prove that the lines  $3x-2y-1=0$  and  $9x-6y+5=0$  are parallel.



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**Second Sessional Theory Examination-(ODD SEM) 2022-2023**

<b>Subject</b>	: Remedial Mathematics (BP106RMT)	<b>Day &amp; Date</b>	: MONDAY 24/04/2023	
<b>Class</b>	: First Year B. Pharmacy/DSY B. Pharma	<b>Semester</b>	: I	<b>Write Your Seat No. Here</b>
<b>Time</b>	: 10:30 am – 11:30 am	<b>Max. Marks</b>	: 30	<input type="text"/>

<b>Instructions:</b>	1. All questions are compulsory	2. Draw a well labeled diagram wherever necessary
	3. Right hand side number indicates full marks	4. Do not write/tick on the question paper

<b>QI</b>	<b>Long Answers Question (Answer any 1 out of 2)</b>	<b>10 M</b>
1.	If $y=e^x \log x$ find $\frac{d^2y}{dx^2}$ or $y_2$	(LL3, CO3)
2.	Find the equation of line passing through the point (-2, 0) and perpendicular to line $4x-3y=2$ .	(LL3, CO4)
<b>QII</b>	<b>Short Answers Question</b>	
	<b>Section – A (Answer any ONE)</b>	<b>05 M</b>
1.	Differentiate $x^2 e^x \log x$ with respect to $x$	(LL2, CO3)
2.	Find equation of line Passing through the point (2,2) and its slope is 1	(LL2, CO4)
	<b>Section – A (Answer any Three)</b>	<b>15 M</b>
1.	Find double derivative of $y= e^x \cos x$ with respect to $x$ .	(LL2, CO3)
2.	Find equation of line Passing through the point (0,0) and perpendicular to line $x + y = 2$	(LL2, CO4)
3.	Find maximum and minimum value of $f(x)=x^3 - 6x^2 + 9x + 15$ .	(LL2, CO3)
4.	Find equation of line passing through the point (-2,3) and (2,7)	(LL2, CO4)



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**Second Sessional Theory Examination-(ODD SEM) 2022-2023**

<b>Subject</b> :	Remedial Mathematics (BP106RMT)	<b>Day &amp; Date</b> :	MONDAY 24/04/2023	
<b>Class</b> :	First Year B. Pharmacy/DSY B. Pharma	<b>Semester</b> :	<b>I</b>	<i>Write Your Seat No. Here</i>
<b>Time</b> :	10:30 am – 11:30 am	<b>Max. Marks</b> :	30	<input type="text"/>

<b>Instructions:</b>	1. All questions are compulsory	2. Draw a well labeled diagram wherever necessary
	3. Right hand side number indicates full marks	4. Do not write/tick on the question paper

<b>QI</b>	<b>Long Answers Question (Answer any 1 out of 2)</b>	<b>10 M</b>
1.	If $y=e^x \log x$ find $\frac{d^2y}{dx^2}$ or $y_2$	(LL3, CO3)
2.	Find the equation of line passing through the point (-2, 0) and perpendicular to line $4x-3y=2$ .	(LL3, CO4)
<b>QII</b>	<b>Short Answers Question</b>	
	<b>Section – A (Answer any ONE)</b>	<b>05 M</b>
1.	Differentiate $x^2 e^x \log x$ with respect to $x$	(LL2, CO3)
2.	Find equation of line Passing through the point (2,2) and its slope is 1	(LL2, CO4)
	<b>Section – A (Answer any Three)</b>	<b>15 M</b>
1.	Find double derivative of $y= e^x \cos x$ with respect to $x$ .	(LL2, CO3)
2.	Find equation of line Passing through the point (0,0) and perpendicular to line $x + y = 2$	(LL2, CO4)
3.	Find maximum and minimum value of $f(x)=x^3 - 6x^2 + 9x + 15$ .	(LL2, CO3)
4.	Find equation of line passing through the point (-2,3) and (2,7)	(LL2, CO4)



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**Second Sessional Theory Examination-(Odd SEM) 2023-2024**

<b>Subject</b> : Remedial Mathematics	<b>Day &amp; Date</b> : Saturday 23/12/2023	
<b>Class</b> : First Year B. Pharmacy	<b>Semester</b> : I/III	<b>Write Your Seat No. Here</b> <input type="text"/>
<b>Time</b> : 10:30 am to 11:30 am	<b>Max. Marks</b> : 30	

**Instructions:** 1. All questions are compulsory 2. Draw a well-labeled diagram wherever necessary  
3. Right-hand side number indicates full marks 4. Do not write/tick on the question paper

<b>Q. I: Long Answers Question (Answer any 1 out of 2)</b>		10 M
a)	Verify Cayley –Hamilton theorem: $A = \begin{bmatrix} 3 & 1 & 2 \\ 4 & 3 & 1 \\ 2 & 5 & 4 \end{bmatrix}$	
b)	i) Find the slope and intercept of the line $2x + 3y = 6$ ii) Evaluate $\int x \cdot \sin x \, dx$ .	
<b>Q. II: Short Answers Question (Answer any 4 out of 6)</b>		20 M
a)	Find $\frac{dy}{dx}$ if $y = 10^x + x^{10} + e^x + \sin x + \tan x$	
b)	Evaluate $\int m^x + x^m + m^m \, dx$ .	
c)	Find the equation of line which is passing through the points $(-4,6)$ , & $(8, -3)$ .	
d)	Show that the points are the vertices of $(-1,5)$ , $(4,3)$ , $(7, -2)$ isosceles triangle.	
e)	If $A = \begin{bmatrix} 1 & 3 & 2 \\ -1 & 0 & 2 \\ 3 & 1 & -1 \end{bmatrix}$ then find adjoint A	
f)	If $A = \begin{bmatrix} 2 & 4 & 4 \\ 4 & 2 & 4 \\ 4 & 4 & 2 \end{bmatrix}$ then find $A^2$ .	





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**Second Sessional Theory Examination-(Odd SEM) 2023-2024**

<b>Subject</b> : Remedial Mathematics	<b>Day &amp; Date</b> : Saturday 23/12/2023	
<b>Class</b> : First Year B. Pharmacy	<b>Semester</b> : I/III	<b>Write Your Seat No. Here</b> <input type="text"/>
<b>Time</b> : 10:30 am to 11:30 am	<b>Max. Marks</b> : 30	

**Instructions:** 1. All questions are compulsory 2. Draw a well-labeled diagram wherever necessary  
3. Right-hand side number indicates full marks 4. Do not write/tick on the question paper

<b>Q. I: Long Answers Question (Answer any 1 out of 2)</b>		10 M
a)	Verify Cayley –Hamilton theorem: $A = \begin{bmatrix} 3 & 1 & 2 \\ 4 & 3 & 1 \\ 2 & 5 & 4 \end{bmatrix}$	
b)	i) Find the slope and intercept of the line $2x + 3y = 6$ ii) Evaluate $\int x \cdot \sin x \, dx$ .	
<b>Q. II: Short Answers Question (Answer any 4 out of 6)</b>		20 M
a)	Find $\frac{dy}{dx}$ if $y = 10^x + x^{10} + e^x + \sin x + \tan x$	
b)	Evaluate $\int m^x + x^m + m^m \, dx$ .	
c)	Find the equation of line which is passing through the points $(-4,6)$ , & $(8, -3)$ .	
d)	Show that the points are the vertices of $(-1,5)$ , $(4,3)$ , $(7, -2)$ isosceles triangle.	
e)	If $A = \begin{bmatrix} 1 & 3 & 2 \\ -1 & 0 & 2 \\ 3 & 1 & -1 \end{bmatrix}$ then find adjoint A	
f)	If $A = \begin{bmatrix} 2 & 4 & 4 \\ 4 & 2 & 4 \\ 4 & 4 & 2 \end{bmatrix}$ then find $A^2$ .	



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**First Sessional Theory Examination 2020-2021**

**Subject: Pharmaceutics I BP103T      Day & Date: Friday, 12<sup>th</sup> March 2021**

**Class: F Y B. Pharm                      Semester: I**

**Time: 90 min                                  Max. Marks: 30**

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- Instructions:* 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks

**Q.No.1 Multiple choice questions (10 X 1=10)**

1. .... is known as 'Father of Pharmacy education in India'.  
a) M. L. Shroff, b) R. N. Chopra, c) B. N. Ghosh, d) None of the above
2. The fourth edition of IP was published in ....  
a) 1985 b) 1996 c) 1960 d) 19963.
3. Pharmacy Council of India (PCI) was established in ....  
a) 1947 b) 1948 c) 1949 d) 1950.
4. Which of the following excipient act as preservative.....  
a) Sorbitan monolactrate b) Tyloxopol c) Benzethonium chloride d) Thiomersal
5. When two or more drugs are used in combination to increase the pharmacological action, the phenomenon is known as .....  
a) Synergism b) Tolerance c) Potentiation c) Idiosyncrasy
6. Nasal drops should have pH .....  
a) 5-6 b) 6-7 c) 7-8 d) 8-9
7. Which of the following is natural colorant .....  
a) Amaranth b) Erythrosine c) Carotenoids d) Eosin
8. The dosage form intended to be rubbed on the skin for rubeficient and counterirritant effect are called  
a) Lotions b) Liniments c) Gels d) Dentifrices
9. 1 $\bar{3}$  (apothecaries) =  
a) 437.5 grains b) 480 grains c) 7000 grains d) 5760 grains
10. Daily adult dose of drug is 120 mg. How much of it should be given to a child of 11 years?  
a) 57.39, b) 60, c) 50.39, d) 37

**Q. No. 2 Long Answers (Answer 1 out of 2)**

**(1 x 10 = 10)**

1. Define dosage form. Explain need of dosage form. Describe various solid dosage forms.
2. Define posology. Write a note on factors affecting posology

**Q. No. 3 Short Answers (Answer 2 out of 3)**

**(2 x 5 = 10)**

1. Write down parts of prescription
2. Discuss various excipient used in the formulation of liquid dosage forms.
3. Write short note on (any one)
  - a) Syrups
  - b) Topical oral preparations

**Improvement Sessional Theory Examination 2018-2019**

**Subject: Human Anatomy and Physiology-I**

**Day & Date:**

Monday, 26.11.18

**Class: F. Y. B. Pharma**

**Semester: I**

**Time: 11:00-12:30**

**Max. Marks: 30**

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- Instructions:*
1. All questions are compulsory
  2. Draw a well labeled diagram wherever necessary
  3. Right hand side number indicates full marks

**Q1. Solve the following MCQs**

**10**

1. Name the arteries and veins facilitating blood flow to and from the lower limbs:

- a) Brachial    b) Mesenteric    c) Femoral    d) Carotid

2. Blood flows out of the ventricles when?

- a) The atrioventricular valves are open    b) The semi-lunar valves are open  
c) The bicuspid valves are open    d) The mitral valves are closed

3. Cardiac output is the amount of blood pumped by:

- a) 1 ventricle in 1 minute    b) 1 atrium in 1 minute  
c) Both ventricles in 1 minute    d) Both atria in 1 minute

4. Blood from the heart flows through the blood vessels in which order?

- a) Veins, arterioles, capillaries, venules, arteries    b) Capillaries, arterioles, arteries, venules, veins  
c) Arteries, arterioles, capillaries, venules, veins    d) Arteries, capillaries, arterioles, venules, veins

5. Blood pressure is highest when leaving which heart chamber?

- a) Right atrium    b) Right ventricle  
c) Left atrium    d) Left ventricle

d) People who first identified them

6. Which of the following structures is the primary organ of the integumentary system?

- a) Sensory receptors    b) Sweat glands  
c) Bones    d) Skin

7. Which cells produce the pigment that contributes to hair color?

- a) Keratinocytes    b) Melanocytes  
c) Langerhans cells    d) Merkel cells

8. The main components of the skeletal system are:

- a) Blood, bones, cartilage and tendons
- b) Bones, cartilage, joints and tendons
- c) Bones, cartilage, joints and ligaments
- d) Cartilage, joints, ligaments and tendons

9. Muscle action can be classified as:

- a) Sympathetic and parasympathetic
- b) Voluntary and involuntary
- c) Autonomic and sympathetic
- d) Autonomic and voluntary

10. Which blood group can be transfused into an O Rh- patient?

- a) O Rh-
- b) O Rh+
- c) AB Rh+
- d) AB Rh-

**QII Long Answers (Answer any 1 out of 2)**

**10**

1. Define tissue. Explain in detail various types of tissue with classification
2. Explain in details ABO and RH blood grouping system.

**QIII Short Answers (Answer any 2 out of 3)**

**10**

1. Define joints. Draw a neat labelled diagram of synovial joints
2. Write the various functions of skin with a diagram
3. Draw a neat labelled diagram of internal structure of eye.



**Improvement Sessional Theory Examination 2018-2019**

**Subject:** Pharmaceutical Inorganic Chemistry (BP104T)

**Day & Date:** Thu, 29/11/2018

**Class:** First Year B. Pharm.

**Semester:** I

**Time:** 10.30 am – 12.00 pm

**Max. Marks:** 30

- Instructions:*
1. All questions are compulsory
  2. Draw a well labeled diagram wherever necessary
  3. Right hand side number indicates full marks

**Q.I. Multiple Choice Questions (MCQs)**

**10**

1. The only water soluble compound used as effective antacid is .....  
a) Magnesium trisilicate      b) Sodium bicarbonate      c) Milk of magnesia  
d) sodium hydroxide
2. Calcium carbonate is used as.....  
a) cleaning agent in dentifrice      b) an antacid      c) both a & b      d) none of the above
3. Assay of ammonium chloride is carried out by ..... titration  
a) acid base      b) precipitation      c) complexometric      d) redox
4. Antacid should not be .....  
a) constipative      b) laxative      c) both a & b      d) protective
5. assay of sodium bicarbonate is carried out by ..... titration  
a) acid base      b) precipitation      c) complexometric      d) redox
6. Drug that induces vomiting is called as .....  
a) Emetic      b) Antacid      c) Astringent      d) Cathartic
7. ammonium chloride can be used as  
a) expectorant      b) systemic acidifier      C) diuretic      d) all of the above
8. which of the following antidote act by adsorption?  
a) activated charcoal      b) sodium thiosulphate      c) sodium nitrite      d) none of the above
9. which of the following is a dental desensitizer?  
a) Calcium carbonate      b) strontium chloride      c) sodium fluoride      d) potassium chloride
10. Aluminium hydroxide gel is.....  
a) absorbable antacid      b) non-absorbable antacid      c) soluble in water      d) soluble in alcohol

**Q. II. Solve any One**

**10**

1. what are dental caries? Explain the role of fluoride in the treatment of dental caries. Write a note on sodium fluoride.
2. Define and classify antacids. Enlist the ideal properties and write a note on Aluminium hydroxide gel.

**Q. III. Solve any Two (5 Marks each)**

**10**

1. Define and classify antidotes with examples. Write in short on cyanide poisoning and its treatment.
2. Write a note on activated charcoal.
3. Define expectorants. Explain the properties, assay and uses of ammonium chloride.





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**Improvement Sessional Theory Examination 2019-2020**

**Subject:** Pharmaceutical Inorganic Chemistry (BP104T)

**Day & Date:** Fri, 15/11/2019

**Class:** First Year B. Pharm.

**Semester:** I

**Time:** 10.30 am – 12.00 pm

**Max. Marks:** 30

- Instructions:** 1. All questions are compulsory  
2. Right hand side number indicates full marks

**Q.I. Multiple Choice Questions (MCQs)**

10

- Aluminium hydroxide gel is.....  
a) absorbable antacid    b) non-absorbable antacid    c) soluble in water    d) soluble in alcohol
- Which of the following is an example of astringent?  
a) Zinc sulphate    b) Potash alum    c) both a & b    d) none of these
- Tincture iodine solution is .....  
a) aqueous iodine solution    b) strong iodine solution    c) Weak iodine solution    d) all of these
- Antacid should not be .....  
a) constipative    b) laxative    c) both a & b    d) protective
- Unit of radioactivity is .....  
a) Calorie    b) Newton    c) Curie    d) Joules
- Saline cathartics act by.....  
a) local irritation of the intestinal tract    b) swells in intestine  
c) acting as lubricants    d) increasing osmotic load of intestine
- Ammonium chloride can be used as.....  
a) expectorant    b) systemic acidifier    c) diuretic    d) all of the above
- .....radiations are deflected in magnetic field  
a)  $\alpha$     b)  $\beta$     c)  $\gamma$     d) both a & b
- Milk of magnesia IP is an aqueous suspension of hydrated .....  
a)  $MgSO_4$     b)  $Mg(OH)_2$     c)  $MgCO_3$     d) Mg-trisilicate
- Which of the following antidote act by adsorption?  
a) activated charcoal    b) sodium thiosulphate    c) sodium nitrite    d) none of the above

**Q. II. Solve any One**

10

- What is radioactive decay? Describe properties of  $\alpha$ ,  $\beta$ ,  $\gamma$  radiations. Write a note on storage and handling of radioactive substances.
- Define and classify antidotes with examples. Write in short on cyanide poisoning and its treatment. Explain the properties, assay and uses of sodium thiosulphate

**Q. III. Solve any Two (5 Marks each)**

10

- Define expectorants. Explain the properties, assay and uses of ammonium chloride.
- Explain the mechanisms by which antimicrobials act. Write any one method of preparation and uses of hydrogen peroxide.
- Define and classify antacids. write a note on properties and uses of Aluminium hydroxide gel.



**Improvement Sessional Theory Examination 2019-2020**

**Subject:** Communication Skills (BP 105T)

**Day & Date:** Saturday, 16/11/2019

**Class:** F. Y. B. Pharmacy

**Semester:** I

**Time:** 1.30 pm to 03.00 pm

**Max. Marks:** 30

- Instructions:**
1. All questions are compulsory
  2. Draw a well labeled diagram wherever necessary
  3. Right hand side number indicates full marks

**QI. Long Answers (Answer any ONE)**

**1\*10**

- (1) Define Communication and explain Communication Process in detail.
- (2) What is Group Discussion? State its objective, purpose & importance in selection procedure.

**QII. Short Answers (Answer any FOUR)**

**4\*5**

1. What is Psychological Barrier? Explain it in detail.
2. Explain Verbal & Non-Verbal Communication in detail.
3. Explain Do's and Don'ts of Interviews.
4. What are the 4P's of Presentation? Explain them in detail.
5. Explain the listening subtypes with techniques to improve Listening.
6. State the importance of Body Language & Verbal Comm. in Interviews.



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**Improvement Sessional Theory Examination 2018-2019**

**Subject: Human Anatomy and Physiology-I**

**Day & Date:** Thrusday, 14.11.19

**Class: F. Y. B. Pharma**

**Semester: I**

**Time: 10:30-12:00**

**Max. Marks: 30**

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*Instructions: 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks*

**Q1. Solve the following MCQs**

**10**

1. Example of ball-and-socket joint is
  - a. vertebral disc joint
  - b. shoulder and hip joints
  - c. knee and elbow joints
  - d. spinal cord joints
2. The ribs, sternum and spine protect these.
  - a. Kidneys, bladder, urethra
  - b. Heart, lungs, blood vessels
  - c. Small intestine, large intestine
  - d. all the above
3. Tissue is a
  - a. Group of organs
  - b. Group of cells
  - c. Group of tissues
  - d. Group of organisms
4. Blood from the heart flows through the blood vessels in which order?
  - a) Veins, arterioles, capillaries, venules, arteries
  - b) Capillaries, arterioles, arteries, venules, veins
  - c) Arteries, arterioles, capillaries, venules, veins
  - d) Arteries, capillaries, arterioles, venules, veins
5. Blood pressure is highest when leaving which heart chamber?
  - a) Right atrium
  - b) Right ventricle

- c) Left atrium
  - d) Left ventricle
6. Which of the following structures is the primary organ of the integumentary system?
- a) Sensory receptors
  - b) Sweat glands
  - c) Bones
  - d) Skin
7. Which cells produce the pigment that contributes to hair color?
- a) Keratinocytes
  - b) Melanocytes
  - c) Langerhans cells
  - d) Merkel cells
8. The main components of the skeletal system are:
- a) Blood, bones, cartilage and tendons
  - b) Bones, cartilage, joints and tendons
  - c) Bones, cartilage, joints and ligaments
  - d) Cartilage, joints, ligaments and tendons
9. Muscle action can be classified as:
- a) Sympathetic and parasympathetic
  - b) Voluntary and involuntary
  - c) Autonomic and sympathetic
  - d) Autonomic and voluntary
10. Which blood group can be transfused into an O Rh- patient?
- a) O Rh-
  - b) O Rh+
  - c) AB Rh+
  - d) AB Rh-

**QII Long Answers (Answer any 1 out of 2)**

**10**

1. Define tissue. Explain in detail various types of tissue with classification
2. Explain in details conduction system of heart.

**QIII Short Answers (Answer any 2 out of 3)**

**10**

1. Define joints. Draw a neat labelled diagram of synovial joints
2. Write the various functions of skin with a diagram
3. Write the composition and function of blood



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**Improvement Sessional Theory Examination 2019-2020**

Subject : Pharmaceutical Analysis Day & Date : 16/11/2019  
Class : F. Y. B. Pharmacy Semester : I  
Time : 10:30 am to 12:00 noon Max. Marks : 30

Instructions: 1. All questions are compulsory  
2. Draw a well labelled diagram wherever necessary  
3. Right hand side number indicates full marks

- Q.1 Multiple Choice Questions [10]
- a In polarographic method DME is used as \_\_\_\_\_ electrode  
a. reference b. indicator  
c. partial reference d. None of the above
- b Potential of standard hydrogen electrode (SHE) is \_\_\_\_\_  
a. 0.298 b. 0.0  
c. 1.0 d. 0.0298
- c Ilkovic equation can be used in \_\_\_\_\_  
a. Conductometry b. Potentiometry  
c. Polarography d. a & b both
- d In \_\_\_\_\_ titration salicylic acid is used as a specific indicator  
a. Redox b. complexometric  
c. acid base d. precipitation
- e Polarography can be utilizes following except \_\_\_\_\_  
a. migration b. residual  
c. diffusion d. potential difference
- f \_\_\_\_\_ is a self-indicator  
a. Potassium permanganate b. Ceric ammonium sulfate  
c. Iodine solution d. All of the above
- g In gravimetric analysis the term 'sol' refer to \_\_\_\_\_  
a. precipitate b. solution  
c. dispersion d. none of the above
- h \_\_\_\_\_ are involved in Bromatometry type of titrations  
a.  $\text{Br}^-$  b.  $\text{BrO}_4^-$   
c.  $\text{Br}_2$  d. All of the above
- i Conductance can be expressed as \_\_\_\_\_  
a. S b.  $\Omega^{-1}$   
c.  $\Omega$  d. a & b both
- j The unit of current is.....  
a. Amperes b. Volts  
c. Ohm d. mhos
- Q.2. Answer the following short questions (any two) [10]
- a Write a short not on masking and demasking agents. Enlist the factors affecting selectivity of EDTA.
- b Describe the construction, working and application of DME.
- c Explain principle involved in various types of iodine titration.
- d Explain the principle and application of diazotization titration
- Q.3. Answer in detail of following (any one) [10]
- a Explain the term *specific* and *molar conductance*. Describe in detail about various conductometric titration with suitable example.
- b Discuss the steps involved in gravimetric analysis? Comment on the differences between lyophilic and lyophobic sols obtained during precipitation
- c Explain the principle behind potentiometer. Write a construction and working of reference electrode and indicator electrode
- \*\*\*\*\*





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**Improvement Sessional Theory Examination 2019-2020**

Subject : Pharmaceutical Analysis Day & Date : 04/09/2019  
Class : F. Y. B. Pharmacy Semester : I  
Time : 10:00 - 11:30 Max. Marks : 30

Instructions: 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks

Q.1 Multiple Choice Questions [10]

- a The difference between residual and limiting current is called as \_\_\_\_\_  
a. Migration  b. Diffusion   
c. Kinetic  d. None of these
- b Metalochromic indicator used in \_\_\_\_\_ titration.  
a. Argentometric  b. Complexometric   
c. Redox  d. Gravimetric
- c Assay of MgSO<sub>4</sub> is \_\_\_\_\_ type of titration  
a. Redox  b. Acid-Base   
c. Complexometric  d. b & c both
- d The conductance is reciprocal of \_\_\_\_\_  
a. Resistance  b. Specific Resistance   
c. Specific conductance  d. Eq. Conductance
- e The unit of current \_\_\_\_\_  
a. Amperes  b. Volt   
c. Ohm  d. Mhos
- f The no. of gram eq. of solute/lit is called as \_\_\_\_\_  
a. Normality  b. Molarity   
c. Molality  d. Formality
- g \_\_\_\_\_ is a compound of sufficient purity from which a standard solution can be prepared by direct weighing of a quantity of it, followed by dilution to give a defined volume of solution.  
a. External standard  b. Primary standard   
c. Secondary standard  d. a & c both
- h \_\_\_\_\_ are the methods of minimisation of error  
a. Amplification methods  b. Standard addition   
c. Internal standards  d. All of the above
- i Dioxane is a \_\_\_\_\_ solvent  
a. Protophillic  b. Protogenic   
c. Amphiprotic  d. Aprotic
- j Crystal violet is \_\_\_\_\_ indicator  
a. Acidic  b. Basic   
c. Neutral  d. Acidic & Basic both

Q.2. Answer the following short questions (any two) [10]

- a What are primary and secondary standards? Explain their types with suitable example
- b Write short note on construction and working of SHE
- c What are conductometric titrations? Explain its types.
- d Define non aqueous titrations? Explain the types of solvent used.

Q.3. Answer in detail of following (any one) [10]

- a Explain the various neutralisation titration curves with suitable example
- b Explain types of error with suitable example. Describe the various methods used for minimisation of errors
- c Describe in details about the various steps involved in gravimetric analysis

\*\*\*\*\*

b  
b  
c  
a  
a  
a  
a  
b  
d  
d  
a





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**Improvement Sessional Theory Examination 2019-2020**

**Subject:** Pharmaceutics IBP103T

**Day & Date:** 14.11.2019

**Class:** F Y B. Pharm **Semester:** I

**Time:** 90 min

**Max. Marks:** 30

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- Instructions:* 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks

**Q.No.1 Multiple choice questions**

**(10 X 1=10)**

1. Bottle method is used for formulation of emulsion containing ..... oil.  
a) Fixed, b) Mineral, c) Volatile, d) None of the above.
2. .... is used as lubricant for gelato-glycerin suppository bases.  
a) Glycerin, alcohol and liquid paraffin/ arachis oil, b) soap solution, c) mineral oil, d) None of the above.
3. In suspensions, ..... are added for suspending insoluble powder in a vehicle.  
a) Suspending agents, b) emulsifying agents, c) surfactants, d) electrolytes
4. .... is used as o/w emulsifiers.  
a) Acacia, b) Span, c) Magnesium oxide, d) None of the above.
5. HLB value of acacia is.....  
a) 8.0, b) 13.2, c) 4.7, d) 16.7
5. .... are viscous semisolid emulsion containing one or more drugs in dissolved state meant for external application to the skin.  
a) Pastes, b) Creams, c) Gels, d) Suppositories.
6. .... is a widely used synthetic thickening agent in suspensions.  
a) Carbopols, b) Acacia, c) Sodium CMC, d) Bentonite.
7. .... form is most stable form of cocoa butter.  
a)  $\alpha$  b)  $\beta$  c)  $\gamma$  d)  $\sigma$
8. Which of the following is hydrocarbon base  
a) Lanolin, b) White soft paraffin, c) Macrogols, d) Non of the above
9. Emulsifying agents which reduces interfacial tension are classified as.....  
a) Sufactants, b) Finely divided solids, c) Polysaccharides, d) Sterols
10. .... are semisolid unit dosage forms and externally inserted into body cavity to promote defecation.  
a) Enemas b) Pessaries c) Suppositories d) Bougies

**Q. No. 2 Long Answers (Answer 1 out of 2)**

**(1 x 10 = 10)**

1. Discuss stability problems of suspension. Write a note on Suspending agents
2. Define suppository. Write down ideal properties of suppository bases. Discuss various suppository bases along with examples.

**Q. No. 3 Short Answers (Answer 2 out of 3)**

**(2 x 5 = 10)**

1. Write a note on ointment bases.
2. Explain various test used for identification of emulsions.
3. Discuss therapeutic incompatibility

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**Improvement Sessional Examination 2018-19**

**Subject:** BP106RBT Remedial Biology

**Class:** B.Pharm (Sem-I)

**Total Marks:** 30

**Time:** 60 min

**Roll. No.:**

**Date:** 1 Dec 2018

**Q1. Answer any one**

(10 marks)

1. Explain in details about cardiovascular system. Explain Cardiac cycle
2. Explain the photosynthesis process in details

**Q2. Answer any fours**

(20 marks)

1. What is blood and explain its Composition and function
  2. Short note on breathing process and its regulation
  3. Describe Role of digestive enzymes in digestion process
  4. Explain different functions of micro and macro nutrients in plant growth.
  5. Short note on ECG
  6. Explain Pituitary gland and explain different function of hormone release from it
-



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**Second Sessional Theory Examination 2020-2021**  
**(Odd SEM)**

**Subject:** Pharmaceutical Analysis I (BP102T)

**Day & Date:** 22-04-2021

**Class:** First Year B. Pharm

**Semester:** I

**Time:** 10:30 am – 12:00 pm

**Max. Marks:** 30

*Instructions: 1. All questions are compulsory*

*2. Draw a well labeled diagram wherever necessary*

*3. Right hand side number indicates full marks*

**QI. Multiple Choice Questions (MCQs)**

**10**

1. Closeness of observed value with true value is .....

  - a) accuracy
  - b) linearity
  - c) precision
  - d) relative error

2. Given the following are the type of systemic error except

  - a) Error of methods
  - b) Instrumental method
  - c) Personal error
  - d) Random error

3. In which step, ions or elements are aggregated in Gravimetric analysis?

  - a) Supersaturation
  - b) Nucleation
  - c) Particle growth
  - d) None of the above

4. When more than two ions in solution and simultaneously second ion is also precipitated it is called as .....

  - a) Co precipitation
  - b) Post precipitation
  - c) A and B
  - d) None of the above

5. The reagent which undergoes reduction is an ..... agent and the reagent which undergoes oxidation is a .....

  - a) Oxidizing, Reducing
  - b) Reducing, Oxidizing
  - c) Complexing, Reducing
  - d) None of these

6. Which method are used to determination of primary amine?



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- a) Diazotization Titration  
b) Karl fischer titration  
c) Redox titration  
d) All of the above
7. Indicator used in standardization of sodium thiosulphate is  
a) Starch  
b) Phenolphthaline  
c) ceric sulphate  
d) Methylene red
8. When an analyte that is a reducing agent is titrated directly with a standard iodine solution, the method is called  
a) iodimetry  
b) iodometry  
c) bromatometry  
d) cerimetry
9. Complexing agent is .....  
a) Electron donating ion  
b) Electron accepting ion  
c) Both a & b  
d) None of the above
- 10 The diffusion current in polarography is expressed by the equation  
a) Beer's law  
b) Nernst equation  
c) Bragg's equation  
d) Ilkovic equation

**QII. Solve the following (any 1)**

**10**

1. Define gravimetric analysis. Explain in detail the steps involved in gravimetric analysis
2. State the basic principle involved in polarography. Write down the construction and working of dropping mercury electrode

**QIII Short Answers (Answer any 2 out of 3)**

**10**

1. Explain the titration involving iodine
2. Write down the construction and working of normal hydrogen electrode
3. Write principle and application of Diazotization titrations.





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**First Sessional Theory Examination 2020-2021**

**Subject:** Pharmaceutical Analysis I (BP102T)

**Day & Date:** 11-03-2021

**Class:** First Year B. Pharm.

**Semester:** I

**Time:** 10:30 am – 12:00 pm

**Max. Marks:** 30

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*Instructions:*

1. All questions are compulsory
2. Draw a well labeled diagram wherever necessary
3. Right hand side number indicates full marks

**Q. I. Solve the following Multiple Choice Questions (MCQs)**

**10**

1. Number of gram equivalents of solute present in one litre of a solution is called as .....
  - a) Molarity
  - b) Normality
  - c) Molality
  - d) Percentage weight by volume
2. The features of Primary standard substance includes:
  - a) High purity
  - b) High stability
  - c) Low hygroscopicity
  - d) All of the above
3. Indian pharmacopoeia is published by .....
  - a. World health organization
  - b. Ministry of health and family welfare
  - c. Ministry of science and technology
  - d. Food and drug administration
4. Limit test of iron is based upon the reactions between iron and ....
  - a. Dithiazone
  - b. Thioglycolic acid
  - c. Hydrogen sulphide
  - d. Barium chloride
5. The color of phenolphthalein in alkaline medium is .....
  - a) Pink
  - b) Yellow
  - c) Brown
  - d) Colorless

6. Benzene is an example of .....solvent.
- Aprotic
  - Protogenic
  - Protophilic
  - Amphiprotic
7. Solubility of precipitate is ..... on addition of common ion.
- Decreased
  - Increased
  - Remains Constant
  - None of These
8. In Mohr's method .....is used as an indicator.
- Potassium chromate
  - Ferric ammonium sulphate
  - Ammonium thiocyanate
  - Silver nitrate
9. In Volhard's method .....is used as a titrant
- Potassium chromate
  - Ferric ammonium sulphate
  - Ammonium thiocyanate
  - Silver nitrate
10. The end-point detection in Fajan's method involves color change due to.....
- Formation of colored precipitate
  - Formation of soluble colored complex
  - Adsorption of indicator on the surface of precipitate
  - None of the above

**Q. II. Solve the following (any 1)**

**10**

1. Explain the neutralization curves in acid-base titrations for i) strong acid vs strong base ii) weak acid vs strong base iii) strong acid vs weak base iv) weak acid vs weak base.
2. Explain in detail Mohr's method and Fajan's method.

**Q. III. Solve the following (any 2)**

**10**

1. Explain principle involved in limit test for Chloride and sulphate.
2. Explain in detail the theories of acid-base indicators.
3. Explain the types of solvents used in non-aqueous titrations with suitable examples.





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**Second Re-Sessional Practical Examination 2020-2021**

**Subject:** Pharmaceutical Analysis - I **Class:** F. Y. B. Pharm. Semester: I

**Batch :** --

**Date** 04/05/2021

**Max. Marks:** 40

Q.N.		Marks
1	Synopsis (any four)	10
	<i>i</i> Predict the method and write the principle suitable for assay of the following:	06
	a. Ammonium Chloride	
	b. Ferrous sulfate	
	c. Copper Sulphate	
	<i>ii</i> Explain the principle and reaction used in limit test for	04
	a. Iron	
	b. Arsenic	
2	To report the % content of given Hydrogen peroxide by permanganometry	15
3	To perform the limit test of Sulfate from the given sample and report whether it passes or fails the test	10
4	Viva-voce	05

(C.B.R. Dravyakar)

Continuous Assessment Theory Examination 2020-2021

Subject: Pharmaceutical Analysis I

Class: First Year B. Pharm.

Semester: I

Max. Marks: 03

Division A and B

Open Book Test	Date: 14-04-2021
Q.1 Write a note on Dropping Mercury electrode. ....3M	
Assignment	Date: 12-04-2021
Q.1. Write a note on construction and working of standard hydrogen electrode. ....3M	
Quiz	Date: 25-03-2021
<b>Q.1. Solve any Three multiple-choice questions. ....3M</b>	
1. Complexing agent is .....	
a. Electron donating ion	
b. Electron accepting ion	
c. Both a & b	
d. None of the above	
2. Which sentence is <b>false</b> about ligand?	
a. It acts like Lewis base	
b. $\text{CN}^-$ is example of unidentate.	
c. It is electron donor group.	
d. EDTA is example of unidentate.	
3. EDTA is a .....	
a. Tetradentate ligand	
b. Hexadentate ligand	
c. Octadentate ligand	
d. Pentadentate ligand	
4. All of the following acts as masking agents, except .....	
a. Ammonium fluoride	
b. Iodide	
c. Iron	
d. Triethanolamine	
5. Which of the following acts as demasking agent?	
a. Ascorbic acid	
b. Magnesium	
c. Triethanolamine	
d. Formaldehyde - acetic acid	



## Preliminary Theory Examination 2020-21

Questions Responses 108

Total points: 75

## Preliminary Theory Examination 2020-21

FY B. Pharm SEM I

Subject : Pharmaceutics I

Maximum Marks 75

Date: 12.5.2021

Name of student



Short answer

Short-answer text

 (0 points)Required 

Roll No:

Short-answer text

PRN NO:

Short-answer text

Class &amp; Division

Short-answer text

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1/21

Wetting agent

- Suspending agent
- b and c

5. When the action of the drug is opposed by the other drug, the phenomenon is known as \*

- Antagonism
- Analgesics
- Addition
- Antioxidant

6. A suitable ..... is added to maintain the powders until the mass becomes coherent but not too damp \*

- Colouring agent
- Flavouring agent
- Sweetening agent
- Granulating agent

7. Enemas are also called as..... \*

- Elixir
- Suspension
- Emulsion
- Clystric

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3/21

Short-answer text

1. The first pharmacy shop was opened in ..... \*

- London
- Damascus
- Bagdad
- Chennai

2. The chairman of the first edition of IP was.... \*

- Dr. B.N. Ghosh
- Dr. B. Mukherjee
- Dr. Nityanand
- Mr. PrasanaTotta

3. Every dosage form is a combination of drug and different kind of non-drug components called..... \*

- Additives
- Non-Additives
- New chemical entity
- All of these

4. Diffusible drugs are formulated without the use of ..... \*

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2/21

Saccharine

- Sucrose
- Fructose
- Glucose

9. Disodium EDTA is an example of ... \*

- Sweetening agent
- Flavouring agent
- Chelating agent
- Preservative

10. Which of the following is anti-foaming agent \*

- Methyl cellulose
- Aspartame
- Sucrose
- Simethicone

11. The part of prescription called subscription contains \*

- Direction to the pharmacist
- Direction to the patient
- Direction to the patient's relatives
- All of these

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4/21

- Silps
- Lapse
- Knowledge
- Reports

13. Douches are meant for application in.....\*

- Buccal cavity
- Rectal cavity
- Vaginal cavity
- Nasal cavity

14. A genetically determined abnormal or unusual response to a drug is.....\*

- Idiosyncrasy
- Tolerance
- Genetic polymorphism
- Salicylism

15. Powders used for external use are ...\*

- Dusting powder
- Compound powder
- Divided powders
- Effervescent powders



16. In purified water the concentration of sugar according to IP is ....\*

- 85% W/W
- 66.67 % W/W
- 85%W/V
- 66.67%W/V

17. Who is the father of medicine.....\*

- Ebers
- Hippocrates
- Egyptian
- Pontus

18. Use of formulations made up of numerous plants referred as.....\*

- Galenicals
- Parenteral
- Plant Vehicles
- Generics

19. The seventh edition of IP was published in ....\*

- 2014
- 2015
- 2013
- 2011



20. Pharmacy Council of India (PCI) was established in ....\*

- 1947
- 1946
- 1949
- 1950

21. The First US pharmacopoeia was published in ....\*

- 1820
- 1822
- 1823
- 1821

22. The first National formulary was published in ....\*

- 1860
- 1885
- 1888
- 1890

23. The ..... Samhita includes reference to drugs of animals, plants and mineral origin used until the first century AD \*

- Ebers
- Charaka
- Hippocrates



Shushruta

24. The "Pharmacy Act" came in force in ....\*

- 1947
- 1948
- 1949
- 1950

25. Drugs converted to suitable form are known as....\*

- Excipient
- Dosage form
- API

26. Acute tolerance is also known as.....\*

- Addiction
- Idiosyncrasy
- Tachyphylaxis
- Habituation

27. What will increase the bulk of product.....\*

- Processing aid
- Diluent
- Granulating Agent



28. Which powder must be sterilized before their use.....\*

- Medicated dusting powders
- Non medicated powders
- Surgical dusting powders
- Natural dusting
- powder

29. The monophasic liquid dosage form is ..... Solutions \*

- True
- Course
- Fine
- Micro

30. Nasal drops should have pH..... \*

- 5-6
- 6-7
- 7-8
- 8-9

31. Which of the following is not monophasic liquid dosage form..... \*

- Solution



Benzethonium chloride

Thiomersal

36. Cochineal is a ..... Agent \*

- Flavouring agent
- Sweetening agent
- Colouring agent
- Thickening agent

37. Which of the following is anti-foaming agent \*

- Methyl cellulose
- Aspartame
- Sucrose
- Simethicone

38. The superscription is represented by a symbol.... \*

- Rx
- Dc
- Doc
- Ph

39. CPT consists of..... \*

20% acacia, 20% tragacanth, 15% starch, 45% Sucrose



Suspension

Enemas

32. Elixirs are ..... \*

- Hydro alcoholic liquid
- Aqueous
- Viscous
- Semi solids

33. Liniments must not be applied on the .....skin \*

- Swelled
- Broken
- Painful
- Normal

34. .... drops are used in otic cavity \*

- Eye
- Ear
- Nose
- Rectum

35. Which of the following excipient act as preservative..... \*

Sorbitan monoactrate



20% acacia, 15% tragacanth, 30% starch, 35% sucrose

20% acacia, 25% tragacanth, 30% starch, 40% Sucrose

20% acacia, 15% tragacanth, 20% starch, 45% Sucrose

40. 1mL = ..... \*

- 16.23 ml
- 24 ml
- 15 ml
- 7 ml

41. 24 3 = \*

- 1lb
- 2lb
- 6lb
- 0.5lb

42. Agents are added to prevent coalescence of the dispersed globules by forming barriers at interface or reduces interfacial tension. \*

- Emulsifying agents
- Suspending agents
- Flocculating agents
- none of above

43. .... is known as 'Father of Pharmacy education in India' \*





M. L. Shroff

- R. N. Chopra
- B. N. Ghosh
- None of the above

44. .... are viscous liquid used in cough that relieves irritation of the mucous membranes in the throat/mouth by forming a protective film. \*

- Demulcent
- Sedatives
- Elixirs
- throat paints

45. A drug that neutralizes excess gastric acid are ..... \*

- Antacid
- Analgesic
- Antiamebic
- Antiepileptic

46. The powders containing solids that absorb moisture from air and dissolve to form solution are called \*

- Deliquescent powders
- Efflorescent powders
- Eutectic mixtures
- Compound powders



47. Daily adult dose of drug is 120mg. How much of it should be given to a child of 11 years? \*

- 57.39
- 60
- 50.39
- 37

48. Which of the following co-solvents are used to increase the solubility of a drug? \*

- Ehanol
- Sorbitol
- Glycerine
- All the above

49. Which part of prescription contains direction to the patient regarding how, how much and when the drug should be taken? \*

- Superscription
- Subscription
- Inscription
- Signatura

50. How will you mix potent drug with large amount of diluent? \*

- Spatulation
- Geometric dilution
- Sifting



51. The dosage form intended to be rubbed on the skin for rubeficient and counterirritant effect are called \*

- Lotions
- Liniments
- Gels
- Dentifrices

52. The powders on exposure to humid atmosphere or during trituration, liberates wholly or partly, the water of crystallization are called \*

- Hygroscopic powders
- Efflorescent powders
- Eutectic mixtures
- Compound powders

53. Drugs that destroys or eliminates intestinal parasitic worms are called \*

- Analgesic
- Anthelmintic
- Antibiotic
- Antifilarial

54. The semisolid unit dosage form intended for vaginal administration are called \*

- Suppositories



Bougies

Cones

55. 1j (apothecaries) = \*

- 437.5 grains
- 480 grains
- 7000 grains
- 5760 grains

56. Which of the following is not used as an emulsifying agent? \*

- Surfactant
- Hydrophilic colloid
- Electrolytes
- Finely divided solids

57. Which of the following hydrocarbon waxes are employed in the manufacture of creams and ointments? \*

- Paraffin wax
- Ceresin
- Both
- none

58. In the preparation of cold creams, which types of bases are used generally? \*



- Water renmovable bases
- Hydrocarbon bases
- None

59. Which of the following is most commonly used suppository base \*

- Cocoa butter
- PEG 1000
- PEG + Hexanetriol
- None

60. Chemical incompatibility in which change in formulation is needed is called as. \*

- Tolerated incompatibility
- Adjusted incompatibility.

61. Tetracycline interacts with calcium to form chelates is an example of ..... \*

- Chemical incompatibility
- Therapeutic incompatibility
- Physical incompatibility
- Tolerance incompatibility.

62. In conductivity test, if bulb glows on passing electric current so what would be the type of emulsion? \*

- W/O emulsion



- De-flocculated

67. Which of the following is an example of Hydrophilic bases? \*

- Hydrogenated oils
- Emulsified cocoa butter
- Glycero-gelatin base
- None of the above

68. Which of the following is substitute for Theobroma oil? \*

- Hydrogenated oils
- Emulsified cocoa
- Polyethylene-Glycol
- None

69. A wetting agent is included in the formulation of a suspension, particularly when the suspended particles: \*

- are hydrophobic
- are more denser than the vehicle
- are water soluble
- have lesser Interfacial tension

70. Suspended particles become flocculated in a suspension, because: \*

- attractive forces between particles are appreciable



- Micro emulsion
- Multiple emulsion

63. In case of O/W emulsion, creaming takes place at \_\_\_\_\_ \*

- Down side
- Upside
- At interface between two phase
- None of the above

64. What is the dispersion of a liquid in another liquid called? \*

- Gel
- Suspension
- Emulsion
- Aerosol

65. For Ideal suspension ,the sedimentation volume should be \*

- Zero
- Equal to one
- More than one
- Less than one

66. Rate of sedimentation is high in? \*



- repulsive forces between particles are appreciable
- vehicle rejects the particles

71. Which type of colloidal material is used commonly in the preparation of a structured vehicle? \*

- association
- hydrophilic
- hydrophobic
- inorganic

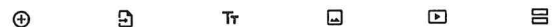
72. The ratio of the ultimate volume of sediment to the actual volume of sediment before settling is called \*

- Sedimentation volume
- Degree of flocculation
- Emulsification volume
- phase volume ratio

73. The HLB range of an emulsifier employed in the preparation of water-in-oil emulsion is:

- 3 to 6
- 7 to 12
- 13 to 15
- more than 15

74. Auxiliary emulsifying agents are used to stabilize the emulsion. They act on the principle: \*



- strengthening the nonpolar tails of the emulsifier \*
- strengthening the polar heads of the emulsifier
- thickening the continuous phase

75. In wet gum method, 4:2:1 consist of \*

- 4 parts oil, 2 parts water, and 1 part emulsifier
- 4 parts water, 2 parts oil, and 1 part emulsifier
- 4 parts emulsifier, 2 parts water, and 1 part emulsifier
- 4 parts oil, 2 parts emulsifier and 1 part water



- Slips
- Lapse
- Knowledge
- Reports

13. Douches are meant for application in.....\*

- Buccal cavity
- Rectal cavity
- Vaginal cavity
- Nasal cavity

14. A genetically determined abnormal or unusual response to a drug is.....\*

- Idiosyncrasy
- Tolerance
- Genetic polymorphism
- Salicylism

15. Powders used for external use are ...\*

- Dusting powder
- Compound powder
- Divided powders
- Effervescent powders



20. Pharmacy Council of India (PCI) was established in ....\*

- 1947
- 1948
- 1949
- 1950

21. The First US pharmacopoeia was published in ....\*

- 1820
- 1822
- 1823
- 1821

22. The first National formulary was published in ....\*

- 1860
- 1885
- 1888
- 1890

23. The ..... Samhita includes reference to drugs of animals, plants and mineral origin used until the first century AD \*

- Ebers
- Charaka
- Hippocrates



16. In purified water the concentration of sugar according to IP is .....

- 85% W/W
- 66.67 % W/W
- 85%W/V
- 66.67%W/V

17. Who is the father of medicine.....\*

- Ebers
- Hippocrates
- Egyptian
- Pontus

18. Use of formulations made up of numerous plants referred as.....\*

- Galenicals
- Parenteral
- Plant Vehicles
- Generics

19. The seventh edition of IP was published in ....\*

- 2014
- 2015
- 2013
- 2011



Shushruta

24. The "Pharmacy Act" came in force in ....\*

- 1947
- 1948
- 1949
- 1950

25. Drugs converted to suitable form are known as.....\*

- Excipient
- Dosage form
- API

26. Acute tolerance is also known as.....\*

- Addiction
- Idiosyncrasy
- Tachyphylaxis
- Habituation

27. What will increase the bulk of product.....\*

- Processing aid
- Diluent
- Granulating Agent







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**Remedial Sessional Theory Examination 2019-2020**

**Subject:** P'ceutical Inorganic Chemistry (BP104T) **Day & Date:** Fri, 06/09/2019

**Class:** First Year B. Pharm.

**Semester:** I

**Time:** 10.30 am – 12.00 pm

**Max. Marks:** 30

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*Instructions: 1. All questions are compulsory*

**Q.I. Multiple Choice Questions (MCQs)**

**10**

1. Lewis base is.....
  - a) electron pair donor
  - b) electron pair acceptor
  - c) proton donor
  - d) proton acceptor
2. Citric acid is used in limit test for iron to prevent.....
  - a) precipitation of iron by ammonia
  - b) reduction of iron
  - c) precipitation of iron by thioglycollic acid
  - d) complexation with thioglycollic acid
3. Indian pharmacopoeia is published by.....
  - a) Ministry of health and family welfare
  - b) food and drug administration
  - c) world health organization
  - d) ministry of science and technology
4. Lead acetate cotton plug is used in limit test for arsenic to.....
  - a) trap hydrogen sulphide
  - b) trap thiosulphate
  - c) trap hydrogen sulphate
  - d) trap arsine gas
5. Limit test for heavy metals is based on reaction between metallic impurities and .....
  - a) hydrogen sulphide
  - b) hydrogen sulphate
  - c) acetic acid
  - d) ammonia
6. The first Indian pharmacopoeia was published by Indian Pharmacopoeia committee in .....
  - a) 1948
  - b) 1955
  - c) 1960
  - d) 1966
7. gamma rays are electromagnetic radiations with a wavelength .....
  - a) much larger than those of light
  - b) much shorter than those of light
  - c) Equal to light
  - d) none of the above
8. Unit of radioactivity is .....
  - a) Calorie
  - b) Newton
  - c) Curie
  - d) Joules
9. The sources of impurities include .....
  - a) raw materials
  - b) manufacturing process
  - c) instability of product
  - d) all of these
10. The limit test for chlorides involves reaction between soluble chlorides and .....
  - a) silver nitrate
  - b) nitric acid
  - c) barium sulphate
  - d) HCl

**Q. II. Solve any One**

**10**

1. Explain different theories of acid and base. Explain the mechanism of buffer action.
2. What is radioactivity? Explain the properties of alpha, beta and gamma radiations.

**Q. III. Solve any Two (5 Marks each)**

**10**

1. Explain the principle behind limit test of Iron and Sulphate.
2. Explain the principle of arsenic limit test. Draw a neat labelled diagram of Gutzeit apparatus.
3. Explain the sources of impurities.



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Institute of Pharmacy, Dhule**

Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

**Remedial Sessional Theory Examination 2018-2019**

**Subject: Pharmaceutics I BP103T**

**Day & Date: 05.09.2018**

**Class: F Y B. Pharm**

**Semester: I**

**Time: 10 am to 11.30 pm**

**Max. Marks: 30**

**Instructions:** 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks

**Q.No.1 Multiple choice questions (10 X 1=10)**

- ..... are used to deodorize oral cavity.  
a) Throat paint, b) Mouth washes, c) Elixirs, d) Syrups.
- Suspension containing loose aggregation of particles is called .....  
suspension. a) Flocculated b) Deflocculated, c) Mixture, d) None of the above.
- ..... form is most stable form of cocoa butter. a)  $\alpha$  b)  $\beta$  c)  $\gamma$  d)  $\sigma$
- According to Stokes's equation rate of sedimentation will ..... with increase in  
viscosity of dispersion phase.  
a) Decrease, b) Increase, c) Remain unaffected, d) None of the above.
- ..... are semisolid unit dosage forms and externally inserted into body cavity to  
promote defecation. a) Enemas b) Pessaries c) Suppositories d) Bougies.
- ..... is used as lubricant for gelato-glycerin suppository bases.  
a) Glycerin, alcohol and liquid paraffin/ arachis oil, b) soap solution, c) mineral oil, d) None of  
the above.
- If water soluble dye (amaranth) is mixed with the emulsion and examined under the  
microscope, the dispersed globules appears red and ground colorless indicates.....  
a) o/w type, b) w/o type, c) w/o/w, d) None of the above.
- Emulsifying agents which reduces interfacial tension are classified as.....  
a) Sufactants, b) Finely divided solids, c) Polysaccharides, d) Sterols.
- In wet gum method containing fixed oil ratio of oil: water:acacia to formulate primary  
emulsion is..... a) 4:2:1, b) 3:2:1, c) 2:2:1, d) 2:2:1
- ..... are clear, pleasantly flavored, sweetened hydroalcoholic monophasic liquid  
containing dissolved active ingredients intended for oral use.  
a) Syrups b) Elixirs c) Linctus d) Suspension.

**Q. No. 2 Long Answers (Answer 1 out of 2) (1 x 10 = 10)**

- Define emulsion. Write down ideal properties of emulsion. Write a note on emulsifying agents.
- Define dosage form. Discuss dosage form in detail.

**Q. No. 3 Short Answers (Answer 2 out of 3) (2 x 5 = 10)**

- Write a note on lotion and liniment.
- Define suspension. Discuss ideal properties of suspension.
- Write short note on parts of prescription.



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## INSTITUTE OF PHARMACY, DHULE

Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai-Agra National Highway

### Re-Sessional Theory Examination 2020-2021

**Subject: Comm. Skills**

**Day & Date: Friday, 07 May 2021**

**Class: FY/ DSA**

**Semester: First**

**Time: 10.30 to 12.00**

**Max. Marks: 30**

*Instructions: 1. All questions are compulsory  
2. Draw a well labeled diagram wherever necessary  
3. Right hand side number indicates full marks*

#### QI Long Answers (Answer any ONE)

1\*10

- (1) Explain the purpose and process of Selection Interview. What are the Dos and Don'ts of Selection Interview? (Unit 04)
- (2) Explain Group Interaction in detail? Elucidate the salient characteristics of Group Discussion. (Unit 05)

#### QII Short Answers (Answer any FOUR)

4\*5

1. Explain the role of Effective Writing in your profession in detail. (Unit 03)
2. Explain the elements involved to make writing more effective. (Unit 03)
3. How will you be Effective Listener? State the strategies and methods you will adopt for Effective Listening.(Unit 03)
4. How will you prepare and deliver your presentation more effectively? (Unit 04)
5. What are the Dos and Don'ts of Group Discussion? (Unit 05)
6. What are the ways of individual contribution in GD? (Unit 05)

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