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First Sessional Theory Examination 2020-2021 (Odd SEM) Subject & Subject Code: Remedial Mathematics (BP106RMB)

Day & Date: Tuesday 16/3/2021

Class: F.Y/S.Y. B. Pharmacy

Semester: I/III

Time:1 hr

Max. Marks: 30

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks

QI Multiple Choice Questions (MCQs)

1. If A and B are symmetric matrices of the same order, then (a) AB is a symmetric matrix (b) A – B is askew-symmetric matrix (c) AB + BA is a symmetric matrix (d) AB – BA is a symmetric matrix 2. If $A = \begin{bmatrix} 2 & 4 \\ x & 0 \end{bmatrix}$ is a symmetric matrix, then x = (a) 4 (b) 3 (c) -4 (d) -3 **3.** If A is a square matrix, then $A - A^T$ is a (a) diagonal matrix (b) skew-symmetric matrix (c) symmetric matrix (d) none of these 4. If A is any square matrix, then which of the following is skew-symmetric? (a) $A + A^{T}$ (b) $A - A^T$ (c) AA^T (d) $A^{T}A$ 5. Let A is any square Matrix Then A^{-1} exist if (a) $|A| \neq 0$ (b) |A| = 0(c) A is singular Matrix (d) None 6. The value of $\begin{vmatrix} 1 & 2 & 3 \\ -3 & -6 & -9 \\ 2 & 1 & 0 \end{vmatrix}$ is (a) 4 (b) 3 (c) -4 (d) 0



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7. The value of
$$\begin{vmatrix} 1 & 1 & 1 \\ 1 & 1 & 1 \\ 1 & 1 & 1 \end{vmatrix}$$
 is
(a) 2
(b) 3
(c) 0
(d) 1
8 A square Matrix A is said to Symmetric if
(a) A = A^T
(b) A = - A^T
(c) A + A^T is identity
(d) A^TA exist.
9. If $\begin{vmatrix} 1 & 2 \\ 3 & x \end{vmatrix} = 0$ then x is
(a) 2
(b) 3
(c) 6
(d) 1
10 If A= $\begin{bmatrix} 1 & 2 \\ 3 & 7 \end{bmatrix}$ then A^{-1} is
(a) $\begin{bmatrix} 7 & -2 \\ -3 & 1 \end{bmatrix}$
(b) $\begin{bmatrix} 1 & -2 \\ 3 & 7 \end{bmatrix}$
(c) $\begin{bmatrix} 1 & 2 \\ -3 & 7 \end{bmatrix}$
(d) $\begin{bmatrix} 1 & -2 \\ -3 & 7 \end{bmatrix}$

QII Long Answers (Answer any 1 out of 2)

1. Solve the following System of equations: 2x + 3y - z = 14x + y - 3z = 113x-2y+5z = 21

2. Solve the following System of equations: -x + 3y-2z = 54x - y - 3z = -82x + 2y - 5z = 7

QIII Short Answers (Answer any 2 out of 3)

QIII Short Answers (Answer x_{1} , x_{2} , x_{2} , x_{3}) 1. If $A = \begin{bmatrix} 1 & 2 & 3 \\ -1 & 0 & 2 \\ -2 & 4 & 1 \end{bmatrix}$ Find A^{-1} using adjoint method. 2. Find the value of x if $\begin{vmatrix} 1 & 2 & 3 \\ 1 & x & 4 \\ 2 & -3 & -1 \end{vmatrix} = 0.$ 3. Find minors and cofactors of each elements of A= $\begin{bmatrix} 0 & 1 & 2 \\ 3 & -1 & 6 \\ 2 & -4 & -8 \end{bmatrix}$



Shri. Vile Parle Kelavani Mandal's Institute of Pharmacy, Dhule

Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

First Sessional Theory Examination 2018-2019

Subject: Comm. Skills	Day & Date: Saturday, 6 th October 2018
Class: FY/ DSA	Semester: First
Time: 10:30 am to 12:00 pm	Max. Marks: 30

Instructions: 1. All questions are compulsory

Draw a well labeled diagram wherever necessary
 Right hand side number indicates full marks

QI Long Answers (Answer any 1 out of 2)

(1) Define Communication and illustrate its importance or objectives.

(2) What are the elements of Communication? Explain them in detail with Comm. Cycle.

QII Short Answers (Answer any 4 out of 6)

1. Explain Physiological barrier with its all subtypes in detail with examples.

- 2. What is linguistic barrier? Explain it in detail with all its sub-types.
- 3. What is Comm. Styles Matrix? Explain Direct and Spirited Comm. Styles with example.
- 4. State the importance of Body language with its constituents in detail.
- 5. Explain Psychological barrier with its all subtypes in detail with examples.
- 6. What are Systematic and Considerate Comm. Styles? Write a note on factors affecting our perspectives.

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First Sessional Theory Examination 2020-2021

Subject: Comm. Skills	Day & Date: Monday, 15 March. 2021
Class: FV/ DSA	Semester: First
Time: 10 30 to 11.30	Max. Marks: 30

Instructions: 1. All questions are compulsory

2. Draw a well labeled diagram wherever necessary

3. Right hand side number indicates full marks

QI Long Answers (Answer any ONE)

- (1) Non-verbal communication enhances the efficacy of professional communication. Explain the statement in relation with all non-verbal codes of communication. (Unit 02)
- (2) Explain the process of Communication with its elements in detail. (Unit 01)

QII Short Answers (Answer any FOUR)

- 1. What is Communication? Define it in detail. (Unit 01)
- Explain Psychological barrier with all its sub-types and suggest ways to overcome it. (Unit 01)
- 3. What are the types and methods of Communication? (Unit 01)
- 4. State the importance of Body language with its constituents in detail. (Unit 02)
- 5. Explain in detail the factors that affect perspective in day-to-day communication. (Unit 02)
- 6. What are the different communication styles? What do they focus on? (Unit 02)

4*5

1*10



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First Sessional Theory Examination 2020-2021

Subject: Comm. Skills	Day & Date: Monday, 15 March. 2021
Class: FY/ DSA	Semester: First
Time: 10.30 to 12.00	Max. Marks: 30

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks

QI Long Answers (Answer any ONE)

- (1) *Non-verbal communication enhances the efficacy of professional communication*. Explain the statement in relation with all non-verbal codes of communication.
- (2) Explain the process of Communication with its elements in detail.

QII Short Answers (Answer any FOUR)

- 1. What is Communication? Define it in detail.
- 2. Explain Psychological barrier with all its sub-types and suggest ways to overcome it.
- 3. What are the types and methods of Communication?
- 4. State the importance of Body language with its constituents in detail.
- 5. Explain in detail the factors that affect perspective in day-to-day communication.
- 6. What are the different communication styles? What do they focus on?

1*10

4*20



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First Sessional Theory Examination 2020-2021

Subject: Comm. Skills	Day & Date: Monday, 15 March. 2021
Class: FY/ DSA	Semester: First
Time: 10.30 to 12.00	Max. Marks: 30

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks

QI Long Answers (Answer any ONE)

- (1) *Non-verbal communication enhances the efficacy of professional communication*. Explain the statement in relation with all non-verbal codes of communication.
- (2) Explain the process of Communication with its elements in detail.

QII Short Answers (Answer any FOUR)

- 1. What is Communication? Define it in detail.
- 2. Explain Psychological barrier with all its sub-types and suggest ways to overcome it.
- 3. What are the types and methods of Communication?
- 4. State the importance of Body language with its constituents in detail.
- 5. Explain in detail the factors that affect perspective in day-to-day communication.
- 6. What are the different communication styles? What do they focus on?

1*10

4*20



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First Sessional Theory Examination 2021-2022

Subject: Communication Skills (BP105T)	Day & Date: Friday, 11/2/2022						
Class: FY/ DSA	Semester: I & III						
Time: 10.30 am to 11.30 am	Max. Marks: 30						

Instructions: 1. All questions are compulsory

2. Draw a well labeled diagram wherever necessary

3. Right hand side number indicates full marks

QI) Long Answer Question (Any One) (CO1)	1*10

- 1. Explain the process of Communication with its elements in detail.
- 2. What is communication? Explain communication perspective.

QII) Short Answer Questions.

Section A: (Any One) (CO1) 1*5

3*5

- 1. Explain Psychological barrier with all its sub-types and suggest ways to overcome it.
- 2. What are the types and methods of Communication? Illustrate them in detail.

Section B: (Any Three) (CO2)

- 1. Is Hearing be Listening? Explain how to be effective listener.
- 2. What are non-verbal codes of communication? Explain them in detail.
- 3. What is Communication Style Matrix? Explain it in detail.
- 4. What is listening? What are its types?

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First Sessional Theory Examination 2021-2022

Subject: Communication Skills (BP105T)	Day & Date: Friday, 11/2/2022						
Class: FY/ DSA	Semester: I & III						
Time: 10.30 am to 11.30 am	Max. Marks: 30						

Instructions: 1. All questions are compulsory

2. Draw a well labeled diagram wherever necessary

3. Right hand side number indicates full marks

QI) Long Answer Question (Any One) (CO1)	1*10

- 1. Explain the process of Communication with its elements in detail.
- 2. What is communication? Explain communication perspective.

QII) Short Answer Questions.

Section A: (Any One) (CO1) 1*5

3*5

- 1. Explain Psychological barrier with all its sub-types and suggest ways to overcome it.
- 2. What are the types and methods of Communication? Illustrate them in detail.

Section B: (Any Three) (CO2)

- 1. Is Hearing be Listening? Explain how to be effective listener.
- 2. What are non-verbal codes of communication? Explain them in detail.
- 3. What is Communication Style Matrix? Explain it in detail.
- 4. What is listening? What are its types?

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			First Sessional Theory Examination	ation-(ODD S	SE	M) 202	22-2	023	
Subje	ect	:	Communication skills (BP105T)	Day & Date	:	WEDN	ESD	AY,08/03	/2023
Class	5	:	First Year B. Pharmacy	Semester	:	Ι	Writ	te Your Seat I	Vo. Here
Time	:	:	10:30 am – 11:30 am	Max. Marks	:	30			
Instruc	tions	1. 3.	All questions are compulsory Right hand side number indicates full marks	2. Draw a well l 4. Do not write/t	labe tick	eled diagro on the que	ım wh estion	erever neces paper	sary
QI	Lor	ng A	Answers Question (Answer any 1 out of 2)						10 M
1.	Illu	stra	ate the process of Communication with its eler	nents in detail. (I	LL3	3, CO1)			
2.	Ex	pla	in the concept of communication along with co	ommunication pe	ersp	ectives.	(LL3,	CO1)	
QII	Sho	ort A	Answers Question						
	Sec	tior	n – A (Answer any ONE)						05 M
1.	Exp	lair	the psychological barrier in detail along with your v	vays to overcome it	. (L	LL2, CO1)		
2.	Inte	rpre	et the linguistic barrier in detail.						
	Section – A (Answer any Three) 15 M								
1. Explain Verbal and Non-verbal communication in detail. (LL2, CO2)									
2.	2. Compare any two communication styles. (LL2, CO2)								
3.	3. Interpret the process of Listening in detail. (LL2, CO2)								
4.	4. Illustrate the significance of Body Language at workplace. (LL2, CO2)								



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			First Sessional Theory Examinatio	n-(ODD SEI	M) :	2023	3-20)24	
Subje	ect	:	Communication skills (BP105T)	Day & Date	:	FRI	DAY	,03/11/2	2023
Class	5	:	First Year B. Pharmacy	Semester	:	Ι	Writ	te Your Seat	No. Here
Time	:	:	10:30 am – 11:30 am	Max. Marks	:	30			
Instruc	tions	1. 3.	All questions are compulsory2.Right hand side number indicates full marks4.1	Draw a well label Do not write/tick o	ed d on th	liagran .e ques	ı whe tion p	rever neces aper	sary
QI	Lor	ng A	nswers Question (Answer any 1 out of 2)						10 M
1.	Illu	stra	ate the different perspectives in Communication in	detail along wit	h ex	ample	s. (L	L3, CO1)	
2.	Ex	pla	in Verbal and Non-verbal communication in details	s. (LL3, CO1)					
QII	Sho	ort A	Answers Question						
	Sec	tior	A – A (Answer any ONE)						05 M
	Exp	olai	n the process of communication along with its varie	ous elements. G	ive	an exa	mple	e to illustra	ate.
1.	(LL	.2,	CO1)						
2.	. Suggest remedies to the barriers to communication								
QIII	I Section – A (Answer any Three) 15						15 M		
1.	Explain the barriers of communication. (LL2, CO2)								
2.	2. Compare any two communication styles. (LL2, CO2)								
3.	Explain Gestalt's theory of Visual perception (LL2, CO2)								
4.	Illu	stra	ate the concept of Face- to- Face communication in	details. (LL2, C	CO2)			



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			First Sessional Theory Examinatio	n-(ODD SEI	M) :	2023	3-20)24	
Subje	ect	:	Communication skills (BP105T)	Day & Date	:	FRI	DAY	,03/11/2	2023
Class	5	:	First Year B. Pharmacy	Semester	:	Ι	Writ	te Your Seat	No. Here
Time	:	:	10:30 am – 11:30 am	Max. Marks	:	30			
Instruc	tions	1. 3.	All questions are compulsory2.Right hand side number indicates full marks4.1	Draw a well label Do not write/tick o	ed d on th	liagran .e ques	ı whe tion p	rever neces aper	sary
QI	Lor	ng A	nswers Question (Answer any 1 out of 2)						10 M
1.	Illu	stra	ate the different perspectives in Communication in	detail along wit	h ex	ample	s. (L	L3, CO1)	
2.	Ex	pla	in Verbal and Non-verbal communication in details	s. (LL3, CO1)					
QII	Sho	ort A	Answers Question						
	Sec	tior	A – A (Answer any ONE)						05 M
	Exp	olai	n the process of communication along with its varie	ous elements. G	ive	an exa	mple	e to illustra	ate.
1.	(LL	.2,	CO1)						
2.	. Suggest remedies to the barriers to communication								
QIII	I Section – A (Answer any Three) 15						15 M		
1.	Explain the barriers of communication. (LL2, CO2)								
2.	2. Compare any two communication styles. (LL2, CO2)								
3.	Explain Gestalt's theory of Visual perception (LL2, CO2)								
4.	Illu	stra	ate the concept of Face- to- Face communication in	details. (LL2, C	CO2)			



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First Sessional Theory Examination 2019-2020 Subject: Comm. Skills (BP105T)

Day & Date: Monday, 30 Sept. 2019

Class: FY/ DSA

Semester: First

Time: 10.30 am to 12.00 pm

Max. Marks: 30

10

20

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks

QI Long Answers (Answer any 1 out of 2)

(1) State the importance and objectives of Communication.

(2) Explain the process of Communication with its elements in detail.

QII Short Answers (Answer any 4 out of 6)

- 1. What is Communication? Define it in detail.
- 2. Explain any two barriers that affect the process of

Communication.,

- 3. What are the types and methods of Communication?
- 4. State the importance of Body language with its constituents in detail.
- 5. What is paralanguage? State its importance in the efficacy of Communication.
- 6. Write a note on Graphical Language.

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First Sessional Theory Examination 2018-2019

Subject: Human Anatomy and Physiology-I Class: F. Y. B. Pharma Time: 10:30-12:00	Day & Date: Monday, 1.10.2018 Semester: I Max. Marks: 30
QI Multiple Choice Questions (MCQs)	10
1. Examples of slightly moveable joints is joints	а. А.
A. between the vertebrae B. between the skull bo	nes
C. between ear ossicles D. between invertebral di	scs
2. Bands of connective tissues that are made of coll	agen are called
A. osteocytes B. chondrocytes C. ligaments and	tendons D. cartilages
3. Joints that move like a hinge on door or back and	1 forth are called
A. slightly moveable joints B. hinge joints C. ba	ll-and-socket joints D. fixed joint
4. Example of ball-and-socket joint is	
A. vertebral disc joint B. shoulder and hip joints	
C. knee and elbow joints D. spinal cord joint	
5. The power house of cell is called	
A) Cell wall B) Mitochondria C) Ribosomes D) Nucl	eus
6. The functional unit of life is called	
A) Cell B) Egg C) Nucleus D) None of these	
7. Tissue is a	
A) Group of organs B) Group of cells C) Group	of tissues D) Group of organisms
8 The control unit of cell is	
A) Nucleus B) Cell wall C) Cytoplasm D) All of t	these
9. The part of the bone where blood cells are made.	
A. Marrow B. Joint C. Ligament D. none of t	he above
10. The ribs, sternum and spine protect these.	
A. Kidneys, bladder, urethra B. Heart, lungs, blo	od vessel
C. Small intestine, large intestine D. all the above	e · · ·
QII Long Answers (Answer any 1 out of 2)	10
1. Draw a neat diagram of skin. Explain in detail ar	atomy and functions of skin
2. Define transport. Explain in details various mech	anism of transport
QIII Short Answers (Answer any 2 out of 3)	10
1. Write a note of cell with a neat labelled diagram	· *
2. Define homeostasis explain in detail mechanism	of homeostasis with suitable illustration.
3. Define joints. Explain in detail synovial joint.	
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First Sessional Theory Examination 2019-2020

Subject: Human Anatomy and Physiology – I (BP101T)	Day & Date: Monday, 26.9.19
Class: F.Y. B. Pharmacy	Semester: I
Time: 10.30 am - 12.00 pm	Max. Marks : 30

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary

3. Right hand side number indicates full marks

QI Multiple Choice Questions (MCQs) 10 1. The spherical structured organelle that contains the genetic material is A Cell walls **B** Ribosomes C Nucleus D Mitochondria 2. Which one of the following organelles digests the old organelles that are no longer useful to the cells? Ribosomes **B** Mitochondria D Chromatin A C Lysosomes 3. The Tissue that lines and covers the body is A. Epithelial C. Nervous D. Muscle B. Connective 4. Blood is classified as a(n) tissue. A. Nervous B. Epithelial C. Muscle D. Connective 5. Contractility is a property of tissue. A. Muscle B. Nervous C. Epithelial D. Connective 6. The tissue that is thin enough to allow diffusion is A. Simple cuboidal epithelium B. Simple squamous epithelium C. Areolar connective tissue D. Hyaline cartilage 7. The intercellular material surrounding connective tissue cells is A. Ground substance B. Matrix C. Adipose D. Bone 8. The cell responsible to the formation of cartilage is the A. Fibroblast B. Chondroblast C. Osteoblast D. Megablast 9. Which cells produce the pigment that contributes to hair color? A) Keratinocytes B) Melanocytes C) Langerhans cells D) Merkel cells 10. Example of flat bone A) Sternum B) Clavicle C) A and B D) None of the above QII Long Answers (Answer any 1 out of 2) 10 1. Explain in details anatomy and functions of skin with a neat labelled diagram. 2. Define homeostasis. Write in details components and mechanism of homeostasis with suitable examples

QIII Short Answers (Answer any 2 out of 3)

1. Draw a neat labelled diagram of cell.

2. Define transport, explain any two mechanism of transport across cell membrane.

3. Define skeleton. Write the various bones of axial and appendicular skeleton

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First Sessional Theory Examination 2020-2021 (Odd SEM)

Subject & Subject Code: Human Anatomy & Physiology I (BP101T) Day & Date: Monday, 07.02.2022 Class: First Year B. Pharm Time: 10.30 AM to 12.30 PM Semester: I Max. Marks: 30 Instructions: 1. All questions are compulsory. 2. Draw a well labelled diagram wherever necessary. 3. Right hand side number indicates full marks. Q.I Multiple-choice questions (MCQs) 10 1. Examples of positive feedback mechanism EXCEPT a. Blood clotting b. Childbirth c. Nerve impulse generation d.CO₂ regulation 2. A sperm cell tail is an example of ---Flagellac. Cilia b. Microvili d. Filaments a. 3. Cells communicate over relatively short distances is known as Autocrine signalling b. Paracrine signalling a. Synaptic signalling Endocrine signalling d. c. Spot weld like junctions of epidermis is an example of 4. Tight junctions b. Adherence junction a. Desmosomes Hemidesmosomes d. c. 5. Following is an example of branched tubular glandular epithelium tissue Large intestine b. Sebaceous gland a. Sweat gland Gastric gland d. c. 6. Which of the following epidermal layer is not found in thin skin Stratum corneum Stratum lucidum b. a. Stratum spinosum d. Stratum granulosum c. Bone tissue stores about --- % of the body's ---7. a. 85 % & phosphorous b. 85% & calcium c. 99% & calcium d. 99% & phosphorous 8. With increasing age, much of the bone marrow changes from ---a. Red to yellow b. Yellow to red b. Red to blue d.Yellow to blue 9. The human body contains --- pectoral or shoulder girdles One Two d. Four b. c. Three a. 10. Following is an example of non-nucleated cell **RBCs** b. WBCs d. Lymphocytes c. Monocytes a.

QII Long Answers (Answer any 1 out of 2)

- **1.** Draw a neat and well labelled diagram of skin. Explain the structure and functions of skin.
- 2. Explain the functions of skeletal system. Explain the different types of bones.

QIII Short Answers (Answer any 2 out of 3)

- **1.** Define and classify tissue with suitable examples. Describe the location, structure and functions of squamous epithelial tissue.
- 2. Draw a neat well labelled diagram of cell. Explain the structure and function of mitochondria.
- 3. Explain the levels of structural organization of human body.

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First Sessional Theory Examination 2020-2021 (Odd SEM)

Subject & Subject Code: Human Anatomy & Physiology I (BP101T) Day & Date: Wednesday, 10.03.2021 Class: First Year B. Pharm Max. Marks: 30 Semester: I Time: 10.30 AM to 12.30PM Instructions: 1. All questions are compulsory. 2. Draw a well labelled diagram wherever necessary. 3. Right hand side number indicates full marks. Q.I Multiple-choice questions (MCQs) 10 1. Integumentary system contains skin and associated structures EXCEPT a. Hair b. Sweat glands c. Bones d. Oil glands 2. Following is an example of negative feedback mechanism a. Regulation of carbon dioxide conc. b. Blood Clotting d. Generation of Nerve Signals c. Childbirth 3. Epithelial tissues that line the stomach, intestines, and urinary bladder contains b. Tight Junctions a. Adherence Junctions c. Gap Junctions d. Desmosomes 4. Somatic cells contain two sets of chromosomes, they are symbolized as c. 4n d. 3n b. 2n a. n 5. Following statements are correct for paracrine signaling a. Communicate through release of chemical signals over short distances b. Communicate through release of ligand that binds to its own receptors c. Communicate through nerve cells over short distances d. Transmits signals over long distances through release of hormones 6. Entire outside surface of the cell often has a loose carbohydrate coat called as c. Glycolipids d. Cholesterol a. Phospholipids b. Glycocalyx

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7. Axial skeleton have	number of bone	es			
a. 126	b. 80	c. 100	d. 110		
8. Only bone cells to undergo cell division					
a. Osteoclasts	b. Osteocytes	c. Osteoblasts	d. Osteogenic Cells		
9. Is the longest and la	rgest bone of the up	per limb			
a. Femur	b. Radius & Ulna	c. Tibia & Fibu	la d. Humerus		
10. Is the largest organ	of the body in both	surface area and v	weight		
a. Skin	b. Liver	c. Lungs	d. Heart		

Q.II Long Answers (Answer any 1 out of 2)

1. Define and classify tissue. Explain the structure, location and functions of any three types of epithelial tissue.

10

10

2. Explain the functions of skeletal system. Write in detail about bones of axial skeleton.

Q.III Short Answers (Answers any 2 out of 3)

1. Draw a neat well labelled diagram of skin. Explain the functions of skin.

2. Draw a neat well labelled diagram of cell. Explain the structure and functions of mitochondria.

3. Classify and explain the different types of bones.

mr.sherikar A-K. Csubject In-charge



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First Sessional Theory Examination 2020-2021 (Odd SEM)

Subject & Subject Code: Human Anatomy & Physiology I (BP101T) Day & Date: Wednesday, 10.03.2021 Class: First Year B. Pharm Max. Marks: 30 Semester: I Time: 10.30 AM to 12.30PM Instructions: 1. All questions are compulsory. 2. Draw a well labelled diagram wherever necessary. 3. Right hand side number indicates full marks. Q.I Multiple-choice questions (MCQs) 10 1. Integumentary system contains skin and associated structures EXCEPT a. Hair b. Sweat glands c. Bones d. Oil glands 2. Following is an example of negative feedback mechanism a. Regulation of carbon dioxide conc. b. Blood Clotting d. Generation of Nerve Signals c. Childbirth 3. Epithelial tissues that line the stomach, intestines, and urinary bladder contains b. Tight Junctions a. Adherence Junctions c. Gap Junctions d. Desmosomes 4. Somatic cells contain two sets of chromosomes, they are symbolized as c. 4n d. 3n b. 2n a. n 5. Following statements are correct for paracrine signaling a. Communicate through release of chemical signals over short distances b. Communicate through release of ligand that binds to its own receptors c. Communicate through nerve cells over short distances d. Transmits signals over long distances through release of hormones 6. Entire outside surface of the cell often has a loose carbohydrate coat called as c. Glycolipids d. Cholesterol a. Phospholipids b. Glycocalyx

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7. Axial skeleton have	number of bone	es			
a. 126	b. 80	c. 100	d. 110		
8. Only bone cells to undergo cell division					
a. Osteoclasts	b. Osteocytes	c. Osteoblasts	d. Osteogenic Cells		
9. Is the longest and la	rgest bone of the up	per limb			
a. Femur	b. Radius & Ulna	c. Tibia & Fibu	la d. Humerus		
10. Is the largest organ	of the body in both	surface area and v	weight		
a. Skin	b. Liver	c. Lungs	d. Heart		

Q.II Long Answers (Answer any 1 out of 2)

1. Define and classify tissue. Explain the structure, location and functions of any three types of epithelial tissue.

10

10

2. Explain the functions of skeletal system. Write in detail about bones of axial skeleton.

Q.III Short Answers (Answers any 2 out of 3)

1. Draw a neat well labelled diagram of skin. Explain the functions of skin.

2. Draw a neat well labelled diagram of cell. Explain the structure and functions of mitochondria.

3. Classify and explain the different types of bones.

mr.sherikar A-K. Csubject In-charge



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			First Sessional Theory Examination	n	- (Odd SEI	VI)	202	3-2024
Sub	ject	:	Human Anatomy & Physiology - I (BP1017	Г)	Day & Date	:	Mor	nday, 30/10/2023
Clas Tim	ss Ie	:	First Year B. Pharmacy 10.30 am to 12.00 pm		Semester Max. Marks	:	I 30	Write Your Seat No. Here
Instri	uctions	s: 1 3	. All questions are compulsory2.2. Right-hand side number indicates full marks4	. D1 . D0	raw a well-labele o not write/tick o	ed å on tl	liagrar ne que	n wherever necessary stion paper
Q. I:	Objec	tive	e Type Questions (5 X 2)					10 M
1.	Enlist	the	functions of muscular tissue (CO1, LL2)					
2.	Write	the	functions of Golgi Complex (CO1, LL2)					
3.	Define	e the	e term					
			a. Passive diffusion (CO1, LL2)b. Anemia (CO2, LL2)					
4.	Enlist	the	functions of blood (CO2, LL2)					
5.	Draw	a ne	eat and well-labelled diagram of skin (CO2. LL2)					
Q. II	: Long	Aı	nswers Questions (Answer any 1 out of 2)					10 M
1.	Draw memb	a n ran	eat and well-labelled diagram of plasma membrane. Expected (CO1, LL2)	pla	in in detail abou	t st	ructure	e and functions of plasma
2.	Define	e an	d classify tissue. Explain the structure and functions of ep	oith	elial tissue (CO1	, <i>LL</i>	.2)	
Q. II	I: Sho	rt A	Answers Questions (Answer any 2 out of 3)					10 M
1.	Enlist	the	bones of appendicular system. Write in detail about struc	ture	e and functions o	f pe	lvic gi	rdle. (CO2, LL2)
2.	Define	e an	d classify joints. Explain the structure and functions of sy	nov	vial joints. (CO2,	LL	2)	
3.	Explai	n tł	ne ABO system of blood group. (CO2, LL2)					

Prepared by Mr. Abdulla Sherikar **Verified by** Mr. Purvaj Barote **Approved by** Dr. Kartik Nakhate



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			First Sessional Theory Examination	n	- (Odd SEI	N)	202	3-2024
Sub	ject	:	Human Anatomy & Physiology - I (BP1017	Г)	Day & Date	:	Mor	nday, 30/10/2023
Clas Tim	ss Ie	:	First Year B. Pharmacy 10.30 am to 12.00 pm		Semester Max. Marks	:	I 30	Write Your Seat No. Here
Instri	uctions	s: 1 3	. All questions are compulsory2.2. Right-hand side number indicates full marks4	. D1 . D0	raw a well-labele o not write/tick o	ed å on tl	liagrar ne que	n wherever necessary stion paper
Q. I:	Objec	tive	e Type Questions (5 X 2)					10 M
1.	Enlist	the	functions of muscular tissue (CO1, LL2)					
2.	Write	the	functions of Golgi Complex (CO1, LL2)					
3.	Define	e the	e term					
			a. Passive diffusion (CO1, LL2)b. Anemia (CO2, LL2)					
4.	Enlist	the	functions of blood (CO2, LL2)					
5.	Draw	a ne	eat and well-labelled diagram of skin (CO2. LL2)					
Q. II	: Long	Aı	nswers Questions (Answer any 1 out of 2)					10 M
1.	Draw memb	a n ran	eat and well-labelled diagram of plasma membrane. Expected (CO1, LL2)	pla	in in detail abou	t st	ructure	e and functions of plasma
2.	Define	e an	d classify tissue. Explain the structure and functions of ep	oith	elial tissue (CO1	, <i>LL</i>	.2)	
Q. II	I: Sho	rt A	Answers Questions (Answer any 2 out of 3)					10 M
1.	Enlist	the	bones of appendicular system. Write in detail about struc	ture	e and functions o	f pe	lvic gi	rdle. (CO2, LL2)
2.	Define	e an	d classify joints. Explain the structure and functions of sy	nov	vial joints. (CO2,	LL	2)	
3.	Explai	n tł	ne ABO system of blood group. (CO2, LL2)					

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Institute of Pharmacy, Dhule Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

First Sessional Theory Examination 2018-2019					
Subje	ect : Pharmaceutical Analysis	Day & Date : 3/10/	2018		
Class	F. Y. B. Pharmacy	Semester : I			
Time	: 10:30 to 12:00 pm	Max. Marks : 30			
	Instructions: 1. All questions are compuls	sory			
	2. Draw a well labeled diagra	im wherever necessary			
	3. Right hand side number ind	licates full marks	[10]		
Q.1	Multiple Choice Questions		[10]		
a	According to Lewis, bases are				
	a. Electron pair donor	b. Electron pair acceptor			
Ъ	c. Proton donor f_{int}	d. Proton acceptor			
D	$pH = pKIII \pm I$ is called as of II.	h Constant interval			
	a. Transmorent interval	d None of the above			
0	If $K_0 > K_0$ then the given salt produces	solution			
C	a Neutral	b Acidic			
	c Basic	d, b & c both			
Ь	Molarity [M] lis				
u	a. 1 mole / lit of solution	b. 1 g / lit. of solution			
	c. M.W. of solute in g / lit. of solution	d. 1 g / lit. of solution			
e	Argentimetric titrations mainly utilizes	as a titrant.			
	a. 0.02M Disodium EDTA	b. 0.1 M sodium hydroxide			
	c. 0.1 N Silver chloride	d. 0.1 N perchloric acid			
f	Precision is expressed by				
	a. Correctness of measurement	b. Standard value			
к 2	c. Reproducibility of measurement	d. True value			
g	is a compound of sufficient pu	urity from which a standard solutio	n can		
	be prepared by direct weighing of a quanti	ity of it, followed by dilution to give	а		
	defined volume of solution.				
	a. External standard	b. Primary standard			
-	c. Secondary standard	d. a & c both			
h	are the methods of minimisa	ation of error			
	a. Amplification methods	b. Standard addition			
:	C. Internal standards	u. All of the above			
1	Dioxane is a solvent	h Protogenic '			
	a. Amphiprotic	d Aprotic			
i	Crystal violet is indicator	u. Aprolie			
J	a. Acidic	b. Basic			
	c. Neutral	d. Acidic & Basic both			
Q.2.	Answer the following short questions (any	two)	[10]		
a	What are primary and secondary standars	s? Explain theirs types with suitabl	e		
	example				
b	Write short note on Mohr's method				
с	What are indicators? Explain it with respe	ect to their various theories.			
d	Define non aqueous titrations? Explain th	e types of solvent used.			
			_ 6 X212		
Q.3.	Answer in detail of following (any one)		[10]		
a	Explain the various titration curves with s	suitable example			
b	Explain types of error with suitable examp	ple. Describe the various methods			
	used for minimisation of errors				



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Institute of Pharmacy, Dhule Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

	First Sessional Theory Example	mination 2019-2020	
Subje	ct : Pharmaceutical Analysis	Day & Date : 01/10/20)19
Class	: FY/DSY B. Pharmacy	Semester : I/III*	
Time	: 10:30 - 12:00 noon	Max. Marks : 30	
	Instructions: 1. All questions are compulsor	y Im uiberever necessary	•
	3. Right hand side number indi	icates full marks	
0.1	Multiple Choice Questions		[10]
1.	According to Lewis bases are	2	
	a. Electron pair donor	b. Electron pair acceptor	
	c. Proton donor	d. Proton acceptor	
2.	Amphiprotic Solvents are both	d Characters	
	a. Aprotic. Protophillic	b. Protophillic. Protogenic	
	c Protogenic Aprotic	d None of the above	
3.	is most suitable method to analyse	d Weak acid and Weak base	
0.	a Aqueous titration	b Non-aqueous titration	
	c Precipitation titration	d Complexometric titration	
4	Molarity [M] is	a. complexometric attation	
	a 1 mole / lit of solution	b $1 \sigma / lit$ of solution	
	$C = M W$ of solute in σ / lit. of solution	d l g / lit of solution	
5	Argentometric titrations mainly utilizes	as a titrant	
0.	a 0.02M Disodium EDTA	b 0 1 M sodium hydroxide	
	c 0.1 N Silver chloride	d 0.1 N perchloric acid	
6	Precision is expressed by		
0.	a Correctness of	b Standard value	
	measurement	5. Standard Value	
	c Reproducibility of measurement	d True value	
7	is a compound of sufficient	t purity from which a standard	•
	solution can be prepared by direct weight	which of a quantity of it followed h	v
	dilution to give a defined volume of solu	ution	
	a External standard	b Primary standard	
	c Secondary standard	d a & c both	
8.	are the methods of minin	nisation of error •	
0.	a. Amplification methods	b. Standard addition	
	c Internal standards	d All of the above	
9	Dioxane is a solvent		
21	a. Protophillic	b. Protogenic	
	c Amphiprotic	d Aprotic	
10	Crystal violet is indicator	a. riprodo	
10.	a. Acidic	b. Basic	
	c Neutral	d Acidic & Basic both	
	o. noutiu		
0.2.	Answer the following short questions	s (any two)	[10]
2.1	What are primary and secondary stand	lards? Explain theirs types	
u	with suitable example	arus. Explain meno types	
h	Write short note on Mohr's method	5	
0	Define non aqueous titrations? Evolain	the types of solvent used	
C	Denne non aqueous unations: Explan	i the types of solvent used.	
0.3	Answer in detail of following (any on		[10]
2.0.	Explain the various titration curves wit	th suitable example	[10]
	Emplain the various thatfor curves with	male Describe the mariane	

Explain types of error with suitable example. Describe the various methods used for minimisation of errors



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First Sessional Theory Examination	2020-2021
Subject: Pharmaceutical Analysis I (BP102T)	Day & Date: 11-03-2021
Class: First Year B. Pharm.	Semester: I
Time: 10:30 am – 12:00 pm	Max. Marks : 30
Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks	
Q. I. Solve the following Multiple Choice Questions (MCQs)	10
1. Number of gram equivalents of solute present in one litre of a solution	is called as
a) Molarity	
b) Normality	
c) Molality	
d) Percentage weight by volume	
2. The features of Primary standard substance includes:	
a) High purity	
b) High stability	
c) Low hygroscopicity	
d) All of the above	
3. Indian pharmacopoeia is published by	
a. World health organization	
b. Ministry of health and family welfare	
c. Ministry of science and technology	
d. Food and drug administration	
4. Limit test of iron is based upon the reactions between iron and	
a. Dithiazone	
b. Thioglycolic acid	
c. Hydrogen sulphide	
d. Barium chloride	
5. The color of phenolphthalein in alkanine medium is	

a) Pink

 \bigcirc

- b) Yellow
- c) Brown
- d) Colorless

- 6. Benzene is an example ofsolvent.
 - a) Aprotic
 - b) Protogenic
 - c) Protophilic
 - d) Amphiprotic
- 7. Solubility of precipitate is on addition of common ion.
 - a) Decreased
 - b) Increased
 - c) Remains Constant
 - d) None of These
- 8. In Mohr's methodis used as an indicator.
 - a) Potassium chromate
 - b) Ferric ammonium sulphate
 - c) Ammonium thiocyanate
 - d) Silver nitrate
- 9. In Volhard's methodis used as a titrant
 - a) Potassium chromate
 - b) Ferric ammonium sulphate
 - c) Ammonium thiocyanate
 - d) Silver nitrate

10. The end-point detection in Fajan's method involves color change due to.....

- a) Formation of colored precipitate
- b) Formation of soluble colored complex
- c) Adsorption of indicator on the surface of precipitate
- d) None of the above

Q. II. Solve the following (any 1)

- 1. Explain the neutralization curves in acid-base titrations for i) strong acid vs strong base ii) weak acid vs strong base iii) strong acid vs weak base iv) weak acid vs weak base.
- 2. Explain in detail Mohr's method and Fajan's method.

Q. III. Solve the following (any 2)

- 1. Explain principle involved in limit test for Chloride and sulphate.
- 2. Explain in detail the theories of acid-base indicators.
- 3. Explain the types of solvents used in non-aqueous titrations with suitable examples.

10



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First Sessional Theory Examination 2021-2022 (Odd SEM)

Subject: Pharmaceutical Analysis-I (BP102-T)	Day & Date: Tuesday 08/02/2022
Class: F.Y. B. Pharm & D.S.Y.B.Pharm	Semester: I & III
Time: 10.30 am- 12.00 pm	Max. Marks: 30

Instructions: 1. All questions are compulsory 2. Draw a well labelled diagram wherever necessary 3. Right hand side number indicates full marks **QI** Multiple Choice Questions (MCQs) 10 1. Potassium chromate is used as indicator in -----a) Fajan's method b) Gay Lussacs method c) Mohrs method d) Volhards method 2. ----- error cannot be minimized or identified a) Random error b) Determinate error d) All of the above c) Indeterminate error 3. reagent is used for limit test for chloride a) BaCl2 b) AgNO3 c) Methyl Orange d) EDTA 4. Measurement which is close to true value is known as..... a) Accuracy b) Precision c) Error d) Average 5. A measurement which on repetition gives same or nearly same results are known as..... a) Accuracy b) Precision c) Error d) Average 6. The acid which dissociate completely in aqueous media is known as a) Weak acid b) Strong acid c) Strong base d) Weak base 7. The number of gm Equivalents to solute in 1 litre of solvent is known as a) Normality c) Molality b) Molarity d) Formality 8. is an example of protophilic solvents? a) HCl b) Acetic acid c) KOH d) Methanol 9. Crystal Voilet is used as indicator in titration b) non-aqueous c) Both d) None of the above a) Aqueous **10**) A substance which can donate proton in aqueous medium is known as a) Acid b) base c) Neutral d) All of the above

QII Long Answers Question (Answer any 1 out of 2)

1. What do you mean by Errors? Explain the types and sources of error in details?

2. Define precipitation titrations? Explain Mohr's and Volhard's method?

QIII Short Answers Question (Answer any 2 out of 3)

1. Discuss various types of solvents used in non-aqueous titration?

- 2. Write down the preparation and standardization of 0.1 N NaOH as per IP?
- 3. i) Write down the neutralization curve for SA v/s SB?

ii) Enlist the sources of impurities in medicinal agents?

10



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Mohd Usman	Azim Ansari	Girija Bhavar
Prepared By	Verified By	Approved By

Mapping of Course Outcome with Second Sessional Theory Examination (2021-2022) Subject: Pharmaceutical Analysis-I

Course Outcome	Question	Marks	
	Q1 – 2,3,4,5,7		
CO 1021:	Q2 – 1	22.5	
	Q3- 2, 3-ii		
	Q1 – 1,6,8,9,10		
CO 1022:	Q2 -2	22.5	
	Q3 – 1, 3-i		

Mohd Usman	Azim Ansari	Girija Bhavar
Prepared By	Verified By	Approved By



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First Sessional Theory Examination 2021-2022 (Odd SEM)

Subject: Pharmaceutical Analysis-I (BP102-T)	Day & Date: Tuesday 08/02/2022
Class: F.Y. B. Pharm & D.S.Y.B.Pharm	Semester: I & III
Time: 10.30 am- 12.00 pm	Max. Marks : 30

Instructions: 1. All questions are compulsory 2. Draw a well labelled diagram wherever necessary 3. Right hand side number indicates full marks **QI** Multiple Choice Questions (MCQs) 10 1. Potassium chromate is used as indicator in -----a) Fajan's method b) Gay Lussacs method c) Mohrs method d) Volhards method 2. ----- error cannot be minimized or identified a) Random error b) Determinate error d) All of the above c) Indeterminate error 3. reagent is used for limit test for chloride a) BaCl2 b) AgNO3 c) Methyl Orange d) EDTA 4. Measurement which is close to true value is known as..... d) Average a) Accuracy b) Precision c) Error 5. A measurement which on repetition gives same or nearly same results are known as..... a) Accuracy b) Precision c) Error d) Average 6. The acid which dissociate completely in aqueous media is known as b) Strong acid a) Weak acid c) Strong base d) Weak base 7. The number of gm Equivalents to solute in 1 litre of solvent is known as a) Normality b) Molarity c) Molality d) Formality **8.** is an example of protophilic solvents? a) HCl b) Acetic acid c) KOH d) Methanol 9. Crystal Voilet is used as indicator in titration b) non-aqueous c) Both d) None of the above a) Aqueous **10**) A substance which can donate proton in aqueous medium is known as a) Acid b) base c) Neutral d) All of the above

QII Long Answers Question (Answer any 1 out of 2)

1. What do you mean by Errors? Explain the types and sources of error in details?

2. Define precipitation titrations? Explain Mohr's and Volhard's method?

QIII Short Answers Question (Answer any 2 out of 3)

- 1. Discuss various types of solvents used in non-aqueous titration?
- 2. Write down the preparation and standardization of 0.1 N NaOH as per IP?
- 3. i) Write down the neutralization curve for SA v/s SB?
 - ii) Enlist the sources of impurities in medicinal agents?

10



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Mohd Usman	
Prepared By	

Azim Ansari Verified By Girija Bhavar Approved By

Mapping of Course Outcome with Second Sessional Theory Examination (2021-2022)

Course Outcome	Question	Marks	
	Q1 – 2,3,4,5,7		
CO 1021:	Q2 - 1	22.5	
	Q3- 2, 3-ii		
	Q1 – 1,6,8,9,10		
CO 1022:	Q2 -2	22.5	
	Q3 – 1, 3-i		

Subject: Pharmaceutical Analysis-I

Mohd Usman	Azim Ansari	Girija Bhavar
Prepared By	Verified By	Approved By



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FIRST	SI	ESSIONAL THEORY EXAMIN	ATION-(OD)	D	SEN	I) 2022-2023
Subject	:	Pharmaceutical Analysis	Day &	:	Sat	,11/03/2023
		(BP102T)	Date			. , ,
Class	:	First Year B. Pharmacy	Semester	:	Ι	Write Your Seat No.
Time	:	10:30 am – 12:00 pm	Max.	:	30	here
		1	Marks			

Instructions:

- 1. All questions are compulsory
- 2. Draw a well-labeled diagram wherever necessary
- 3. Right-hand side number indicates full marks

4. Do not write/tick on the question paper

OI **Objective Type Questions**

- Define: Accuracy and Precision 1.
- How to prepare 0.1 N Sodium hydroxide solution 2.
- Write a reaction for weak acid and strong base titration 3.
- Enlist type of solvents in Non-aqueous titration with one example each. 4.
- 5. Define acidimetry and limit test

QII Long Answers Question (Answer any 1 out of 2) 1. Define Pharmaceutical analysis and describe its scope. Enlist techniques of analysis and

- describe any three of them.
- 2. What are the sources of errors and describe methods of minimizing errors.

OIII Short Answers Question (Answer any 2 out of 3)

- 1. Enlist the theories of acid-base indicators and explain any one.
- 2. Describe Volhard's methods for precipitation titration.
- 3. Explain masking & de-masking reagents. Give examples of metal ion indicators

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FIRST	SI	ESSIONAL THEORY EXAMIN	ATION-(OD)	D	SEN	I) 2022-2023
Subject	:	Pharmaceutical Analysis	Day &	:	Sat	,11/03/2023
		(BP102T)	Date			. , ,
Class	:	First Year B. Pharmacy	Semester	:	Ι	Write Your Seat No.
Time	:	10:30 am – 12:00 pm	Max.	:	30	here
		1	Marks			

Instructions:

- 1. All questions are compulsory
- 2. Draw a well-labeled diagram wherever necessary
- 3. Right-hand side number indicates full marks

4. Do not write/tick on the question paper

OI **Objective Type Questions**

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- Write a reaction for weak acid and strong base titration 3.
- Enlist type of solvents in Non-aqueous titration with one example each. 4.
- 5. Define acidimetry and limit test

QII Long Answers Question (Answer any 1 out of 2) 1. Define Pharmaceutical analysis and describe its scope. Enlist techniques of analysis and

- describe any three of them.
- 2. What are the sources of errors and describe methods of minimizing errors.

OIII Short Answers Question (Answer any 2 out of 3)

- 1. Enlist the theories of acid-base indicators and explain any one.
- 2. Describe Volhard's methods for precipitation titration.
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First Sessional Theory Examination-(ODD SEM) 2023-2024							
Subject	: Pharmaceutical Analysis (BP102T)	Day & Date	:	Tue	,31/10/20)23	
Class	: First Year B. Pharmacy	Semester	:	Ι	Write Your S	Seat No.	Here
Time	: 10:30 am – 12:00 pm	Max. Marks	:	30			
Instruction	1. All questions are compulsory23. Right hand side number indicates full marks4	. Draw a well labeled Do not write/tick on	d dia 1 the	agram ques	t wherever ne tion paper	cessarį	y
Q. 1	Objective Type Questions					10 N	M
1.	Define: i) Accuracy; ii) Primary Standard (CO1, LL)	1)					
2.	What is Impurity? Enlist their types (CO1, LL3)						
3.	Give any two difference between Acidimetric and A	kalimetry type of	titra	tion	(CO2, LL1)		
4.	Enlist any four indicators used in non-aqueous titrati	on. (CO2, LL1)					
5.	Write the principle and reaction for limit test for chlo	oride (CO1, LL1)					
Q 2	Long Answers Question (Any 1 x 10 Marks)					10	Μ
1	Elaborate on Acids and Bases with respect to differe	nt theories. Discus	s dif	fferer	nt neutraliza	tion cu	irves
1.	with suitable examples. (CO2, LL4)						
2	What are errors? Discuss their sources and method	ods of minimizing	erra	ors w	ith suitable	exam	ples.
2.	Comment on relationship between accuracy and pre-	ecision. (CO1, LL3	3)				
Q 3	Short Answer Question (Any 2 x 05= 10 M)					10	Μ
1.	Write a note on Acid-Base indicator theories. (CO2	, LL1)				1	
2.	Describe Volhard's methods in precipitation titration. (CO2, LL2)						
3.	Discuss different Pharmacopoeia and its types (CO1, LL2)						



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First Sessional Theory Examination-(ODD SEM) 2023-2024				
Subject	: Pharmaceutical Analysis (BP102T)	Day & Date : Tue,31/10/20)23	
Class	: First Year B. Pharmacy	Semester : I Write Your S	Seat No. Here	
Time	: 10:30 am – 12:00 pm	Max. Marks : 30		
Instruction	1. All questions are compulsory23. Right hand side number indicates full marks4.	. Draw a well labeled diagram wherever ne Do not write/tick on the question paper	cessary	
Q. 1	Objective Type Questions		10 M	
1.	Define: i) Accuracy; ii) Primary Standard (CO1, LL)	1)		
2.	What is Impurity? Enlist their types (CO1, LL3)			
3.	Give any two difference between Acidimetric and A	lkalimetry type of titration (CO2, LL1)		
4.	Enlist any four indicators used in non-aqueous titrati	on. (CO2, LL1)		
5.	Write the principle and reaction for limit test for chlo	oride (CO1, LL1)		
Q 2	Long Answers Question (Any 1 x 10 Marks)		10 M	
1	Elaborate on Acids and Bases with respect to differe	nt theories. Discuss different neutraliza	tion curves	
	with suitable examples. (CO2, LL4)			
2. What are <i>errors</i> ? Discuss their <i>sources</i> and <i>methods of minimizing errors</i> with suitable examples.				
	Comment on relationship between accuracy and pro-	ecision. (CO1, LL3)		
Q 3	Short Answer Question (Any 2 x 05= 10 M)		10 M	
1.	Write a note on Acid-Base indicator theories. (CO2	, LL1)	·	
2.	Describe Volhard's methods in precipitation titratio	n. (CO2, LL2)		
3.	Discuss different Pharmacopoeia and its types (CO	I, LL2)		



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First Sessional Theory Examination 2020-2021 Subject: Pharmaceutics I BP103T Day & Date: Friday, 12th March 2021

Class: F Y B. Pharm	Semester: 1
Time: 90 min	Max. Marks: 30

Instructions: 1. All questions are compulsory

Draw a well labeled diagram wherever necessary
 Right hand side number indicates full marks

Q.No.1 Multiple choice questions

(10 X 1=10)

1. is known as 'Father of Pharmacy education in India'.

a) M. L. Shroff, b) R. N. Chopra, c) B. N. Ghosh, d) None of the above

2. The fourth edition of IP was published in

a) 1985 b) 1996 c) 1960 d) 19963.

3. Pharmacy Council of India (PCI) was established in

a) 1947 b) 1948 c) 1949 d) 1950.

4. Which of the following excipient act as preservative.....

a) Sorbitan monolactrate b) Tyloxopol c) Benzethenium chloride d) Thiomersal

5. When two or more drugs are used in combination to increase the pharmacological action, the phenomenon is known as

a) Synergism b) Tolerance c) Potentiation c) Idiosyncrasy

6. Nasal drops should have pH

- a) 5-6 b) 6-7 c) 7-8 d) 8-9
- 7. Which of the following is natural colorant
- a) Amaranth b) Erythrosine c) Carotenoids d) Eosin

8. The dosage form intended to be rubbed on the skin for rubeficient and counterirritant effect are called

a) Lotions b) Liniments c) Gels d) Dentifrices

9. 13 (apothecaries) =

a) 437.5 grains b) 480 grains c) 7000 grains d) 5760 grains

10. Daily adult dose of drug is 120 mg. How much of it should be given to a child of 11 years? a) 57.39, b) 60, c) 50.39, d) 37

Q. No. 2 Long Answers (Answer 1 out of 2)

$(1 \times 10 = 10)$

,

1. Define dosage form. Explain need of dosage form. Describe various solid dosage forms.

2. Define posology. Write a note on factors affecting posology

Q. No. 3 Short Answers (Answer 2 out of 3)

$(2 \times 5 = 10)$

. .

1. Write down parts of prescription

2. Discuss various excipient used in the formulation of liquid dosage forms.

3. Write short note on (any one)

- a) Syrups
- b) Topical oral preparations

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First Sessional Theory Examination 2019-2020

Subject: Pharmaceutics I BP103TDay & Date: Friday, 27th Sept 2019Class: F Y B. PharmSemestor: ITime: 10.30 pm to 12.00 pmMax. Marks: 30

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary

3. Right hand side number indicates full marks

(10 X 1=10)Q.No.1 Multiple choice questions 1. The fifth edition of Indian Pharmacopoeia was published in (a)2007 c) 1985 d) 1966 b) 1996 2. Which part of prescription contains instructions for the patients? c) Inscription, ut) Signa a) Superscription. b) Subscription. 3. Which of the following is antifoaming agent? a) Simethicone, b) Polysorbate, c) Methylcellulose, d) saccharin 4. is commonly used viscous base in throat paint. c) Alcohol, d) Hydro-alcoholic mixture. b) Water, a) Glycerin, 5. A drug that neutralizes excess gastric acid are c) Antiamebic, d) Antiepileptic a) Antacid, b) Analgesic, 6. Daily adult dose of drug is 120mg. How much of it should be given to a child of 11 vears? d) 37 a) 57.39, b) 60. c) 50.39, 7. 60 gr = \dots a) 1 9. b) 2 9. ~e)3 D. d) 4 9 8. The semisolid unit dosage form intended for vaginal administration are called c) Bougies a) Suppositories b) Pessaries d) Cones 9. How will you mix potent drug with large amount of diluent? b) Geometric dilution - c) Sifting d) Tumbling. a) Spatulation 10. Which of the following dosage form is not recommended for ocular administration d) Solutions a) Lotions, ,b) Liniments. c) Suspensions,

Q. No. 2 Long Answers (Answer 1 out of 2)

1. Define dosage form. Explain need of dosage form. Describe various liquid dosage forms.

2. Define prescription. Write down parts of prescription and methods of handling of prescription.

Q. No. 3 Short Answers (Answer 2 out of 3)

1. Write a note on factors affecting posology.

2. Discuss methods of dispensing of powders involving special problems.

3. Write short note on (any one)

a) Write down silent features of Indian Pharmacopoeia.

b) Differentiate between: 1) Lotions and liniments. 2) Syrups and elixirs.

. _ _ ** _ _

 $(1 \times 10 = 10)$

 $(2 \times 5 = 10)$






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First Sessional Theory Examination-(ODD SEM) 2022-2023

		•	•		•	
Subject	:	Pharmaceutics-I (BP103T)	Day & Date	:	Moi	nday, 13/03/2023
Class	:	First Year B. Pharmacy	Semester	:	Ι	Write Your Seat No.
Time	:	10.30 am to 12.00 noon	Max. Marks	:	30	Here

Instructions:

- 1. All questions are compulsory
- 2. Draw a well-labeled diagram wherever necessary
- 3. Right-hand side number indicates full marks
- 4. Do not write/tick on the question paper

Q. I: Objective Type Questions (5 X 2)

- **1** Define following:
 - a) Avoirdupois system
 - b) Apothecaries system
- 2 Define and Classify the dosage form.
- **3** Define the term Posology. Enlist the various factors affecting the dose of the drug.
- 4 Define solid dosage form. Gives advantages and disadvantages of solid dosage forms.
- 5 What is the dose for a 9-month-old infant if the average adult dose of a drug is 500 mg? (By using Fried's formula)

Q. II: Long Answers Questions (Answer any 1 out of 2)

- **1** Define prescription. Explain the parts of the prescription and methods of handling the prescription.
- 2 Discuss the historical development of the pharmacy profession in India considering education and industry.

Q. III: Short Answers Questions (Answer any 2 out of 3)

- **1** Define and classify powders. Discuss various methods for mixing powders.
- 2 Write a note on various excipients used in the formulation of liquid dosage forms.
- 3 Explain in brief about solubility enhancement techniques.

10 M

10 M

10 M



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First Sessional Theory Examination-(ODD SEM) 2022-2023

		•	•		•	
Subject	:	Pharmaceutics-I (BP103T)	Day & Date	:	Moi	nday, 13/03/2023
Class	:	First Year B. Pharmacy	Semester	:	Ι	Write Your Seat No.
Time	:	10.30 am to 12.00 noon	Max. Marks	:	30	Here

Instructions:

- 1. All questions are compulsory
- 2. Draw a well-labeled diagram wherever necessary
- 3. Right-hand side number indicates full marks
- 4. Do not write/tick on the question paper

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First Sessional Theory Examination-(ODD SEM) 2023-2024 Pharmaceutics-I (BP103T) Subject Day & Date Wednesday, 01/11/2023 : : Write Your Seat No. Here Class First Year B.Pharm Semester : Ι : Max. Marks 30 Time : 10.30 am to 12.00 pm : 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary Instructions *Right hand side number indicates full marks* 4. Do not write/tick on the question paper **Objective Type Questions** $(5 \times 2) = 5 \times 2 = 10$ (Answer all the questions) QI 10 M Give salient features of first edition of Indian pharmacopoeia (LL1, CO1) 1. 2. Give formula for Pediatric dose calculations based on age (LL1, CO1) 3. Define a. Pharmacopoeia b. Mouthwashes (LL1, CO1) Convert the following imperial values to metric system (*LL1*, *CO2*) 4. b. 1 Pound (Avoir) = _ 1 fluid ounce = a. ml gm. 5. What are effervescent granules (LL1, CO2) QII Long Answers Question (Answer any 1 out of 2) 10 M Classify dosage forms and describe all with examples (LL1, CO2) 1. 2. What is Prescription? Explain in detail Parts of Prescription and Errors in Prescription. (LL1, CO2) Short Answers Question (Answer any 2 out of 3) -**Q.III 10M** Definition and classify Powders. Write a short note on Dusting powder. (LL2, CO1) 1. In what proportion should alcohol 95% and 65% strengths be mixed to make 200 ml of 80% alcohol? (LL2, CO1) 2. What are liquid dosage form? Give classification, advantages and disadvantages of liquid dosage form. (LL2, CO1) 3.



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First Sessional Theory Examination-(ODD SEM) 2023-2024 Pharmaceutics-I (BP103T) Subject Day & Date Wednesday, 01/11/2023 : : Write Your Seat No. Here Class First Year B.Pharm Semester : Ι : Max. Marks 30 Time : 10.30 am to 12.00 pm : 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary Instructions *Right hand side number indicates full marks* 4. Do not write/tick on the question paper **Objective Type Questions** $(5 \times 2) = 5 \times 2 = 10$ (Answer all the questions) QI 10 M Give salient features of first edition of Indian pharmacopoeia (LL1, CO1) 1. 2. Give formula for Pediatric dose calculations based on age (LL1, CO1) 3. Define a. Pharmacopoeia b. Mouthwashes (LL1, CO1) Convert the following imperial values to metric system (*LL1*, *CO2*) 4. b. 1 Pound (Avoir) = _ 1 fluid ounce = a. ml gm. 5. What are effervescent granules (LL1, CO2) QII Long Answers Question (Answer any 1 out of 2) 10 M Classify dosage forms and describe all with examples (LL1, CO2) 1. 2. What is Prescription? Explain in detail Parts of Prescription and Errors in Prescription. (LL1, CO2) Short Answers Question (Answer any 2 out of 3) -**Q.III 10M** Definition and classify Powders. Write a short note on Dusting powder. (LL2, CO1) 1. In what proportion should alcohol 95% and 65% strengths be mixed to make 200 ml of 80% alcohol? (LL2, CO1) 2. What are liquid dosage form? Give classification, advantages and disadvantages of liquid dosage form. (LL2, CO1) 3.

First Sessional Theory Exami	nation 2018-2019
Subject: Pharmaceutics I BP103T	Day & Date: Thu, 4 th Oct2018
Class: F Y B. Pharm	Semester: I
Time: 90 min	Max. Marks: 30

Instructions: 1. All questions are compulsory

2. Draw a well labeled diagram wherever necessary

3. Right hand side number indicates full marks

Q.No.1 Multiple choice questions

1. The third edition of Indian Pharmacopoeia was published in

a) 2007 b) 1996 c) 1985 d) 1966

2. The semisolid unit dosage form intended for vaginal administration are called

a) Suppositories b), Pessaries c) Bougies d) Cones

3. Drugs that suppresses nausea and vomiting are called

a) Analgesic b) Anthelmintic c) Antibiotic,d) Antiemetic

4.13 (apothecaries) =

a) 437.5 grains b) 480 grains c) 7000 grains d) 5760 grains

5. The powders on mixing turn to liquid due to depression of melting point are called

a) Hygroscopic powders b) Efflorescent powders c) Eutectic mixtures d) Compound powders

6. The adult dose of drug is 100mg, what will be the dose for a child having body surface area 0.57m²?

a) 33 mg b) 66 mg c) 50 mg d) 100mg.

7. Which of the following co-solvents are used to increase the solubility of a drug?

a) Ehanol, b) Sorbitol, c) Glycerine, d) All the above

8. Which of the following dosage form is not recommended for ocular administration

a) Lotions, b) Liniments, c) Suspensions, d) Solutions

9. Which part of prescription contains name and quantities of the prescribed ingredients?

a) Superscription, b) Subscription, c) Inscription, d) Signa

10. is known as 'Father of Medicine'.

a) Hippocrates, b) Theophrastus, c) Paracelsus, d) Benjamin Franklin

Q. No. 2 Long Answers (Answer 1 out of 2)

1. Define dosage form. Classify dosage form. Describe various solid dosage forms.

2. Define the term Posology. Explain in detail various factors affecting dose of drug.

Q. No. 3 Short Answers (Answer 2 out of 3)

1. Define prescription. Write a note on parts of prescription.

2. a) Explain various methods of powders mixing.

b) Differentiate between lotions and liniments.

3. Write short note on (any one)

a) History of Indian Pharmacopoeia.

b) Excipient used in liquid dosage form.

 $(1 \times 10 = 10)$

$(2 \times 5 = 10)$

(10 X 1=10)

Survey No. 4991, 1101 No.5,	
First Sessional The Subject: P'ceutical Inorganic Chemistry (BP1047 Class: First Year B. Pharm. Time: 10.30 am – 12.00 pm	Beory Examination 2018-2019 T) Day & Date: Fri, 05/10/2018 Semester: I Max. Marks: 30
Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wh	herever necessary
O.I. Multiple Choice Questions (MCOs)	s juii marks
Lewis base is	
a) electron pair donor b) electron pair accepto	or c) proton dopor d) proton acceptor
2. The principal function of chloride is	
a) maintenance of proper hydration	b) maintenance of osmotic pressure
c) normal electrolyte balance	d) all of the above
3 Citric acid is used in limit test for iron to pre	a) an of the above
a) precipitation of iron by ammonia	b) reduction of iron
a) precipitation of iron by thiorhood	id d) complexition with this glycollic acid
4 Indian pharmacanacia is published by	ia a) complexation with thiogiyconic acia
4. Indian pharmacopoela is published by	h) food and down a durinistration
a) Ministry of health and family welfare	b) rood and drug administration
c) world nearth organization	a) ministry of science and technology
5. Lead acetate cotton plug is used in limit test	for arsenic to
a) trap hydrogen sulphide b) trap thiosulphate	c) trap hydrogen sulphate d) trap arsine gas
6. High serum potassium level is called as	
a) hypocalcemia b) hypokalemia c) hype	ercalcemia d) hyperkalemia
7. Choose the correct composition of ORS as pe	er current WHO formula
i) NaCl ii) KCl iii) Sodium citrate iv) sodium	n bicarbonate v) anhydrous glucose
L a) i, ii, iii, v b) i, ii iv, v	C) I, ii, iii, iv d) ii, iii, iv, v
8. calcium gluconate is prepared from	•
a) Lactic acid and CaCO ₃ b) Oxalic acid and Ca	aCO ₃ c) Gluconic acid and CaCO ₃
d) Gluconic acid and Ca(OH) ₂	· ·
9. Limit test for heavy metals is based on reactive	ion between
${\mathfrak t}$ a) metallic impurities and hydrogen sulphide	b) metallic impurities and hydrogen sulphate
c) metallic impurities and acetic acid	d) metallic impurities and ammonia
10. The first Indian pharmacopoeia was publish	hed by Indian Pharmacopoeia committee in
o a) 1948 b) 1955 c) 1960 d) 1960	6
Q. II. Solve any One	10
1. Explain different theories of acid and base. E	Explain the mechanism of buffer action.
2. Discuss the method of preparation, uses and	l assay of sodium chloride and calcium gluconate.
Q. III. Solve any Two (5 Marks each)	10
1. Explain the principle behind limit test of Iron	and Sulphate.
2. Explain the principle of arsenic limit test. Dra	aw a neat labelled diagram of Gutzeit apparatus.
3. Explain the sources of impurities.	

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Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway First Sessional Theory Examination 2019-2020 Subject: Pharmaceutical Inorganic Chemistry (BP104T) Day & Date: 28-09-2019 Class: First Year B. Pharm. Semester: I Time: 10:30 am – 12:00 pm Max. Marks: 30 Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks QI. Solve the following Multiple Choice Questions (MCQs) 10 1. The sources of impurities include b) manufacturing process c) instability of product d) all of these a) raw materials 2. Lewis base is a. electron pair acceptor b. electron pair donor c. proton acceptor d. proton donor 3. Indian pharmacopoeia is published by a) world health organization b) ministry of health and family welfare c) ministry of science and technology d) food and drug administration 4. Citric acid is used in the limit test for iron to prevent a) precipitation of iron by ammonia b) precipitation of iron by thioglycolic acid c) oxidation of iron d) reduction of iron 5. Limit test for heavy metals is based on reaction between heavy metal and a) hydrogen sulphide b) hydrogen sulphate c) ammonia d) dithiazone 6. Lead acetate cotton plug is used in arsenic limit test to trap a) arsine gas b) hydrogen sulphide gas c) thiosulphate d) hydrogen sulphate 7. The major functions of electrolytes in body are a) to maintain osmotic pressure b) to maintain acid-base balance c) to maintain fluid volume d) all of the above 8. The concentration of sodium chloride solution that is isotonic with body fluids is a) 0.9 % w/y b) 1.6 % w/v c) 0.7 % w/v d) 2% w/v 9. Good dental cleaning agent generally have..... a) coarse to fine particle size b) good abrasive property c) both a and b d) none of these 10. Assay of calcium gluconate is performed by titration a) Redox b) complexometric c) acid-base d) precipitation QII. Solve any one of the following 10 1. Explain the different theories of acids and bases. Define Buffer. Explain the mechanism of

buffer action and buffer capacity.2. Explain the role of sodium and potassium in the body. Write a note on regulation of acid-base balance in body fluids.

QIII Solve the following (any 2)

1. What are dental caries? Explain the role of fluoride in the treatment of dental caries.

- 2. Write a note on ORS. Explain the preparation, properties, assay and uses of Sodium chloride
- 3. Explain the principle involved in limit test of Arsenic with sequence of reactions involved. Draw a neat labelled diagram of Gutzeit apparatus for limit test of Arsenic.



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First Sessional Theory Examination 2020-2021

Subject: Pharmaceutical Inorganic Chemistry (BP104T) Day & Date: 13-03-2021

Class: First Year B. Pharm.

Time: 10:30 am – 12:00 pm

Semester: I Max. Marks: 30

Instructions: 1. All questions are compulsory

2. Draw a well labeled diagram wherever necessary

3. Right hand side number indicates full marks

Q. I. Solve the following Multiple Choice Questions (MCQs)

10

- 1. The first Indian pharmacopoeia was published by Indian Pharmacopoeia committee in
 - a. 1948
 - b. 1955
 - c. 1960
 - d. 1966

2. Limit test for sulphate involves reaction between soluble sulphate and

- a. Silver nitrate
- b. Acetic acid
- c. Barium chloride
- d. Hydrochloric acid
- 3. Which of the following is a major extracellular cation?
 - a. Sodium
 - b. Chloride
 - c. Potassium
 - d. Phosphate
- 4. The role of sodium in body is
 - a. To maintain osmotic pressure
 - b. To transmission of nerve impulses
 - c. To maintain fluid volume
 - d. All of the above
- 5. Calcium gluconate is prepared from.....
 - a. Lactic acid and CaCO3
 - b. Oxalic acid and CaCO3
 - c. Gluconic acid and CaCO3
 - d. Gluconic acid and Ca(OH)₂

- 6. Fluoride inhibits caries formation via
 - a. Decreasing acid solubility of enamel
 - b. Bacterial inhibition
 - c. Both a and b
 - d. Increasing acid solubility of enamel
- 7. Absence of HCl in gastric secretion is called as
 - a. Hyperacidity
 - b. Achlorhydria
 - c. Constipation
 - d. Dirrhoea
- 8. Antacid should not be
 - a. Constipative
 - b. Laxative
 - c. Both a and b
 - d. Protective
- 9. What is incorrect about sodium bicarbonate?
 - a. It is systemic antacid
 - b. It is called as baking soda
 - c. Produces effervescence with acid
 - d. Does not get absorbed systemically
- 10. The sources of impurities include
 - a. Raw materials
 - b. Manufacturing process
 - c. Instability of product
 - d. All of the above

Q. II. Solve any one of the following

1. Explain various sources of impurities. Explain the principle involved in limit test for Arsenic with reactions. Draw a neat-labelled diagram of Gutzeit apparatus

2. Explain different theories of acids and bases. Add a note of mechanism of buffer action.

Q. III. Solve the following (any 2)

- 1. What are ideal properties of antacids? Classify antacid and write a note on Aluminium hydroxide gel.
- 2. Write a note on ORS. Explain the preparation, properties, assay and uses of Sodium chloride
- 3. Explain the principle and reaction involved in limit test of Iron

10



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First Sessional Theory Examination-(ODD SEM) 2022-2023

Subject	:	Pharmaceutical Inorganic Chemistry (BP104T)	Day & Date	:	Tue	es, 14/03/2023
Class	:	First Year B. Pharmacy	Semester	:	Ι	Write Your Roll No.
Time	:	10:30 am – 12:00 pm	Max. Marks	:	30	Here

Instructions:

- 1. All questions are compulsory
- 2. Draw a well-labeled diagram wherever necessary
- 3. Right-hand side number indicates full marks
- 4. Do not write/tick on the question paper

Q. I: Objective Type Questions

- 1. Define the following term. (CO1, CO2, LL1)
 - Pharmacopoeia ii) Desensitizing agents
- 2. Define Achlorhydria and gives its treatment. (CO2, LL1)
- 3. What is the role of citric acid and ammonia in the limit test of iron? (CO1, LL1)
- 4. Enlist the composition of ORS (CO2, LL1)
- 5. Write the principle and reaction behind the limit test of chloride and sulphate. (CO1, LL1)

Q. II: Long Answers Questions (Answer any 1 out of 2)

1. Define Buffers. Explain the mechanism of buffer action with an example. Briefly discuss the role of Buffers in Pharmacy. (CO1, LL3)

2. Define Impurities. Enlist and explain the various sources and types of impurities. (CO1, LL3)

Q. III: Short Answers Questions (Answer any 2 out of 3)

- 1. Explain the role of fluoride in the treatment of Tooth Decay. (CO2, LL3)
- 2. Define and discuss on Antacid used in Combination Therapy. (CO2, LL3)
- 3. Write note on Major Extracellular Physiological ions. (Any Two) (CO2, LL3)

10 M

10 M

10 M



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First Sessional Theory Examination-(ODD SEM) 2022-2023

Subject	:	Pharmaceutical Inorganic Chemistry (BP104T)	Day & Date	:	Tue	es, 14/03/2023
Class	:	First Year B. Pharmacy	Semester	:	Ι	Write Your Roll No.
Time	:	10:30 am – 12:00 pm	Max. Marks	:	30	Here

Instructions:

- 1. All questions are compulsory
- 2. Draw a well-labeled diagram wherever necessary
- 3. Right-hand side number indicates full marks
- 4. Do not write/tick on the question paper

Q. I: Objective Type Questions

- 1. Define the following term. (CO1, CO2, LL1)
 - Pharmacopoeia ii) Desensitizing agents
- 2. Define Achlorhydria and gives its treatment. (CO2, LL1)
- 3. What is the role of citric acid and ammonia in the limit test of iron? (CO1, LL1)
- 4. Enlist the composition of ORS (CO2, LL1)
- 5. Write the principle and reaction behind the limit test of chloride and sulphate. (CO1, LL1)

Q. II: Long Answers Questions (Answer any 1 out of 2)

1. Define Buffers. Explain the mechanism of buffer action with an example. Briefly discuss the role of Buffers in Pharmacy. (CO1, LL3)

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Q. III: Short Answers Questions (Answer any 2 out of 3)

- 1. Explain the role of fluoride in the treatment of Tooth Decay. (CO2, LL3)
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	First Sessional Theory Examinatio	n-(ODD SEM) 2023-2024	
Sul	bject : Pharmaceutical Inorganic Chemistry (BP104T)	Day & : Thu, 02/11/2 Date	023
Cla	Class : First Year B. Pharmacy Semester : I Write Your Rol Here		
Tin	ne : 10:30 am – 12:00 pm	Max. : 30 Marks	
Instr	1. All questions are compulsory 2.	Draw a well-labeled diagram wherever neces	ssary
11000	<i>3. Right-hand side number indicates full marks 4.</i>	Do not write/tick on the question paper	
Q.	I: Objective Type Questions		10 M
1.	Define the following term. (CO1, CO2, LL1)		
	i) Buffers ii) Antacid		
2.	Write the chemical reaction for limit test of lead. (CO1, LL3	3)	
3.	Define acid and base according to Lowry Bronsted theory w	ith example. (CO1, LL1)	
4.	Enlist the composition of Zinc-Eugenol Cement (CO2, LL3)	
5.	Define Hyponatremia and Hypernatremia. (CO2, LL1)		
Q.	II: Long Answers Questions (Answer any 1 out of	· 2)	10 M
1. E	Explain the types of tonicity and methods of adjusting isotonic	ity. Describe the mechanism of action o	f
buf	fers (CO1, LL3)		
2. E	Define limit test and describe the principle, apparatus and proc	redure for the limit test of arsenic. (CO1	, <i>LL3)</i>
Q.	III: Short Answers Questions (Answer any 2 out	of 3)	10 M
1. E	Define dentifrices and describe the method of M.W., M.F., pro-	perties, preparation and uses of CaCO ₃ .	(<i>CO2</i> ,
LL3	3)		
2. E	Define and discuss on Acidifiers with example. (CO2, LL3)		
3. V	Vhat is electrolyte combination therapy? Explain a note on Ol	RS. (Any Two) (<i>CO2, LL3</i>)	



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	First Sessional Theory Examinatio	n-(ODD SEM) 2023-2024	
Sul	bject : Pharmaceutical Inorganic Chemistry (BP104T)	Day & : Thu, 02/11/2 Date	023
Cla	Class : First Year B. Pharmacy Semester : I Write Your Rol Here		
Tin	ne : 10:30 am – 12:00 pm	Max. : 30 Marks	
Instr	1. All questions are compulsory 2.	Draw a well-labeled diagram wherever neces	ssary
11000	<i>3. Right-hand side number indicates full marks 4.</i>	Do not write/tick on the question paper	
Q.	I: Objective Type Questions		10 M
1.	Define the following term. (CO1, CO2, LL1)		
	i) Buffers ii) Antacid		
2.	Write the chemical reaction for limit test of lead. (CO1, LL3	3)	
3.	Define acid and base according to Lowry Bronsted theory w	ith example. (CO1, LL1)	
4.	Enlist the composition of Zinc-Eugenol Cement (CO2, LL3)	
5.	Define Hyponatremia and Hypernatremia. (CO2, LL1)		
Q.	II: Long Answers Questions (Answer any 1 out of	· 2)	10 M
1. E	Explain the types of tonicity and methods of adjusting isotonic	ity. Describe the mechanism of action o	f
buf	fers (CO1, LL3)		
2. E	Define limit test and describe the principle, apparatus and proc	redure for the limit test of arsenic. (CO1	, <i>LL3)</i>
Q.	III: Short Answers Questions (Answer any 2 out	of 3)	10 M
1. E	Define dentifrices and describe the method of M.W., M.F., pro-	perties, preparation and uses of CaCO ₃ .	(<i>CO2</i> ,
LL3	3)		
2. E	Define and discuss on Acidifiers with example. (CO2, LL3)		
3. V	Vhat is electrolyte combination therapy? Explain a note on Ol	RS. (Any Two) (<i>CO2, LL3</i>)	



Shri. Vile Parle Kelvani Mandal's Institute of Pharmacy, Dhule Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

First Sessional Theory Examination 2020-2021

Subject: Pharmaceutical Analysis I (BP102T)	Day & Date: 11-03-2021
Class: First Year B. Pharm.	Semester: I
Time: 10:30 am – 12:00 pm	Max. Marks : 30

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary

3. Right hand side number indicates full marks

Q. I. Solve the following Multiple Choice Questions (MCQs)

- 1. Number of gram equivalents of solute present in one litre of a solution is called as
 - a) Molarity
 - b) Normality
 - c) Molality
 - d) Percentage weight by volume
- 2. The features of Primary standard substance includes:
 - a) High purity
 - b) High stability
 - c) Low hygroscopicity
 - d) All of the above
- 3. Indian pharmacopoeia is published by
 - a. World health organization
 - b. Ministry of health and family welfare
 - c. Ministry of science and technology
 - d. Food and drug administration
- 4. Limit test of iron is based upon the reactions between iron and
 - a. Dithiazone
 - b. Thioglycolic acid
 - c. Hydrogen sulphide
 - d. Barium chloride
- 5. The color of phenolphthalein in alkanine medium is
 - a) Pink
 - b) Yellow
 - c) Brown
 - d) Colorless

6. Benzene is an example ofsolvent.

- a) Aprotic
- b) Protogenic
- c) Protophilic
- d) Amphiprotic

7. Solubility of precipitate is on addition of common ion.

- a) Decreased
- b) Increased
- c) Remains Constant
- d) None of These
- 8. In Mohr's methodis used as an indicator.
 - a) Potassium chromate
 - b) Ferric ammonium sulphate
 - c) Ammonium thiocyanate
 - d) Silver nitrate
- 9. In Volhard's methodis used as a titrant
 - a) Potassium chromate
 - b) Ferric ammonium sulphate
 - c) Ammonium thiocyanate
 - d) Silver nitrate

10. The end-point detection in Fajan's method involves color change due to.....

- a) Formation of colored precipitate
- b) Formation of soluble colored complex
- c) Adsorption of indicator on the surface of precipitate
- d) None of the above

Q. II. Solve the following (any 1)

- 1. Explain the neutralization curves in acid-base titrations for i) strong acid vs strong base ii) weak acid vs strong base iii) strong acid vs weak base iv) weak acid vs weak base.
- 2. Explain in detail Mohr's method and Fajan's method.

Q. III. Solve the following (any 2)

- 1. Explain principle involved in limit test for Chloride and sulphate.
- 2. Explain in detail the theories of acid-base indicators.
- 3. Explain the types of solvents used in non-aqueous titrations with suitable examples.

10



Shri. Vile Parle Kelvani Mandal's Institute of Pharmacy, Dhule Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

First Sessional Theory Examination 2020-2021

Subject: Pharmaceutical Analysis I (BP102T)	Day & Date: 11-03-2021
Class: First Year B. Pharm.	Semester: I
Time: 10:30 am – 12:00 pm	Max. Marks : 30

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary

3. Right hand side number indicates full marks

Q. I. Solve the following Multiple Choice Questions (MCQs)

- 1. Number of gram equivalents of solute present in one litre of a solution is called as
 - a) Molarity
 - b) Normality
 - c) Molality
 - d) Percentage weight by volume
- 2. The features of Primary standard substance includes:
 - a) High purity
 - b) High stability
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 - d) All of the above
- 3. Indian pharmacopoeia is published by
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 - c. Hydrogen sulphide
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7. Solubility of precipitate is on addition of common ion.

- a) Decreased
- b) Increased
- c) Remains Constant
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- 8. In Mohr's methodis used as an indicator.
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 - a) Potassium chromate
 - b) Ferric ammonium sulphate
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- a) Formation of colored precipitate
- b) Formation of soluble colored complex
- c) Adsorption of indicator on the surface of precipitate
- d) None of the above

Q. II. Solve the following (any 1)

- 1. Explain the neutralization curves in acid-base titrations for i) strong acid vs strong base ii) weak acid vs strong base iii) strong acid vs weak base iv) weak acid vs weak base.
- 2. Explain in detail Mohr's method and Fajan's method.

Q. III. Solve the following (any 2)

- 1. Explain principle involved in limit test for Chloride and sulphate.
- 2. Explain in detail the theories of acid-base indicators.
- 3. Explain the types of solvents used in non-aqueous titrations with suitable examples.

10



First Sessional Theory Examination 2020-2021

Subject: Pharmaceutical Inorganic Chemistry (BP104T) Day & Date: 13-03-2021

Class: First Year B. Pharm.

Semester: I

Time: 10:30 am – 12:00 pm

Max. Marks: 30

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks

Q. I. Solve the following Multiple Choice Questions (MCQs)

10

- 1. The first Indian pharmacopoeia was published by Indian Pharmacopoeia committee in
 - a. 1948
 - b. 1955
 - c. 1960
 - d. 1966

2. Limit test for sulphate involves reaction between soluble sulphate and

- a. Silver nitrate
- b. Acetic acid
- c. Barium chloride
- d. Hydrochloric acid
- 3. Which of the following is a major extracellular cation?
 - a. Sodium
 - b. Chloride
 - c. Potassium
 - d. Phosphate
- 4. The role of sodium in body is
 - a. To maintain osmotic pressure
 - b. To transmission of nerve impulses
 - c. To maintain fluid volume
 - d. All of the above
- 5. Calcium gluconate is prepared from.....
 - a. Lactic acid and CaCO3
 - b. Oxalic acid and CaCO3
 - c. Gluconic acid and CaCO3
 - d. Gluconic acid and Ca(OH)₂

- 6. Fluoride inhibits caries formation via
 - a. Decreasing acid solubility of enamel
 - b. Bacterial inhibition
 - c. Both a and b
 - d. Increasing acid solubility of enamel
- 7. Absence of HCl in gastric secretion is called as
 - a. Hyperacidity
 - b. Achlorhydria
 - c. Constipation
 - d. Dirrhoea
- 8. Antacid should not be
 - a. Constipative
 - b. Laxative
 - c. Both a and b
 - d. Protective
- 9. What is incorrect about sodium bicarbonate?
 - a. It is systemic antacid
 - b. It is called as baking soda
 - c. Produces effervescence with acid
 - d. Does not get absorbed systemically
- 10. The sources of impurities include
 - a. Raw materials
 - b. Manufacturing process
 - c. Instability of product
 - d. All of the above

Q. II. Solve any one of the following

1. Explain various sources of impurities. Explain the principle involved in limit test for Arsenic with reactions. Draw a neat-labelled diagram of Gutzeit apparatus

2. Explain different theories of acids and bases. Add a note of mechanism of buffer action.

Q. III. Solve the following (any 2)

- 1. What are ideal properties of antacids? Classify antacid and write a note on Aluminium hydroxide gel.
- 2. Write a note on ORS. Explain the preparation, properties, assay and uses of Sodium chloride
- 3. Explain the principle and reaction involved in limit test of Iron

10



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First Sessional Theory Examination 2021-2022 (Odd SEM)

Subject & Subject Code: Pharmaceutical Inorganic Chemistry (BP104T)

Day & Date: Thursday 10/02/2022

Class: F.Y. B. Pharmacy

Time: 10:30 AM to 12:00 PM

Semester: I

Max. Marks: 30

Instructions: 1. All questions are compulsory
2. Draw a well labeled diagram wherever necessary
3. Right hand side number indicates full marks
QI Multiple Choice Questions (MCQs) 10
1. Ammonium chloride is an example of
A) Stimulant expectorant B) Sedative expectorant
C) Both a and b D) Antacid
2. Citric acid is added in limit test of iron to
A) Produce pink colour B) To form complex with iron and prevent the
precipitation by ammonia Solution
C) React with iron D) Produce acidity
3. Dental cement consists of
A) Zn + Eugenol B) Na + Mannitol C) Al + Cresol D) None of above
4. Alcohol is used in barium sulphate reagent to prevent
A) Saturation B) Supersaturation C) Turbidity D) Opalascence
5. Impurities in pharmaceutical preparation may be due to following
A) Raw material B) Manufacturing process
C) Chemical instability D) All of the above
6. In which method tonicity is calculated by adding water to the drugs to
make an isotonic solution
A) Sodium chloride equivalent method B) Cryoscopic method
C) White Vincent method D) Potentiometric method
7. First edition of IP was published in
A) 1900 B) 1948 C) 1955 D) 1900
0. The body huld within the cens is called as
9 To prevent dental caries it becomes necessary to use
A) Sodium Fluoride B) Gold particles () Silver nitrate D) Conner sulphate
10 Which of the following is ideal properties of dental product
A) Easy to use B) Must not produce any stain on teeth
C) Must not produce nausea or vomiting D) All of the above
OII Long Answers (Answer any 1 out of 2) 10
1. Define the limit test and discuss in detail the limit test for arsenic with
suitable diagram of gutzeit apparatus.
2. Define the term isotonicity and discuss in detail different methods used
to measure the isotonicity
OIII Short Answers (Answer any 2 out of 3) 10
1 Define anti - carries agents Discuss how fluoride produces anti-carries
activity
activity O Define and Classific optacide with socially account and size (1) (1) (1)
2. Denne and Classify antacids with suitable example and give the ideal
properties of Antacid.

3. Discuss in detail Metabolic Acidosis and Alkalosis.



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First Sessional Theory Examination 2021-2022 (Odd SEM)

Subject & Subject Code: Pharmaceutical Inorganic Chemistry (BP104T)

Day & Date: Thursday 10/02/2022

Class: F.Y. B. Pharmacy

Time: 10:30 AM to 12:00 PM

Semester: I

Max. Marks: 30

Instructions: 1. All questions are compulsory
2. Draw a well labeled diagram wherever necessary
3. Right hand side number indicates full marks
QI Multiple Choice Questions (MCQs) 10
1. Ammonium chloride is an example of
A) Stimulant expectorant B) Sedative expectorant
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First Sessional Theory Examination 2021-2022 (Odd SEM)

Subject & Subject Code: Pharmaceutical Inorganic Chemistry (BP104T)

Day & Date: Thursday 10/02/2022

Class: F.Y. B. Pharmacy

Time: 10:30 AM to 12:00 PM

Semester: I

Max. Marks: 30

Instructions: 1. All questions are compulsory
2. Draw a well labeled diagram wherever necessary
3. Right hand side number indicates full marks
QI Multiple Choice Questions (MCQs) 10
1. Ammonium chloride is an example of
A) Stimulant expectorant B) Sedative expectorant
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A) Produce pink colour B) To form complex with iron and prevent the
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A) Zn + Eugenol B) Na + Mannitol C) Al + Cresol D) None of above
4. Alcohol is used in barium sulphate reagent to prevent
A) Saturation B) Supersaturation C) Turbidity D) Opalascence
5. Impurities in pharmaceutical preparation may be due to following
A) Raw material B) Manufacturing process
C) Chemical instability D) All of the above
6. In which method tonicity is calculated by adding water to the drugs to
make an isotonic solution
A) Sodium chloride equivalent method B) Cryoscopic method
C) White Vincent method D) Potentiometric method
7. First edition of IP was published in
A) 1960 B) 1948 C) 1955 D) 1966
8. The body fluid within the cells is called as
A) plasma B) Intracellular fluid C) Extracellular fluid D) Water
9. To prevent dental carles it becomes necessary to use
A) Sodium Fluoride B) Gold particles C) Silver nitrate D) Copper sulphate
A) Fase to use B) Must not produce any stain on teeth
C) Must not produce pouses or vomiting D All of the above
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2. Define the term isotonicity and discuss in detail different methods used
to measure the isotonicity.
QIII Short Answers (Answer any 2 out of 3) 10
1. Define anti - carries agents. Discuss how fluoride produces anti-carries
activity
2. Define and Classify antacids with suitable example and give the ideal
properties of Antacid.

3. Discuss in detail Metabolic Acidosis and Alkalosis.



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First Sessional Theory Examination-(ODD SEM) 2022-2023

Subje	jectRemedial Biology (BP106RBT)Day & Date :THURSDAY,09/03						3/20)23		
Class	;	:	First Year B. Pharmacy	Semester	:	Ι	Write	Your	Seat	No.
Time		:	10:30 am – 11:30 am	Max. Marks	••	30	неге			
Instruc	tions	s: 1. 3.	All questions are compulsory Right hand side number indicates full marks	2. Draw a well lab 4. Do not write/tick	elea k on	l diagran the ques	ı wherei tion pap	ver nec er	essarı	ý
QI	Long Answers Question (Answer any 1 out of 2)10									
1.	De	escr	ibe kingdom of life with suitable example (LL3,	CO1)						
2.	Elaborate types of Inflorescences with suitable example (LL3, CO1)									
QII	Short Answers Question									
	Sec	tion	n – A (Answer any ONE)						05	Μ
1.	Wr	ite	a note on classification and uses of Fruit (LL2, C	201)						
2.	Illuminate the general histological characters of Dicot and Monocot root. (LL2, CO1)									
	Section – B (Answer any Three) 15							5 M		
1.	Write a note on composition and functions of the Lymphatic system (LL2, CO2)									
2.	Comment on Digestive enzymes (LL2, CO2)									
3.	. Explain the mechanism of breathing (LL2, CO2)									
4.	Draw well labeled structure of Human Urinary System. (LL2, CO2)									



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First Sessional Theory Examination-(ODD SEM) 2022-2023

Subje	bject:Remedial Biology (BP106RBT)Day & Date:THURSDAY,09				<u>09/</u> 0	3/20	23			
Class	;	:	First Year B. Pharmacy	Semester	:	Ι	Write	Your	Seat	No.
Time	Cime : 10:30 am - 11:30 am Max. Marks : 30									
Instruc	I. All questions are compulsory2. Draw a well labeled diagram wherever necessary3. Right hand side number indicates full marks4. Do not write/ tick on the question paper									1
QI	Long Answers Question (Answer any 1 out of 2) 10									
1.	De	escr	ibe kingdom of life with suitable example (LL3,	CO1)						
2.	Ela	boı	ate types of Inflorescences with suitable example	e (LL3, CO1)						
QII	Short Answers Question									
	Sec	tior	n – A (Answer any ONE)						05]	Μ
1.	Wr	ite	a note on classification and uses of Fruit (LL2, C	01)						
2.	Illu	mi	nate the general histological characters of Dicot a	and Monocot roo	t. (1	LL2, CC	01)			
	Section - B (Answer any Three)15							; M		
1.	Write a note on composition and functions of the Lymphatic system (LL2, CO2)									
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First Sessional Theory Examination-(ODD SEM) 2022-2023

Subje	bject:Remedial Biology (BP106RBT)Day & Date:THURSDAY,09				<u>09/</u> 0	3/20	23			
Class	;	:	First Year B. Pharmacy	Semester	:	Ι	Write	Your	Seat	No.
Time	Cime : 10:30 am - 11:30 am Max. Marks : 30									
Instruc	I. All questions are compulsory2. Draw a well labeled diagram wherever necessary3. Right hand side number indicates full marks4. Do not write/ tick on the question paper									1
QI	Long Answers Question (Answer any 1 out of 2) 10									
1.	De	escr	ibe kingdom of life with suitable example (LL3,	CO1)						
2.	Ela	boı	ate types of Inflorescences with suitable example	e (LL3, CO1)						
QII	Short Answers Question									
	Sec	tior	n – A (Answer any ONE)						05]	Μ
1.	Wr	ite	a note on classification and uses of Fruit (LL2, C	01)						
2.	Illu	mi	nate the general histological characters of Dicot a	and Monocot roo	t. (1	LL2, CC	01)			
	Section - B (Answer any Three)15							; M		
1.	Write a note on composition and functions of the Lymphatic system (LL2, CO2)									
2.	Comment on Digestive enzymes (LL2, CO2)									
3.	Explain the mechanism of breathing (LL2, CO2)									
4.	Draw well labeled structure of Human Urinary System. (LL2, CO2)									



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First Sessional Theory Examination 2020-2021 (Odd SEM) Subject & Subject Code: Remedial Biology (BP106RBT)

Day & Date: 16.03.2021 Tuesday

Class: F.Y. B. Pharm

Time: 10.30 to 11.30

Max. Marks: 30 marks

Semester: I

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks

Q1. Answer any one

(10 marks)

(20 marks)

Explain the morphology and anatomy of flower with well labelled diagram.
 Give structure of human heart and explain in brief about cardiac cycle

Q2. Answer any fours

- 1. Enlist difference between Eukaryotic and Prokaryotic cells
- 2. Write a note on binominal nomenclature
- 3. Explain about ECG
- 4. Give a short note on mechanism of breathing
- 5. Explain the role of Digestive Enzymes
- 6. Give composition of blood

<u>Kahod</u> [Is-s-Rathod) Subject Incharge

Institute of Pharmacy, Dhule

Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

First Sessional Theory Examination 2018-2019

Subject: Remedial Biology	Day & Date: Monday, 8th October 2018
Class: F. Y. B. Pharma	Semester: I
Time: 10:30 am -12:00 pm	Max. Marks: 30

Instructions: 1. All questions are compulsory

2. Draw a well labeled diagram wherever necessary

3. Right hand side number indicates full marks

QI Answer any one

- 1. Short note on human heart and blood vessels. Describe in details Cardiac cycle
- 2. Explain the kingdoms of life and basis of classification. Short note on Salient features of

Plantae and Virus.

QII Answer any four

- 1. Short note on breathing process and its regulation
- 2. What is blood and explain its Composition and function.
- 3. Describe Role of digestive enzymes in digestion process
- 4. Explain General Anatomy of Root, stem, leaf of monocotyledons & Dicotylidones.
- 5. Short note on Electrocardiogram
- 6. Definition and characters of organisms.

10

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First Sessional Theory Examination 2018-2019 Class: F.Y.B.Pharmacy Semester: First Subject : Remedial Mathematics Max. Marks: 30 Date: October. 8th, 2018 Time : 10.30 am - 12.09 pm. • Instructions : 1. Both the questions are compulsory. 2. Figures to the right indicates full marks. 1. Attempt any two $\oint f$ the following. (5+5=10)(a) Solve by Cramer's rule x + y + z = 3,x - y + z = 1, x + y - 2z = 0.(b) Using log-tables find the value of $(30.2153)^{\frac{3}{2}}(0.00516)^{\frac{2}{3}}$ $\sqrt{0.02157}$ (c) Resolve into partial fractions $\frac{x}{(x+1)^3(x-2)^2}$. 2. Attempt any five of the following. (4x5=20)(a) Show that $7\log_{15}^{16} + 5\log_{24}^{25} + 3\log_{80}^{81} + 21\log_{10}^{81} = \log_{10}^{81}$ (b) Resolve into partial fractions $\frac{x^2+1}{x(x-1)(x-3)}$. (c) Solve the equation, $\begin{vmatrix} x+a & b & c \\ a & x+b & c \\ a & b & x+c \end{vmatrix} = 0.$ (d) Solve $log_{10}(x+1) - log_{10}(x-1) = 1$. (e) Without expanding the determinant show that $\begin{vmatrix} 0 & b & -c \\ -b & 0 & a \\ c & -a & 0 \end{vmatrix} = 0.$ (f) Is the matrix $\begin{bmatrix} 1 & 2 & 3 \\ 0 & 1 & 2 \\ 0 & 0 & 1 \end{bmatrix}$ idempotent? Justify. (g) If $A = \begin{bmatrix} 3 & 5 & -1 \\ 6 & 4 & 2 \end{bmatrix}$ and $B = \begin{bmatrix} 2 & 5 & -1 \\ 1 & 3 & -2 \end{bmatrix}$, verify that $(A+B)^t = A^t + B^t$. *****



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First Sessional Theory Examination 2019-2020

Subject: Remedial Mathematics (BP106RMT)Day & Date: Tuesday, 01/10/19Class: First Year B. Pharm,Semester: I/III*Time: 01-30 pm to 03.00 pmMax. Marks: 30

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks QI Attempt any two of the following (5+5=10)(1).Find the value of x if $\begin{vmatrix} 1 & 2 & 3 \\ 1 & x & 4 \\ 2 & -3 & -1 \end{vmatrix} = 0.$ $\mathcal{R} = 3$ (2).Find the value of $\begin{vmatrix} 2 & -3 & 4 \\ 6 & 1 & -2 \\ 0 & -2 & -6 \end{vmatrix} = -176$. (3).Find inverse of $A = \begin{vmatrix} 7 & -3 & 2 \\ 4 & 0 & 1 \\ 1 & 2 & 3 \end{vmatrix}$ (4×5=20) QII Attempt any five of the following (1).Define the following terms with example . $O_{7} = \begin{vmatrix} 1 & 0 & -1 \\ 2 & 9 & 3 \\ 1 & 2 & 1 \end{vmatrix} = 3 - (4 - 9.) = 3 = 3 - (-5) = 8$ (i) Diagonal Matrix. (ii) Symmetric Matrix (iii) Skew symmetric Matrix.(iv)Transpose of a matrix. (2). Define the following terms with example $D_{2} = \begin{vmatrix} 1 & 1 & 0 \\ 2 & 1 & 9 \\ -1 & 6 \end{vmatrix} = 11 - (-5) = 16$ (i) Singular Matrix (ii) Non singular Matrix. (iii)Transpose of a matrix.(iv) Identity Matrix. (3). Find minors and cofactors of each elements of A=3(4).Solve by Cramer's rule x + y-z=0 $2x + y + 3z = 9 \quad D_{\pi} = 8 \qquad \chi = 1$ $X - y + z = 2 \qquad D_{\pi} = 8 \qquad y = 1$ $P_{2} = 16 \qquad z = 2$ (5) If $A = \begin{bmatrix} 3 & 1 & 1 \\ -1 & 5 & -1 \\ -1 & -1 & 3 \end{bmatrix}$ And $B = \begin{bmatrix} 1 & 3 & 6 \\ 1 & 4 & 5 \\ 1 & 5 & 4 \end{bmatrix}$ Verify that $(A+B)^{T} = A^{T} + B^{T}$ (6). Find The value of Determinant Of The Matrix $\begin{bmatrix} 2 & 5 & 7 \\ 8 & 1 & 0 \\ 12 & 11 & 14 \end{bmatrix} = 0$ (7). Verify that A(B+C)=AB+AC. Where $A = \begin{bmatrix} 1 & 2 \\ 3 & -1 \end{bmatrix}$, $B = \begin{bmatrix} -1 & 2 \\ 1 & 0 \end{bmatrix}$, $C = \begin{bmatrix} 2 & -3 \\ 4 & -8 \end{bmatrix}$.



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First Sessional Theory Examination 2021-2022 (Odd SEM) Subject & Subject Code: Remedial Mathematics (BP106RMT)

Day & Date: Saturday, 12/02/2022

Class: F.Y/S.Y. B. Pharmacy

Semester: I/III

Time:1 hr

Max. Marks: 30

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks

QI Multiple Choice Questions (MCQs)

1. If $5^3 = 125$ Then its logarithmic form is (a) $Log_5 125 = 3$ (b) $Log_5 25 = 3$ (c) $Log_3 125 = 5$ (d) $Log_{125}5 = 3$ 2. If $A = \begin{bmatrix} 2 & 4 \\ X & 0 \end{bmatrix}$ is a symmetric matrix, then x = (a) 4 (b) 3 (c) -4 (d) -3 3. If $A^T = -A$ Then A is (a) diagonal matrix (b) skew-symmetric matrix (c) symmetric matrix (d) none of these **4.** If $Log_{\left(\frac{1}{2}\right)}\left(\frac{1}{8}\right) = 3$ then its exponential form is (a) $\left(\frac{1}{3}\right)^3 = \frac{1}{27}$ (b) $\left(\frac{1}{2}\right)^3 = 8$ (c) $\left(\frac{1}{2}\right)^3 = \frac{1}{8}$ (d) $\left(\frac{1}{2}\right)^4 = \frac{1}{8}$ (d) 5. Let A is any square Matrix Then A^{-1} exist if (a) $|A| \neq 0$ (b) |A| = 0(c) A is singular Matrix (d) None 6. The value of $\begin{vmatrix} 1 & 2 & 3 \\ -3 & -6 & -9 \\ 2 & 1 & 0 \end{vmatrix}$ is (a) 4 (b) 3 (c) -4 (d) 0



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7. Lim
$$\frac{x^2-5x+6}{x^2-4} =$$

(a) $\frac{1}{8}$ (b) $\frac{3}{4}$
(c) $\frac{1}{4}$ (d) $-\frac{1}{4}$
(d) $-\frac{1}{4}$
8. Lim $\frac{x^5-3^5}{x-3} =$
(a) 410
(b) 405
(c) 415
(d) 420
9. If $\begin{vmatrix} 1 & 2 \\ 3 & x \end{vmatrix} = 0$ then x is
(a) 2
(b) 3
(c) 6
(d) 1
10 If $A = \begin{bmatrix} 1 & 2 \\ 3 & 7 \end{bmatrix}$ then A^{-1} is
(a) $\begin{bmatrix} 7 & -2 \\ -3 & 1 \end{bmatrix}$
(b) $\begin{bmatrix} 1 & -2 \\ 3 & 7 \end{bmatrix}$
(c) $\begin{bmatrix} 1 & 2 \\ -3 & 7 \end{bmatrix}$
(d) $\begin{bmatrix} 1 & -2 \\ -3 & 7 \end{bmatrix}$

QII Long Answers (Answer any 1 out of 2)

1. Solve the following System of equations: x + y + z = 6 2x + 3y - z = 5 6x - 2y - 3z = -7Using crammers rule. 2. Solve the following System of equations: -x + 3y-2z = 5

4x - y - 3z = -82x+2y-5z = 7

Using crammers rule.

QIII Short Answers (Answer any 2 out of 3)

1. Resolve into the partial fraction: $\frac{x^2-3x+1}{(x-1)^2(x-2)}$ 2. Resolve into the partial fraction: $\frac{x-6}{(x^2+9)(x+2)}$

3. Prove that:
$$2\log\left(\frac{15}{18}\right) - \log\left(\frac{25}{162}\right) + \log\left(\frac{4}{9}\right) = \log 2.$$

10



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	First/Second Sessional Theory Examination-(ODD SEM) 2022-2023								
Subje	ect	:	Remedial Mathematics(BP106RMT)	Day & Date	:	Thu	ırsda	y,09/03	/2023
Class	5	••	First Year B. Pharmacy	Semester	:	Ι	Writ	te Your Seat	No. Here
Time		••	10:30 am – 11:30 am	Max. Marks	:	30			
Instruc	tions	s:1. 3.	All questions are compulsory2.Right hand side number indicates full marks4.	Draw a well label Do not write/tick o	ed á on th	liagrai .e que	n whe stion p	rever neces aper	sary
QII	Lor	ng A	Answers Question (Answer any 1 out of 2)						10 M
1.	Res	solv	we into Partial Fraction $\frac{9x-7}{(x+3)(x^2+1)}$				(LL	3, CO1)	
2.	Usi	ing	the adjoint Method, Find inverse of A= $\begin{bmatrix} 1 & -1 \\ 2 & -1 \\ 1 & 1 \end{bmatrix}$	$\begin{bmatrix} 2\\1\\-1 \end{bmatrix}$			(LL	3, CO2)	
QIII	Sho	ort 4	Answers Question (Answer any 4 out of 6)						20 M
1.	Res	solv	the into the partial fraction: $\frac{x^2 - 3x + 1}{(x-1)^2(x-2)}$				(LL.	3, CO1)	
2.	Fine	d th	e value of x, if $log(x+5)+log(x-5)=4log2+2log3$				(LL3	3, CO1)	
3.	Prove that : $2\log(\frac{15}{18}) - \log(\frac{25}{162}) + \log(\frac{4}{9}) = log2$ (LL3, CO1)								
4.	Sol	ve t	the following System of equations: $x + y + z = 2x + 3y - z = 6x - 2y - 3z = 6x - 2y - 3z = 2x + 3y - 2z = 2x + 3y - 3z = 2x + 3y - 3x + 3y - 3x + 3y + 3x + 3y - 3x + 3y + 3y + 3y + 3y + 3y + 3x + 3y + 3y$	6 5 -7			(LL	3, CO2)	
5.	Fine	d C	haracteristics equation of of A= $\begin{bmatrix} 8 & -6 & 2 \\ -6 & 7 & -4 \\ 2 & -4 & 3 \end{bmatrix}$				(LL	3, CO2)	
6.	Ver	ify	Cayley Hamilton Theorem for $A = \begin{bmatrix} 1 & 1 \\ 1 & 0 \end{bmatrix}$				(LL	3, CO2)	



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	First/Second Sessional Theory Examination-(ODD SEM) 2022-2023								
Subje	ect	:	Remedial Mathematics(BP106RMT)	Day & Date	:	Thu	ırsda	y,09/03	/2023
Class	5	••	First Year B. Pharmacy	Semester	:	Ι	Writ	te Your Seat	No. Here
Time		••	10:30 am – 11:30 am	Max. Marks	:	30			
Instruc	tions	s:1. 3.	All questions are compulsory2.Right hand side number indicates full marks4.	Draw a well label Do not write/tick o	ed á on th	liagrai .e que	n whe stion p	rever neces aper	sary
QII	Lor	ng A	Answers Question (Answer any 1 out of 2)						10 M
1.	Res	solv	we into Partial Fraction $\frac{9x-7}{(x+3)(x^2+1)}$				(LL	3, CO1)	
2.	Usi	ing	the adjoint Method, Find inverse of A= $\begin{bmatrix} 1 & -1 \\ 2 & -1 \\ 1 & 1 \end{bmatrix}$	$\begin{bmatrix} 2\\1\\-1 \end{bmatrix}$			(LL	3, CO2)	
QIII	Sho	ort 4	Answers Question (Answer any 4 out of 6)						20 M
1.	Res	solv	the into the partial fraction: $\frac{x^2 - 3x + 1}{(x-1)^2(x-2)}$				(LL.	3, CO1)	
2.	Fine	d th	e value of x, if $log(x+5)+log(x-5)=4log2+2log3$				(LL3	3, CO1)	
3.	Prove that : $2\log(\frac{15}{18}) - \log(\frac{25}{162}) + \log(\frac{4}{9}) = log2$ (LL3, CO1)								
4.	Sol	ve t	the following System of equations: $x + y + z = 2x + 3y - z = 6x - 2y - 3z = 6x - 2y - 3z = 2x + 3y - 2z = 2x + 3y - 3z = 2x + 3y - 3x + 3y - 3x + 3y + 3x + 3y - 3x + 3y + 3y + 3y + 3y + 3y + 3x + 3y + 3y$	6 5 -7			(LL	3, CO2)	
5.	Fine	d C	haracteristics equation of of A= $\begin{bmatrix} 8 & -6 & 2 \\ -6 & 7 & -4 \\ 2 & -4 & 3 \end{bmatrix}$				(LL	3, CO2)	
6.	Ver	ify	Cayley Hamilton Theorem for $A = \begin{bmatrix} 1 & 1 \\ 1 & 0 \end{bmatrix}$				(LL	3, CO2)	



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		First Sessional Theory Examination	-(Odd SEM) 2023-2	2024				
Subject	:	Remedial Mathematics	Day & Date	:	Saturday, 4/11/2023			
Class	:	First Year B. Pharmacy	Semester	:	I/III	Write Your Seat No.		
Time	:	10:30 am to 11:30 am	Max. Marks	:	30	Here		

Instru	<i>uctions:</i> 1. All questions are compulsory 3. Right-hand side number indicate	es full marks	2. Draw a well-labeled diag 4. Do not write/tick on the q	gram wherever necessary westion paper
Q. I	I: Long Answers Question (Answer	any 1 out of	(2)	10 M
a)	Solve the equations using Cramer $x = y + z = 0$, $2x = y + z = 1$	r's Rule of d	eterminants:	
b)	i) Find the value of x if : $\begin{vmatrix} 1 & 1 & 1 \\ 3 & x & 3 \\ 1 & x & 2 \end{vmatrix}$	=0		
	ii) If $f(x) = \frac{x+5}{3x-4}$ and $t = \frac{5+4x}{3x-2}$ sh	ow that f(t) = x	
Q. I	II: Short Answers Question (Answe	r any 4 out o	of 6)	20 M
a)	Resolve into partial fraction $(x+3)$	13x + 19 (x - 2)(x + 1)	
b)	Resolve into partial fraction $1 x^3 + x^3 + 3 x^3 $	1		
c)	Express the following as a sing $\log_2 4 + 2 \log_2 3 - 3 \log_2 2$	le logarith	m using Properties of lo	garithms3
d)	Find the value of $f(1/2) \& f(1/4)$	$f(x) = \int f(x) = \int f(x) dx$	$16^x - \log_2 x$	
e)	$i) \lim_{x \to 0} \frac{3^x - 2^x}{\tan 4x}$	$ii) \lim_{x\to 0} 3^t$	$\frac{2^{2n}x}{2x}$	
f)	$ \begin{array}{cccccccccccccccccccccccccccccccccccc$			



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Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

Second Sessional Theory Examination 2020-2021 Subject: Pharmaceutics I BP103T Day & Date: Tuesday, 27th April 2021

Class: FYB. Pharm Semester: I Max. Marks: 30

Time: 90 min

1 1

Instructions: 1. All questions are compulsory

2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks

Q.No.1 Multiple choice questions

1. In the mixing of thymol and menthol the following type of incompatibility occurs,

b. Therapeutic incompatibility a. Chemical incompatibility

c. Physical incompatibility

d. None of these.

2. Upward creaming is observed in?

a. W/O b. O/W c. Both d. Micro

3. For ideal suspension, the sedimentation volume should be

a. Zero b. Equal to one c. More than one d. Less than one

4. Which of the following is an example of Hydrophilic bases?

a. Hydrogenated oils b. Emulsified cocoa butter c. Glycero-gelatin base d. None of these

5. In stokes' relationship, a parameter that greatly changes the velocity of settling is: a. Density of the liquid b. Density of particle c. Radius of the particle d. Viscosity of the medium

6. Structured vehicle is included in the formulation of a suspension, in order to:

b. Prevents the caking of the sediment a. Decreases the interfacial tension

d. Reduces the size by chemical means c. Prevents the sedimentation of particles

7. In the preparation of vanishing creams, which types of bases are used generally? a. Absorption bases b. Water removable bases c. Hydrocarbon bases d. None

8. The melting range of cocoa butter lies between a. 12 – 13°C b. 20 – 30°C c. 5 – 10°C d. 30 – 36° C (10 X 1=10)
9. Tetracycline interacts with calcium to form chelates is an example of

a. Chemical incompatibility b. Therapeutic incompatibility

c. Physical incompatibility d. Tolerance incompatibility.

10. What is the difference between vanishing cream and cold cream?

- a. Both are examples of oil-in-water emulsions
- b. Vanishing cream is an oil-in-water emulsion whereas cold cream is a water-in-oil emulsion

c. Vanishing cream is a water-in-oil emulsion whereas cold cream is an oil-in-water emulsion

d. Both are examples of water-in-oil emulsions

Q. No. 2 Long Answers (Answer 1 out of 2)

1. Discuss stability problems of emulsion and methods to overcome.

2. Write in detail various suppositories bases used in the formulation of suppositories.

Q. No. 3 Short Answers (Answer 2 out of 3)

1. Define pharmaceutical incompatibility. Write in short about physical incompatibility.

2. Discuss various emulsifying agents.

3. Write short note on (any one)

- a) Paste
- b) Gels

$(2 \times 5 = 10)$

 $(1 \times 10 = 10)$



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Second Sessional Theory Examination-(ODD SEM) 2023-2024

Subject		:	Pharmaceutical Inorganic Chemistry (BP104T)	Day & Date	: Thursday 21/12/2023		rsday 12/2023		
Cla	SS	:	First Year B. Pharmacy	Semester	:	Ι	Write Your Her	ır Roll No. ere	
Tin	ne	:	10:30 am – 12:00 pm	Max. Marks	:	30			
Instructions:1. All questions are compulsory2. Draw a well-labeled diagram wherever necessor3. Right-hand side number indicates full marks4. Do not write/tick on the question paper							ssary		
Q. I: Objective Type Questions							10 M		
1.	Defin	e tl	he following term. (CO3, CO4, LL3)						
	i) Em	eti	cs ii) alpha particle decay						
2.	Draw	the	e structure and uses of sodium potassium tartrate	e. (CO3, LL3)					
3.	3. Define hematinic with example. (CO3, LL3)								
4.	Write any four therapeutic application of radiopharmaceuticals. (CO4, LL3)								
5.	Defin	ie A	Antidote and enlist the types of antidote with exa	mple. (CO4, LL3)					
Q. II: Long Answers Questions (Answer any 1 out of 2)						10 M			
1. D	Define 1	rad	ioactivity and describe in detail methods for mea	surement of radioad	ctivity	у. (<i>CO</i> -	4, LL3)		
2. V	Vrite a	not	te on a) Heavy metal poisoning b) Astringent with	th one example. (Co	04, L	L3)			
Q. III: Short Answers Questions (Answer any 2 out of 3)						10 M			
1. Define and classify antimicrobial agents and describe M.W., M.F., properties, preparation and uses of hydrogen peroxide or boric acid. (CO3, LL3)									
2. D	2. Define and classify expectorants with example. (CO3, LL3)								
3. D	3. Describe and classify the cathartics based on their mechanism of action. (CO3, LL3)								

Prepared By

Checked By

Verified By



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Second Resessional 7 Subject: Pharmaceutics I BP103T	Theory Examination 2020-2021 Day & Date: Wednesday, 05 th May 2021
Class: FYB. Pharm	Semester: I
Time: 90 min	Max. Marks: 30
Instructions: 1. All questions are compulse 2. Draw a well labeled diagre 3. Right hand side number inc	ory am wherever necessary dicates full marks
Q.No.1 Multiple choice questions	(10 X 1=10)
1. Which of the following is not used as an em a. Surfactant b. Hydrophilic colloid c. Electrol	ulsifying agent? ytes d. Finely divided solids
2. Vaginal suppositories also called as a. Pessaries b. Simple suppositories c. Bougies	s d. None
3is stable form of Cocoa butter. a. α -form b. β -form c. γ -form d. All	
4. The tail of a surfactant molecule is oil-lovir a. Hydrophilic b. Lipophobic c. Lipophilic d. l	ng or known as? Hydrophilic
 5. In case of coalescence a. Dispersed droplet does not fuse b. Globules c. No of globules increase d. Dispersed drople 	size decrease ts tend to fuse
6. Upward creaming is observed in? a. W/O b. O/W c. Both d. Micro	
7. Which of the following is an example of Sya. Witespol b. Emulsified cocoa butterc. Glycero-gelatin base d. Hydrogenated oils	nthetic base?
8. In the preparation of a structured vehicle, while is used?	nich one of the following substances

.

a. bismuth subnitrate b. ethyl alcohol c. glycerin d. methyl cellulose

9. The value of sedimentation volume (F) gives a knowledge about the ______ of the suspension a. physical stability b. Incompatibility c. Solubility d. All of the above

10. In Dry gum method, 4:2:1 consist ofa. 4 parts oil, 2 parts water, and 1 part emulsifierb. 4 parts water, 2 parts oil, and 1 part emulsifierc. 4 parts emulsifier, 2 parts water, and 1 part emulsifierd. 4 parts oil, 2 parts emulsifier and 1 part water

Q. No. 2 Long Answers (Answer 1 out of 2)

1. Classify and discuss with advantages and disadvantages various semisolid bases used in the preparation of semisolid dosage form.

2. Define suppositories. Write down types and advantages and disadvantages of suppositories. Explain method of preparation.

Q. No. 3 Short Answers (Answer 2 out of 3)

1. Write down various test used for identification of emulsion.

2. Discuss physical incompatability.

3. Write short note on emulsifying agents.

 $(1 \times 10 = 10)$

 $(2 \times 5 = 10)$

* * ^{**} *



Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai-Agra National Highway

Second Sessional Theory Examination 2020-2021

Subject: Comm. Skills	Day & Date: Friday, 23 April 2021
Class: FY/ DSA	Semester: First
Time: 10.30 to 12.00	Max. Marks: 30
Time: Topo to Theor	

Instructions: 1. All questions are compulsory

2. Draw a well labeled diagram wherever necessary

3. Right hand side number indicates full marks

QI Long Answers (Answer any ONE)

- (1) Explain in detail how to make presentation effective w.r.t. four P's of effective presentation and how to deal with fear and structure the presentation. (Unit 04)
- (2) How to make Writing and Listening more effective? Enumerate the techniques and methods required to make Writing and Listening more productive. (Unit 03)

QII Short Answers (Answer any FOUR)

- 1. Explain the Selection GD in detail with all its characteristics. (Unit 05)
- 2. What is Listening? Explain its sub-types with their purpose in detail. (Unit 03)
- 3. Explain the role of Team member and Leader wrt GD. (Unit 05)
- 4. What strategies will you adopt to contribute in Group Interaction? (Unit 02)
- How will you prepare for Selection Interview? Explain the Interview process in detail. (Unit 04)
- 6. What are the Dos and Don'ts of Group Discussion? (Unit 05)

-----END-----END------



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Institute of Pharmacy, Dhule Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

Second Sessional Theory Examination 2018-2019

Subject: Comm. Skills	Day & Date: Thursday, 22 Nov. 2018
Class: FY/ DSA	Semester: First
Time: 10.30 to 12.00	Max. Marks: 30

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks

QI. Long Answers (Answer any 1 out of 2) 10

 What is Group Discussion? State its objective, purpose & importance in selection procedure.

20

(2) Explain effective Listening & Writing skills in detail.

QII. Short Answers (Answer any 4 out of 6)

- 1. Explain Do's and Don'ts of Interviews.
- 2. What are the 4Ps of Presentation? Explain them in detail.
- 3. Explain the role of 'know your audience' in Presentation & Writing effectively.
- 4. Explain the listening subtypes with techniques to improve Listening.
- 5. What is 'writing effectively'? Explain the elements involved in it.
- 6. State the importance of Body Language & Verbal Comm. in Interviews.



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Second Sessional Theory Examination 2019-2020

Subject: Comm. Skills	Day & Date: Saturday, 09 Nov. 2019
Class: FY/ DSA	Semester: First
Time: 10.30 am to 12.00 pm	Max. Marks: 30
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Instructions: 1. All questions are compulsory

2. Draw a well labeled diagram wherever necessary

3. Right hand side number indicates full marks

QI. Long Answers (Answer any ONE)

- (1) What is Group Discussion? State its objective, purpose & importance in selection procedure.
- (2) Explain Dr Eileen Russo's Communication Style Matrix in detail with different Communication Styles.

QII. Short Answers (Answer any FOUR)

- 1. Explain Do's and Don'ts of Interviews.
- 2. What are the 4Ps of Presentation? Explain them in detail.
- 3. Explain the role of 'know your audience' in Presentation & Writing effectively.
- 4. Explain the listening subtypes with techniques to improve Listening.
- 5. What are the Perspectives? Explain their importance in Communication.
- 6. State the importance of Body Language & Verbal Comm. in Interviews.

****BEST OF LUCK****

1*10

4*5



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Subje	Second Sessional Theory Examination 2021-2022 Subject: Communication Skills (BP105T) Day & Date: Thursday, 07 April 2022							
Class:	First Year and Direct Scond Year B.Pharm	Semester: I & III						
Time:	02.00PM to 03.00PM	Max. Marks: 30						
Instruc	ctions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever nece 3. Right hand side number indicates full marks	ssary						
QI Lo	ng Answers (Answer any ONE) (CO3)	1*10						
(1) Explain in detail how to make presentation effect (Unit 04)	ctive w.r.t. four P's of effective presentation.						
(2	2) Enumerate the factors that assist to prepare and	appear for Interview. (Unit 04)						
QII SI	hort Answer Questions							
	Section A: (Any One) (CO3)	1*5						
1. 2.	Explain the role on non-verbals in Selection Internation How will you deal with performance anxiety due 04)	erview. (Unit 04) Iring Interviews and Presentation? (Unit						
	Section B: (Any One) (CO4)	3*5						
1.	Explain the Selection GD in detail with all its cl	haracteristics. (Unit 05)						
2.	Explain the role of Team member and Leader w	rrt GD. (Unit 05)						
3.	What strategies will you adopt to contribute in 0	Group Interaction? (Unit 05)						
4.	What are the Dos and Don'ts of Group Discussion	on? (Unit 05)						
	END							



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Mapping of Course Outcome with First Sessional Theory Examination (AY 2021-22)

Subject: Communication Skills

Course Outcome	Questions	Marks
CO 1053: Identify interview-process and effective presentation. (<i>Level: 1</i>)	QI: 1, 2, 3, 4	15
CO 1054: Recognize contribution in Group Discussion. <i>(Level: 1)</i>	QII: 1, 2, 3, 4	15

Prepared by

Verified by

Approved by



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Second Sessional Theory Examination-(ODD SEM) 2022-2023

Subject	:	Communication skills (BP105T)	Day & Date	:	Frid	ay, 2	21/04/2023
Class		First Year B. Pharmacy & DSA	Semester	:	Ι	Writ	te Your Seat No. Here
Time	••	10:30 am – 11:30 am	Max. Marks	:	30		
Instructions	1. 3.	All questions are compulsory Right hand side number indicates full marks	2. Draw a well labele 4. Do not write/tick o	ed d n th	liagran e ques	ı whe tion p	rever necessary aper

QI Long Answers Question (Answer any 1 out of 2) 10 M Illustrate how to deliver and design an effective presentation. (LL3, CO3) 1. 2. Explain interview process with your ways to prepare and appear of an interview. (LL3, CO3) **Short Answers Question** QII Section – A (Answer any ONE) 05 M Illustrate Do's and Don'ts of an interview. (LL3, CO3) 1. Interpret your ways to cope-up with stage fear and anxiety during presentation. (LL3, CO3) 2. Section – A (Answer any Three) 15 M 1. Interpret the role of Non-verbal communication during Group Discussion. (LL3, CO4) Explain your methods of contribution in Group Discussion. (LL3, CO4) 2. Illustrate Do's and Don'ts of Group Discussion. (LL3, CO4) 3. Summarize the significance of tone, voice and language in Group Discussion. (LL3, CO4) 4.



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Second Sessional Theory Examination-(ODD SEM) 2023-2024

Subject		:	Communication skills (BP105T)	Day & Date	:	FRI	DAY,	22/12/2	023
Class		:	First Year B. Pharmacy	Semester	:	Ι	Write	e Your Seat I	lo. Here
Time		:	10:30 a.m. – 11:30 a.m.	Max. Marks	:	30			
Instruc	nstructions 1. All questions are compulsory2. Draw a well labeled diagram wherever necessary3. Right hand side number indicates full marks4. Do not write/tick on the question paper								
QI	Long Answers Question (Answer any 1 out of 2)10 M								10 M
1.	Def	fine	e self-awareness. Explain in detail the ways of becc	oming an active	liste	ner. (l	LL2, (CO4)	
2.	Sta	te t	he purpose of an interview and enlist the do's and	don'ts of facing	the	interv	iew. (LL2, CO4))
QII	Sho	rt A	Answers Question						
	Section – A (Answer any ONE) 05 M							05 M	
	Def	fend	d the following sentence: Knowing your audience/	readers is of pr	ime	impor	tance	in order to	make
1.	you	ır w	vriting effective. (LL2, CO4)						
2.	Sug	gge	st the remedies of dealing with stage fright/ perform	mance anxiety.	(LL2	2, CO4	4)		
QII	Sec	tion	a – B (Answer any Three)						15 M
1.	Write a short note on listening in difficult situations. (LL2, CO3)								
2.	Explain in brief about the techniques of presentation delivery. (LL2, CO3)								
3.	Enlist the tips for framing a good e-mail subject line. (LL2, CO3)								
4.	Illustrate the concept of group discussion. State do's and don'ts of GD (two each). (LL3, CO3)								



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Second Sessional Theory Examination 2020-2021 (Odd SEM)

	X	/						
Subject & Subject Code: Human Anatomy & Physiology I (BP101T)								
4Day & Date: Monday, 26.04.2021 Class: First Year B. Pharm								
Semester: I	Time: 10.30 AM t	to 12.30PM	Max. Marks: 30					
Instructions: 1. All questions are compulsory.								
2. Draw	a well labelled diagram	m wherever necessary.						
3. Right	hand side number indi	cates full marks.						
Q.I Multiple-choice	questions (MCQs)		10					
1. The temperature of	f blood is							
a. 35 ⁰ C	b. 36 ⁰ C	c. 37 ⁰ C	d. 38 ⁰ C					
2. Blood is about	% formed elements an	nd % blood plasma	respectively					
a. 75 & 25	b. 50 & 50	c. 45 & 55	d. 60 & 40					
3 contributes	to higher hematocrits	in males						
a. Testosterone	b. Estrogen	c. Progesterone	d. Prolactin					
4. About 90 per cent	of all erythropoietin is	formed in the						
a. Liver	b. Kidney	c. Lungs	d. Heart					
5. Tissues that lack ly	ymphatic capillaries ind	clude avascular tissues	such as					
a. Liver	b. Kidney	c. Red bone marrow	d. Lungs					
6. The superior vena	cava enters the	of the heart.						
a. Right ventricle	b. Right Atrium	c. Left atrium	d. Left ventricle					



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7. Which of the following contains blood with a relatively low oxygen content?

a. Pulmonary artery	b. Pulmonary vein	c. Left ventricle	d. Aorta
5 5	2		

8. The time taken for an impulse to travel from the S-A to the A-V node is evidenced in the ---

a. QRS complex b. S-T interval c. P-Q interval d. QRS-T interval

9. The amount of blood ejected by the left ventricle each minute is calculated from which formula?

a. Heart rate + stroke volume b. Heart rate X stroke volume

c. Heart rate divided by stroke volume d. Heart rate - stroke volume

10. In an ECG, where does the depolarization of the atria occur?

a. P wave	b. QRS complex	c. T wave	d. U wave
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Q.II Long Answers (Answer any 1 out of 2)

1. Draw a neat well labelled diagram of heart. Explain the structure and function of heart.

2. Write the functions of blood. Explain about the ABO system of blood group.

Q.III Short Answers (Answers any 2 out of 3)	
--	--

10

10

1. Differentiate between parasympathetic and sympathetic nervous system

2. Write the location and functions of cranial nerves.

3. Draw a neat well labelled diagram of spinal cord. Explain about mechanism of reflex arc.



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Second Sessional Theory Exam	ination 2018-2019
Subject: Human Anatomy and Physiology-	I
Day & Date:	
Class: F. Y. B. Pharma	Semester: I
Time: 11:00-12:30	Max. Marks: 30
Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram whereve 3. Right hand side number indicates full t	er necessary marks
QI Write a short note on (Any five)	10
1. Classify WBCs	
2. Write a note on heart valves	n na tan
3. Define the following term a) Blood pr	ressure b) Heart rate
4. Write the composition of blood	
5. Differentiate between artery and veir	n
6. Write a note on various types of ane	mia
OII Long Answers (Answer any 1 out of	2) 10
1. Explain in detail ABO and RH blood	grouping and its
significance. Enlist various coagulat	tion factors.
2. Draw a neat labelled diagram of inte	ernal structure of eye. Add
a note on physiology of vision	0
QIII Short Answers (Answer any 2 out	t of 3) 10
1. Write a note on ECG with neat labe	elled diagram
2. Explain blood coagulation pathway	7.
3.Write the anatomy and physiology of	of RBCs.
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Second Sessional Theory Examination 2020-2021 (Odd SEM)

Subject & Subject Code: Human Anatomy & Physiology I (BP101T)			
4Day & Date: Monday, 26.04.2021 Class: First Year B. Pharm			
Semester: I	Time: 10.30 AM 1	to 12.30PM	Max. Marks: 30
Instructions: 1. All q	uestions are compulsor	y.	
2. Draw	v a well labelled diagra	m wherever necessary.	
3. Right	hand side number indi	cates full marks.	
Q.I Multiple-choice	questions (MCQs)		10
1. The temperature o	f blood is		
a. 35 ⁰ C	b. 36 ⁰ C	c. 37 ⁰ C	d. 38 ⁰ C
2. Blood is about	% formed elements an	nd % blood plasma	respectively
a. 75 & 25	b. 50 & 50	c. 45 & 55	d. 60 & 40
3 contributes	to higher hematocrits	in males	
a. Testosterone	b. Estrogen	c. Progesterone	d. Prolactin
4. About 90 per cent	of all erythropoietin is	formed in the	
a. Liver	b. Kidney	c. Lungs	d. Heart
5. Tissues that lack l	ymphatic capillaries in	clude avascular tissues	such as
a. Liver	b. Kidney	c. Red bone marrow	d. Lungs
6. The superior vena	cava enters the	of the heart.	
a. Right ventricle	b. Right Atrium	c. Left atrium	d. Left ventricle
R	- >> 0		

mr.S.Lerikar A.K.

(Subject Tn-charge)



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7. Which of the following contains blood with a relatively low oxygen content?

a. Pulmonary artery b. Pulmonary vein c. Left ventricle d. Aorta

8. The time taken for an impulse to travel from the S-A to the A-V node is evidenced in the ---

a. QRS complex b. S-T interval c. P-Q interval d. QRS-T interval

9. The amount of blood ejected by the left ventricle each minute is calculated from which formula?

a. Heart rate + stroke volume b. Heart rate X stroke volume

c. Heart rate divided by stroke volume

d. Heart rate - stroke volume

10. In an ECG, where does the depolarization of the atria occur?

a. P wave b. QRS complex c. T wave d. U wave

Q.II Long Answers (Answer any 1 out of 2)

1. Draw a neat well labelled diagram of heart. Explain the structure and function of heart.

2. Write the functions of blood. Explain about the ABO system of blood group.

Q.III Short Answers (Answers any 2 out of 3)

10

10

1. Differentiate between parasympathetic and sympathetic nervous system

2. Write the location and functions of cranial nerves.

3. Draw a neat well labelled diagram of spinal cord. Explain about mechanism of reflex arc.

Mr. Sherikar A-K. (Subject In-charge) *** END***



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Second Sessional Theory Examination 2020-2021 (Odd SEM)

Subject & Subject Code: Human Anatomy & Physiology I (BP101T)			
4Day & Date: Monday, 26.04.2021 Class: First Year B. Pharm			
Semester: I	Time: 10.30 AM 1	to 12.30PM	Max. Marks: 30
Instructions: 1. All q	uestions are compulsor	y.	
2. Draw	v a well labelled diagra	m wherever necessary.	
3. Right	hand side number indi	cates full marks.	
Q.I Multiple-choice	questions (MCQs)		10
1. The temperature o	f blood is		
a. 35 ⁰ C	b. 36 ⁰ C	c. 37 ⁰ C	d. 38 ⁰ C
2. Blood is about	% formed elements an	nd % blood plasma	respectively
a. 75 & 25	b. 50 & 50	c. 45 & 55	d. 60 & 40
3 contributes	to higher hematocrits	in males	
a. Testosterone	b. Estrogen	c. Progesterone	d. Prolactin
4. About 90 per cent	of all erythropoietin is	formed in the	
a. Liver	b. Kidney	c. Lungs	d. Heart
5. Tissues that lack l	ymphatic capillaries in	clude avascular tissues	such as
a. Liver	b. Kidney	c. Red bone marrow	d. Lungs
6. The superior vena	cava enters the	of the heart.	
a. Right ventricle	b. Right Atrium	c. Left atrium	d. Left ventricle
R	- >> 0		

mr.S.Lerikar A.K.

(Subject Tn-charge)



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7. Which of the following contains blood with a relatively low oxygen content?

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a. QRS complex b. S-T interval c. P-Q interval d. QRS-T interval

9. The amount of blood ejected by the left ventricle each minute is calculated from which formula?

a. Heart rate + stroke volume b. Heart rate X stroke volume

c. Heart rate divided by stroke volume

d. Heart rate - stroke volume

10. In an ECG, where does the depolarization of the atria occur?

a. P wave b. QRS complex c. T wave d. U wave

Q.II Long Answers (Answer any 1 out of 2)

1. Draw a neat well labelled diagram of heart. Explain the structure and function of heart.

2. Write the functions of blood. Explain about the ABO system of blood group.

Q.III Short Answers (Answers any 2 out of 3)

10

10

1. Differentiate between parasympathetic and sympathetic nervous system

2. Write the location and functions of cranial nerves.

3. Draw a neat well labelled diagram of spinal cord. Explain about mechanism of reflex arc.

Mr. Sherikar A-K. (Subject In-charge) *** END***



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Second Sessional Theory Examination 2021-2022 (Odd SEM)

Subject & Subject Code: Human Anatomy & Physiology I (BP101T)Day & Date: Monday, 4.04.2022Class: First Year B. PharmSemester: ITime: 2:00PM to 3.30PMMax. Marks: 30Instructions: 1. All questions are compulsory.

- 2. Draw a well labelled diagram wherever necessary.
- 3. Right hand side number indicates full marks.

Q.I Multiple-choice questions (MCQs)

10

1.	The components first line defence innate immunity is					
	a. Fever	b. NK	cells	c. Inflammation	d. skir	1
2.	In skin, the	lymphatic	vessels lie in	tissue		
	a. Subcutan	eous	b. Epidermis	c. Dermis		d. Keratin
3.	The membr	ane that su	rrounds and pr	otects the heart is the	e –	
	a. Myocard	ium	b. Endocardium	m c. Epicardiu	m	d. Pericardium
4.	is also ca	alled the ri	ght atrioventric	ular valve		
	a. Bicuspid		b. Tricuspid	c. Mitral		d. Semilunar valve
5.	The surf	ace of hear	rt is deep to the	sternum and ribs		
	a. Inferior		b. Anterior	c. Right bor	der	d. Left border
6.	All somatic	motor neu	rons release or	nly as their neuro	otransmit	ter
	a. Norepine	phrine	b. Epinephrine	c. Acetylcho	oline	d. Both a and b
7.	The cell bo	dies of pre	ganglionic neu	rons are located in	-	
	a. Brain	b. Auto	onomic ganglia	c. Visceral organs	d. Ver	tebral column
8.	is the la	rgest autor	nomic plexus			
	a. Superior	mesenteric	b. Infer	rior mesenteric	c. Ren	al d. Celiac
9.	There are	pairs of	spinal nerves			
	a. 12	b. 21	c. 31	d. 13		
10.	Somatic ner	rvous syste	m consists of -	axons		
	a. Single &	myelinate	d	b. Multiple and my	elinated	
	c. Single &	unmyelina	ated	d. Multiple & unmy	elinated	

QII Long Answers (Answer any 1 out of 2)

10

10

- **1.** Explain the functions of lymphatic system. Explain the structure and functions of lymph nodes.
- 2. Draw a neat labelled diagram of heart. Explain the anatomical structure of heart.

QIII Short Answers (Answer any 2 out of 3)

- 1. Differentiate between sympathetic and parasympathetic nervous system.
- 2. Write the number, name, location and function of cranial nerves.
- 3. Draw a neat labelled diagram of eye. Explain the physiology of vision.



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Second Sessional Theory Examination 2021-2022 (Odd SEM)

Subject & Subject Code: Human Anatomy & Physiology I (BP101T)Day & Date: Monday, 4.04.2022Class: First Year B. PharmSemester: ITime: 2:00PM to 3.30PMMax. Marks: 30Instructions: 1. All questions are compulsory.

- 2. Draw a well labelled diagram wherever necessary.
- 3. Right hand side number indicates full marks.

Q.I Multiple-choice questions (MCQs)

10

1.	The components first line defence innate immunity is					
	a. Fever	b. NK	cells	c. Inflammation	d. skir	1
2.	In skin, the	lymphatic	vessels lie in	tissue		
	a. Subcutan	eous	b. Epidermis	c. Dermis		d. Keratin
3.	The membr	ane that su	rrounds and pr	otects the heart is the	e –	
	a. Myocard	ium	b. Endocardium	m c. Epicardiu	m	d. Pericardium
4.	is also ca	alled the ri	ght atrioventric	ular valve		
	a. Bicuspid		b. Tricuspid	c. Mitral		d. Semilunar valve
5.	The surf	ace of hear	rt is deep to the	sternum and ribs		
	a. Inferior		b. Anterior	c. Right bor	der	d. Left border
6.	All somatic	motor neu	rons release or	nly as their neuro	otransmit	ter
	a. Norepine	phrine	b. Epinephrine	c. Acetylcho	oline	d. Both a and b
7.	The cell bo	dies of pre	ganglionic neu	rons are located in	-	
	a. Brain	b. Auto	onomic ganglia	c. Visceral organs	d. Ver	tebral column
8.	is the la	rgest autor	nomic plexus			
	a. Superior	mesenteric	b. Infer	rior mesenteric	c. Ren	al d. Celiac
9.	There are	pairs of	spinal nerves			
	a. 12	b. 21	c. 31	d. 13		
10.	Somatic ner	rvous syste	m consists of -	axons		
	a. Single &	myelinate	d	b. Multiple and my	elinated	
	c. Single &	unmyelina	ated	d. Multiple & unmy	elinated	

QII Long Answers (Answer any 1 out of 2)

10

10

- **1.** Explain the functions of lymphatic system. Explain the structure and functions of lymph nodes.
- 2. Draw a neat labelled diagram of heart. Explain the anatomical structure of heart.

QIII Short Answers (Answer any 2 out of 3)

- 1. Differentiate between sympathetic and parasympathetic nervous system.
- 2. Write the number, name, location and function of cranial nerves.
- 3. Draw a neat labelled diagram of eye. Explain the physiology of vision.



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Second Sessional Theory Examination-(ODD SEM) 2022-2023 Human Anatomy & Physiology - I Day & Date : Monday, 17/04/2023 Subject : (BP101T) Write Your Seat No. Here Class First Year B. Pharmacy Semester Ι : :

10.30 am to 12.00 pm Max. Marks 30 Time **Instructions:** 1. All questions are compulsory 2. Draw a well-labeled diagram wherever necessary 3. Right-hand side number indicates full marks 4. Do not write/tick on the question paper

- **O. I:** Objective Type Questions (5 X 2)
- 1. Write about structure of lymph node. (CO3, LL3)
- 2. Draw a neat and well labelled diagram of interior of heart. (CO3, LL3)
- 3. Write location and functions of heart valves. (CO3, LL3)
- 4. Draw a neat and well labelled diagram of eye. (CO2, LL2)
- 5. Differentiate between somatic and autonomic motor neurons. (CO2. LL2)

Q. II: Long Answers Questions (Answer any 1 out of 2)

Write in detail about name, number, nature, location and functions of cranial nerves. (CO2, LL2)

Write in detail about comparison of sympathetic and parasympathetic divisions of autonomic 2. nervous system. (CO2, LL2)

Q. III: Short Answers Questions (Answer any 2 out of 3) **10 M** Explain the composition, formation and circulation of lymph. (CO3, LL3) 1. Write about electrocardiogram. (CO3, LL3) 2. Explain about cardiac cycle. (CO3, LL3) 3.

10 M

10 M



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			Second Sessional Theory Examinatio	n - (Odd S	EM	i) 20	23-2024
Sul	oject	:	Human Anatomy & Physiology - I (BP101T)	Day & Date	:	Moi	nday, 18/12/2023
Cla Tin	lss ne	:	First Year B. Pharmacy 10.30 am to 12.00 pm	Semester Max. Marks	:	I 30	Write Your Seat No. Here
Instr	uction	s: 1 3	. All questions are compulsory 2. D . Right-hand side number indicates full marks 4. D	raw a well-labei o not write/tick	led c on ti	liagrar he que	n wherever necessary estion paper
Q. I	: Objec	tive	e Type Questions (5 X 2)				10 M
1.	Write	the	functions of lymphatic system (CO4, LL2)				
 2. 3. 4. 	 Draw diagram of spleen. Enlist the functions of spleen (CO4, LL2) Answer the followings a. Enlist the phases of cardiac muscle action potential (CO4, LL2) b. Enlist the functions of ear (CO3, LL2) 						
5.	Draw	a ne	eat and well-labelled diagram of eye (CO3. LL2)				
Q. I	I: Long	Ar	nswers Questions (Answer any 1 out of 2)				10 M
1.	Draw	a ne	eat and well-labelled diagram of heart. Explain in detail abou	t circulatory sys	tem	of hea	rt (<i>CO4</i> , <i>LL2</i>)
2.	Explai	n al	bout electrocardiogram and cardiac cycle (CO4, LL2)				
Q. I	II: Sho	rt A	Answers Questions (Answer any 2 out of 3)				10 M
1.	Differ	enti	ate between sympathetic and parasympathetic nervous system	m. (<i>CO3</i> , <i>LL2</i>)			
2.	2. Write the name, number, location and functions of cranial nerves. (CO3, LL2)						
3.	3. Draw neat and well-labelled diagram of ear. Explain the mechanism of hearing. (CO3, LL2)						



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	Second Sessional Theory Ex	xamination 2019-202	20
Subj	ect : Pharmaceutical Analysis (BP102T)	Day & Date :	13/11/2019
Class	F. Y. B. Pharmacy	Semester :	I
Time	: 10:30 am to 12:00 noon	Max. Marks :	30
	Instructions: 1. All questions are compulsor	Ŋ	
	2. Draw a well labelled diagra	m wherever necessary	
	3. Right hand side number indi	cates full marks	
Q.1	Multiple Choice Questions		[10]
а	In polarographic method DME is used as	electrode	
	a. reference	b indicator	
·1	c. partial reference	d. None of the above	
b	Standarad hydrogen electrode (SHE / NHE	E) is electrode.	
	a. Indicator	b. Reference	
	c. Second Kind	d. None of the above	
C	Nearnist equation can be used to calculate	h. Desistance	
	a. Conductance	D. Resistance	1
Ь	In titration solicylic sold is used as	a. Residual current	
u	n utration salicylic acid is used as	a specific indicator	io
	a. Action	d precipitation	IC .
e	Potentiometry can be utilized in below type	e of reactions except	
	a. Neutralization	b Redox	
	c. Precipitation	d Gravimetric analysis	
f	is not a self-indicator		
	a. Potassium permanganate	b. Ceric ammonium sul	fate
	c. Iodine solution	d. Potassium dichromat	e
g	In gravimetric analysis the term 'digestion'	carried out to obtained	
	a. pure precipitate	b. small crystals	6
	c. impure precipitate	d. a & c both	
h	are involved in Bromatometr	y type of titrations	
	a. Br	b. BrO ₄	
	c. Br_2^{\uparrow}	d. All of the above	
1	Resistance can be expressed as		
	a. S	b. Ω ⁻¹	
1		d. a & b both	11 1
J	The difference between the residual curren	it and limiting current is c	alled as
		h maidual	
	a. lingiation	d h & a both	
	c. kilictic		
0.2.	Answer the following short questions (any	two)	[10]
a	Write a short not on masking and demask	ing agents. Enlist the facto	rs
	affecting selectivity of EDTA.		
b	Describe the construction, working and ap	plication of DME.	
С	Explain principle involved in various types	of iodine titration.	
d	Explain the principle and application of dia	azotization titration	
Q.3.	Answer in detail of following (any one)		[10]
а	Explain the term specific and molar conduc	<i>ctance.</i> Describe in detail a	bout
	various conductometric titration with suita	able example.	
b	Discuss the steps involved in gravimetric a	nalysis? Comment on the	
	unierences between lyophilic and lyophobic	c sols obtained during	
0	Function the principle behind not on the principle behind	Write a construction	1
U	working of reference electrode and indicate	relectrode	L
	*****	JI CICCUOUC	



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	Second Sessional Theory Examination 2020-2021					
Subje	ect: Pharmaceutical Analysis I (BP102T)	Day & Date: 22-04-2021				
Class	First Year B. Pharm	Semester: I				
Time	: 10:30 am – 12:00 pm	Max. Marks : 30				
Instruc	tions: 1. All questions are compulsory					
	2. Draw a well labeled diagram wherever necessa	ry				
	3. Right hand side number indicates full marks					
QI. M	ultiple Choice Questions (MCQs)	10				
1. Clo	oseness of observed value with true value is					
a)	accuracy					
b)	linearity					
c)	precision					
d)	relative error					
2. Give	en the following are the type of systemic error except					
a)]	Error of methods					
b)	Instrumental method					
c)]	Personal error					
d)	Random error					
3. In w	hich step, ions or elements are aggregated in Gravimetric	e analysis?				
a)	Supersaturation					
b)	Nucleation					
c)	Particle growth					
d)	None of the above					
4. Whe	en more than two ions in solution and simultaneously second	ond ion is also precipitated it is called				
as						
a)	Co precipitation					
b)	Post precipitation					
c)	A and B					
d)	None of the above					
5. The	reagent which undergoes reduction is an age	ent and the reagent which undergoes				
oxidati	on is a					
-)	Oridining Deducing					

- a) Oxidizing, Reducing
- b) Reducing, Oxidizing
- c) Complexing, Reducing
- d) None of these
- 6. Which method are used to determination of primary amine?



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- a) Diazotization Titration
- b) Karl fischer titration
- c) Redox titration
- d) All of the above
- 7. Indicator used in standardization of sodium thiosulphate is
 - a) Starch
 - b) Phenolpthaline
 - c) ceric sulphate
 - d) Methyle red

8. When an analyte that is a reducing agent is titrated directly with a standard iodine solution, the

method is called

- a) iodimetry
- b) iodometry
- c) bromatometry
- d) ceriometry
- 9. Complexing agent is
 - a) Electron donating ion
 - b) Electron accepting ion
 - c) Both a & b
 - d) None of the above

10 The diffusion current in polarography is expressed by the equation

- a) Beer's law
- b) Nernst equation
- c) Bragg's equation
- d) Ilkovik equation

QII. Solve the following (any 1)

1. Define gravimetric analysis. Explain in detail the steps involved in gravimetric analysis

2. State the basic principle involved in polarography. Write down the construction and working of dropping mercury electrode

QIII Short Answers (Answer any 2 out of 3)

- 1. Explain the titration involving iodine
- 2. Write down the construction and working of normal hydrogen electrode
- 3. Write principle and application of Diazotization titrations.

10



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	Second Sessional Theory Ex	kamination 2018-2	01	9	
Subje	ct : Pharmaceutical Analysis	Day & Date	:	17/11/	2018
Class	: F. Y. B. Pharmacy	Semester	:	I	
Time	: 10:30 am to 12:00 noon	Max. Marks	:	30	
	Instructions: 1. All questions are compuls	sory			
	2. Draw a well labelled diagra	am wherever necessary			
0.1	3. Right hand side number ind	licates full marks			[10]
Q.1	Multiple Choice Questions				[10]
a	Indometric titration utilizes	h Liberated indin	9		
	a. Iounic solution c Both a & b	d None of the abo	ve		
h	Hydrogen electrode is electrode	d. None of the abe	VC		
D.	a Indicator	b. Reference			
	c. Second kind	d. None of the abo	ve		
с	Mordant Black II mixture is used in	titration			
	a. Precipitation	b. Neutralization			
	c. Redox	d. Complexometrie	С		
d	EDTA is type of ligand				
	a. Unidentate	b. Bidentate			
	c. Tetradendentate	d. Hexadenta	te		
e	DME stands for	· · ·		. <u>-</u>	
	a. Diffusion mercury	b. Dropping mercury	ele	ctrode	
	c. Diffusion measuring	d. Direct mercury			
1	is self-indicator		10	ia nas	
	a. Potassium permanganate	b. Ceric ammonium	sulla	ate	
~	C. Iodine solution	d. a & D Doth	toto		
g	in gravinetric analysis the term ignition i	b 250-500°	laic		
	a. $250-1200^{\circ}$	d b & c both			
h	are involved in Bromatomet	ry type of titrations			
	a. Br	b. BrO_4			
	c. Br_2^{\uparrow}	d. All of the above			
i	Conductance can be expressed as				
	a. S	b. Ω ⁻¹			
	c. Ω	d. a & b both			
j	Diffusion current (<i>id</i>) is defined as	·			
	a. ii	b. ir			
	C. İl-İr	d. im			
					(10)
Q.2.	Answer the following short questions (any	two)	1-1-		[10]
а	What do you mean by permanganate titra	tions explain with suita	DIE		
h	Write about note on principle of polargraph	h			
D	write short note on principle of polargraph	ny			
с	Write down the construction and working	of conductivity meter			
0.2					[10]
Q.3.	Answer in detail of following (any one)	stance Describe in data	J1 ~	hout	[10]
а	explain the term specific and molar conau	able example	шa	bout	
h	Write down the construction and working	of dronning mercury el	ectr	ode	
U	involved in polarography.	or aropping mercury cit		ouc	

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S C T	ubje lass ime	 Second Sessional Theory Exc Pharmaceutical Analysis-I F. Y. B. Pharmacy 10:30 am to 12:00 noon 	amination 2018- Day & Date Semester Max. Marks	2019 : 17/1 : I : 30	1/2018
		Instructions: 1. All questions are compuls	ory		
		2. Draw a well labelled diagra	m wherever necessary	J	
		3. Right hand side number ind	icates full marks		[10]
Ç).1	Multiple Choice Questions			[10]
,	a	Iodometric titration utilizes			
b		a. Iodine solution	b. Liberated lodi	ne	
	н.,	c. Both a & b	d. None of the at	oove	
	b	Standarad hydrogen electrode (SHE / NHE) is electrode.		
h		a. Indicator	d None of the of		
/		C. Second Kind	titration	1000	
x	С	Mordant Black II mixture is used in	unanon b Neutralization		
d		a. Precipitation	d Complexomet	ric	
	А	FDTA is type of ligand	a. complexonice.		
	u	a unidentate	b. bidentate		
C		c tetradendentate	d. hexadenta	ate	
,	е	Potentiometric titrations can be utilizes in	type of react	ions.	
	Ū	a. Neutralization	b. Redox		
d-		c. Precipitation	d. All of the above		
	f	is self-indicator			
		a. Potassium permanganate	Ceric ammonium	ı sulfate	
d		c. Iodine solution	d. a & b both		
	g	In gravimetric analysis the term 'drying' re	fers to heating precipi	itate	
Ci.		a. upto 250°	b. 250-500°		
q		c. 250-1200°	d. b & c both		
X	h	are involved in Bromatometr	y type of titrations		
0		a. Br	b. BrO ₄		
		c. Br_2^{\uparrow}	d. All of the above		
1	1	Conductance can be expressed as	1 01		
0		a. S	$D. \Omega^{-1}$		
	:	C. Ω			
	J	Diffusion current (a) is defined as			
C			d im		
		V li-ir	a. in		
C).2.	Answer the following short questions (any	two)		[10]
	a.	Explain the principle involved in iodimetri	c and iodometric titra	tions? Why	
	u	starch indicator added towards the end po	vint only in iodometric	and not in	
		indimention	fint only in louomouro	unu not m	
	1	logimetric utration.			
	D	write short note on principle of polargraph	iy		
	с	Explain the various types of EDTA fitration	is.		
	d	Define any five of the following:			
		i. Precipitation, ii. Co-precipitation, iii. Pos	t precipitation, iv. Tyn	dall effect.	
		v. Peptisation, vi. Coagulation value			
		international and a second sec			
(Q.3.	Answer in detail of following (any one)			[10]

a Explain the term *specific* and *molar conductance*. Describe in detail about various conductometric titration with suitable example.

b Discuss the steps involved in gravimetric analysis? Comment on the differences between lyophilic and lyophobic sols obtained during precipitation



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Second Sessional Theory Examination 2020-2021 (Odd SEM)

Subject: Pharmaceutical Analysis I (BP102T)	Day & Date: 22-04-2021
Class: First Year B. Pharm	Semester: I
Time: 10:30 am – 12:00 pm	Max. Marks : 30

Instructions: 1. All questions are compulsory

2. Draw a well labeled diagram wherever necessary

3. Right hand side number indicates full marks

QI. Multiple Choice Questions (MCQs)

- 1. Closeness of observed value with true value is
 - a) accuracy
 - b) linearity
 - c) precision
 - d) relative error

2. Given the following are the type of systemic error except

- a) Error of methods
- b) Instrumental method
- c) Personal error
- d) Random error

3. In which step, ions or elements are aggregated in Gravimetric analysis?

- a) Supersaturation
- b) Nucleation
- c) Particle growth
- d) None of the above

4. When more than two ions in solution and simultaneously second ion is also precipitated it is called

as

- a) Co precipitation
- b) Post precipitation
- c) A and B
- d) None of the above

5. The reagent which undergoes reduction is an agent and the reagent which undergoes oxidation is a

- a) Oxidizing, Reducing
- b) Reducing, Oxidizing
- c) Complexing, Reducing
- d) None of these

6. Which method are used to determination of primary amine?



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- a) Diazotization Titration
- b) Karl fischer titration
- c) Redox titration
- d) All of the above

7. Indicator used in standardization of sodium thiosulphate is

- a) Starch
- b) Phenolpthaline
- c) ceric sulphate
- d) Methyle red

8. When an analyte that is a reducing agent is titrated directly with a standard iodine solution, the

method is called

- a) iodimetry
- b) iodometry
- c) bromatometry
- d) ceriometry

9. Complexing agent is

- a) Electron donating ion
- b) Electron accepting ion
- c) Both a & b
- d) None of the above

10 The diffusion current in polarography is expressed by the equation

- a) Beer's law
- b) Nernst equation
- c) Bragg's equation
- d) Ilkovik equation

QII. Solve the following (any 1)

1. Define gravimetric analysis. Explain in detail the steps involved in gravimetric analysis

2. State the basic principle involved in polarography. Write down the construction and working of dropping mercury electrode

QIII Short Answers (Answer any 2 out of 3)

- 1. Explain the titration involving iodine
- 2. Write down the construction and working of normal hydrogen electrode
- 3. Write principle and application of Diazotization titrations.

10



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Second Sessional Theory Examination 2021-2022 (Odd SEM)

Subject: Pharmaceutical Analysis-I (BP102-T)	Day & Date: Friday 01/04/2022
Class: F.Y. B. Pharm	Semester: I
Time: 02.00 pm – 03.30 pm	Max. Marks : 30

Instructions: 1. All questions are compulsory

2. Draw a well labelled diagram wherever necessary

3. Right hand side number indicates full marks

QI Multiple Choice Questions (MCQs)

- 1. is a chelating agent
- a) Salicylic acid b) Glycerol c) Benzoic acid d) EDTA
- 2. ----- is used as a titrant in conductometry
- a) Sodium thiosulphate b) Oxalic acid c) Perchloric acid d) None of these
- 3. The complexometric titration where EDTA is used carried out at basic pH. Why?
- a) For the stability of complex formed b) Reaction rate is optimum in basic pH
- c) There is less number of side reaction d) All

4 is the weight of analyte per unit weight of the precipitate.

a) Gravimetric factor b) Precipitation factor c) Electrogravimetry factor d) None of the above

5. Following are properties of good precipitate except

a) Easily filtered and washed free of contamination. b) Significant loss of the analyte occurs during filtration and washing. c) Unreactive with constituents of the atmosphere. d) both a and b

6. Which of the following is used as indicator electrode in polarography

a) Glass b) Silver c) Dropping Murcury d) platinum

- **7**. SI unit of conductance is
- a) Mho b) Seimens c) Volt d) None of the above
- 8. Example for reference electrode except

a) Antimony electrode b) Silver-silver electrode c) Calomel electrode d) None of the above

- 9. Each electrochemical cell is composed of
- a) Two half cells b) Half cells c) Both A and B d) None of the above
- 10) For Qualitative analysis by polarography, characteristic parameter used is
- a) Diffusion current b) Half wave potential c) Voltage d) None of the above

QII Long Answers Question (Answer any 1 out of 2)

1. What is redox titration? Enlist its type? Explain iodometry and iodometric titration?

2. Write a short note on conductometry titration?

QIII Short Answers Question (Answer any 2 out of 3)

- 1. Discuss diazotization titration?
- 2. What are masking and demasking agents?
- 3. Describe the construction and working of Dropping Mercury Electrode?

10



Mapping of Course Outcome with Second Sessional Theory Examination (2021-2022)

Course Outcome	Question	Marks	
	Q1 – 1,2,3,4,5		
CO 1023:	Q2 – 1	25	
	Q3-1, 2,		
	Q1 – 6,7,8,9,10		
CO 1024:	Q2 -2	20	
	Q3 – 3		

Subject: Pharmaceutical Analysis-I

Mohd Usman	Azim Ansari	Girija Bhavar
Prepared By	Verified By	Approved By



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	Second Sessional Theory Examinat	ion-(ODD SEM) 2023-2024	
Subject	: Pharmaceutical Analysis (BP102T)	Day & Date :	Tuesday, 19/12	2/2023
Class	: First Year B. Pharmacy	Semester :	I Write Your Se	eat No. Here
Time	: 10:30 am – 12:00 pm	Max. Marks :	30	
Instructions	1. All questions are compulsory2.3. Right hand side number indicates full marks4.	. Draw a well labeled o Do not write/tick on th	liagram wherever neo ne question paper	cessary
QI	Objective Type Questions			10 M
1.	Define Masking and Demasking with its example (C	CO3, LL1)		
2.	Describe co-precipitation in gravimetric analysis (CO	03, LL2)		
3.	Define: Self Indicators and Half wave potential (CO	3, LL1; CO4, LL1)		
4.	Draw a well labelled diagram of Calomel electrode ((CO4, LL1)		
5.	Write Ilkovic equation and explain all terms (CO4, I	LL3)		
QII	Long Answers Question (Answer any 1 out of 2) ((CO3, LL2)		10 M
1. Explain in detail the principle involved in Iodometry and cerimetry titrations with its applications				
2.	Explain the principle and steps involved in gravimet	ric analysis and desc	cribe post-precipitat	tion.
QIII	Short Answers Question (Answer any 2 out of 3)			10 M
1.	Explain conductivity cells and write applications of c	conductometric titra	tion (CO4, LL2)	
2	Enumerate the indicator electrodes used in potention	netric titration and e	xplain the Glass ele	ectrode
2.	(CO4, LL1)			
3	Describe rotating dropping mercury electrode and its	s applications (CO4	III)	
5.	beserie rotating dropping increary electrode and its			



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Second Sessional Theory Examination-(ODD SEM) 2022-2023

Subject	:	Pharmaceutics-I (BP103T)	Day & Date	:	Wedn	esday,19/04/2023
Class	:	First Year B. Pharmacy	Semester	:	Ι	Write Your Seat No.
Time	:	10.30 pm to 12.00 noon	Max. Marks	:	30	Here

Instructions:

- 1. All questions are compulsory
- 2. Draw a well-labeled diagram wherever necessary
- 3. Right-hand side number indicates full marks
- 4. Do not write/tick on the question paper

Q. I: Objective Type Questions (5 X 2)

- 1 Differentiate between flocculated and deflocculated Suspension (CO3, LL2)
- 2 State the ideal properties of suppositories bases (CO3, LL2)
- **3** Define displacement value and cream.(CO3 and CO4, LL1)
- 4 Enlist the factor affecting on dermal penetration of semisolid dosage forms.(CO4, LL1)
- 5 Define paste and Enlist the method of prepation of ointment(CO4, LL1)

Q. II: Long Answers Questions (Answer any 1 out of 2)

- 1 Define emulsion. Explain different methods of preparation and identification tests.(CO3, LL3)
- 2 Define the term pharmaceutical incompatibility and discuss physical, chemical, and therapeutic incompatibility with examples. (CO3, LL3)

Q. III: Short Answers Questions (Answer any 2 out of 3)

- 1 Discuss in brief about Jellies and add note on types of Jellies (CO4, LL3)
- 2 Differentiate between ointment and paste. Add note on ointment bases.(CO4, LL3)
- **3** Give the evaluation of semisolid dosage forms.(CO4, LL3)

10 M

10 M



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Second Sessional Theory Examination-(ODD SEM) 2022-2023

Subject	:	Pharmaceutics-I (BP103T)	Day & Date	:	Wedn	esday,19/04/2023
Class	:	First Year B. Pharmacy	Semester	:	Ι	Write Your Seat No.
Time	:	10.30 pm to 12.00 noon	Max. Marks	:	30	Here

Instructions:

- 1. All questions are compulsory
- 2. Draw a well-labeled diagram wherever necessary
- 3. Right-hand side number indicates full marks
- 4. Do not write/tick on the question paper

Q. I: Objective Type Questions (5 X 2)

- 1 Differentiate between flocculated and deflocculated Suspension (CO3, LL2)
- 2 State the ideal properties of suppositories bases (CO3, LL2)
- **3** Define displacement value and cream.(CO3 and CO4, LL1)
- 4 Enlist the factor affecting on dermal penetration of semisolid dosage forms.(CO4, LL1)
- 5 Define paste and Enlist the method of prepation of ointment(CO4, LL1)

Q. II: Long Answers Questions (Answer any 1 out of 2)

- 1 Define emulsion. Explain different methods of preparation and identification tests.(CO3, LL3)
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Q. III: Short Answers Questions (Answer any 2 out of 3)

- 1 Discuss in brief about Jellies and add note on types of Jellies (CO4, LL3)
- 2 Differentiate between ointment and paste. Add note on ointment bases.(CO4, LL3)
- **3** Give the evaluation of semisolid dosage forms.(CO4, LL3)

10 M

10 M

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Second Sessional Theor Subject: Pharmaceutics I BP103T Class: FYB. Pharm Time: 90 min	y Examination 2018-2019 Day & Date: 19.11.2018 Semester: I Max. Marks: 30
Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherea 3. Right hand side number indicates full Q.No.1 Multiple choice questions 1. The amount of alcohol in elixirs ranges from	ver necessary 1 marks (10 X 1=10)
 (a) 4-40 % b) 76% c) 70% d) 5-50% 2 are liquid dosage forms and externall (a) Enemas b) Pessaries c) Suppositories d) Bougies 	y inserted into body cavity to promote defecation.
3. Regarding the rate of sedimentation of pharma- which of the following statements are true?	ceutical suspensions designed for oral administration,
 a. The rate of sedimentation is increased as the diam b. The rate of sedimentation is increased as the visc c. The rate of sedimentation is affected by the conce 	neter of the dispersed drug particles is increased. osity of the continuous phase is increased. entration of buffer salts.
 d. None of the above. 4. Bottle method is used for formulation of emulsio a) Fixed, b) Mineral, C Volatile, d) None of the above. 5. The desage form intended to be rubbed on the second secon	n containing oil. ove.
 a) Lotions b) Liniments c) Gels d) Dentifrices a is used as lubricant for gelato-glyce 	erin suppository bases.
 a) Glycerin, alcohol and liquid paraffin/ arachis oil, 7 form is most stable form of cocoa butter. a) α b) β c) x d) σ 	b) soap solution, c) mineral oil, d) None of the above.
 8. Calculate the displacement value of zinc oxide in oxide prepared in a 1g mould (weight of 10 suppose 14 72g) 	theobroma oil suppositories containing 40% of zinc itories is 10g and weight of medicated suppositories is
a) $(5, d) (4, c) (6, d) (2)$	
 a) mouth-washes, b) gargles, c) throat-paints, d) No 10. Tetracycline taken with milk, is a type of a) Therapeutic, b) Physical, c) Chemical, d) Adjust 	one of the above incompatibility. ed.
 Q. No. 2 Long Answers (Answer 1 out of 2) 1. Discuss stability problems of emulsion and 2. Define suspension. Write down ideal propulsion. 	(1 x 10 = 10) methods to overcome. perties of suspension. Write a note on formulation of
 Q. No. 3 Short Answers (Answer 2 out of 3) 1. Define emulsion. Write down ideal properties of identification emulsion. 2. Discuss various bases used in the formulation of the f	$(2 \times 5 = 10)$ of emulsion. Write down any four test used for of suppositories.
3. Write short note on (any one)a) Elixir.b) Physical incompatibility.	



IN- SEMESTER THEORY ASSESSMENT 2018 -2019

Class: First Year B. Pharmacy	Semester: I	Division : A
Subject: Pharmaceutical Analysis-I		Total Marks: 05
Assessment Type: Quiz		Date: 25-10-2018

1. Write the important feature of ethylene diamne tetra acetic acid which make it choice of chelating agent

Assessment Type: Class Test

Date: 26-10-2018

1. Discuss in details the steps involved in gravimetric analysis. Add note on difference between lyophilic and lyophobic colloids.

Assessment Type: Assignment

Date: 27-10-2018

1. Explain in detail the neutralization curves of various acidbase titration
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Second Sessional Theory Subject: Pharmaceutics I BP103T Class: FYB. Pharm Time: 90 min	r Examination 2018-2019 Day & Date: 19.11.2018 Semester: I Max. Marks: 30						
<i>Instructions:</i> 1. All questions are compulso 2. Draw a well labeled diagram	ry n wherever necessary						
3. Right hand side number inc	licates full marks						
Q.No.1 Multiple choice questions	(10 X 1=10)						
1 are diluted with warm water before use.							
a) Gargles b) Throat paints c) Mouthwashes d) All of abo	ve						
2. In suspensions, are added for suspending in	soluble powder in a vehicle.						
a) Suspending agents, b) emulsifying agents, c) surfactant	s, d) electrolytes						
3 Is used as a flocculating agent.							
a) Tweens, b) Butylated hydroxyl anisole, c) glycerine, d)	None of the above.						
4 are semisolid unit dosage forms and extern	ally inserted into body cavity to promote defecation.						
a) Enemas b) Pessaries c) Suppositories d) Bougies							
5. Emulsifying agents which reduces interfacial tension ar	e classified as						
a) Sufactants, b) Finely divided solids, c) Polysaccharides	, d) Sterols						
6. Regarding the rate of sedimentation of pharmaceutical	suspensions designed for oral administration, which of the						
following statements are true?							
1. The rate of sedimentation is increased as the diameter of	the dispersed drug particles is increased.						
ii. The rate of sedimentation is increased as the viscosity of the continuous phase is increased.							
iii. The rate of sedimentation may be increased by centrifu	igation.						
iv. The rate of sedimentation may be increased by decreas	ing density difference.						
a) i &iii b) i ⅈ, c) ii &iii, d) ii& iv							
7 is used as o/w emulsifiers.							
a) Acacia, b) Span, c) Magnesium oxide, d) None of the a	bove.						
8 is used as lubricant for gelato-glycerin su	pository bases.						
a) Glycerin, alcohol and liquid paraffin/ arachis oil, b) soa	p solution, c) mineral oil, d) None of the above.						
9 form is most stable form of cocoa butter.							
a) α b) β c) γ d) σ							
10. Tetracycline taken with milk, is a type of	incompatibility.						
a) Therapeutic, b) Physical, c) Chemical, d) Adjusted.							
Q. No. 2 Long Answers (Answer 1 out of 2)	$(1 \times 10 = 10)$						
3. Define emulsion. Write down ideal properties of emu	lsion. Write a note on emulsifying agents.						
4. Define suppositories. Write down types and advantage preparation.	ges and disadvantages of suppositories. Explain method of						
	n in sec Tradition in the second second second second second second second second second second second second second sec						
Q. No. 3 Short Answers (Answer 2 out of 3)	$(2 \times 5 = 10)$						
1. Define suspension. Write down ideal properties of susp	ensions.						
2. Discuss therapeutic incompatibility.							
3. Write short note on Syrup.							

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Second Sessional Theor	ry Examination 2018-2019
Subject: Pharmaceutics I BP103T	Day & Date: 19.11.2018
Class: FYB. Pharm	Semester: 1 More Morela: 20
Time: 90 min	Max. Marks. 50
2 Drow a well labeled diagra	am wherever necessary
2. Draw a well labeled diagre	ndicates full marks
5. Aight hand side humber h	
O No.1 Multiple choice questions	(10 X 1=10)
1	
a) Gargles b) Throat paints c) Mouthwashes d) All of at	oove
2. In suspensions, are added for suspending	insoluble powder in a vehicle.
a) Suspending agents, b) emulsifying agents, c) surfacta	nts, d) electrolytes
3 Is used as a flocculating agent.	
a) Tweens, b) Butylated hydroxyl anisole, c) glycerine,	d) None of the above.
4 are semisolid unit dosage forms and exte	rnally inserted into body cavity to promote defecation.
a) Enemas b) Pessaries c) Suppositories d) Bougies	
5. Emulsifying agents which reduces interfacial tension	are classified as
a) Sufactants, b) Finely divided solids, c) Polysaccharid	es, d) Sterols
6. Regarding the rate of sedimentation of pharmaceutic	al suspensions designed for oral administration, which of the
following statements are true?	
i. The rate of sedimentation is increased as the diameter	of the dispersed drug particles is increased.
ii. The rate of sedimentation is increased as the viscosit	y of the continuous phase is increased.
iii. The rate of sedimentation may be increased by central	ifugation.
iv. The rate of sedimentation may be increased by decre	easing density difference.
a) i &iii b) i ⅈ, c) ii &iii, d) ii& iv	
7 is used as o/w emulsifiers.	
a) Acacia, b) Span, c) Magnesium oxide, d) None of the	e above.
8 is used as lubricant for gelato-glycerin	suppository bases.
a) Glycerin, alcohol and liquid paraffin/ arachis oil, b)	soap solution, c) mineral oil, d) None of the above.
9 form is most stable form of cocoa butter.	
a) α b) β c) γ d) σ	
10. Tetracycline taken with milk, is a type of	incompatibility.
a) Therapeutic, b) Physical, c) Chemical, d) Adjusted.	
Q. No. 2 Long Answers (Answer 1 out of 2)	$(1 \times 10 = 10)$
3. Define emulsion. Write down ideal properties of e	mulsion. Write a note on emulsifying agents.
4. Define suppositories. Write down types and advan	ntages and disadvantages of suppositories. Explain method of
preparation.	
	a de tre
Q. No. 3 Short Answers (Answer 2 out of 3)	$(2 \times 5 = 10)$
1. Define suspension. Write down ideal properties of s	uspensions.
2. Discuss therapeutic incompatibility.	
3. Write short note on Syrup.	

Shri, Vile Parle Kelavani Mandal's **Institute of Pharmacy, Dhule** Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway Second Sessional Theory Examination 2019-2020 Day & Date: Thursday, 7th Nov 2019 Subject: Pharmaceutics I BP103T Class: FYB. Pharm Semester: I Time: 90 min Max. Marks: 30 Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks **Q. No.1 Multiple choice questions** (10 X 1=10)1. In dry gum method containing fixed oil ratio of oil: water: acacia to formulate primary emulsion is..... (a) 4:2:1, b) 3:2:1, c) 2:2:1, d) 2:2:1 2. form is most stable form of cocoa butter. a) α vb) B c) γ d) σ 3. If water soluble dye (amaranth) is mixed with the emulsion and examined under the microscope, the dispersed globules appears red and ground colorless indicates a) o/w type, b) w/o type, c) w/o/w, d) None of the above. 4. Glycero-gelatin suppository USP contains...... a) 70% glycerin, 20 % gelatin and 10 % water, b) 70% glycerin, 14 % gelatin and 16 % water, c) 40% glycerin, 32.5 % gelatin and 28.5 % water, d) 60% glycerin, 20% gelatin and 20% water. 5. If two drugs having opposing pharmacological effect are prescribed together results in..... a) Physical incompatibility, b) Chemical incompatibility, c) Therapeutic incompatibility, d) None of the above. 6. Regarding the rate of sedimentation of pharmaceutical suspensions designed for oral administration, which of the following statements are true? i. The rate of sedimentation is increased as the diameter of the dispersed drug particles is increased. ii. The rate of sedimentation is increased as the viscosity of the continuous phase is increased. iii. The rate of sedimentation may be increased by centrifugation. iv. The rate of sedimentation may be increased by decreasing density difference. a) i &iii b) i &ii. c) ii &iii, d) ii& iv 7. Which of the following is hydrocarbon base a) Lanolin, b) White soft paraffin, c) Macrogols, d) Non of the above 8. 2% dispersion is used as suspending agent for indiffusible solids. a) Bentonite. b) Benzoic acid, c) Acacia, d) Sodium acetate. 9. Emulsifying agents which reduces interfacial tension are classified as..... a) Sufactants, b) Finely divided solids, c) Polysaccharides, d) Sterols 10. are semisolid unit dosage forms and externally inserted into body cavity to promote defecation. a) Enemas b) Pessaries c) Suppositories d) Bougies Q. No. 2 Long Answers (Answer 1 out of 2) $(1 \times 10 = 10)$ 1. Discuss stability problems of emulsion and methods to overcome. 2. Define ointment. Write down ideal properties of ointment. Discuss various ointment bases along with examples. Q. No. 3 Short Answers (Answer 2 out of 3) $(2 \times 5 = 10)$ 1. Write a note on suppository bases. 2. Write down various test used for identification of emulsions. 3. Write a note on (any one) a. Suspending agents b. Therapeutic incompatibility

Second Sessional Theory Examination 2018-2019 Subject: P'ceutical Inorganic Chemistry (BP104T) Day & Date: Wed, 21/11/2018 Class: First Year B. Pharm. Semester: I Time: 10.30 am – 12.00 pm Max. Marks: 30
Instructions:1. All questions are compulsory2. Draw a well labeled diagram wherever necessary3. Right hand side number indicates full marksQ.I. Multiple Choice Questions (MCQs)10
1. Fluoride inhibits caries formation via
a) decreasing acid solubility of enamel b) bacterial inhibition
c) both a & b d) increasing acid solubility of enamel
2. Calcium carbonate is used as
a) cleaning agent in dentifrice b) an antacid
c) both a & b d) none of the above
3. Absence of HCl in gastric secretions is called as
a) Hyperacidity b) Achlorhydria c) Constipation d) Diarrhoea
4. Antacid should not be
a) constipative b) laxative c) both a & b d) protective
5. assay of sodium bicarbonate is carried out by titration
a) acid base b) precipitation c) complexometric d) redox
6. saline cathartics act by
a) local irritation of the intestinal tract b) swells in intestine
c) acting as lubricants d) increasing osmotic load of intestine
7. ammonium chloride can be used as
a) expectorant b) systemic acidifier C) diuretic d) all of the above
8. which of the following antidote act by adsorption?
a) activated charcoal b) sodium thiosulphate c) sodium nitrite d) none of the above
9. the drug used in the treatment of cyanide poisoning is
a) sodium nitrite b) sodium thiosulphate c) both a & b d) none of these
10. Aluminium hydroxide gel is
a) absorbable antacid b) non-absorbable antacid
c) soluble in water d) soluble in alcohol
Q. II. Solve any One
1. What are dental caries? Explain the role of fluoride in the treatment of dental caries. Write a note on
sodium fluoride.
2. Define and classify antacids. Enlist the ideal properties and write a note on Aluminium hydroxide gel.
Q. III. Solve any Two (5 Marks each) 10
1. Define and classify antidotes with examples. Write in short on cyanide poisoning and its treatment.
2. What are emetics? Explain copper sulphate with respect to properties and assay.
3. Define expectorants. Explain the properties, assay and uses of ammonium chloride.

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Second Sessional Theory Examination 2019-2020
Subject: P'ceutical Inorganic Chemistry (BP104T) Day & Date: Fri, 08/11/2019
Class: First Year B. Pharm. Semester: I
Time: 10.30 am - 12.00 pm Max. Marks: 30
Instructions: 1. All questions are compulsory 2. Right hand side number indicates full marks
Q.I. Multiple Choice Questions (MCQs)101radiations are deflected in magnetic field
a) α b) β c) γ d) both a & b
2. Which of the following is an example of astringent?
a) Zinc sulphate b) Potash alum c) both a & b d) none of these
3. Tincture iodine solution is
a) aqueous iodine solution b) strong iodine solution c) Weak iodine solution d) all of these
4. Antacid should not be
a) constipative b) laxative c) both a & b d) protective
5. Unit of radioactivity is
a) Calorie b) Newton c) Curie d) Joules
6. Saline cathartics act by
a) local irritation of the intestinal tract (b) swells in intestine
c) acting as lubricants d) increasing osmotic load of intestine
7. Ammonium chloride can be used as
a) expectorant b) systemic acidifier C) diuretic d) all of the above
8. The drug used in the treatment of cyanide poisoning is
a) sodium nitrite b) sodium thiosulphate c) both a & b d) none of these
9. Milk of magnesia IP is an aueous suspension of hydrated
a) MgSO ₄ b) Mg(OH) ₂ c) MgCO ₃ d) Mg ⁻ trisillicate
10. which of the following antidote act by adsorption?
a) activated charcoal b) sodium thiosulphate c) sodium nitrite d) none of the above
 Q. II. Solve any One 10 1. Define and classify antacids. Enlist the ideal properties and write a note on Aluminium hydroxide gel. 2. Define and classify antidotes with examples. Write in short on example poisoning and its
treatment. Explain the properties, assay and uses of sodium thiosulphate
 Q. III. Solve any Two (5 Marks each) 1. Explain the mechanisms by which antimicrobials act. Write any one method of preparation and assay of hydrogen peroxide. 2. What is radioactive decay? Describe properties of α, β, γ radiations.

3. Define expectorants. Explain the properties, assay and uses of ammonium chloride.



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Second Sessional Theory Examination 2020-2021

(Odd SEM)

Subject: Pharmaceutical Inorganic Chemistry (BP104T)	Day & Date: 28-04-2021
Class: First Year B. Pharm.	Semester: I
Time: 10:30 am – 12:00 pm	Max. Marks : 30

Instructions: 1. All questions are compulsory

- 2. Draw a well labeled diagram wherever necessary
- 3. Right hand side number indicates full marks

Q.I. Multiple Choice Questions (MCQs)

- 1. Saline cathartics act by
 - a. Local irritation of intestinal tract
 - b. Swelling in intestine
 - c. Acting as a lubricant
 - d. increasing osmotic load of intestine
- 2. Hydrogrn peroxide act as an antimicrobial agent by mechanism
 - a. Protein precipitation
 - b. Oxidation
 - c. Halogenation
 - d. All of the above
- 3. Which of the following is an example of astringent?
 - a. Zinc sulphate
 - b. Potash alum
 - c. Both a and b
 - d. None of these
- 4. Drug that is used to clear mucus from respiratory tract is called as
 - a. Expectorant
 - b. Antacid
 - c. Astringent
 - d. Cathartic
- 5. Which of the following is an example of hematinic?
 - a. Ferrous gluconate
 - b. Zinc sulphate
 - c? · Copper sulphate
 - d. Magnesium sulphate
- 6. Which of the following antidote act by adsorption?
 - a. Sodium thiosulphate
 - b. Activated charcoal

10



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- c. Sodium nitrite
- d. None of the above
- 7. Unit of radioactivity is
 - a. Calorie
 - b. Newton
 - c. Curie
 - d. Joules

8.Radiations are deflected in magnetic field

- a. Alpha
- b. Beta
- c. Gamma
- d. Both alpha and beta

9. The radioisotope used in studies of thyroid gland is

- a. Iodine-131
- b. Sodium-22
- c. Iron-59
- d. Gold-198
- 10. Chemically potash alum is
 - a. Potassium sulphate
 - b. Potassium aluminium sulphate
 - c. Aluminium chloride
 - d. Potassium permanganate

Q. II. Solve the following (any 1)

- 1. Define and classify antidotes. Write in short about cyanide poisoning. Write a note on sodium thiosulphate.
- 2. What is radioactive decay? Describe the properties of radioactive decay particles. Write a note on handling and storage of radiopharmaceuticals.

Q. III. Solve the following (any 2)

1. Explain the mechanism of action of antimicrobial agents with examples.

- 2. What are expectorants? Explain the method for preparation, properties, Assay and uses of Ammonium chloride.
- 3. What are hematinics? Explain the method for preparation, properties, Assay and uses of Ferrous sulphate

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Second Sessional Theory Examination 2021-2022 (Odd SEM)

Subject & Subject Code: Pharmaceutical Inorganic Chemistry (BP104T)

Day & Date: Wednesday 06/04/2022

Class: F.Y. B. Pharmacy

Time: 2:00 PM to 3:30 PM

Semester: I

Max. Marks: 30

10

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks **QI Multiple Choice Questions (MCQs)** 10 1. When cyanide poisoning treated with sodium nitrite it converts blood hemoglobin to ____ A) Oxyhemoglobin B) Hemoglobin C) Carboxyhemoglobin D) Methemoglobin 2. Universal Antidote consist of A) Tannic acid B) Charcoal C) MgO D) All of the above 3. Agent which inhibit growth of microorganism living tissue is called as _____ A) Disinfectant B) Antiseptic C) Germicide D) Sanitizers 4. Which of the following devices are used for the measurement of radio- activity? A) Geiger - Muller counter B) Ionization chamber C) Proportional counter D) All of the above 5. Delayed evacuation of the bowels / Faeces is called B) Constipation C) Peristalsis D) Cathartics A) Acidity 6. Radiopharmaceuticals use in study of thyroid uptake_ A) K-42 B) Cr-51 C) I-131 D) S-357 7. The pharmaceutical aid used as astringent and clearing agent is C) Benzyl Alcohol A) Agar B) Alum D) All 8. Which of following is called Rochelle salt? _ A) Sodium Potassium Tartrate B) Potassium Bitartrate C) Potassium citrate D) Potassium oxalate 9. The mechanism of halogenation occurring with antiseptic are_____. A) Sulfhydryl B) Halogen C) Hypohalite D) Oxidation 10. All of the following drugs used for iron deficiency anaemia EXCEPT: A) Ferrous Sulphate B) Ferrous Fumarate C) Calcium Gluconate D) Ferrous Gluconate QII Long Answers (Answer any 1 out of 2) 10

- 1. Describe in detail Geiger Muller counters and Scintillation counters for the measurement of Radio activity.
- 2. Define & classify Antidote. Write a note on cyanide poisoning and discuss the physical and chemical properties and assay of sodium thiosulphate.

QIII Short Answers (Answer any 2 out of 3)

- 1. Define Haematinics, give synonym physical properties and uses of FeSO4.
- 2. Define and classify expectorant, give assay, physical properties and uses of Ammonium chloride.
- 3. Define Antimicrobial agent. Discuss mechanism of action of inorganic antimicrobial agent.



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Second Sessional Theory Examination 2020-2021

(Odd SEM)

Subject: Pharmaceutical Inorganic Chemistry (BP104T)	Day & Date: 28-04-2021
Class: First Year B. Pharm.	Semester: I
Time: 10:30 am – 12:00 pm	Max. Marks : 30

Instructions: 1. All questions are compulsory

- 2. Draw a well labeled diagram wherever necessary
- 3. Right hand side number indicates full marks

Q.I. Multiple Choice Questions (MCQs)

- 1. Saline cathartics act by
 - a. Local irritation of intestinal tract
 - b. Swelling in intestine
 - c. Acting as a lubricant
 - d. increasing osmotic load of intestine
- 2. Hydrogrn peroxide act as an antimicrobial agent by mechanism
 - a. Protein precipitation
 - b. Oxidation
 - c. Halogenation
 - d. All of the above
- 3. Which of the following is an example of astringent?
 - a. Zinc sulphate
 - b. Potash alum
 - c. Both a and b
 - d. None of these
- 4. Drug that is used to clear mucus from respiratory tract is called as
 - a. Expectorant
 - b. Antacid
 - c. Astringent
 - d. Cathartic
- 5. Which of the following is an example of hematinic?
 - a. Ferrous gluconate
 - b. Zinc sulphate
 - c. Copper sulphate
 - d. Magnesium sulphate
- 6. Which of the following antidote act by adsorption?
 - a. Sodium thiosulphate
 - b. Activated charcoal



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- c. Sodium nitrite
- d. None of the above
- 7. Unit of radioactivity is
 - a. Calorie
 - b. Newton
 - c. Curie
 - d. Joules

8.Radiations are deflected in magnetic field

- a. Alpha
- b. Beta
- c. Gamma
- d. Both alpha and beta
- 9. The radioisotope used in studies of thyroid gland is
 - a. Iodine-131
 - b. Sodium-22
 - c. Iron-59
 - d. Gold-198
- 10. Chemically potash alum is
 - a. Potassium sulphate
 - b. Potassium aluminium sulphate
 - c. Aluminium chloride
 - d. Potassium permanganate

Q. II. Solve the following (any 1)

- 1. Define and classify antidotes. Write in short about cyanide poisoning. Write a note on sodium thiosulphate.
- 2. What is radioactive decay? Describe the properties of radioactive decay particles. Write a note on handling and storage of radiopharmaceuticals.

Q. III. Solve the following (any 2)

- 1. Explain the mechanism of action of antimicrobial agents with examples.
- 2. What are expectorants? Explain the method for preparation, properties, Assay and uses of Ammonium chloride.
- 3. What are hematinics? Explain the method for preparation, properties, Assay and uses of Ferrous sulphate

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Subject	:	Pharr	nace	utical	Anal	ysis (l	BP102'	T)	Da	ay & I	Date	:	Tue	sday	,18/04	/202	23
Class	:	First	Year	B. Pl	narma	acy			Se	emest	er	:	Ι	Writ	e Your Se	eat No.	Here
Time	:	10:30	am	- 12:0	00 pm	1			M	ax. M	arks	:	30				
Instruction	1. 1.	All ques Right ho	tions (Ind sid	are com de num	pulsorı ber indi	y icates f	full mark:	S	2. Dro 4. Do 1	uw a we not writ	ell label e/tick o	ed á n th	liagran .e ques	n whe tion p	rever nec aper	xessarı	y
QI		Objectiv	ve Ty	pe Qu	estion	S										10	M
1.	W	rite the	appli	cations	s of so	dium r	nitrite tit	tration (C	03, I	LL3)							
2.	D	escribe	co-pr	ecipita	tion in	gravii	metric a	nalysis (203,	LL2)							
3.	Define: Self Indicators and Half wave potential (CO3, LL1; CO4, LL1)																
4.	Draw a well labelled diagram of Calomel electrode (CO4, LL1)																
5.	W	/rite Ilko	onic e	quatio	n and e	explair	n all terr	ms (CO4	LL3)							
QII	Lon	g Answ	ers Q	uestio	n (Ans	swer a	nny 1 ou	nt of 2) (C O3 ,	LL2)						10 M	
1.	Expl	lain in d	etail t	he prii	nciple	involv	e in Idio	ometry ar	nd cer	imetry	titrati	ons	with	its ap	plicatio	ns	

- **2.** Explain principle and steps involved in gravimetric analysis and describe post-precipitation.
- QIII Short Answers Question (Answer any 2 out of 3)
 - 1. Explain conductivity cells and write applications of conductometric titration (CO4, LL2)
 - Enumerate the indicator electrodes used in potentiometric titration and explain Glass electrode
 (CO4, LL1)
 - **3.** Describe rotating platinum electrode and its applications (CO4, LL1)

10 M



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	Se	cond	Ses	ssior	1al T	`heo	ry Ex	amin	atio	n-(0	DD S	SE	M) 2	202	2-202	23	
Subject	:	Pharr	nace	utical	Anal	ysis (l	BP102'	T)	Da	ay & I	Date	:	Tue	sday	,18/04	/202	23
Class	:	First	Year	B. Pl	narma	acy			Se	emest	er	:	Ι	Writ	e Your Se	eat No.	Here
Time	:	10:30	am	- 12:0	00 pm	1			M	ax. M	arks	:	30				
Instruction	1. 1.	All ques Right ho	tions (Ind sid	are com de num	pulsorı ber indi	y icates f	full mark:	S	2. Dro 4. Do 1	uw a we not writ	ell label e/tick o	ed á n th	liagran .e ques	n whe tion p	rever nec aper	xessarı	y
QI		Objectiv	ve Ty	pe Qu	estion	S										10	M
1.	W	rite the	appli	cations	s of so	dium r	nitrite tit	tration (C	03, I	LL3)							
2.	D	escribe	co-pr	ecipita	tion in	gravii	metric a	nalysis (203,	LL2)							
3.	Define: Self Indicators and Half wave potential (CO3, LL1; CO4, LL1)																
4.	Draw a well labelled diagram of Calomel electrode (CO4, LL1)																
5.	W	/rite Ilko	onic e	quatio	n and e	explair	n all terr	ms (CO4	LL3)							
QII	Lon	g Answ	ers Q	uestio	n (Ans	swer a	nny 1 ou	nt of 2) (C O3 ,	LL2)						10 M	
1.	Expl	lain in d	etail t	he prii	nciple	involv	e in Idio	ometry ar	nd cer	imetry	titrati	ons	with	its ap	plicatio	ns	

- **2.** Explain principle and steps involved in gravimetric analysis and describe post-precipitation.
- QIII Short Answers Question (Answer any 2 out of 3)
 - 1. Explain conductivity cells and write applications of conductometric titration (CO4, LL2)
 - Enumerate the indicator electrodes used in potentiometric titration and explain Glass electrode
 (CO4, LL1)
 - **3.** Describe rotating platinum electrode and its applications (CO4, LL1)

10 M



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	Se	cond	Ses	ssior	1al T	`heo	ry Ex	amin	atio	n-(0	DD S	SE	M) 2	202	2-202	23	
Subject	:	Pharr	nace	utical	Anal	ysis (l	BP102'	T)	Da	ay & I	Date	:	Tue	sday	,18/04	/202	23
Class	:	First	Year	B. Pl	narma	acy			Se	emest	er	:	Ι	Writ	e Your Se	eat No.	Here
Time	:	10:30	am	- 12:0	00 pm	1			M	ax. M	arks	:	30				
Instruction	1. 1.	All ques Right ho	tions (Ind sid	are com de num	pulsorı ber indi	y icates f	full mark:	S	2. Dro 4. Do 1	uw a we not writ	ell label e/tick o	ed á n th	liagran .e ques	n whe tion p	rever nec aper	xessarı	y
QI		Objectiv	ve Ty	pe Qu	estion	S										10	M
1.	W	rite the	appli	cations	s of so	dium r	nitrite tit	tration (C	03, I	LL3)							
2.	D	escribe	co-pr	ecipita	tion in	gravii	metric a	nalysis (203,	LL2)							
3.	Define: Self Indicators and Half wave potential (CO3, LL1; CO4, LL1)																
4.	Draw a well labelled diagram of Calomel electrode (CO4, LL1)																
5.	W	/rite Ilko	onic e	quatio	n and e	explair	n all terr	ms (CO4	LL3)							
QII	Lon	g Answ	ers Q	uestio	n (Ans	swer a	nny 1 ou	nt of 2) (C O3 ,	LL2)						10 M	
1.	Expl	lain in d	etail t	he prii	nciple	involv	e in Idio	ometry ar	nd cer	imetry	titrati	ons	with	its ap	plicatio	ns	

- **2.** Explain principle and steps involved in gravimetric analysis and describe post-precipitation.
- QIII Short Answers Question (Answer any 2 out of 3)
 - 1. Explain conductivity cells and write applications of conductometric titration (CO4, LL2)
 - Enumerate the indicator electrodes used in potentiometric titration and explain Glass electrode
 (CO4, LL1)
 - **3.** Describe rotating platinum electrode and its applications (CO4, LL1)

10 M

Second Sessional Theory Examination 2018-2019

Subject: Remedial Biology	Day & Date: Saturday, 24/11/2018
Class: F. Y. B.Pharm	Semester: I
Time: 10:30 am-12:00 pm	Max. Marks : 30

Q1. Answer any one

(10 marks)

1. Elucidate the functions of various hormones secreted by pituitary gland.

2. Explain importance of Rennin angiotensin aldosterone system in regulation of homeostasis

Q2. Answer any fours

1. Explain structure and functions of cell.

2. Classification of nervous system and short note on structure of brain and spinal cord

3. Describe macro and micronutrients

4. Explain photosynthesis process

5. Short note on Plant Respiration

6. Short note on Spermatogenesis

(20 marks)



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Second Sessional Theory Examination 2020 2021 (Odd SEM) Subject & Subject Code: Remedial Biology (BP106RBT) Day & Date: 24.04.2021 Saturday Class: F.Y. B. Pharm Semester: I

Time: 10.30 to 11.30 Max.

Marks: 30 marks

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks

Q1. Answer any one (10 marks)

1. Explain in detail about human excretory system. Add a note on Renin-Angiotensin System (RAS)

2. What is respiration? Give its classification and explain in detail about anaerobic respiration.

Q2. Answer any fours (20 marks)

1. Write a short note on Peripheral nervous system

2. Enlist various endocrine glands along with their functions

3. Add a note on spermatogenesis

4. Add a note on nitrogen metabolism

5. Add a short note on cell division

6. Give the phases of plant growth

- (S.S.Rathod)

Name and Signature Subject-Teacher



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Second Sessional Theory Examination 2020-2021 (Odd SEM)

Subject & Subject Code: Remedial Mathematics (BP106RMB)

Day & Date: Tuesday 24/4/2021

Semester: I/III

Class: F.Y/S.Y. B. Pharmacy

Time:1 Hour 30 Minutes

Max. Marks: 30

Instructions: 1. All questions are compulsory

2. Draw a well labeled diagram wherever necessary3. Right hand side number indicates full marks

QI Multiple Choice Questions (MCQs)

(b) 3 (c) 6

10 Marks

1. If 3x-4y+2=0 is given equation of line. Then its slope is

(a)
$$\frac{3}{4}$$
 (b) $\frac{1}{4}$
(c) $-\frac{3}{4}$ (d) $\frac{3}{2}$
2. If $A = \begin{bmatrix} 1 & 2 & 3 \\ 0 & 1 & -1 \\ 2 & 4 & x \end{bmatrix}$ is a singular matrix, then $x =$

3. If 2, 3, -1 are eigen values of matrix A Then eigen values of adjA are

(d) 2

(a) 1, 3, 6	(b) -2,-4, 6
(c) 3, 2, 1	(d) -3,-2, 6

(a) 4

(b) 3, 5

4. If line L passes through the points (1, 2) and (-2, 3) then slope of line L is (a) $-\frac{1}{3}$ (b) $\frac{2}{3}$ (c) $\frac{1}{3}$ (d) $-\frac{2}{3}$

5. θ is inclination of line parallel to line y=x then θ = (a) $\frac{\pi}{2}$ (b) $\frac{\pi}{3}$ (c) $\frac{\pi}{4}$ (d) $\frac{\pi}{6}$

6. If y=sin2x then
$$\left(\frac{dy}{dx}\right)$$
 at x = $\frac{\pi}{4}$ is

(d) 1, 2

7. If α and β are eigen values of Matrix $A = \begin{bmatrix} 1 & 2 \\ 2 & -1 \end{bmatrix}$ then $\alpha + \beta$ and $\alpha\beta$ are respectively (a) 0, -5 (c) 0, 5



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8 If
$$y = x^3 - x^2 + 2x - 1$$
. Then $\left(\frac{dy}{dx}\right)_{x=2}$ is
(a) -10 (b) 14 (b) 10 (d) 12

9. If
$$If y = a^{2x} \sqrt{x}$$
 then $\left(\frac{dy}{dx}\right)_{x=1}$ is

(a) $\frac{1}{2}a^2(4loga + 1)$ (b) $-\frac{1}{2}a^2(4loga + 1)$ (c) $\frac{1}{2}(4loga + 1)$ (d) $a^2(4loga + 1)$

10 If y=mx and y= - x are perpendicular to each other then m=

(a) -1 (b) 1 (c) 2 (d) -2

QII Long Answers (Answer any 1 out of 2) 10 Marks 1. Solve By Crammers Rule. x + y - z=0 2x + y + 3z=9 X - y + z=2

2. Find
$$\frac{dy}{dx}$$
 if $y = \frac{a^{2x}}{\log x} + \sin 2x \cdot \cos 4x + e^{-2x} \cdot \sec 2x$

QIII Short Answers (Answer any 2 out of 3)10 Marks1. Find Eigen values and Eigen vectors of $A = \begin{bmatrix} 5 & 4 \\ 1 & 2 \end{bmatrix}$ 10 Marks

2. Find equation of line passing through (3, 2) and perpendicular to the line x-2y-3=0

3. Find equation of line passing through (1, 2) and Parallel to the line 3x+4y+10=0

S. V. K. M's. INSTITUTE OF PHARMACY, DHULE Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

Second Sessional Theory Examination 2018-2019

Class: F.Y.B.Pharmacy	Semester: First
Subject : Remedial Mathematics	Max. Marks : 30
Date: November 24^{th} , 2018	Time : 1 hr. 30 min.

- Instructions :
 - 1. Both the questions are compulsory.
 - 2. Figures to the right indicates full marks.
- 1. Attempt any two of the following.
 - (a) Solve by matrix method
 - (b) Show that $\lim_{x \to a} \left(\frac{x^n a^n}{x a}\right) = na^{n-1}$. Using the same, evaluate $\lim_{x \to 2} \left(\frac{x^5 32}{x 2}\right)$. (c) Find A^{-1} , if $A = \begin{bmatrix} 3 & 1 & 2 \\ 2 & 4 & 1 \\ -1 & 2 & 2 \end{bmatrix}$.

2. Attempt any five of the following.

- (a) Without expanding the determinant, show that $\begin{vmatrix} 1 & 2 & 3 \\ 4 & 5 & 6 \\ 7 & 8 & 0 \end{vmatrix} = 0.$
- (b) If $f : \mathbf{R} \to \mathbf{R}$ such that f(x) = 2x + 1 and $g : \mathbf{R} \to \mathbf{R}$ such that $g(x) = x^2 2$. Find i) gofii) fof.
- (c) Evaluate : $\lim_{x \to 3} \left(\frac{x^2 5x + 6}{x^2 9} \right)$.
- (d) Evaluate : $\lim_{x \to 0} \left(\frac{3^x 5^x}{3^x 2^x} \right)$.
- (e) If $A = \begin{bmatrix} 1 & 5 \\ 2 & 1 \end{bmatrix}$ and $B = \begin{bmatrix} 2 & 4 \\ 1 & 6 \end{bmatrix}$, verify that $(AB)^{-1} = B^{-1}A^{-1}$.
- (f) Using Cramer's rule, solve x + y = 3, x y = 2.
- (g) If $A = \begin{bmatrix} 1 & 2 \\ 3 & 4 \end{bmatrix}$ and $B = \begin{bmatrix} 1 & 5 \\ 1 & 6 \end{bmatrix}$, verify that $(A + B)^t = A^t + B^t$.

(5+5=10)

(4x5=20)



Shri. Vile Parle Kelavani Mandal's Institute of Pharmacy, Dhule

Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

Second Sessional Theory Examination 2019-2020

Subject: Remedial Mathematics (BP106RMT) Day & Date: Tuesday,

11/11/19

Class: First Year B. Pharm Time: 10-30 am to 12.00 pm

Semester: I/III* Max. Marks: 30

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks

QI Attempt any two of the following

1. Solve By Crammers Rule. x + y- z=0

2x + y +3z=9

X - y + z = 2

	٢1	-2	07	
2. If $A =$	0	1	6	Find A ⁻¹
	2	-3	4	

3. if $y(x) = e^{\tan 4x} \cos(\log x)$ Find $\frac{dy}{dx}$

QII Attempt any five of the following

(4×5=20)

(5+5=10)

1.(i) Define Rectangular Matrix. (ii) Write definition of Logarith.

(iii) Nonsingular Matrix. (iv) Write equation of line in Double intercept form. 2. Find equation Of a line with inclination $\frac{\pi}{3}$ and passes through the point (2, 3).

3.If $y(x) = e^{4x}\log x + \sin(e^{4x}) \cdot x^3$ Find $\frac{dy}{dx}$

4. Find equation Of a line which passes through the point (-2, -3) and (-7,-9).

5. Find
$$\begin{vmatrix} 5 & -4 & 9 \\ -6 & -8 & -2 \\ 0 & -2 & 12 \end{vmatrix}$$

6. If A = $\begin{bmatrix} 2 & 1 & 1 \\ -3 & 4 & -3 \\ -1 & -5 & -5 \end{bmatrix}$ And $\mathcal{B} = \begin{bmatrix} 0 & 6 & -6 \\ -2 & -4 & 0 \\ -8 & 5 & -5 \end{bmatrix}$ Verify that $(A+B)^{T} = A^{T} + B^{T}$
7. If y(x) = $\frac{\log(\cot x)}{\sin(4x).x^{3}}$ Find $\frac{dy}{dx}$.





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Second Sessional Theory Examination 2021-2022 (Odd SEM)

Subject & Subject Code: Remedial Mathematics (BP106RMT)

Day & Date: Friday 08/04/2022

Class: F.Y/S.Y. B. Pharmacy.

Semester: I/III

Time:1 hr

Max. Marks: 30

Instructions: 1. All questions are compulsory

Draw a well labeled diagram wherever necessary
 Right hand side number indicates full marks

QI Multiple Choice Questions (MCQs)

1. If y = sin 2x Then $\frac{dy}{dx} =$ (a) 2sin2x(b) 2cos2x(c) 2cosx(d) 2sinx2. If y = tanx Then $\frac{dy}{dx} =$ (a) $-sec^2x$ (b) $cosec^2x$ (c) sec^2x (d) $-cosec^2x$ 3. Laplace transform of t^2 is (a) $\frac{1}{s}$ (b) $\frac{1}{s^2}$ (c) $\frac{2}{s^2}$ (d) $\frac{4}{s^2}$

4. If inclination of line L is $\frac{\pi}{4}$ Then slop of line L is

(a) -1 (b) 1 (c) 0 (d) $\frac{1}{2}$ (d) 5. If Line passes through points (1 2) and (2 4) then slop of line L is (a) -1 (b) 0 (c) 1 (d) 2 6. If y = secx Then $\frac{dy}{dx} =$ (a) sec2x tan2x(b) secx tan2x(c) secx tanx(d) sec2x tanx



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- 7. If $y = x^4$ Then $\frac{dy}{dx} =$
- (a) $4x^3$ (b) x^3 (c) $4x^2$ (d) 0

8. Slop of line y=4 is

(a) -2 (b) 3 (c) -3 (d) 0 9. Laplace Transform of sin2t is (a) $\frac{1}{s^2-4}$ (b) $\frac{2}{s^2-4}$ (c) $\frac{s}{s^2+4}$ (d) $\frac{2}{s^2+4}$

10. Laplace Transform of cos2t is

(a) $\frac{1}{s^2+4}$ (b) $\frac{s}{s^2-4}$ (c) $\frac{1}{s^2-4}$ (d) $\frac{s}{s^2+4}$

QII Long Answers (Answer any 1 out of 2)

- 1. Using adjoint method, Find inverse of the matrix $A = \begin{bmatrix} 1 & -1 & 2 \\ 2 & -1 & 1 \\ 1 & 1 & -1 \end{bmatrix}$
- 2. Find equation of line passing through the point (-2, 0) and perpendicular to the line to the line 4x-3y=2

QIII Short Answers (Answer any 2 out of 3)

- 1. If $y = cosx. logx. find \frac{dy}{dx}$
- 2. Verify Cayley Hamilton theorem for $A = \begin{bmatrix} 1 & 1 \\ 1 & 0 \end{bmatrix}$ hence find its inverse.
- 3. Prove that the lines 3x-2y-1=0 and 9x-6y+5=0 are parallel.

10



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			Second Sessional Theory Examination	n-(ODD SE	CM) 20	22-2	2023			
Subje	ect	:	Remedial Mathematics (BP106RMT) D a	ay & Date	:	MC	NDA	Y 24,	/04/	2023	
Class	;	:	First Year B. Pharmacy/DSY B. Pharma	emester	:	Ι	Wri	te Your	Seat	No. Here	
Time	i 10:30 am - 11:30 am Max. Marks : 30										
Instructions1. All questions are compulsory 3. Right hand side number indicates full marks2. Draw a well labeled diagram wherever necessary 4. Do not write/ tick on the question paper										sary	
QI	Lo	Long Answers Question (Answer any 1 out of 2)10 1									
1.	If y	'=e	$x \log x find \frac{d^2 y}{dx^2}$ or y_2						(LL3	8, CO3)	
2.	Fin	d tl	the equation of line passing through the point $(-2, 0)$ and	d perpendicula	ar to	o line	4x-3y	v=2.	(LL	3, CO4)	
QII	She	ort	Answers Question								
	Section – A (Answer any ONE) 05 M								05 M		
	Dif	fer	entiate								
1.	$x^2\epsilon$	e ^x l	ogx with respect to x						(LL2	2, CO3)	
2.	Fir	ıd e	equation of line Passing through the point (2,2))and its slop	e is	:1			(LL2	, CO4)	
	Sec	tio	n – A (Answer any Three)							15 M	
	Fin	d d	ouble derivative of $y = e^x \cos x$ with respect to x.						(LL2,	
1.	CO	3)									
	Fir	ıd e	equation of line Passing through the point (0,0)) and perpen	dic	ular	to lir	1e x +	- y =	2	
2.											
	(LI	(LL2, CO4)									
3.	Fin	d n	naximum and minimum value of $f(x)=x^3 - 6x^2 + 9x$: + 15.					(LL2	2, CO3)	
4.	Fin	d e	quation of line passing through the point (-2,3) and (2	2,7)					(LL2	, CO4)	



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			Second Sessional Theory Examination	n-(ODD SE	CM) 20	22-2	2023			
Subje	ect	:	Remedial Mathematics (BP106RMT) D a	ay & Date	:	MC	NDA	Y 24,	/04/	2023	
Class	;	:	First Year B. Pharmacy/DSY B. Pharma	emester	:	Ι	Wri	te Your	Seat	No. Here	
Time	i 10:30 am - 11:30 am Max. Marks : 30										
Instructions1. All questions are compulsory 3. Right hand side number indicates full marks2. Draw a well labeled diagram wherever necessary 4. Do not write/ tick on the question paper										sary	
QI	Lo	Long Answers Question (Answer any 1 out of 2)10 1									
1.	If y	'=e	$x \log x find \frac{d^2 y}{dx^2}$ or y_2						(LL3	8, CO3)	
2.	Fin	d tl	the equation of line passing through the point $(-2, 0)$ and	d perpendicula	ar to	o line	4x-3y	v=2.	(LL	3, CO4)	
QII	She	ort	Answers Question								
	Section – A (Answer any ONE) 05 M								05 M		
	Dif	fer	entiate								
1.	$x^2\epsilon$	e ^x l	ogx with respect to x						(LL2	2, CO3)	
2.	Fir	ıd e	equation of line Passing through the point (2,2))and its slop	e is	:1			(LL2	, CO4)	
	Sec	tio	n – A (Answer any Three)							15 M	
	Fin	d d	ouble derivative of $y = e^x \cos x$ with respect to x.						(LL2,	
1.	CO	3)									
	Fir	ıd e	equation of line Passing through the point (0,0)) and perpen	dic	ular	to lir	1e x +	- y =	2	
2.											
	(LI	(LL2, CO4)									
3.	Fin	d n	naximum and minimum value of $f(x)=x^3 - 6x^2 + 9x$: + 15.					(LL2	2, CO3)	
4.	Fin	d e	quation of line passing through the point (-2,3) and (2	2,7)					(LL2	, CO4)	



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	Second Sessional Theory Examin	ation-(Odd SEM) 2	202.	3-2024	ļ –	
Subj	iect : Remedial Mathematics	Day & Date	:	Satu	rday 23/1	12/2023
Clas	s : First Year B. Pharmacy	Semester	:	I/III	Write Y	our Seat No
Time	e : 10:30 am to 11:30 am	Max. Marks	:	30	Here	
Instruc	<i>ctions</i> :1. All questions are compulsory2.3. Right-hand side number indicates full marks4.	Draw a well-labeled di Do not write/tick on the	agra que	m wher stion pc	ever neces. per	sary
Q. I:	: Long Answers Question (Answer any 1 out of 2)					10 M
a)	Verify Cayley –Hamilton theorem: $A = \begin{bmatrix} 3 & 1 & 2 \\ 4 & 3 & 1 \\ 2 & 5 & 4 \end{bmatrix}$					
b)	i) Find the slope and intercept of the line $2x + 3y = 6$					
	ii) Evaluate $\int x \cdot \sin x dx$.					
Q. II	I: Short Answers Question (Answer any 4 out of 6)					20 M
a)	Find $\frac{dy}{dx}$ if $y = 10^x + x^{10} + e^x + sinx + tan x$					
b)	Evaluate $\int m^x + x^m + m^m dx$.					
c)	Find the equation of line which is passing through the j	points (-4,6), & (8, -	-3).			
d)	Show that the points are the vertices of $(-1,5)$, $(4,3)$, $($	7, –2) isosceles tria	ngl	e.		
e)	$If A = \begin{bmatrix} 1 & 3 & 2 \\ -1 & 0 & 2 \\ 3 & 1 & -1 \end{bmatrix} \text{ then } find \ adjoint \ A$					
f)	$If A = \begin{bmatrix} 2 & 4 & 4 \\ 4 & 2 & 4 \\ 4 & 4 & 2 \end{bmatrix} \text{ then } find \ A^2.$					



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	Second Sessional Theory Examin	ation-(Odd SEM) 2	202.	3-2024	ļ –	
Subj	iect : Remedial Mathematics	Day & Date	:	Satu	rday 23/1	12/2023
Clas	s : First Year B. Pharmacy	Semester	:	I/III	Write Y	our Seat No
Time	e : 10:30 am to 11:30 am	Max. Marks	:	30	Here	
Instruc	<i>ctions</i> :1. All questions are compulsory2.3. Right-hand side number indicates full marks4.	Draw a well-labeled di Do not write/tick on the	agra que	m wher stion pc	ever neces. per	sary
Q. I:	: Long Answers Question (Answer any 1 out of 2)					10 M
a)	Verify Cayley –Hamilton theorem: $A = \begin{bmatrix} 3 & 1 & 2 \\ 4 & 3 & 1 \\ 2 & 5 & 4 \end{bmatrix}$					
b)	i) Find the slope and intercept of the line $2x + 3y = 6$					
	ii) Evaluate $\int x \cdot \sin x dx$.					
Q. II	I: Short Answers Question (Answer any 4 out of 6)					20 M
a)	Find $\frac{dy}{dx}$ if $y = 10^x + x^{10} + e^x + sinx + tan x$					
b)	Evaluate $\int m^x + x^m + m^m dx$.					
c)	Find the equation of line which is passing through the j	points (-4,6), & (8, -	-3).			
d)	Show that the points are the vertices of $(-1,5)$, $(4,3)$, $($	7, –2) isosceles tria	ngl	e.		
e)	$If A = \begin{bmatrix} 1 & 3 & 2 \\ -1 & 0 & 2 \\ 3 & 1 & -1 \end{bmatrix} \text{ then } find \ adjoint \ A$					
f)	$If A = \begin{bmatrix} 2 & 4 & 4 \\ 4 & 2 & 4 \\ 4 & 4 & 2 \end{bmatrix} \text{ then } find \ A^2.$					



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Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

First Sessional Theory Examination 2020-2021 Subject: Pharmaceutics I BP103T Day & Date: Friday, 12th March 2021

Class: F Y B. Pharm	Semester: 1
Time: 90 min	Max. Marks: 30

Instructions: 1. All questions are compulsory

Draw a well labeled diagram wherever necessary
 Right hand side number indicates full marks

Q.No.1 Multiple choice questions

(10 X 1=10)

1. is known as 'Father of Pharmacy education in India'.

a) M. L. Shroff, b) R. N. Chopra, c) B. N. Ghosh, d) None of the above

2. The fourth edition of IP was published in

a) 1985 b) 1996 c) 1960 d) 19963.

3. Pharmacy Council of India (PCI) was established in

a) 1947 b) 1948 c) 1949 d) 1950.

4. Which of the following excipient act as preservative.....

a) Sorbitan monolactrate b) Tyloxopol c) Benzethenium chloride d) Thiomersal

5. When two or more drugs are used in combination to increase the pharmacological action, the phenomenon is known as

a) Synergism b) Tolerance c) Potentiation c) Idiosyncrasy

6. Nasal drops should have pH

- a) 5-6 b) 6-7 c) 7-8 d) 8-9
- 7. Which of the following is natural colorant
- a) Amaranth b) Erythrosine c) Carotenoids d) Eosin

8. The dosage form intended to be rubbed on the skin for rubeficient and counterirritant effect are called

a) Lotions b) Liniments c) Gels d) Dentifrices

9. 13 (apothecaries) =

a) 437.5 grains b) 480 grains c) 7000 grains d) 5760 grains

10. Daily adult dose of drug is 120 mg. How much of it should be given to a child of 11 years? a) 57.39, b) 60, c) 50.39, d) 37

Q. No. 2 Long Answers (Answer 1 out of 2)

$(1 \times 10 = 10)$

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1. Define dosage form. Explain need of dosage form. Describe various solid dosage forms.

2. Define posology. Write a note on factors affecting posology

Q. No. 3 Short Answers (Answer 2 out of 3)

$(2 \times 5 = 10)$

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1. Write down parts of prescription

2. Discuss various excipient used in the formulation of liquid dosage forms.

3. Write short note on (any one)

- a) Syrups
- b) Topical oral preparations

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Institute of Pharmacy, Dhule Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

Improvement Sessional Theory Examination 2018-2019							
Subject: Human Anatomy and Physiology-I	Day & Date:						
Class: F. Y. B. Pharma Semester: I							
Time: 11:00-12:30 Max. Marks: 30							
Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks							
QI. Solve the following MCQs101. Name the arteries and veins facilitating blood flow to and from the lower limbs:							
a) Brachial b) Mesenteric c) Femoral d) Carotid							
2. Blood flows out of the ventricles when?							
a) The atrioventricular valves are open b) The semi-lunar valves are open							
c) The bicuspid valves are open d) The mitral valves are closed							
3. Cardiac output is the amount of blood pumped by:							
a) 1 ventricle in 1 minute b) 1 atrium in 1 minute							
c) Both ventricles in 1 minute d) Both atria in 1 minute							
4. Blood from the heart flows through the blood vessels in which order?							
a) Veins, arterioles, capillaries, venules, arteries b) Capillaries, arterioles, arterie	s, venules, veins						
c) Arteries, arterioles, capillaries, venules, veins d) Arteries, capillaries, arteriole	s, venules, veins						
5. Blood pressure is highest when leaving which heart chamber?							
a) Right atrium b) Right ventricle							
c) Left atrium d) Left ventricle							
d) People who first identified them							
6. Which of the following structures is the primary organ of the integumentary syste	m?						
a) Sensory receptors b) Sweat glands							
c) Bones d) Skin							
7. Which cells produce the pigment that contributes to hair color?							
a) Keratinocytes b) Melanocytes							
c) Langerhans cells d) Merkel cells							

8. The main components of the skeletal system are:

a) Blood, bones, cartilage and tendons

b) Bones, cartilage, joints and tendons

c) Bones, cartilage, joints and ligaments

d) Cartilage, joints, ligaments and tendons

9. Muscle action can be classified as:

a) Sympathetic and parasympathetic

b) Voluntary and involuntary

c) Autonomic and sympathetic

d) Autonomic and voluntary

10. Which blood group can be transfused into an O Rh- patient?

a) O Rh-

b) O Rh+

c) AB Rh+

d) AB Rh-

QII Long Answers (Answer any 1 out of 2)

1. Define tissue. Explain in detail various types of tissue with classification

10

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2. Explain in details ABO and RH blood grouping system.

QIII Short Answers (Answer any 2 out of 3)

- 1. Define joints. Draw a neat labelled diagram of synovial joints
- 2. Write the various functions of skin with a diagram

3. Draw a neat labelled diagram of internal structure of eye.

	Shri. Vile Parle Kelavani Mandal's Institute of Pharmacy, Dhule Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway
	Improvement Sessional Theory Examination 2018-2019Subject: P'ceutical Inorganic Chemistry (BP104T)Day & Date: Thu, 20/11/2018Class: First Year B. Pharm.Semester: ITime: 10.30 am – 12.00 pmMax. Marks: 30
	Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary
	3. Right hand side number indicates full marks O.I. Multiple Choice Questions (MCOs)
	1. The only water soluble compound used as effective antacid is
	a) Magnessium trisillicate b) Sodium bicarbonate c) Milk of magnesia
	d) sodium hydroxide
	2. Calcium carbonate is used as
	a) cleaning agent in dentifrice b) an antacid c) both a & b d) none of the above
	3. Assay of ammonium chloride is carried out by titration
	a) acid base b) precipitation c) complexometric d) redox
	4. Antacid should not be
	a) constipative b) laxative c) both a & b d) protective
	5. assay of sodium bicarbonate is carried out by titration
	a) acid base b) precipitation c) complexometric d) redox
	6. Drug that induces vomiting is called as
×	a) Emetic b) Antacid c) Astringent d) Cathartic
	7. ammonium chloride can be used as
	a) expectorant b) systemic acidifier C) diuretic d) all of the above
	8. which of the following antidote act by adsorption?
	a) activated charcoal b) sodium thiosulphate c) sodium nitrite d) none of the above
	9. which of the following is a dental desensitizer?
	a) Calcium carbonate b) strontium chloride c) sodium fluoride d) potassium chloride
	10. Aluminium hydroxide gel is
	a) absorbable antacid b) non-absorbable antacid c) soluble in water d) soluble in alcohol
	Q. II. Solve any One 10
	1. what are dental caries? Explain the role of fluoride in the treatment of dental caries. Write a note on sodium
	fluoride.
	2. Define and classify antacids. Enlist the ideal properties and write a note on Aluminium hydroxide gel.
	Q. III. Solve any Two (5 Marks each) 10
	1. Define and classify antidotes with examples. Write in short on cyanide poisoning and its treatment.
	2. Write a note on activated charcoal.

3. Define expectorants. Explain the properties, assay and uses of ammonium chloride.

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Institute of Pharmacy, Dhule Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

Improvement Sessional Theory Examination 2019-2020	
Subject: P'ceutical Inorganic Chemistry (BP104T) Day & Date: Fri, 15/11/2019	
Class: First Year B. Pharm. Semester: I	
Time: 10.30 am - 12.00 pm Max. Marks: 30	
<i>Instructions:</i> 1. All questions are compulsory 2. Right hand side number indicates full marks	
Q.I. Multiple Choice Questions (MCQs)101. Aluminium hydroxide gel is	
a) absorbable antacid b) non-absorbable antacid c) soluble in water d) soluble in alcol	ıol
2. Which of the following is an example of astringent?	
a) Zinc sulphate b) Potash alum c) both a & b d) none of these	
3. Tincture iodine solution is	
a) aqueous iodine solution b) strong iodine solution c) Weak iodine solution d) all of these	
4. Antacid should not be	
a) constipative b) laxative c) both a & b d) protective	
5. Unit of radioactivity is	
a) Calorie b) Newton c) Curie d) Joules	
6. Saline cathartics act by	
a) local irritation of the intestinal tract b) swells in intestine	
c) acting as lubricants d) increasing osmotic load of intestine	
7. Ammonium chloride can be used as	
a) expectorant b) systemic acidifier C) diuretic d) all of the above	
8radiations are deflected in magnetic field	
a) α b) β c) γ d) both a & b	
9. Milk of magnesia IP is an aueous suspension of hydrated	
a) MgSO ₄ b) Mg(OH) ₂ c) MgCO ₃ d) Mg-trisillicate	
10. Which of the following antidote act by adsorption?	
a) activated charcoal b) sodium thiosulphate c) sodium nitrite d) none of the above	
 Q. II. Solve any One 10 1. What is radioactive decay? Describe properties of α, β, γ radiations. Write a note on storage and handling of radioactive substances. 2. Define and classify antidotes with examples. Write in short on cyanide poisoning and its treatment. Explain the properties, assay and uses of sodium thiosulphate 	
 Q. III. Solve any Two (5 Marks each) 1. Define expectorants. Explain the properties, assay and uses of ammonium chloride. 2. Explain the mechanisms by which antimicrobials act. Write any one method of preparation and uses of hydrogen perovide 	

uses of hydrogen peroxide. 3. Define and classify antacids. write a note on properties and uses of Aluminium hydroxide gel.



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Improvement Se	essional Theory Exam	ination 2019-2020
Subject: Communication Skills	(BP 105T)	Day & Date: Saturday, 16/11/2019
Class: F. Y. B. Pharmacy Time: 1.30 pm to 03.00 pm		Semester: I Max. Marks: 30

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks

QI. Long Answers (Answer any ONE)

1*10

- (1) Define Communication and explain Communication Process in detail.
- (2) What is Group Discussion? State its objective, purpose & importance in selection procedure.

QII. Short Answers (Answer any FOUR)

4*5

- 1. What is Psychological Barrier? Explain it in detail.
- 2. Explain Verbal & Non-Verbal Communication in detail.
- 3. Explain Do's and Don'ts of Interviews.
- 4. What are the 4P's of Presentation? Explain them in detail.
- 5. Explain the listening subtypes with techniques to improve Listening.
- 6. State the importance of Body Language & Verbal Comm. in Interviews.



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Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

Improvement Sessional Theory Examination 2018-2019

Subject: Human Anatomy and Physiology-I

Day & Date: Thrusday, 14.11.19

Class: F. Y. B. Pharma

Time: 10:30-12:00

Max. Marks: 30

Semester: I

Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks

QI. Solve the following MCQs

- 1. Example of ball-and-socket joint is
- a. vertebral disc joint
- b. shoulder and hip joints
- c. knee and elbow joints
- d. spinal cord joints
- 2. The ribs, sternum and spine protect these.
- a. Kidneys, bladder, urethra
- b. Heart, lungs, blood vessels
- c. Small intestine, large intestine
- d. all the above
- 3. Tissue is a
 - a. Group of organs
 - b. Group of cells
 - c. Group of tissues
 - d. Group of organisms
- 4. Blood from the heart flows through the blood vessels in which order?
- a) Veins, arterioles, capillaries, venules, arteries
- b) Capillaries, arterioles, arteries, venules, veins
- c) Arteries, arterioles, capillaries, venules, veins
- d) Arteries, capillaries, arterioles, venules, veins
- 5. Blood pressure is highest when leaving which heart chamber?
- a) Right atrium
- b) Right ventricle

c) Left atrium

d) Left ventricle

6. Which of the following structures is the primary organ of the integumentary system?

a) Sensory receptors

b) Sweat glands

c) Bones

d) Skin

7. Which cells produce the pigment that contributes to hair color?

a) Keratinocytes

b) Melanocytes

c) Langerhans cells

d) Merkel cells

8. The main components of the skeletal system are:

a) Blood, bones, cartilage and tendons

- b) Bones, cartilage, joints and tendons
- c) Bones, cartilage, joints and ligaments

d) Cartilage, joints, ligaments and tendons

9. Muscle action can be classified as:

a) Sympathetic and parasympathetic

b) Voluntary and involuntary

c) Autonomic and sympathetic

d) Autonomic and voluntary

10. Which blood group can be transfused into an O Rh- patient?

a) O Rh-

b) O Rh+

c) AB Rh+

d) AB Rh-

QII Long Answers (Answer any 1 out of 2)

1. Define tissue. Explain in detail various types of tissue with classification

2. Explain in details conduction system of heart.

QIII Short Answers (Answer any 2 out of 3)

1. Define joints. Draw a neat labelled diagram of synovial joints

2. Write the various functions of skin with a diagram

3. Write the composition and function of blood

10

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Institute of Pharmacy, Dhule Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

SVKM

	Improvement Sessional Th	eory Examination 201	9.2020
Su	bject : Pharmaceutical Analysis	Day & Date	16/11/2010
CI	ass : F. Y. B. Pharmacy	Semester	· I
Ti	me : 10:30 am to 12:00 noon	Max. Marks	: 30
	Instructions: 1. All questions are comp	pulsory	
	2. Draw a well labelled d	iagram wherever necessary	
	3. Right hand side number	r indicates full marks	
Q.,	Multiple Choice Questions		[10]
	a In polarographic method DME is used	l as electrode	[10]
	a. reference	b. indicator	
	c. partial reference	 d. None of the abo 	ve
	b Potential of standarad hydrogen electr	rode (SHE)is	
	a. 0.298	b. 0.0	
	C. I.U	d. 0.0298	
	c incovic equation can be used in		
	a. Conductometry	b. Potentiometry	
	d In titration collection activity	d. a & b both	
	a Redox	as a specific indicator	
	C acid base	b. complexome	etric
e	Polarography can be utilizes following	d. precipitation	1 .
	a. migration	except	
	c. diffusion	D. residual	
3	is a self-indicator	d. potential difference	
	a. Potassium permanganate	h Cario ammanian	10 .
	c. Iodine solution	d All of the above	ullate
g	In gravimetric analysis the term 'sol' re	efer to	
	a. precipitate	h solution	
	c. dispersion	d. none of the above	
h	are involved in Bromaton	netry type of titrations	
	a. Br	b. BrO4	
	c. Br_2^{\dagger}	d. All of the above	
1	Conductance can be expressed as		
	a. S	b. Ω ⁻¹	
:	C. 12 The unit of an and the	d. a & b both	
J	The unit of current is		
	a. Amperes	b. Volts	
	c. onin	d. mhos	
02	Answer the following at a to the		
Q.2.	Write a short not on meeting short questions (a	ny two)	[10]
	affecting selectivity of FDTA	isking agents. Enlist the fact	ors
b	Describe the construction working of		
c	Explain principle involved in vorient	application of DME.	
d	Explain the principle and application of	pes of todine titration.	
1.05	Prancipie and application of	diazouzation titration	
Q.3.	Answer in detail of following (any one)		
a	Explain the term specific and molar con	ductanca Deparibe in 1	[10]
	various conductometric titration with su	uitable example	adout
ь	Discuss the steps involved in gravimetri	c analysis? Commont or the	12
	differences between lyophilic and lyophi	bic sols obtained during	
	precipitation	and a second	

precipitation ,*
 c Explain the principle behind potentiometer. Write a construction and working of reference electrode and indicator electrode



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Shri. Vile Parle Kelavani Mandal's

Institute of Pharmacy, Dhule Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

	Subje	ect	: Pharmaceutic	al Analysis		Day &	Date	:-	04/	09/2019	•
	Class		: F. Y. B. Pharn	lacy		Seme	ster	:	I		
	Time		: 10:00 - 11:30	-		Max. I	Marks	:	30		
		Instru	ctions: 1. All quest	ons are comp	ulsory						
			2. Draw a v	vell labeled di	aaram w	hereve	er necess	sary			
			3. Right han	d side numbe	r indicate	es full	marks	0			
	0.1	Multir	le Choice Ouestio	ns		. .					[10]
	a	The di	fference between r	esidual and li	imiting c	urrent	is calle	l as			
•	-	а.	Migration		. 0	br Di	ffusion			•	
		c.	Kinetic			d. No	one of th	ese			
	b	Metall	ochromic indicato:	r used in	t	itratio	n.				
		a.	Argentometric	• 2		b- Co	mplexo	metr	ic		
		с.	Redox		· *	d. Gr	avimetr	ic			
	с	Assav	of MgSO4 is	type of	titration						
		a.	Redox			b. Ac	id-Base				
		Pr	Complexometric			d. ba	& c both	L.			
	d	The co	nductance is recip	procal of							
		a.	Resistance			b.	Specif	ic Re	esista	nce	
		c.	Specific conducta	ance 🛥		d.	Eq. Co	ondu	ctand	ce	
	e	The ur	nit of current	·							
		a.	Amperes		b.	Volt					
		c.	Ohm		d.	Mhos	3				
	f	The no	o. of gram eq. of so	lute/lit is call	led as						
		-8.	Normality		b.	Molai	rity				
		с.	Molality		d.	Form	ality				
	g		is a compou	nd of sufficier	nt purity	from v	which a	stan	dard	solution	can
		be pre	pared by direct we	ighing of a qu	antity of	fit, fol	lowed by	dilu	tion	to give a	
, *		defined	d volume of solution	on.				. 8			
		a.	External standar	d	b -	Prima	ary stan	dard			
		. c.	Secondary stands	ard	d.	a&c	both				
	\mathbf{h}		are the me	thods of mini	misation	of err	or	8 g.	A	·. ·	
		a.	Amplification me	thods	b.	Stand	dard add	litior	1		
		с.	Internal standar	ds	Å.	All of	the abo	ve			
	i	Dioxar	ne is a	solvent	- 1 x						
		a.	Protophillic	5	b.	Proto	genic				
•		с.	Amphiprotic		d.	' Aprot	ic			•	
	j	Crysta	l violet is	indicator		97 2 10110 - 104				· · · ·	
		p.	Acidic		b.	Basic	(
		с.	Neutral		d.	Acidi	c & Basi	c bo	th		
					· · · ·						
	Q.2.	Answe	r the following sho	ort questions ((any two)	1. <u>1. 1</u> .	1. A.				[10]
	a	What a	are primary and se	condary stan	dards? E	xplain	theirs t	ypes	s with	L Z	
		suitab	le example				-				
	ъ	Write s	short note on cons	truction and	working	of SHE	9		a 	· * * *	
	с	What a	are conductometric	c titrations? E	xplain it	s type	s.	÷.,	•	il e e e	
	d	Define	non aqueous titra	tions? Explai	n the typ	bes of s	solvent i	ised	6 S		

Q.3. Answer in detail of following (any one)
a Explain the various neutralisation titration curves with suitable example
b Explain types of error with suitable example. Describe the various methods used for minimisation of errors

[10]

c Describe in details about the various steps involved in gravimetric analysis

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Shri. Vile Parle Kelavani Mandal's Institute of Pharmacy, Dhule

Survey No.:499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

Improvement Sessional Theory Examination 2019-2020 Subject: Pharmaceutics IBP103T Day &Date:14.11.2019 Class: FYB. Pharm Semester: I Time: 90 min Max. Marks: 30 Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks Q.No.1 Multiple choice questions $(10 \times 1=10)$ 1. Bottle method is used for formulation of emulsion containing oil. a) Fixed, b) Mineral, c) Volatile, d) None of the above. 2. is used as lubricant for gelato-glycerin suppository bases. a) Glycerin, alcohol and liquid paraffin/ arachis oil, b) soap solution, c) mineral oil, d) None of the above. 3. In suspensions, are added for suspending insoluble powder in a vehicle. a) Suspending agents, b) emulsifying agents, c) surfactants, d) electrolytes 4..... is used as o/w emulsifiers. a) Acacia, b) Span, c) Magnesium oxide, d) None of the above. 5. HLB value of acacia is...... a) 8.0, b) 13.2, c) 4.7, d) 16.7 5. are viscous semisolid emulsion containing one or more drugs in dissolved state meant for external application to the skin. a) Pastes, b) Creams, c) Gels, d) Suppositories. 6. is a widely used synthetic thickening agent in suspensions. a) Carbopols, b) Acacia, c) Sodium CMC, d) Bentonite. 7. form is most stable form of cocoa butter. a) α b) β c) γ d) σ 8. Which of the following is hydrocarbon base a) Lanolin, b) White soft paraffin, c) Macrogols, d) Non of the above 9. Emulsifying agents which reduces interfacial tension are classified as..... a) Sufactants, b) Finely divided solids, c) Polysaccharides, d) Sterols 10. are semisolid unit dosage forms and externally inserted into body cavity to promote defecation. a) Enemas b) Pessaries c) Suppositories d) Bougies Q. No. 2 Long Answers (Answer 1 out of 2) $(1 \ge 10 = 10)$ 1. Discuss stability problems of suspension. Write a note on Suspending agents 2. Define suppository. Write down ideal properties of suppository bases. Discuss various suppository basesalong with examples. Q. No. 3 Short Answers (Answer 2 out of 3) $(2 \times 5 = 10)$ 1. Write a note on ointment bases. 2. Explain various test used for identification of emulsions. 3. Discuss therapeutic incompatibility

SVKM's Institute of Pharmacy, Dhule Improvement Sessional Examination 2018-19

Subject: BP106RBT Remedial Biology	Class: B.Pharm (Sem-I)
Total Marks: 30	Time: 60 min
Roll. No.:	Date: 1 Dec 2018
Q1. Answer any one 1. Explain in details about cardiovascular system. Explain Cardiac cycle 2. Explain the photosynthesis process in details	(10 marks)
 Q2. Answer any fours 1. What is blood and explain its Composition and function 2. Short note on breathing process and its regulation 3. Describe Role of digestive enzymes in digestion process 4. Explain different functions of micro and macro nutrients in plant groups 5. Short note on ECG 6. Explain Pituitary gland and explain different function of hormone restriction 	(20 marks) rowth. release from it



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Second Sessional Theory Examination 2020-2021 (Odd SEM)

Subject: Pharmaceutical Analysis I (BP102T)	Day & Date: 22-04-2021
Class: First Year B. Pharm	Semester: I
Time: 10:30 am – 12:00 pm	Max. Marks : 30

Instructions: 1. All questions are compulsory

2. Draw a well labeled diagram wherever necessary

3. Right hand side number indicates full marks

QI. Multiple Choice Questions (MCQs)

- 1. Closeness of observed value with true value is
 - a) accuracy
 - b) linearity
 - c) precision
 - d) relative error

2. Given the following are the type of systemic error except

- a) Error of methods
- b) Instrumental method
- c) Personal error
- d) Random error

3. In which step, ions or elements are aggregated in Gravimetric analysis?

- a) Supersaturation
- b) Nucleation
- c) Particle growth
- d) None of the above

4. When more than two ions in solution and simultaneously second ion is also precipitated it is called

as

- a) Co precipitation
- b) Post precipitation
- c) A and B
- d) None of the above

5. The reagent which undergoes reduction is an agent and the reagent which undergoes oxidation is a

- a) Oxidizing, Reducing
- b) Reducing, Oxidizing
- c) Complexing, Reducing
- d) None of these

6. Which method are used to determination of primary amine?



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- a) Diazotization Titration
- b) Karl fischer titration
- c) Redox titration
- d) All of the above

7. Indicator used in standardization of sodium thiosulphate is

- a) Starch
- b) Phenolpthaline
- c) ceric sulphate
- d) Methyle red

8. When an analyte that is a reducing agent is titrated directly with a standard iodine solution, the

method is called

- a) iodimetry
- b) iodometry
- c) bromatometry
- d) ceriometry

9. Complexing agent is

- a) Electron donating ion
- b) Electron accepting ion
- c) Both a & b
- d) None of the above

10 The diffusion current in polarography is expressed by the equation

- a) Beer's law
- b) Nernst equation
- c) Bragg's equation
- d) Ilkovik equation

QII. Solve the following (any 1)

1. Define gravimetric analysis. Explain in detail the steps involved in gravimetric analysis

2. State the basic principle involved in polarography. Write down the construction and working of dropping mercury electrode

QIII Short Answers (Answer any 2 out of 3)

- 1. Explain the titration involving iodine
- 2. Write down the construction and working of normal hydrogen electrode
- 3. Write principle and application of Diazotization titrations.

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First Sessional Theory Examination	2020-2021
Subject: Pharmaceutical Analysis I (BP102T)	Day & Date: 11-03-2021
Class: First Year B. Pharm.	Semester: I
Time: 10:30 am – 12:00 pm	Max. Marks : 30
Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks	
Q. I. Solve the following Multiple Choice Questions (MCQs)	10
1. Number of gram equivalents of solute present in one litre of a solution	is called as
a) Molarity	
b) Normality	
c) Molality	
d) Percentage weight by volume	
2. The features of Primary standard substance includes:	
a) High purity	
b) High stability	
c) Low hygroscopicity	
d) All of the above	
3. Indian pharmacopoeia is published by	
a. World health organization	
b. Ministry of health and family welfare	
c. Ministry of science and technology	
d. Food and drug administration	
4. Limit test of iron is based upon the reactions between iron and	
a. Dithiazone	
b. Thioglycolic acid	
c. Hydrogen sulphide	
d. Barium chloride	
5. The color of phenolphthalein in alkanine medium is	

a) Pink

 \bigcirc

- b) Yellow
- c) Brown
- d) Colorless

- 6. Benzene is an example ofsolvent.
 - a) Aprotic
 - b) Protogenic
 - c) Protophilic
 - d) Amphiprotic
- 7. Solubility of precipitate is on addition of common ion.
 - a) Decreased
 - b) Increased
 - c) Remains Constant
 - d) None of These
- 8. In Mohr's methodis used as an indicator.
 - a) Potassium chromate
 - b) Ferric ammonium sulphate
 - c) Ammonium thiocyanate
 - d) Silver nitrate
- 9. In Volhard's methodis used as a titrant
 - a) Potassium chromate
 - b) Ferric ammonium sulphate
 - c) Ammonium thiocyanate
 - d) Silver nitrate

10. The end-point detection in Fajan's method involves color change due to.....

- a) Formation of colored precipitate
- b) Formation of soluble colored complex
- c) Adsorption of indicator on the surface of precipitate
- d) None of the above

Q. II. Solve the following (any 1)

- 1. Explain the neutralization curves in acid-base titrations for i) strong acid vs strong base ii) weak acid vs strong base iii) strong acid vs weak base iv) weak acid vs weak base.
- 2. Explain in detail Mohr's method and Fajan's method.

Q. III. Solve the following (any 2)

- 1. Explain principle involved in limit test for Chloride and sulphate.
- 2. Explain in detail the theories of acid-base indicators.
- 3. Explain the types of solvents used in non-aqueous titrations with suitable examples.

10



Shri. Vile Parle Kelavani Mandal's

Institute of Pharmacy, Dhule Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

Second Re-Sessional Practical Examination 2020-2021

Subject:	Pharmaceutical Analysis	- I	Class:	F. Y. B. Pharm.	Semester: I	
Batch :	Da	ate	04/05/202	21	Max. Marks:	40

Q.	N.		Marks
1		Synopsis (any four)	10
	i	Predict the method and write the principle suitable for assay of the following:	06
		a. Ammonium Chloride	
		b. Ferrous sulfate	
		c. Copper Sulphate	
	ii	Explain the principle and reaction used in limit test for	04
		a. Iron	
		b. Arsenic	
2		To report the % content of given Hydrogen peroxide by permanganometry	15
3		To perform the limit test of Sulfate from the given sample and report whether	10
		it passes or fails the test	
4		Viva-voce	05

(B·R·Dravyakar)

Continuous Assessment Theory Examination 2020-2021

Subject: Pharmaceutical Analysis I

Class: First Year B. Pharm.

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Semester: I

1

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Max. Marks: 03

Division A and B	
Open Book Test	Date: 14-04-2021
Q.1 Write a note on Dropping Mercury electrode	
Assignment	Date: 12-04-2021
Q.1. Write a note on construction and working of standard hydrogen el	ectrode3M
Quiz	Date: 25-03-2021
Q.1. Solve any Three multiple-choice questions	3M
1. Complexing agent is	
a. Electron donating ion	
b. Electron accepting ion	
c. Both a & b	
d. None of the above	
2. Which sentence is false about ligand?	
a. It acts like Lewis base	
b. CN is example of unidentate.	
c. It is electron donor group.	
d. EDTA is example of unidentate.	
3. EDTA is a	
a. Tetradentate ligand	
b. Hexadentate ligand	
c. Octadentate ligand	
d. Pentadentate ligand	
4. All of the following acts as masking agents, except	
a. Ammonium fluoride	
d Tristhanolomina	
 G. The full of the following gate as demosting agent? 	
a Accorbic rold	
a. Ascolute actu	
c. Triethanolomine	
d Formaldehyde - acetic acid	
a. Tormandenyac - accue acia	

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Preliminary Theory Examination 2020-21	Tatal pointe: 75	
		1. The first pharmacy shop was opened in
Preliminary Theory Examinat	tion 2020-21	C London
FY B. Pharm SEM I Subject : I	Pharmaceutics I Maximum Marks 75	O Damascus
Date: 12.5.2021		O Bagdad
		⊖ Chennal
Name of student		
	= Short answer -	2. The chairman of the first edition of IP was
Short-answer text		O Dr. B.N. Ghosh
		O Dr. B. Mukherjee
🗹 (O points)	🔲 🔟 Required	O Dr. Nityanand
	i •	Mr. PrasanaTotta
Roll No:	*	
Short-answer text		 Every dosage form is a combination of drug and different kind of non-drug compone called
		Additives
PRN NO:	*	Non-Additives
Short-answer text		
Class & Division	*	
Short-answer text		
		4. Diffusible drugs are formulated without the use of
(+) (-) (Li Li Ci Millori Mi	Fr Fr https://docs.google.com/forms/d/11B022VgY028pTInko-MANW8GV311zD0clg0o5Htzwmg0/edit 5/12/2021 Preliminary Theory Examination 2020-21 - Google Forms Saccharine
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(L L E 121 with 2020-21 - Google Forms drug, the phenomenon is known as * I the mass becomes coherent but not too *	Frictose Glucose Proteining agent Flavouring agent Chelating agent Preservative Number of the following is anti-foaming agent Methyl cellulose Aspartame
(● () Ir p:/docs.google.com/formal/d118022Vg1025pTh/CoMANNECy211t2D0clg005BHzw 22021 Pretiminary Theory Examination Wetting agent () () b and c 5. When the action of the drug is opposed by the other () Antagonism () Antagonism () Analgesics () Antioxidant 6. A suitable is added to maintain the powders until damp () Colouring agent () Flavouring agent () Sweetening agent	L L C L L L L L L L L L L L L L L L L L	Preservative 9. Disordium EDTA is an example of 9. Disordium EDTA is an example of 9. Sweetening agent 6. Chelating agent 7. Preservative
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	○ 66.67%W/V
3. Douches are meant for application in	
) Buccal cavity	17. Who is the father of medicine
) Rectal cavity	O Ebers
) Vaginal cavity	O Hippocrates
) Nasal cavity	C Egyptian
	O Pontus
 A genetically determined abnormal or unusual response to a drug is 	
) Idiosyncrasy	18. Use of formulations made up of numerous plants referred as
) Tolerance	Galenicals
) Genetic polymorphism	O Parenteral
Salirylism	O Plant Vehicles
y outopoint	Generics
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5. Powders used for external use are	19 The seventh edition of IP was published in
) Dusting powder	
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121 Preliminary Theory Examination 2020-21 - Google Forms	5/12/2021 Preliminary Theory Examination 2020-21 - Google Forms Suspension	
	O Enemas	
Medicated dusting powders		
	32. Elixirs are	
	Hydro alcoholic liquid	
) Surgical dusting powders		
) Natural dusting	◯ Viscous	
) powder	🔿 Semi solids	
29. The monophasic liquid dosage form is Solutions	* 33 Liniments must not be applied on the	
) True	() Swelled	
) Course	O Broken	
) Fine		
) Micro		
	O Normal	
10. Nasal drops should have pH	* 34 drops are used in otic cavity	
2 5-0	C Eye	
۰۶ ر م) Ear	
7-8 (○ Nose	
) 8-9		
11. Which of the following is not monophasic liquid dosage form	* 35. Which of the following excipient act as preservative	
) Solution		
	Sorbitan monolactrate	
Denzethenlum chloride	Sorbitan monolactrate Sorbitan monolactrate Tr D D Tr D	Ξ
Tr Tr Tr Tr Tr Descriptionna/d/118022VgY028pThKc-MANWBCy3112D0clg0o68Hzwmg0/dedit 211 Pretiminary Theory Examination 2020-21 - Google Forms Denzethenlum chloride Thiomersal	Sorbitan monolactrate	Ē
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Tr Tr Tr Tr Tr Tr Tr Tr Tr Tr	Sorbitan monolactrate Performance Sorbitan monolactrate Performance Preliminary Theory Examination 2020-21 - Google Forms 20% acacia, 15% tragacanth, 30% starch, 35% sucrose 20% acacia, 25% tragacanth, 30% starch, 40% Sucrose 20% acacia, 15% tragacanth, 20% starch, 45% Sucrose 20% acacia, 15% tragacanth, 20% starch, 45% Sucrose 16.23 16.23 24 15 7	E
Tr Tr Tr Tr Tr Tr Tr Tr Tr Tr	Sorbitan monolactrate	E
Tr Tr Tr Tr Tr Tr Tr Tr Tr Tr	Sorbitan monolactrate	E
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r remaining theory challmassin court - Stople Forms	5/12/2021 Preliminary Theory Examination 2020-21 - Google Forms
M. L. Shroff	47. Daily adult dose of drug is 120mg. How much of it should be given to a child of 11 years? *
) R. N. Chopra	O 57.39
) B. N. Ghosh	○ 60
) None of the above	O 50.39
4 are viscous liquid used in cough that relieves irritation of the mucous membranes *	O 37
the throat/mouth by forming a protective film.	48. Which of the following co-solvents are used to increase the solubility of a drug? *
	C Ehanol
) Sedatives	⊖ Sorbitol
) Elixirs	
) throat paints	All the above
15. A drug that neutralizes excess gastric acid are *	
) Antacid	49. Which part of prescription contains direction to the patient regarding how, how much and *
) Analgesic	Superscription
) Antiamebic	
Antiepilepilc	
16. The powders containing solids that absorb moisture from air and dissolve to form solution *	U Signatura
re called	To the set of the set
Deliquescent powders	50. How will you mix potent arug with large amount or alluent?
) Efflorescent powders	
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tions, geogle.com/forms/d/18022VgY028pThKc-MANNECy3112D0clg068Hzwmj0/edit 221 Preliminary Theory Examination 2020-21 - Geogle Forms 31. The dosage form intended to be rubbed on the skin for rubeficient and counterirritant effect * ire called Lotions Liniments Gels	https://docs.google.com/forms/d/18022VgY028pThKo-MANMBCy3112D00g0o8Hzwmj0/edit 14/21 Sr122021 Preliminary Theory Examination 2020-21 - Google Forms Bougles Cones 55. 13 (apothecaries) = * 437.5 grains 480 grains 2000 proloco 2000 proloco
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docs.geogle.com/formal/d118022Vg7026pThKc-MANWECy3112D0clg068Hzwmj0Vedt 13/21 21 Preliminary Theory Examination 2020-21 - Google Forms 51. The dosage form intended to be rubbed on the skin for rubeficient and counterirritant effect * are called 1 Lotions 1 Lotions 2 Gets 1 Dentifrices 32. The powders on exposure to humid atmosphere or during trituration, liberates wholly or arity, the water of crystallization are called 1 Hygroscopic powders 2 Efforescent powders 33. Drugs that destroys or eliminates intestinal parasitic worms are called * Analgesic Anthelmintic	https://docs.geogle.com/forms/d/18022VgY028pThK-MAWBCy31t2D0tg006B/ttxm/04eft 1421 Sr122221 Preliminary Theory Examination 2020-21 - Google Forms
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decas_geogle.com/cmms/d118022VgY028pThike-MMMMeGy8112D0clg068Hxmmglided 1271 221 Preliminary Theory Examination 2020-21 - Google Forms 231 Preliminary Theory Examination 2020-21 - Google Forms 241 Preliminary Theory Examination 2020-21 - Google Forms 251. The dosage form intended to be rubbed on the skin for rubeficient and counterirritant effect * 252. The consequence of the skin for rubeficient and counterirritant effect 253. The powders on exposure to humid atmosphere or during trituration, liberates wholly or arthy, the water of crystallization are called 254. The powders on exposure to humid atmosphere or during trituration, liberates wholly or arthy, the water of crystallization are called 255. The powders 256. Efflorescent powders 257. Eucecic mixtures 253. Drugs that destroys or eliminates intestinal parasitic worms are called 353. Drugs that destroys or eliminates intestinal parasitic worms are called 353. Anthelminitic 354. Anthelminitic 355. Antifiarial 354. The semisolid unit dosage form intended for vaginal administration are called	https://docs.geogle.com/doms/018022/01/02/gb1Mx-MANNBC/Q112/D01gb08/B1MxmW04ed1 S122221 Pretiminary Theory Examination 2020-21 - Google Forms Bougles
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Water removable bases	O Micro emulsion
1 Hydrocarbon bases	O Multiple emulsion
() None	
	63. In case of O/W emulsion, creaming takes place at
59. Which of the following is most commonly used suppository base *	O Down side
O Cocoa butter	O Upside
O PEG 1000	At interface between two phase
O PEG + Hexanetriol	
○ None	
	64. What is the dispersion of a liquid in another liquid called?
60. Chemical incompatibility in which change in formulation is needed is called as	O Gel
Adjusted incompatibility.	C Emulsion
	○ Aerosol
61. Tetracycline interacts with calcium to form chelates is an example of *	
Chemical incompatibility	65. For Ideal suspension ,the sedimentation volume should be
O Physical Incompatibility	C Equal to one
O Tolerance incompatibility.	O More than one
	O Less than one
43 In conductivity test if hulk alour on passing electric surrant so what would be the type of	
emulsion?	
○ W/O emulsion	66. Rate of sedimentation is high in?
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	strengthening the nonpolar tails of the emulsifier *
0	strengthening the polar heads of the emulsifier
0	thickening the continuous phase
75.	In wet gum method, 4:2:1 consist of
0	4 parts oil, 2 parts water, and 1 part emulsifier
0	4 parts water, 2 parts oil, and 1 part emulsifier
0	4 parts emulsifier, 2 parts water, and 1 part emulsifier
0	4 parts oil, 2 parts emulsifier and 1 part water

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O Slips '	16. In purified water the concentration of sugar according to IP is *
	○ 85% W/W
	○ 66.67 % W/W
C Reports	
	○ 66.67%W/V
13. Douches are meant for application in	
O Buccal cavity	17. Who is the father of medicine
	() Ebers
Vaginal cavity	() Hibbociates
O Nasal cavity	C Egyptlan
	O Pontus
14 A genetically determined abnormal or unusual response to a drug is	
	18. Use of formulations made up of numerous plants referred as
○ Tolerance	
◯ Genetic polymorphism	O Parenteral
🔿 Salicylism	O Plant Vehicles
	Generics
15. Powders used for external use are	
O Dusting powder	19. The seventh edition of IP was published in
	0 2014
Compound powder	○ 2015
O Divided powders	O 2013
O Effervescent powders	O 2000
5/21	https://docs.doodle.com/ioms/d/ HBUZZVg102op111KC-WANWBCyS112D00g000012Winjorean
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Shri. Vile Parle Kelavani Mandal's Institute of Pharmacy, Dhule

Remedial Sessional Theory Examination 2019-2020					
Subject: P'ceutical Inorganic Chemistry (BP104T) Day & Date: Fri, 06/09/2019					
Class: First Year B. Pharm. Semester: I					
Time: 10.30 am – 12.00 pm Max. Marks: 30					
Instructions: 1. All questions are compulsory					
Q.I. Multiple Choice Questions (MCQs) 10					
1. Lewis base is					
a) electron pair donor b) electron pair acceptor					
c) proton donor d) proton acceptor					
2. Citric acid is used in limit test for iron to prevent					
a) precipitation of iron by ammonia b) reduction of iron					
c) precipitation of iron by thioglycollic acid d) complexation with thioglycollic acid					
3. Indian pharmacopoeia is published by					
a) Ministry of health and family welfare b) food and drug administration					
c) world health organization d) ministry of science and technology					
4. Lead acetate cotton plug is used in limit test for arsenic to					
a) trap hydrogen sulphide b) trap thiosulphate					
c) trap hydrogen sulphate d) trap arsine gas					
5. Limit test for heavy metals is based on reaction between metallic impurities and					
a) hydrogen sulphide b) hydrogen sulphate c) acetic acid d) ammonia					
6. The first Indian pharmacopoeia was published by Indian Pharmacopoeia committee in					
a) 1948 b) 1955 c) 1960 d) 1966					
7. gamma rays are electromagnetic radiations with a wavelength					
a) much larger than those of lightb) much shorter than those of light					
c) Equal to light d) none of the above					
8. Unit of radioactivity is					
a) Calorie b) Newton c) Curie d) Joules					
. The sources of impurities include					
a) raw materials b) manufacturing process					
c) instability of product d) all of these					
10. The limit test for chlorides involves reaction between soluble chlorides and					
a) silver nitrate b) nitric acid c) barium sulphate d) HCl					
Q. II. Solve any One 10					
1. Explain different theories of acid and base. Explain the mechanism of buffer action.					
2. What is radioactivity? Explain the properties of alpha, beta and gamma radiations.					
Q. III. Solve any Two (5 Marks each) 10					
1. Explain the principle behind limit test of Iron and Sulphate.					
2. Explain the principle of arsenic limit test. Draw a neat labelled diagram of Gutzeit apparatus.					
3 Explain the sources of impurities					

3. Explain the sources of impurities.



Shri. Vile Parle Kelavani Mandal's

Institute of Pharmacy, Dhule Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai Agra Highway

12.60

Remedial Sessional Theory Examination 2018-2019				
Subject: Pharmaceutics I BP103T Day & Date: 05.09.2018				
Class: FYB. Pharm Semester: I				
Time: 10 am to 11.30 pm Max. Marks: 30				
Instructions: 1. All questions are compulsory 2. Draw a well labeled diagram wherever necessary 3. Right hand side number indicates full marks Q.No.1 Multiple choice questions (10 X 1=10)				
 are used to deodorize oral cavity. a) Throat paint, b) Mouth washes, c) Elixirs, d) Syrups. 2. Suspension containing -loose aggregation of particles is calledsuspension. a) Flocculated b) Deflocculated, c) Mixture, d) None of the above. 3 form is most stable form of cocoa butter. a) α b) β c) γ d) σ 				
 4. According to Stokes's equation rate of sedimentation will with increase in viscosity of dispersion phase. a) Decrease, b) Increase, c) Remain unaffected, d) None of the above. 5 are semisolid unit dosage forms and externally inserted into body cavity to 				
promote defecation. a) Enemas b) Pessaries c) Suppositories d) Bougies				
6 is used as lubricant for gelato-glycerin suppository bases.				
a) Glycerin, alcohol and liquid paraffin/arachis oil, b) soap solution, c) mineral oil, d) None of				
the above.				
 7. If water soluble dye (amaranth) is mixed with the emulsion and examined under the microscope, the dispersed globules appears red and ground colorless indicates				
a) Sufactants, b) Finely divided solids, c) Polysaccharides, d) Sterols.				
 9. In wet gum method containing fixed oil ratio of oil: water:acacia to formulate primary emulsion is a) 4:2:1, b) 3:2:1, c) 2:2:1, d) 2:2:1 10 are clear, pleasantly flavored, sweetened hydroalcoholic monophasic liquid containing dissolved active ingredients intended for oral use. a) Syrups b) Elixirs c) Linctus d) Suspension. 				
Q. No. 2 Long Answers (Answer 1 out of 2) $(1 \times 10 = 10)$				
 Define emulsion. Write down ideal properties of emulsion. Write a note on emulsifying agents. Define dosage form. Discuss dosage form in detail. Q. No. 3 Short Answers (Answer 2 out of 3) (2 x 5 = 10) 				
1. Write a note on lotion and liniment. A shall a set of a standard and llast a set				
 2 Define suspension. Discuss ideal properties of suspension. 3. Write short note on parts of prescription. 				

Shri. Vile Parle Kelavani Mandal's

INSTITUTE OF PHARMACY, DHULE

Survey No. 499/1, Plot No.3, Behind Gurudwara, Mumbai-Agra National Highway

Re-Sessional Theory Examination 2020-2021

Subject: Comm. Skills	Day & Date: Friday, 07 May 2021	
Class: FY/ DSA	Semester: First	
Time: 10.30 to 12.00	Max. Marks: 30	
Time. 10.30 to 12:00		

Instructions: 1. All questions are compulsory

2. Draw a well labeled diagram wherever necessary

3. Right hand side number indicates full marks

QI Long Answers (Answer any ONE)

- Explain the purpose and process of Selection Interview. What are the Dos and Don'ts of Selection Interview? (Unit 04)
- (2) Explain Group Interaction in detail? Elucidate the salient characteristics of Group Discussion. (Unit 05)

QII Short Answers (Answer any FOUR)

- 1. Explain the role of Effective Writing in your profession in detail. (Unit 03)
- 2. Explain the elements involved to make writing more effective. (Unit 03)
- 3. How will you be Effective Listener? State the strategies and methods you will adopt for Effective Listening.(Unit 03)
- 4. How will you prepare and deliver your presentation more effectively? (Unit 04)
- 5. What are the Dos and Don'ts of Group Discussion? (Unit 05)
- 6. What are the ways of individual contribution in GD? (Unit 05)

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SVKM

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4*5